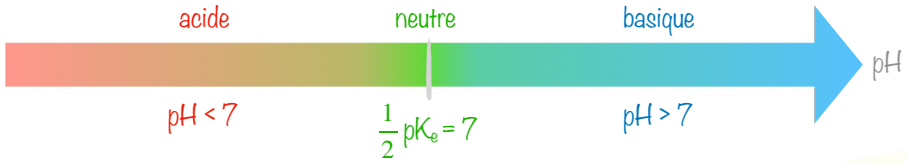


$$2\,H_2O\,(\ell) \longrightarrow H_3O^+\,(aq) + HO^-\,(aq)$$

produit ionique de l'eau

$$K_e = \frac{[H_3O^+]_{eq} \cdot [HO^-]_{eq}}{c^{o2}}$$

$$AH + H_2O \longrightarrow A^- + H_3O^+$$

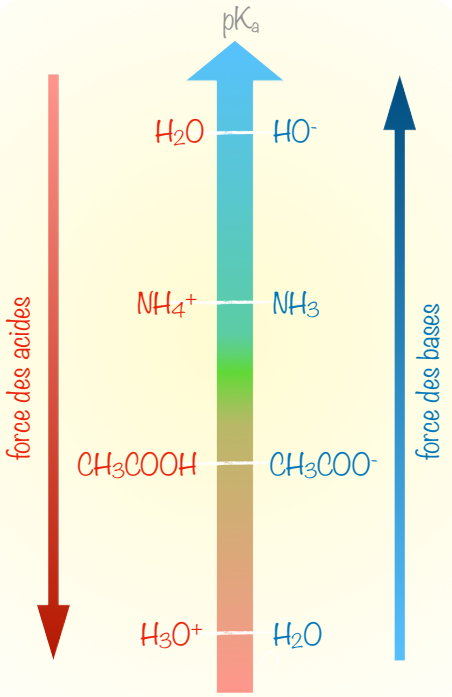
constante d'acidité

$$K_A = \frac{[H_3O^+]_{eq} \cdot [A^-]_{eq}}{[AH]_{eq} \cdot c^o}$$

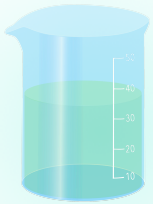
$$pK_A = -\log(K_A)$$

si $\tau = \frac{x_f}{x_{max}} = 1$ acide fort

si $\tau = \frac{x_f}{x_{max}} < 1$ acide faible



solution tampon



pH varie peu
si ajout d'acide, de base ou dilution

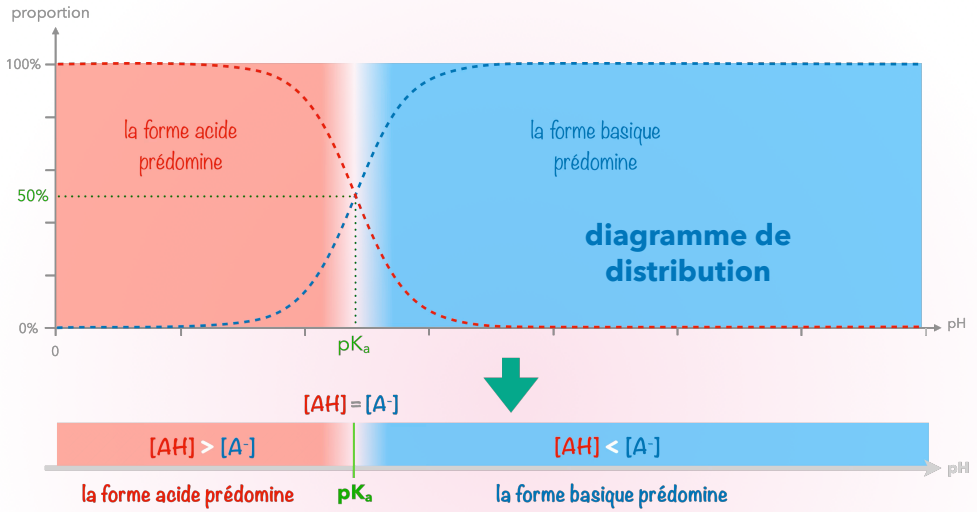
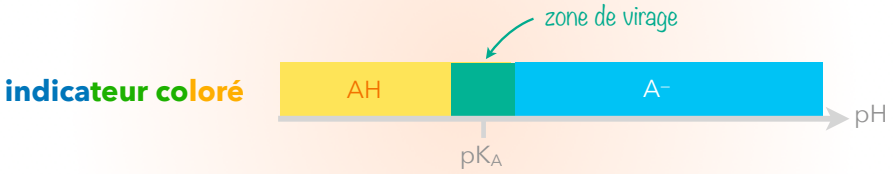


diagramme de prédominance



acides aminés