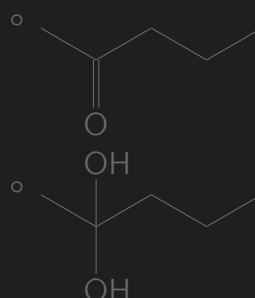
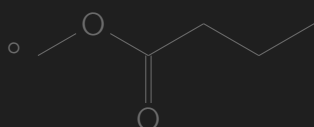
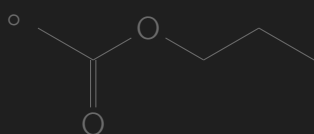


Mini Quizzes

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|-----------------------|---|
| Week 3–4 — Chapter 16 | 2 |
| Week 2 — Chapter 15 | 4 |
| Week 1 — Chapter 14 | 6 |

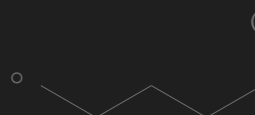
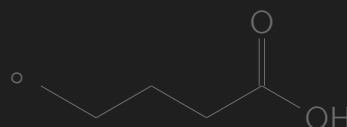
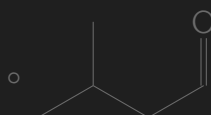
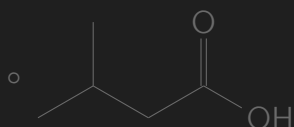
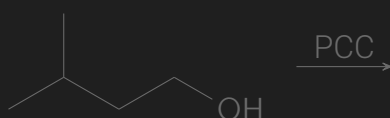
Week 3-4 – Chapter 16

1. What is the major product for the following reaction.



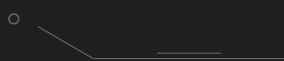
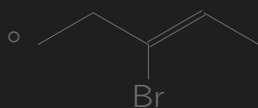
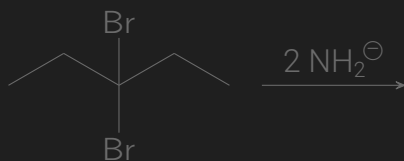
- CrO_3 is an oxidizing agent, which means electron density will be pulled away from the carbon to form a double bonded oxygen.
- Oxygen does not insert itself into the chain.

2. Give the major product for the following oxidation.



- PCC is a mild oxidizing agent that is commonly used for selective oxidation of alcohols to aldehydes or ketones.
- In this case we started with a terminal enol, which would produce an aldehyde.
- The carbonyl group is not affected, it should not change.

3. What is the major product from the following reaction?

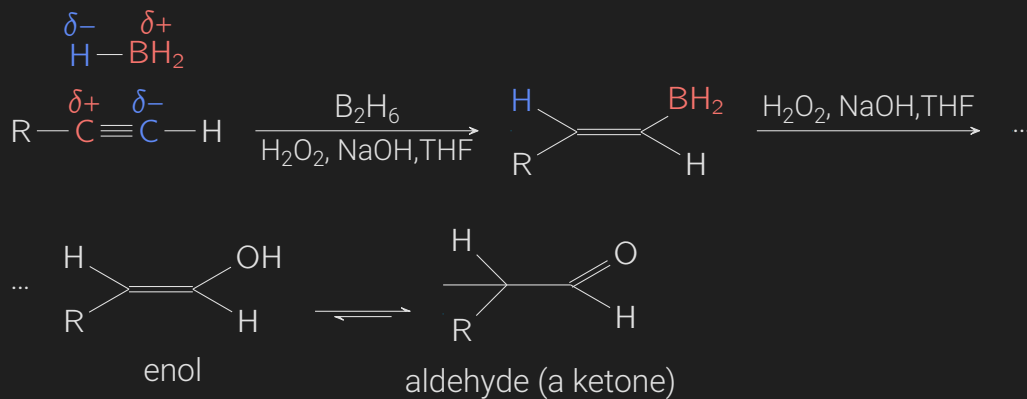


- This looks like dihaloalkane elimination. I took this quiz a week early, so this explanation might not be the best, but looks like NH_2^- is acting as a reducing agent(?); causing the elimination of bromine, leaving the carbon to form a triple bond.

4. The major product of a hydroboration oxidation reaction on a terminal alkyne is

- a carboxylic acid
- alkane
- ketone
- aldehyde

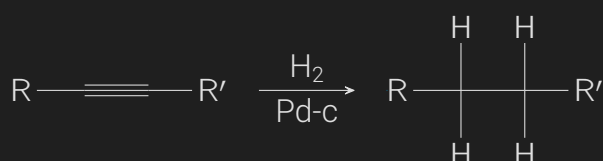
- Example of an alkyne undergoing a hydroboration-oxidation reaction:



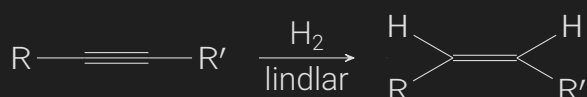
Week 2 – Chapter 15

1. The reagent needed to convert 2-butyne to cis-2-butene is

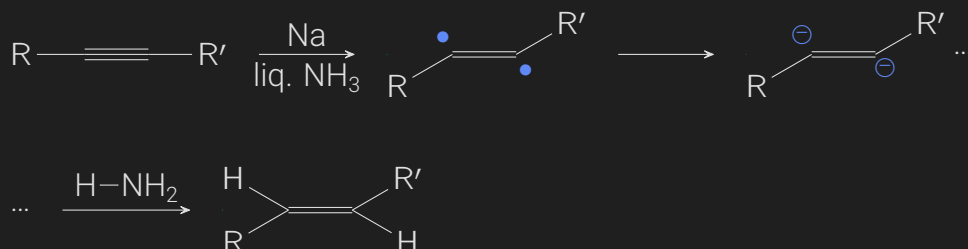
- $\text{H}_2/\text{Pd}-\text{C}$
- Li/NH_3
- Na/NH_3
- $\text{H}_2/\text{LindlarCatalyst}$
- Complete hydrogenation of an alkyne:



- Alkyne \rightarrow cis-alkene; use of lindlar catalyst (Pd-c poisoned with lead) limits further reduction by controlling hydrogens available:



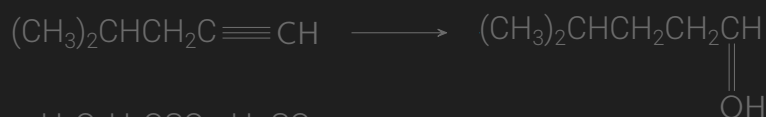
- Alkyne \rightarrow trans-alkene; using generation of free radicals (•, single electron) that pair up with another electron generated by the dissociation of $\text{Na} \rightarrow \text{Na}^+ + \text{e}^-$ to create a free pair of electrons that then receive a hydrogen from NH_3 :



2. A mixture of 1-heptyne, 2-heptyne, and 3-heptyne was hydrogenated in the presence of a palladium catalyst until hydrogen uptake stopped. If one assumes that the hydrogenation went to completion for all the reactants present in the mixture, how many distinct seven-carbon isomers were produced?

- Only 1
- 2
- 4
- 6
- $\text{H}_2/\text{Pd}-\text{c}$ (palladium catalyst) generates completely saturated alkenes, thus the location of the double bond in a heptyne will make no difference overall.

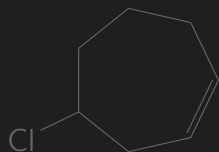
3. Give the best reagents for the reaction



- H_2O , H_2OSO_4 , HgSO_4
 - BH_3 , H_2O_2 , NaOH
 - $\text{K}_2\text{Cr}_2\text{O}_7$
 - H_2 , Lindlar Catalyst
 - First, this is a hydration reaction, so that limits just the first two options.
 - Hydration using H_2O and H_2OSO_4 or HgSO_4 does have difference, but both follow Markovnikov's rule and end produce internal enols and thus internal ketones.
 - Hydroboration-oxidation reaction follows anti-Markovnikov rule and produces a terminal enol and thus an aldehyde, which is the desired product.
4. Which of the alkyne addition reactions below involves an enol intermediate?
- Hydroboration/oxidation
 - dil. H_2SO_4 in HgSO_4
 - Hydrogenation
 - Both hydroboration/oxidation and dil. H_2SO_4 in HgSO_4
 - See question three, both hydroboration/oxidation and dil. H_2SO_4 in HgSO_4 are used in hydration, which have enol intermediates.
 - Hydrogenation only has to do with adding hydrogens to saturate the alkyne through elimination reactions, which question one covers.

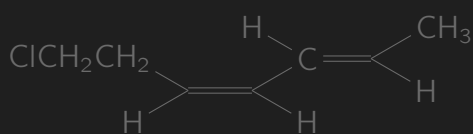
Week 1 – Chapter 14

1. Name the structure:



- 1-chloro-3-cycloheptene
- 4-chloro-1-cycloheptene
- 4-chloro-1-cyclohexene
- 6-chloro-1-cycloheptene
- When numbering the parent chain, the double bond should receive the lowest number possible; **k=1**
 - Note: define the location *k* of the double bond as being the number of its first carbon, not at the end.
- The locant (*k*) of the double bond should be placed right before the suffix of “ene,” though, it was previously recommended before the parent (both are acceptable), e.g., 2-pentene = pent-2-ene; **1-cycloheptene**
- Name and the side groups (other than hydrogen) according to the appropriate rules; **chloro**
- Define the position of each side group as the number of the chain carbon it is attached to; **4-**

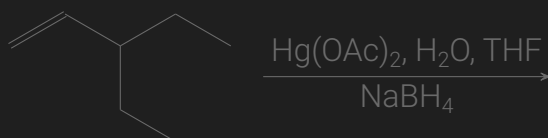
2. Name the structure:



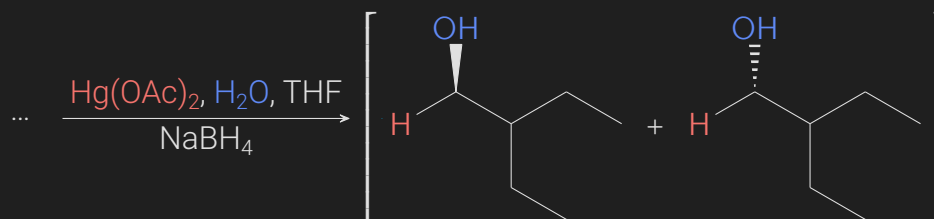
- (2E,4E)-7-chloro-2,4-heptadiene
- (2Z,4Z)-7-chloro-2,4-heptadiene
- (2Z,4E)-7-chloro-2,4-heptadiene
- (2E,4Z)-7-chloro-2,4-heptadiene
- **E-Z notation:** recommended instead of *cis* and *trans* in order to account for cases that has more than two different groups attached to the double bond by first determining the priority using the Cahn-Ingold-Prelog System.
 - **E, entgegen, “opposite”.**
 - **Z, zusammen, “together”; “on ze zame zide.”**

- When numbering the parent chain, the double bond should receive the lowest number possible; $k=2$
 - The two highest priority groups are on **opposite** sides; **2E**
- There is more than one double bond; $k_2 = 4$
 - The two highest priority groups are on **same** side; **4Z**

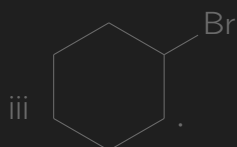
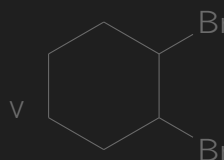
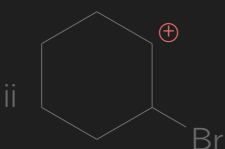
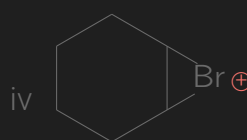
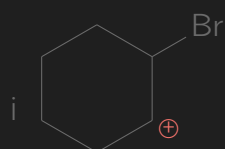
3. How many stereoisomeric product(s) do you get in the reaction below.



- Oxymercuration-demercuration reactions follow Markovnikov's rule, i.e., H^+ is added to the carbon with the **greatest** number of hydrogen atoms while the X^- **component** is added to the carbon with the **fewest** hydrogen atoms.
- Drawing the intermediate is not necessary, and no chiral centers are found in the products:



4. Which reaction intermediate is formed when Br_2/CCl_4 reacts with cyclohexene?



- **Halogenation:** a reaction that involves the addition of one or more halogens to a compound or material.
 - The addition of halogens to alkenes proceeds via **intermediate halonium ions**.

- **Halonium ion:** any onium ion containing a halogen atom carrying a positive charge. This cation has the general structure: $R \cdot +X R'$
- **Onium ion:** a **cation** formally obtained by the protonation of mononuclear parent hydride of a pnictogen (group 15 of the periodic table), chalcogen (group 16), or halogen (group 17); Br^{\oplus} in our case.