## The all-seeing eye of resonant Auger electron spectroscopy: a study on aqueous KCl

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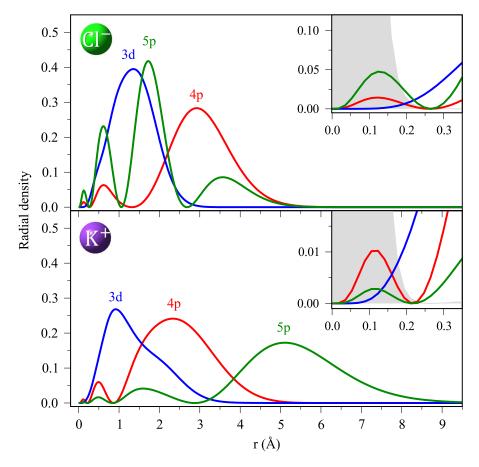


Figure 1: Radial density distributions of the singly-occupied natural orbital occupied by the excited electron corresponding to the  $1s^{-1}4p$ ,  $1s^{-1}3d$  and  $1s^{-1}5p$  core excitations in  $K^+$  (lower panel) and  $Cl^-$  (upper panel). The insets show the region of distances relevant for the overlap with the 1s core orbital whose radial density is shown as a grey shaded area.

In what follows we give a simple explanation of the difference in the radial density distributions of the  $1s^{-1}4p$  and  $1s^{-1}5p$  states in  $K^+$  and  $Cl^-$ . In the case of  $K^+$ , the excited electron mainly sees a 2/r potential. In addition, it sees a short range potential originating from the point-like nucleus and the screening electrons. The influence of the latter can be described by a quantum defect  $\delta \neq 0$ , which is almost constant for the entire infinite Rydberg series. However, in case of  $Cl^-$  the outer electron does not experience a Coulomb potential and the short range potential becomes dominant. As a result of the absence of the Coulomb potential we see a different behavior in the properties of the states, like e.g. only a finite number of bound states (here obviously 4p)<sup>1</sup>. In contrast to this, the 3d and 5p are not bound.

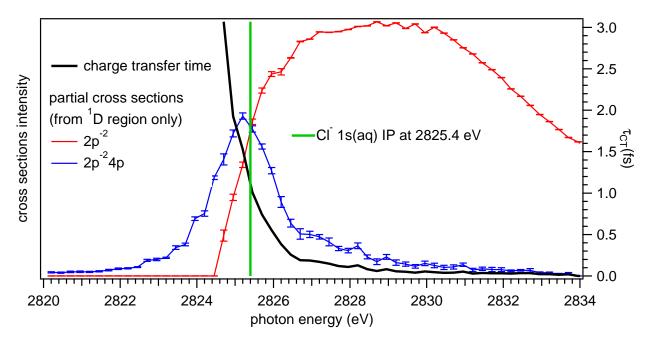


Figure 2: Partial cross sections and charge transfer time extracted from Fig. 3. The blue and red curves are obtained by integrating the area of the  $2p^{-2}$  and  $2p^{-2}4p$  final states (<sup>1</sup>D state region only) at each photon energy step. From these curves we determine the charge transfer time  $\tau_{\rm CT}$  according to the formula  $\tau_{\rm CT} = \tau_c l/d$ , with  $\tau_c$  being the Cl 1s core-hole lifetime and l/d being the intensity ratio of the localized  $(2p^{-2}4p)$  and delocalized  $(2p^{-2})$  states at a given excitation energy. The green line defines the  ${\rm Cl}_{\rm aq}^-(1s)$  ionization potential.

## References

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