

Q1. 21 January Shift 2

The r.m.s. speed of oxygen molecules at 47°C is equal to that of the hydrogen molecules kept at ____ $^{\circ}\text{C}$. (Mass of oxygen molecule/mass of hydrogen molecule = $32/2$)

- (1) -235 (2) -253 (3) -100 (4) -20

Q2. 22 January Shift 1

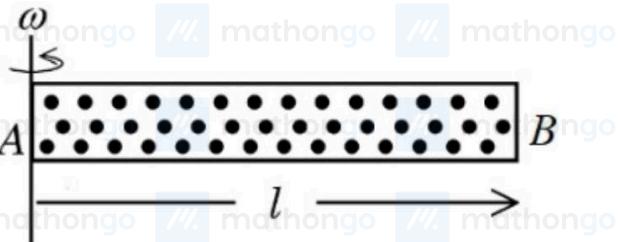
The volume of an ideal gas increases 8 times and temperature becomes $(1/4)^{\text{th}}$ of initial temperature during a reversible change. If there is no exchange of heat in this process ($\Delta Q = 0$) then identify the gas from the following options (Assuming the gases given in the options are ideal gases):

- (1) O_2 (2) NH_3 (3) CO_2 (4) He

Q3. 22 January Shift 1

A cylindrical tube AB of length l , closed at both ends contains an ideal gas of 1 mol having molecular weight M .

The tube is rotated in a horizontal plane with constant angular velocity ω about an axis perpendicular to AB and passing through the edge at end A , as shown in the figure. If P_A and P_B are the pressures at A and B respectively,



then (Consider the temperature is same at all points in the tube)

- (1) $P_B = P_A$ (2) $P_B = P_A \exp(M\omega^2 l^2 / 3RT)$
 (3) $P_B = P_A \exp(M\omega^2 l^2 / RT)$ (4) $P_B = P_A \exp(M\omega^2 l^2 / 2RT)$

Q4. 22 January Shift 2

Consider two boxes containing ideal gases A and B such that their temperatures, pressures and number densities are same. The molecular size of A is half of that of B and mass of molecule A is four times that of B . If the collision frequency in gas B is $32 \times 10^{18}/\text{s}$ then collision frequency in gas A is ____ /s.

- (1) 2×10^8 (2) 8×10^8 (3) 4×10^8 (4) 32×10^8

Q5. 24 January Shift 1

A gas of certain mass filled in a closed cylinder at a pressure of 3.23 kPa has temperature 50°C . The gas is now heated to double its temperature. The modified pressure is ____ Pa.

Q6. 28 January Shift 2 mathongo // mathongo // mathongo // mathongo // mathongo // mathongo

The mean free path of a molecule of diameter 5×10^{-10} m at the temperature 41°C and pressure 1.38×10^5 Pa, is given as _____ m. (Given $k_B = 1.38 \times 10^{-23}$ J/K).

- (1) $10\sqrt{2} \times 10^{-8}$ (2) $2\sqrt{2} \times 10^{-8}$ (3) $2\sqrt{2} \times 10^{-10}$ (4) 2×10^{-8}

ANSWER KEYS

1. (2) 2. (4) 3. (4) 4. (3) 5. 3730 6. (2)

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