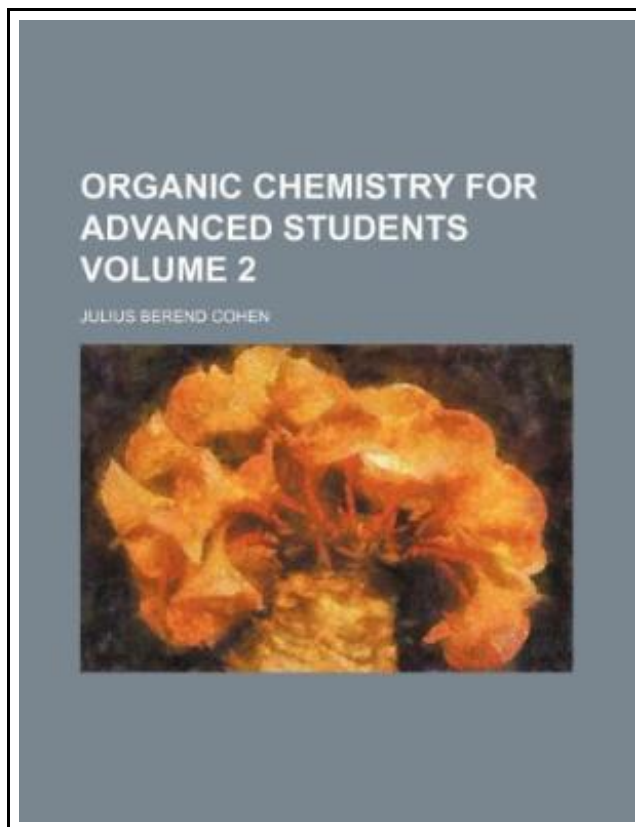


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1913 Excerpt: .stages and is therefore a composite reaction; cf. Orton and King, Trans. Chem. Soc., 1911, 89, 1369; also Orton and Jones, Trans. Chem. Soc., 1909, 95, 1456.  $\text{CH}_3\text{CO}_2\text{C}_2\text{H}_5 + \text{NaOH} = \text{CH}_3\text{CO}_2\text{Na} + \text{C}_2\text{H}_5\text{OH}$  The saponification proceeds at a rate which can be conveniently measured at temperatures between 0 and 25 if dilute solutions are employed. If the original solution contains a grm.-mols. (mole) of ester and  $a$  mols of hydroxide per unit volume, and if  $x$  mols of ester have been saponified after time  $t$ , the concentrations of the reacting substances at this moment will be  $a-x$  and  $b-x$  respectively. According to the mass law, the speed of the change will be given by  $\frac{dx}{dt} = k_2(a-x)(b-x)$  and this on integration becomes,  $\frac{1}{b-a} \ln \frac{a(b-x)}{b(a-x)} = k_2 t$  In the following table are given the data obtained by Reicher for the saponification of ethyl acetate at 15-8, the alkali being present in excess ( $b > a$ ) in the one experiment, whilst the ester predominated in the second ( $a > b$ ). The quantities of saponified ester ( $x$ ) are expressed in terms of the standard acid solution which was used in following the progress of the change. If the reacting substances are present in equivalent proportions ( $a = b$ ), the rate of change at time  $t$  is given by  $\frac{dx}{dt} = k(a-x)^2$  from which  $\frac{1}{a-x} = k_2 t + \frac{1}{a}$  That this is in agreement with the actual course of saponification under these conditions is shown by the...



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