Parishram (2025)

Physical Chemistry

Electrochemistry

DPP: 1

Q1 Galvanic or voltaic cell is a device used to convert chemical energy produced in..... reaction into..... energy.
(A) Chemical, chemical
(B) Electrical, chemical
(C) Redox, electrical
(D) Redox, redox

Q2 In Daniell cell
(A) Zn rod as anode and Cu rod as cathode
(B) Zn rod as cathode and Cu rod as anode
(C) Zn rod is represented by (+ve) electrode
(D) Cu rod is represented by (-ve) electrode

Q3 The potential difference between the two electrodes of a galvanic cell is called
(A) Electrode potential
(B) Zeta potential
(C) Cell potential
(D) Junction potential

Q4 In Galvanic cell, conventional,
(A) Anode is on the left hand side and is (+ve)
(B) Cathode is on the left hand side and is (-ve) electrode
(C) Cathode is on the left hand side and is (+ve) electrode
(D) Anode is on the left hand side and is (-ve) electrode, cathode is on right hand side and is +ve electrode

Q5 A potential develops between the electrode and electrolyte which is called as
(A) Electrode potential
(B) Zeta potential
(C) Cell potential
(D) No potential develops

Q6 An electrochemical cell stops working after some times because
(A) Electrode potential of both the electrodes becomes zero.
(B) Electrode potential of both electrodes becomes equal.
(C) One of the electrode is eaten away.
(D) The reaction starts proceeding in opposite directions.

Q7 Which of following statements is/are correct for a galvanic cell?
I. Reduction at cathode and oxidation at anode.
II. Reduction at anode and oxidation at cathode.
III. Electrons flow from anode to cathode
IV. Electrons flow from cathode to anode
(A) I, IV (B) II, III
(C) II, IV (D) I, III

Q8 When the salt bridge is removed, the voltage:
(A) Does not change.
(B) Increase to maximum.
(C) Decrease to half the value.
(D) Drops to zero.

Answer Key

Q1	(C)	Q5	(A)
Q2	(A)	Q6	(B)
Q3	(C)	Q7	(D)
Q4	(D)	Q8	(D)

Hints & Solutions

Note: scan the QR code to watch video solution

Q1 Video Solution:

Q2 Video Solution:

Q3 Video Solution:

Q4 Video Solution:

Q5 Video Solution:

Q6 Video Solution:

Q7 Video Solution:

Q8 Text Solution:

The salt bridge helps to complete the internal circuit of the cell.

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