

**Section A: Multiple Choice: (1 mark each)****Write your answers on the multiple choice grid on page 2**

Select the alternative A, B, C or D that best answers the question. Fill in the response oval completely.

**Sample:**  $2 + 4 =$  (A) 2 (B) 6 (C) 8 (D) 9  
A ☐ B ☒ C ☐ D ☐

If you think you have made a mistake, put a cross through the incorrect answer and fill in the new answer.

A ☒ B ☒ C ☐ D ☐

If you change your mind and have crossed out what you consider to be the correct answer, then indicate the correct answer by writing the word **correct** and drawing an arrow as follows.

A ☒ B ☒ C ☐ D ☐  
correct

1. Which substance would not be classified as an acid according to Lavoisier?

- (A)  $\text{CH}_3\text{COOH}$
- (B)  $\text{HCl}$
- (C)  $\text{H}_2\text{SO}_4$
- (D)  $\text{H}_2\text{CO}_3$

**Answer:** B

**Outcome:** H1

2. Which equation shows the ionisation of sulfuric acid in water?

- (A)  $\text{H}_2\text{SO}_4 (\text{l}) + \text{H}_2\text{O} (\text{l}) \rightarrow 2\text{H}_2\text{O} (\text{l}) + \text{SO}_3 (\text{aq})$
- (B)  $\text{H}_2\text{SO}_4 (\text{l}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_3\text{O}^+ (\text{aq}) + \text{OH}^- (\text{aq}) + \text{SO}_3 (\text{aq})$
- (C)  $\text{H}_2\text{SO}_4 (\text{l}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_2\text{SO}_3 (\text{aq}) + \text{H}_2\text{O}_2 (\text{aq})$
- (D)  $\text{H}_2\text{SO}_4 (\text{l}) + \text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_3\text{O}^+ (\text{aq}) + \text{HSO}_4^- (\text{aq})$

**Answer:** D

**Outcome:** H8

3. The table shows data about four different acid solutions ...

<i>Solution</i>	<i>Data</i>
A	pH 3.50
B	$[\text{H}^+] = 2.0 \times 10^{-4} \text{ mol L}^{-1}$
C	$0.00025 \text{ mol L}^{-1} \text{ CH}_3\text{COOH}$
D	$0.00025 \text{ mol L}^{-1} \text{ HCl}$

Which solution will have the lowest pH?

- (A) A
- (B) B
- (C) C
- (D) D

**Answer:** A

**Outcome:** H10

4. Rosa Canina wanted to see if a certain potting mix is suitable for cultivating a special type of grass seed requiring a soil pH of 4.5. Given the following indicator properties:

<i>Indicator</i>	<i>Transition pH</i>
Methyl orange (MO)	3.1 – 4.4
Bromothymol blue (BTB)	6.0 – 7.6
Phenolphthalein (Ph)	8.0 – 9.6

Which indicator or combination of indicators will she need to ascertain the pH most accurately.

- (A) BTB and MO
- (B) BTB and Ph
- (C) BTB only
- (D) MO only

**ANS A**

**Outcome: H9**

5. Which of the following is the best description for a  $0.01 \text{ mol L}^{-1}$  solution of citric acid?

- (A) A dilute strong acid solution
- (B) A dilute weak acid solution
- (C) A concentrated strong acid solution
- (D) A concentrated dilute acid solution

**ANS B**

**Outcome: H13**

6. Which of the following is an acidic oxide?

- (A) CoO
- (B) CaO
- (C) MgO
- (D) NO<sub>2</sub>

**ANS D**

**Outcome: H9**

## Section A

### Multiple Choice Answer Grid

1.	A O	B ●	C O	D O
2.	A O	B O	C O	D ●
3.	A ●	B O	C O	D O
4.	A ●	B O	C O	D O
5.	A O	B ●	C O	D O
6.	A O	B O	C O	D ●

## Section B: Questions requiring short answers

### Question 7 (3 marks)

Concentrated sulfuric acid is shipped around Australia in large steel tanks by road and rail transport. Accidents have occurred where the acid tanks have ruptured and the concentrated acid has spilled out into the surrounding environment.

Assess the use of neutralisation reactions as a safety measure to minimise damage in such accidents. (3 marks)

#### Model Answer

A spill of concentrated sulfuric acid is of serious concern due to its highly corrosive nature. It can cause severe damage to a wide variety of materials and organisms near the spill site. The accident damage is best minimised by neutralising the acid.

A neutralisation reaction locks up the acid's corrosive  $H^+$  ions in the form of safe, stable water...  $H^+ + OH^- \rightarrow H_2O$ . The neutralising agent should be a safe, weak base such as  $NaHCO_3$ . Use of a strong base (e.g.  $NaOH$ ) is not suitable as its caustic nature can cause collateral damage at the accident site. The use of water to dilute the acid is not recommended as the dilution reaction produces considerable damaging heat and the acid remains effectively corrosive when diluted.

✓ Stating that a neutralisation reaction involves a base.	1 mark
✓ Citing a suitable base, e.g. $NaHCO_3$ , $Na_2CO_3$ , $Ca(OH)_2$	1 mark

### Question 8 (1 mark)

Compare the relative strengths of  $0.10 \text{ mol L}^{-1}$  solutions of acetic, citric and hydrochloric acids.

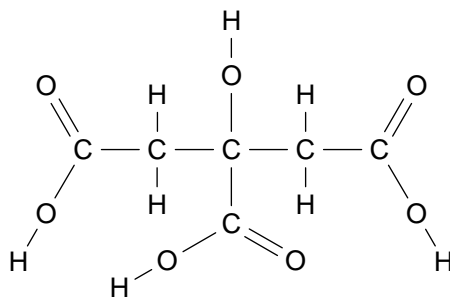
(1 mark)

Outcome: H8

Answer: Acetic < Citric < Hydrochloric

### Question 9

Calculate the mass of citric acid crystals required to prepare 500 mL of 0.10 mol L<sup>-1</sup> solution.  
(1 mark)



**Citric acid**

**Outcome:** H10

**Answer:**  $m = c \times V \times M = 0.10 \text{ mol L}^{-1} \times 0.500 \text{ L} \times 192.124 \text{ g mol}^{-1}$

### Question 10

Identify the need for collaboration between chemists as they collect and analyse data. (2 marks)

**Outcome:** H2, 12, 13, 14, 15

**Model Answer:**

The foundations of modern science are based on the tenets of the scientific method. Hypotheses are formulated and tested by collecting and analysing data. Enlarging the data set increases reliability. If the data is collected and/or reviewed by independent groups of chemists, reliability is further enhanced as methodical human error is reduced.

✓ Citing one valid point.	1 mark
✓ Citing two valid points.	2 marks

## Question 11

Ken Chemiski attempts to prepare propyl butanoate by setting up the apparatus diagrammed.

► For clarity, supporting clamps are not shown on the diagram.

He mixes 25 mL of the alkanoic acid and 25 mL of the alkanol in the flask and adds a few boiling chips. The reaction mixture is heated using a hot plate set on low heat.

- (a) Identify two errors Ken has made in his procedure/equipment. **(2 marks)**

### Answers:

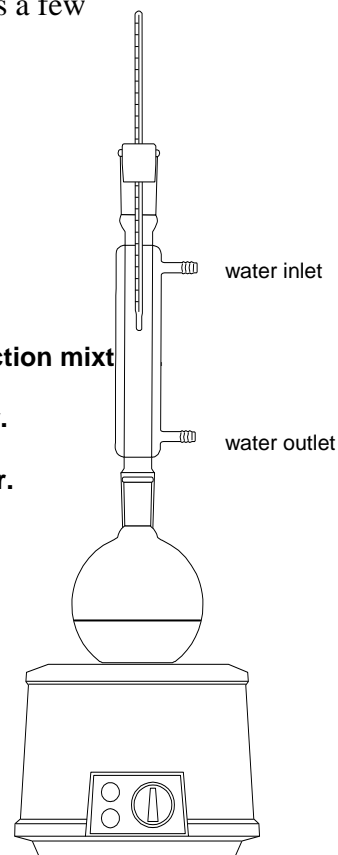
- (a) Ken made three errors in his esterification experiment...

(1) He did not add a small amount of concentrated  $\text{H}_2\text{SO}_4$  catalyst to the reaction mixture.

(2) He should have connected the water inlet to the bottom of the condenser.

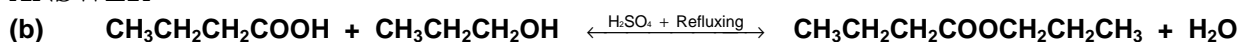
(3) He should not have a thermometer connected to the top of the condenser.

✓ Citing one error.	1 mark
✓ Citing two errors.	2 marks



- (b) Using structural formula write the balanced equation for the esterification reaction. **(1 mark)**

### ANSWER



► No marks deducted for lack of conditions.

**Question 12 (3 marks)**

The mass of a sealed 300 mL soft drink was found to be 400.02 g. The soft drink was then decarbonated by shaking, opening gently to release the gas and recapping, the series of processes being done several times. When no visible sign of gas was being evolved, the capped soft drink weighed 397.08 g.

What volume of carbon dioxide gas was released at 25°C and 100 kPa? Show your working (3 marks)

**Answer:**

**Outcome: H10, H9**

$$\text{mass of CO}_2 = 400.02 - 397.08 = 2.94 \text{ g} \quad 1 \text{ mark}$$

$$\text{moles CO}_2 = \text{mass CO}_2 / \text{MM}_{\text{CO}_2} = \frac{400.02 - 397.08}{12.01 + 2(16.00)} = 0.0668 \quad 1 \text{ mark}$$

$$\begin{aligned} \text{Volume CO}_2 &= \# \text{ moles} \times \text{molar volume} \\ &= 0.0668 \text{ moles} \times 24.79 \text{ L/mole} \\ &= 1.66 \text{ L} \end{aligned} \quad 1 \text{ mark}$$

Criteria	Mark
Mass of CO <sub>2</sub>	1
Calculation of # of moles	1
Volume of CO <sub>2</sub>	1

**Question 13 (3 marks)**

Describe an experiment which will allow you to decide which of two acids, HX and HZ is the stronger. Justify your experimental technique. (3 marks)

**Answer:**

*Using a pH probe, the pH of equal concentrations of acids HX and HZ is determined. The pH measures the free [H<sup>+</sup>] hence, the stronger acid which releases a greater [H<sup>+</sup>] will record a lower pH.*

Criteria	Mark
brief description of technique (must include equal concentrations of acids and pH measurement)	2
justification of technique	1

**Question 14 (2 marks)**

Describe the way in which a medical radioisotope is used and explain its use in terms of its property.

**Possible answer:**

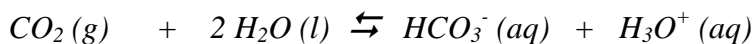
*Iodine-131 is a radioisotope used in medicine. . Because of its structure and reactivity iodine –131 is selectively incorporated in the thyroid gland, when injected into the body. Iodine-131 is a  $\beta$ -emitter and a  $\gamma$ -emitter and hence able to destroy cancerous tissues in the thyroid gland.*

Criteria	Mark
Use	1
Property (nuclear or chemical)	1

**Question 15 (3 marks)**

- (a) Write a net ionic equation to explain the acidity of a solution of carbon dioxide in water. (1 mark)

**Ans:**



The equation must be *ionic*

- (b) Explain what will happen to the pH of carbonated water if the temperature of the solution is increased? (1 mark)

**Ans:**

*When the temperature is increased, the solubility of gaseous carbon dioxide in water decreases. This pushes the equilibrium to the left. More of the hydrogen carbonate ion ( $\text{HCO}_3^-$ ) reacts with the hydronium to produce more  $\text{CO}_2$ . This decreases the  $[\text{H}_3\text{O}^+]$  and hence increases the pH.*

- (c) What happens to the pH of the carbonated water if the pressure of the  $\text{CO}_2$  gas above the solution is reduced? Explain your answer. (1 mark)

**Ans:**

*If the pressure above the solution is decreased, then the solubility of the gas ( $\text{CO}_2$ ) in the solution is decreased. This pushes the equilibrium to the left to maintain equilibrium with the  $\text{CO}_2$  above the solution, resulting in an increase in the pH because of the escape of carbon dioxide gas from the solution (carbonated drink).*

**A END of TEST**