## Ex2 The Ideal Gas Again

Then we have N particles in a box at temperature T and Volume V

$$E = \frac{\vec{p}_1^2}{2m} + \frac{p^2}{2m}$$

= E, + ... + EN

The partition function is

$$\frac{Z}{V_{\text{TOT}}} = \frac{1}{N!} \int \frac{d^3\vec{r}_1 d^3\vec{p}_1}{h^3} \frac{d^3\vec{r}_1 d^3\vec{p}_1}{h^3} e^{-\vec{E}/kT}$$

Since the energy is a sum the integrals factorize

$$Z_{TOT} = \frac{1}{N!} \left[ \int \frac{d^3 \vec{r}_1 d^3 \vec{p}_1}{h^3} e^{-\epsilon_1/kT} \right]$$

$$Z_{1} = \int d^{3}r d^{3}p e^{-\epsilon/kT}$$

$$F = -kT \ln 2$$

$$F = -NkT \ln \left(\frac{eV}{N\lambda^3}\right) = -NkT \left[\ln \left(\frac{V_N}{\lambda^3}\right) + 1\right]$$

The Free energy as a function of Temperature and volume determines everything using thermodynamics

· From its dependence on volume:

$$P = -\partial F = -\partial \left[ -NkT \left( l_n V + const \right) \right]$$

$$P = \frac{NkT}{V}$$

Thus we have recovered the ideal gas law. The entropy follows by differentiation too

$$S = -\partial F = -\partial \left(-NkT\left[\ln\left(V_N/\chi_{t_n}^3\right) + 1\right]\right)$$

$$S = NK \left[ \ln \left( V_N / \chi_{+h}^3 \right) + 1 \right] + NKT \frac{2}{2T} \left( \ln T^{3/2} + const \right)$$

$$S = Nk \left( ln \left( V_N / \frac{3}{\lambda_{+n}} \right) + \frac{5}{2} \right)$$

$$\frac{S}{Nk} = \ln \left( \frac{V_N}{\lambda_{+k}^3} \right) + \frac{5}{2}$$

## · Finally note

Thus

$$U = \frac{3}{2}NkT$$

· An Alternate method starts with Inz

In Z = N In(eZ,/N)

Then only Z, depends on B=1/kT

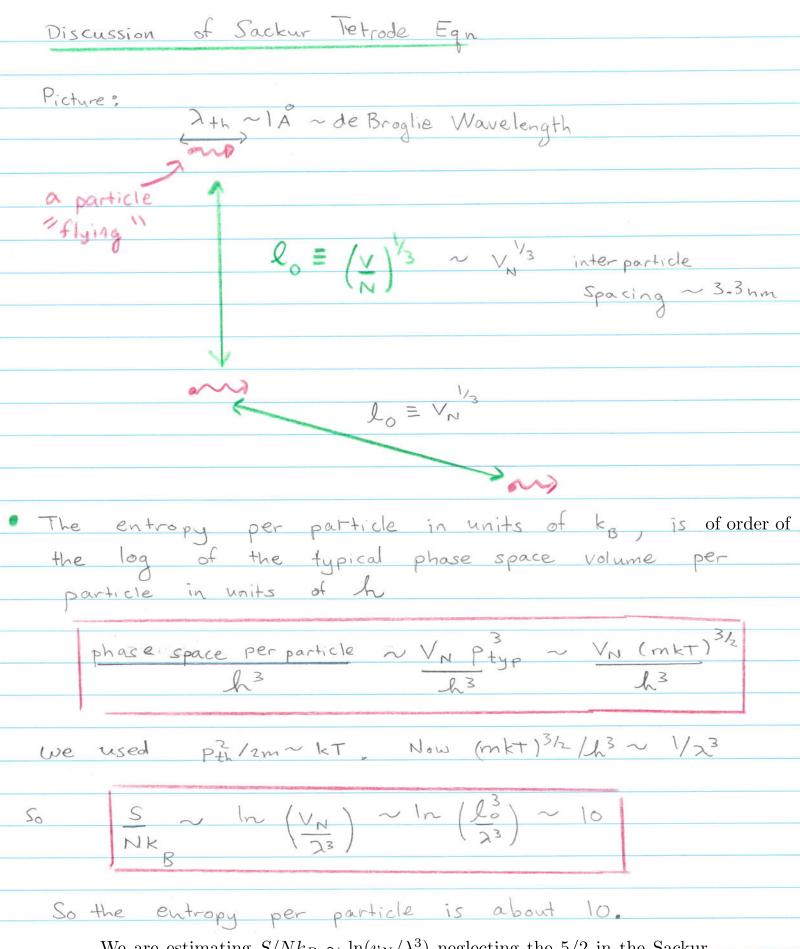
U =- N 2 (In Z, + const)

Now  $2 = \frac{V}{\lambda^3} \propto \beta^{-3/2}$  since  $\lambda \propto T^{-1/2}$ 

So

 $U = -N \partial \left( \ln \beta^{-3/2} + const \right)$ 

 $U = \frac{3N}{2\beta} = \frac{3NkT}{2}$  Conce again



We are estimating  $S/Nk_B \sim \ln(v_N/\lambda^3)$  neglecting the 5/2 in the Sackur-Tetrode formula. The 5/2 is a quantum correction to the classical part. The classical part is the log of the phase space per particle.