

Problem 1. Charge Neutrality

Consider a lattice of protons consisting of a total of N sites. The protons (located at the sites) share the available electrons which can hop from site to site. The total system is neutral so that the number of electrons is equal to the number of protons (or lattice sites). Electrons can hop from site to site, so each site does not need to be neutral. Suppose that each atom can exist in four states.

state	$N_{\text{electrons}}$	Energy
ground state	1	$-\frac{1}{2}\Delta$
positive ion	0	$-\frac{1}{2}\delta$
negative ion	2	$\frac{1}{2}\delta$
excited hydrogen	1	$\frac{1}{2}\Delta$

- (a) Compute the grand potential $\Phi_G = -kT \log \mathcal{Z}_G$ of a site by evaluating the grand partition function, \mathcal{Z}_G . After some algebra you should find

$$\mathcal{Z}_G = 2e^{\beta\mu} \cosh(\beta\Delta/2) + 2e^{\beta\mu} \cosh(\beta(\delta/2 - \mu)) \quad (1)$$

- (b) Determine the mean number of electrons per site as a function of the electron chemical potential and temperature.
- (c) Show that the electron chemical potential is $\delta/2$ and that probability of having neutral hydrogen is

$$\frac{e^{\beta(\Delta+\delta)/2}}{4e^{\beta\delta/2} \cosh^2(\beta\Delta/4)} \quad (2)$$

Hint: The identity $(\cosh(x) + 1)/2 = \cosh^2(x/2)$, is the hyperbolic analog of the cosine identity $(\cos(\theta) + 1)/2 = \cos^2 \theta$.

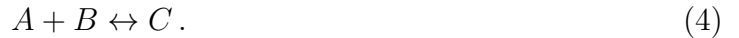
- (d) Determine the the entropy per site. You should find

$$\frac{S}{k_B} = \log [4 \cosh^2(\frac{\beta\Delta}{4})] - \frac{\beta\Delta}{2} \tanh(\frac{\beta\Delta}{4}) \quad (3)$$

Make a sketch of this function as function of $\beta\Delta$. What is the limit of this function as $\beta\Delta \rightarrow 0$? Give a physical interpretation of this limit.

Problem 2. Yields of three species

Consider three ideal gasses in equilibrium. They participate in the following chemical reaction



It is energetically favorable to form atom C , so that the energy of one molecule of C is

$$\epsilon_C = \frac{p^2}{2m_C} - \Delta, \quad (5)$$

where $\Delta > 0$ is the binding energy of C . The molecule C has only one internal state. The other two atoms have energies $\epsilon_A = p^2/2m_A$ and $\epsilon_B = p^2/2m_B$ and form simple ideal gasses

- (a) If the partition function Z_{tot} of a gas of N indistinguishable particles is given by $Z_{\text{tot}} = Z_1^N/N!$, where Z_1 is the single-particle partition function, show that the chemical potential is given by

$$\mu = -kT \log \left(\frac{Z_1}{N} \right) \quad (6)$$

- (b) Assume that at one moment there are N_A , N_B , and N_C particles of type A , B , and C , respectively. Determine the partition function of each species, and find the corresponding chemical potentials.
- (c) Show that

$$n_A = \frac{e^{\mu_A/kT}}{\lambda_A^3} \quad (7)$$

$$n_B = \frac{e^{\mu_B/kT}}{\lambda_B^3} \quad (8)$$

$$n_C = \frac{e^{\mu_C/kT} e^{\beta\Delta}}{\lambda_C^3} \quad (9)$$

Here $n_A = N_A/V$ is the density of species A , and λ_A is the thermal wavelength of A , with an analogous notation for B and C .

- (d) Show that in equilibrium the densities of A , B and C satisfy

$$\frac{n_A n_B}{n_C} = \frac{(2\pi m_{\text{red}} kT)^{3/2}}{h^3} e^{-\beta\Delta} \quad (10)$$

where $m_{\text{red}} = m_A m_B / (m_A + m_B)$ is the reduced mass. Note $m_C = m_A + m_B$.

Problem 3. The Saha Equation

The Saha equation describes the relative abundance of neutral hydrogen to ionized hydrogen at a given temperature. The reaction here is



Read the setup of problem Blundell 22.5, and recognize that the results of the previous problem apply. The only difference is that in Blundell's part (a), they have approximated $m_{\text{red}} = m_e m_p / (m_e + m_p) \simeq m_e$. They (following Saha) also have approximated the internal partition function of the hydrogen atom as a single bound state with binding energy $\Delta = R = 13.6 \text{ eV}$. We evaluate the partition function of hydrogen with the same approximation in a previous problem. Answer the following

- (a) Explain why charge neutrality implies that $n_e = n_p$ and conservation of nucleons implies $n_H + n_p = n$, where n is the total number density of hydrogen (neutral and ionized). Writing $y = n_p/n$ as the degree of ionization, show that

$$\frac{y^2}{1-y} = \frac{e^{-\beta R}}{n \lambda_{\text{th}}^3} \quad (12)$$

where λ_{th} is the thermal wavelength for the electrons. Solve for y in terms of $x(T) = e^{-\beta R}/n\lambda_{\text{th}}^3$ and graph the degree of ionization as a function of temperature in kelvin for a density of 10^{20} m^{-3} . You should find that the temperature where the system becomes fully ionized is approximately 10000 °K.

- (b) Equation 12 shows that the degree of ionization goes up when the density n goes down. Why is that? Answer the following closely related question: The temperature where the hydrogen becomes fully ionized is approximately 10000 °K. But, the Boltzmann factor at this temperature is very small

$$e^{-R/k_B T} = e^{-13.6 \text{ eV}/k_B(10^4 \text{ °K})} \simeq e^{-16} \sim 10^{-7}, \quad (13)$$

Qualitatively explain why the ionized fraction at this temperature is of order unity in spite of this penalizing factor.

Problem 4. Absorption and Oscillations

Consider an ideal mono-atomic gas at temperature T and pressure P in contact with a surface. The atoms of the gas can be absorbed on specific sites on the surface, which are sparsely enough distributed over the surface that they do not interact. There are N_0 such sites and each one can absorb zero, one, or two atoms. If a site is empty, we can take that energy as zero. If the site is singly occupied the energy is ϵ_1 . If it is doubly occupied, the two absorbed atoms interact in a vibrational mode of frequency ω_0 , so that the corresponding energy levels are $\epsilon_2 + n'\hbar\omega_0$. Here ϵ_2 is the energy for absorbing a pair and $n' = 0, 1, \dots$ is the vibrational quantum number, parametrizing the additional energy associated with the vibrations.

- (a) The chemical potential is determined by the temperature and pressure of the surrounding gas. Recall that the single particle partition function of the gas (mono-atomic or even poly-atomic) is

$$Z_1 = \sum_s \int \frac{d^3r d^3p}{h^2} e^{-\beta p^2/2m} e^{-\beta \epsilon_s} \quad (14)$$

Let's strip off the overall dependence on volume and define:

$$Z_1 \equiv V\zeta_1(T) \quad (15)$$

Show that the so-called *fugacity* $z \equiv e^{\beta\mu}$ of the gas is proportional to the pressure, with proportionality constant determined by $\zeta_1(T)$:

$$z \equiv \frac{P}{T\zeta_1(T)} \quad (16)$$

Show that for a mono-atomic gas

$$z = \frac{P}{Tn_Q(T)} \quad (17)$$

- (b) Compute the grand partition function for an absorption site.
- (c) Compute the grand potential Φ_G for an absorption site.
- (d) Compute the mean number of atoms absorbed by a site directly from (b)
- (e) Compute the mean number of atoms absorbed by a site by differentiating (c).
- (f) What is the probability that a site will absorb two atoms? And what is the probability that the site will absorb two atoms, and that the pair will vibrate with three vibrational quanta?

Answers:

- (b) See (c).

- (c)

$$\Phi_G = -kT \ln[1 + z e^{-\beta\epsilon_1} + z^2 e^{-\beta\epsilon_2} Z_{HO}(\beta)] \quad (18)$$

where z is the fugacity of the surrounding gas, and $Z_{HO}(\beta) = 1/(1 - e^{-\beta\hbar\omega_0})$.

- (d)

$$\bar{n} = \frac{z e^{-\beta\epsilon_1} + 2z^2 e^{-\beta\epsilon_2} Z_{HO}}{1 + z e^{-\beta\epsilon_1} + z^2 e^{-\beta\epsilon_2} Z_{HO}} \quad (19)$$

- (e) Same as (d).

- (f)

$$\mathcal{P}_2 = \frac{z^2 e^{-\beta\epsilon_2} Z_{HO}}{1 + z e^{-\beta\epsilon_1} + z^2 e^{-\beta\epsilon_2} Z_{HO}}, \quad \mathcal{P}_{2,3} = \mathcal{P}_2 \cdot \left(\frac{e^{-3\beta\hbar\omega_0}}{Z_{HO}} \right). \quad (20)$$