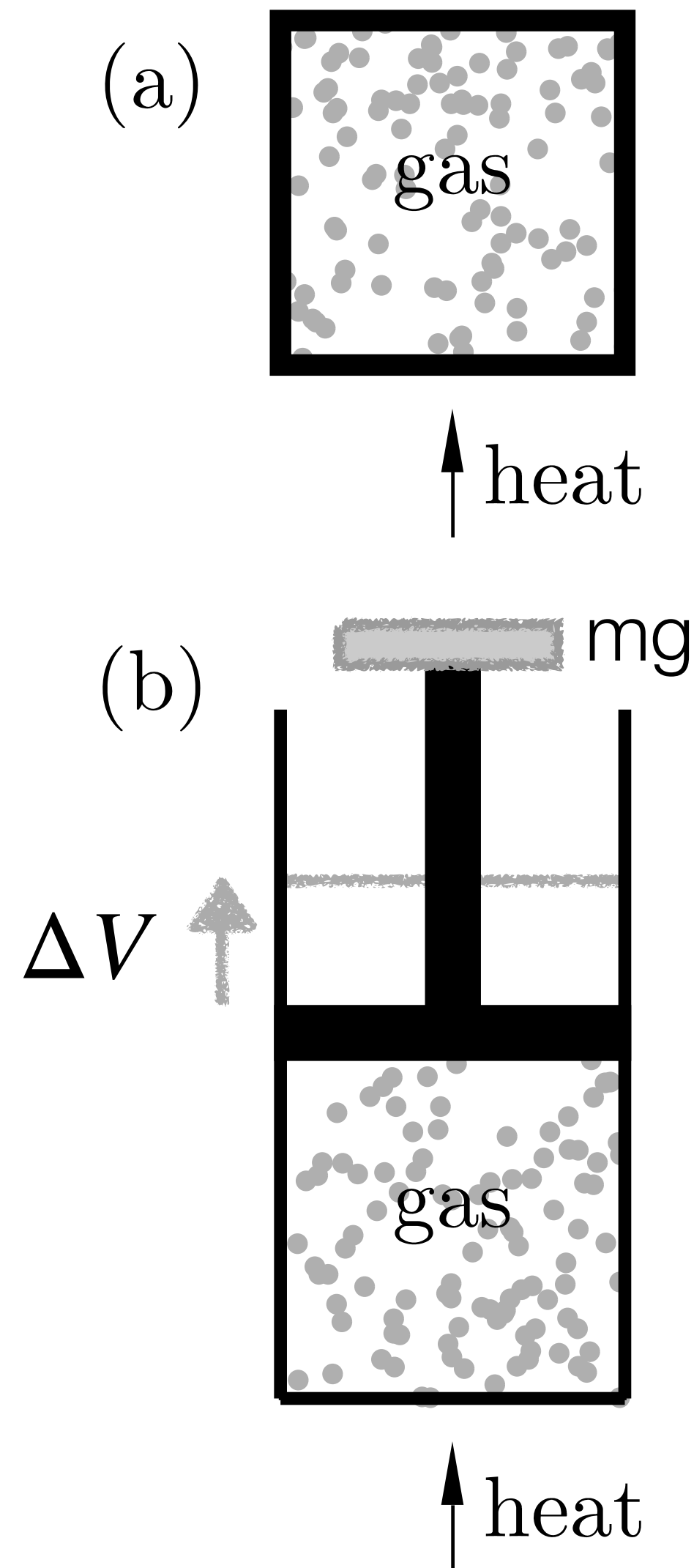


# Specific heats at constant volume and pressure

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Constant volume: add heat,  
and temperature goes up

$$C_v = \left( \frac{dQ}{dT} \right)_v$$

Constant pressure: add heat,  
the gas expands doing work, and  
temperature goes up, but not as much

$$C_p = \left( \frac{dQ}{dT} \right)_p$$

# Specific Heat of $H_2$

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