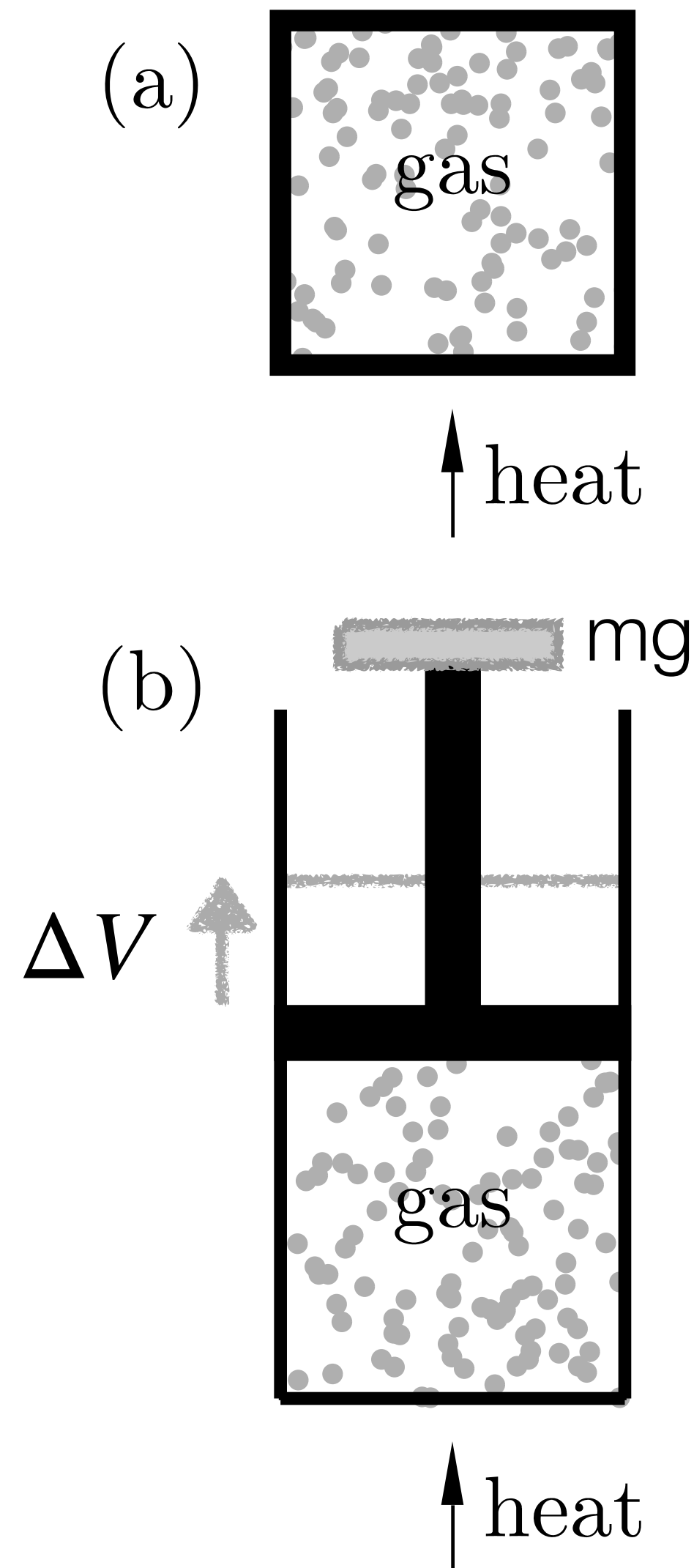


Specific heats at constant volume and pressure



Constant volume: add heat,
and temperature goes up

$$C_v = \left(\frac{dQ}{dT} \right)_v$$

Constant pressure: add heat,
the gas expands doing work, and
temperature goes up, but not as much

$$C_p = \left(\frac{dQ}{dT} \right)_p$$

Specific Heat of H_2

