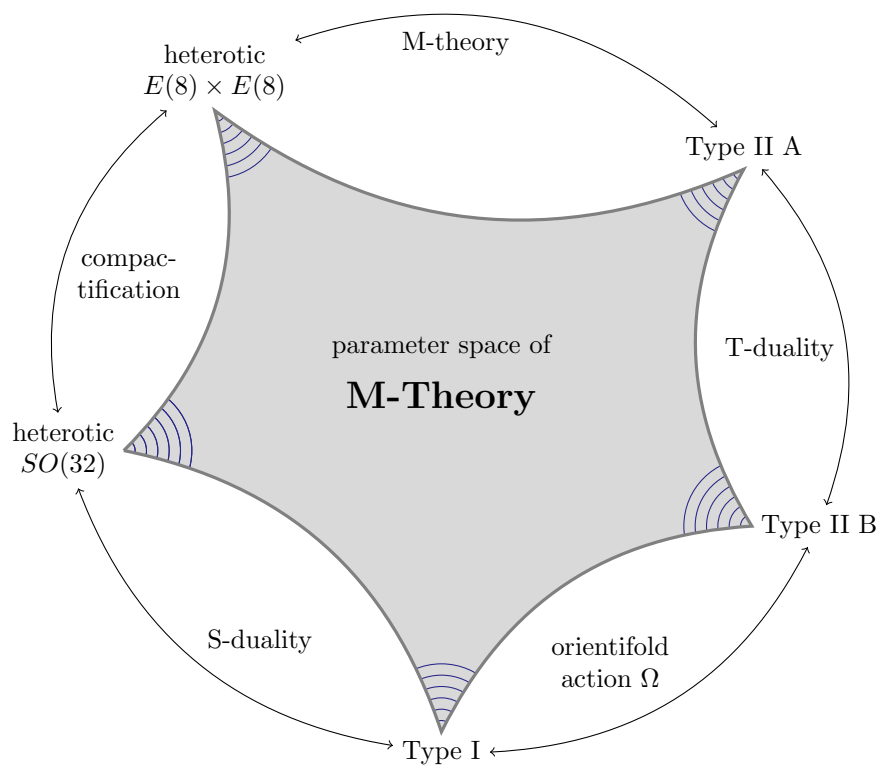


# Introduction to Quantum Mechanics

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# 1 The Wave Function

## 1.1 The Schrödinger Equation

Imagine a particle of mass  $m$ , constrained to move along the  $x$  axis, subject to move to some specified force  $F(x, t)$ . The program of *classical mechanics* is to determine the position of the particle at any given time  $x(t)$ . Once we know that, we can figure out the velocity ( $v = \frac{dx}{dt}$ ), the momentum  $p = mv$ , the kinetic energy ( $T = \frac{1}{2}mv^2$ ), or any other dynamical variable of interest. To determine  $x(t)$ , we apply Newton's Second Law:  $F = ma$ , or more specifically,  $F = -\frac{\partial V}{\partial x}$ , the derivative of a potential energy function, where  $m\frac{\partial^2 x}{\partial t^2} = -\frac{\partial V}{\partial x}$ . This together, with initial conditions, determines  $x(t)$ .

Quantum mechanics approaches this same problem a bit differently. In this case what we're looking for is the particles **wave function**,  $\Phi(x, t)$ , and we get it by solving the **Schrödinger Equation**:

Schrödinger Equation

$$i\hbar \frac{\partial \Psi}{\partial t} = -\frac{\hbar^2}{2m} \frac{\partial^2 \Phi}{\partial x^2} + V\Phi$$

Here  $i$  is the square root of  $-1$ , and  $\hbar$  is Planck's constant - or rather, his *original* constant ( $h$ ) divided by  $2\pi$ :

$$\hbar = \frac{h}{2\pi} = 1.054573 \times 10^{-34} Js.$$

The Schrödinger equation plays a role logically analogous to Newton's second law.

## 1.2 The Statistical Interpretation

But what exactly *is* this wave function, and what does it do for you once you've *got* it? After all a particle by nature is a point, whereas the wave function (as its name suggests) is spread out in space (a function of  $x$ , for any given  $t$ ).

How can such an object represent the state of a *particle*? The answer is provided by Born's **statistical interpretation**, which says that  $|\Psi(x, t)|^2$  gives the *probability* of finding the particle at point  $x$ , at time  $t$  - or, more precisely,

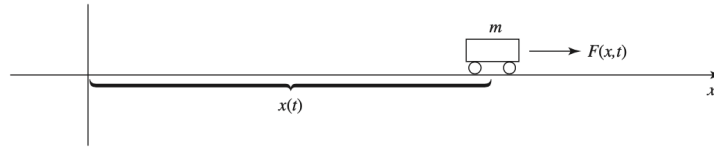


Figure 1: a "particle" constrained to move in one dimension under the influence of a specified force

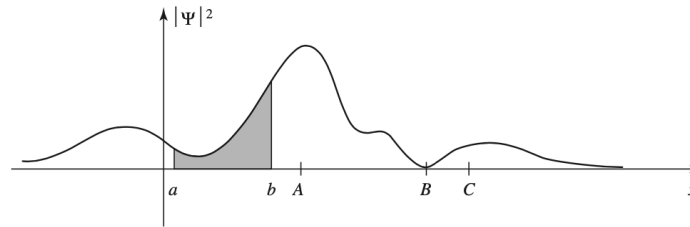


Figure 2: a typical wave function. the shaded area represents the probability of finding the particle between  $a$  and  $b$ . the particle would be relatively likely to be found near  $A$ , and unlikely to be found near  $B$ .

#### Born's Statistical Interpretation

$$\int_a^b |\Psi(x, t)|^2 dx = \{\text{probability of finding particle between } a \text{ and } b, \text{ at time } t\}$$

Probability is the *area* under the graph of  $|\Psi|^2$ . For the wave function in the figure above, you would be quite likely to find the particle in the vicinity of point  $A$ , where  $|\Psi|^2$  is large, and relatively unlikely to find it near point  $B$ .

The statistical interpretation introduces a kind of **indeterminacy** into quantum mechanics, for even if you know everything, the theory has to tell you about the particle, still you can't predict with certainty the outcome of a simple experiment to predict its position - all quantum mechanics has to offer is *statistical* information about *possible* results. It is natural to wonder whether this indeterminacy is a fact of nature, or a defect in the theory.

Suppose I *do* measure the position of the particle, and I find it to be a point  $C$ .

#### Question

Where was the particle just *before* I made the measurement?

### ***Solution***

There are three plausible answers

1. The **realist** position: The *particle was at  $C$* . This certainly seems reasonable, and it is the response Einstein advocated. However, if this is true, quantum mechanics is an *incomplete* theory, since the particle *really was* at  $C$ , and yet quantum mechanics was unable to tell us so.

2. The **orthodox** position: The *particle wasn't really anywhere*. It was the act of measurement that forced it to "take a stand" (though how and why it chose the point  $C$  we dare not ask). This view is associated with Bohr and his followers. Among physicists it is the most widely accepted position. However, if it is correct, a century worth of debate about the act of measurement has done preciously little to illuminate.

The **agnostic** position: *Refuse to answer*. This is not as silly as it sounds - what sense can there be in making assertions about the status of a particle *before* a measurement. For decades this was the "fall-back" position of most physicists: they'd try to sell you the orthodox answer, but if you were persistent they'd retreat to the agnostic response, and terminate the conversation.

However, in 1964, John Bell astonished the physics community by showing that it makes an *observable* difference whether the particle had a precise (though unknown) position prior to the measurement or not. Bell's discovery effectively eliminated agnosticism as a viable option and made it an *experimental* question whether 1 or 2 was the correct choice.

What if I made a *second* measurement, immediately after the first? Would I get  $C$  again, or does the act of measurement cough up a completely new number each time? On this question everyone is in agreement:

A repeated measurement (on the same particle) must return the same value

Indeed, it would be tough to prove the particle was really at  $C$  in the first instance. How does the orthodox interpretation account for the fact that the second measurement is bound to yield the same result  $C$ ? It must be that the first measurement radically alters the wave function, so that it now sharply peaked about  $C$ . We say that a wave function **collapses**, upon measurement, to a spike at the point  $C$  (it soon spreads out again, in accordance with the Schrödinger equation,

so the second measurement must be made quickly. There are then, two entirely distinct kinds of

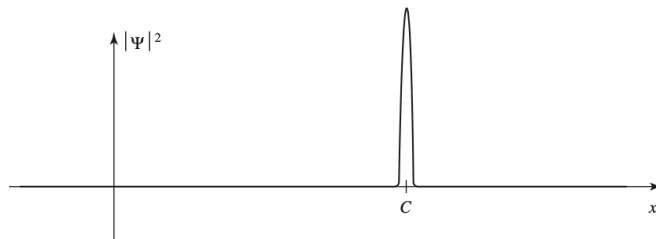


Figure 3: collapse of the wave function: graph of  $|\Psi|^2$  immediately *after* a measurement has found the particle at point  $C$ .

physical processes: "ordinary" ones, in which the wave function evolves in a leisurely fashion under the Schrödinger equation, and "measurements," in which  $\Psi$  suddenly and discontinuously collapses.

### 1.2.1 Example 1.1

**Electron Interference.** How might we check that particles (electrons, for example) have a wave nature, encoded in  $\Psi$ ?

The classic signature of a wave phenomena is *interference*: two waves *in phase* interfere constructively, and out of phase they interfere destructively. The wave nature of light wave confirmed in 1801 by Young's famous double slit experiment, showing interference "fringes" on a distant screen when a monochromatic beam passes through two slits. If essentially the same experiment is done with *electrons*, the same pattern develops, confirming the wave nature of electrons as more and more electrons collide with the screen.

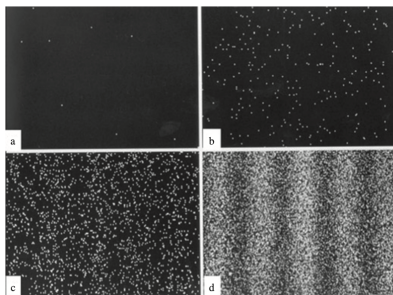


Figure 4: Build up of electron interference pattern. (a) Eight electrons, (b) 270 electrons, (c) 2000 electrons, (d) 160,000 electrons.

Of course, if you close off one slit, or somehow contrive to detect which slit each electron passes through, the interference pattern disappears; the wave function of the emerging particle is now entirely different.

## 2 Probability

### 2.1 Discrete Variables

Because of statistical interpretation, probability plays a central role in quantum mechanics, so I digress now for a brief discussion of probability theory. It is mainly a question of introducing some notation and terminology, and I shall do it in the context of a simple example:

Imagine a room containing 14 people, whose ages are:

one person aged 14  
one person aged 15  
three people aged 16  
two people aged 22  
two people aged 24  
five people aged 25

If we let  $N(j)$  represent the number of people of age  $j$ , then

$N(14) = 1$   
 $N(15) = 1$   
 $N(16) = 3$   
 $N(22) = 2$   
 $N(24) = 2$   
 $N(25) = 5$

while  $N(17)$ , for instance, is 0. The total number of people in the room is

$$N = \sum_{j=0}^{\infty} N(j).$$

**Question** If you selected one individual at random from this group, what is the **probability** that this person's age would be 15? **Solution** One chance in 14, since there are 14 possible choices,

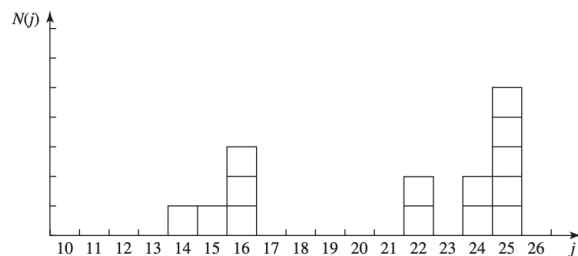


Figure 5: histogram of people

all equally likely, of whom only one is aged 14. If  $P(j)$  is the probability of getting age  $j$ , then  $P(14) = 1/14, P(15) = 1/14, P(16) = 3/14$ , and so on. In general,

$$P(j) = \frac{N(j)}{N}.$$

Notice that the probability of getting *either* 14 or 15 is the *sum* of the individual probabilities ( $1/7$ ). In particular, the sum of *all* probabilities is 1 - the person you select must have *some* age:

$$\sum_{j=0}^{\infty} P(j) = 1$$

#### Question

What is the **most probable** age?

#### Solution

25, obviously, the most probably  $j$  is the  $j$  for which  $P(j)$  is a maximum.

#### Question

What is the **median** age?

#### Solution

23, for 7 people are younger than 23, and 7 people are older

#### Question

What is the **median** age?



Solution

$$\frac{14 + 15 + 3(16) + 2(22) + 2(24) + 5(25)}{14} = \frac{294}{14} = 21.$$