

CHEMISTRY

Metals and Non-Metals

- Q1** Which of the following is used for dissolution of gold?
(A) Hydrochloric acid
(B) Sulphuric acid
(C) Nitric acid
(D) Aqua regia
- Q2** Which among the following is the correct order of rate of reaction of metals with same quantity and concentration of dilute hydrochloric acid?
(A) $Fe > Al > Zn > Pb$
(B) $Mg < Al < Zn < Fe$
(C) $Mg > Al > Zn > Fe$
(D) $Al > Mg > Zn > Fe$
- Q3** When a copper wire is introduced in the colourless solution of silver nitrate, the solution starts turning blue due to the formation of
(A) copper chloride
(B) copper nitrate
(C) copper sulphate
(D) copper carbonate
- Q4** Which of the following can undergo a chemical reaction?
(A) $MgSO_4 + Fe$ (B) $ZnSO_4 + Fe$
(C) $MgSO_4 + Pb$ (D) $CuSO_4 + Fe$
- Q5** The composition of aqua-regia is:
(A) Conc.HCl : Dil.HNO₃, 1: 3
(B) Conc.HCl : Dil. HNO₃, 3: 1
(C) Conc.HCl : Conc.HNO₃, 3: 1
(D) Dil. HCl : Dil. HNO₃, 3:1



Answer Key

Q1 (D)

Q2 (C)

Q3 (B)

Q4 (D)

Q5 (C)



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Hints & Solutions

Q1 Text Solution:

The correct one is a mixture of concentrated nitric and hydrochloric acid in the ratio 1:3.

Video Solution:**Q2 Text Solution:**

Try to land to the correct answer through the mnemonic of the reactivity series.

Video Solution:**Q3 Text Solution:**

Try to think about the product formed and their respective colours.

Video Solution:**Q4 Text Solution:**

More reactive will displace the less reactive metal. Try to land to the correct answer through the knowledge of reactivity series of metals.

Video Solution:**Q5 Text Solution:**

Both the acids will be concentrated. Now, try to land to the correct answer.

Video Solution:[Android App](#)[| iOS App](#)[| PW Website](#)