

**Geoneutrino oscillations
approach to discriminate
distributions of HPE in the
Earth's mantle using the
Monte Carlo Technique.**
Monography for Geoscientist degree ¹

DANIEL FELIPE FORERO SÁNCHEZ²

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¹Universidad de Los Andes - Bogotá, Colombia

²github.com/dforero0896

Dedicated to Ana María, without whom this
thesis would have been finished much, much
earlier.

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Preface

Acknowledgements

1

Introduction

“This is a quote and I don’t know who said this.”

– Author’s name, *Source of this quote*

2

Relevance of the Study

3

Distribution of Radioactive Elements,

3.1 Distribution of Elements

A first, basic approach to understanding the distribution of elements in the Earth was given by Goldschmidt [4], who developed a classification system for the chemical elements. This classification is, in a nutshell, *lithophile* or oxygen-loving elements tend to remain in the uppermost part of the solid planet; *siderophile* or iron-loving elements, tend to be close to the nucleus and are not attracted to oxygen, unlike the lithophiles. The *chalcophile* elements are sulfur-loving and, thus, do not like the deep Earth. Finally, the *atmophile* elements are those who like to be in the atmosphere.

There are many radioactive isotopes in the Earth, which are, basically, atoms so heavy that they are unstable, and go through different processes to reach a more stable state. This will be the topic of chapter 4. Most of these heavy elements are classified as lithophiles. We are particularly interested in some isotopes that dominate the geoneutrino production (99% of production) [7], these are ^{40}K , ^{235}U , ^{238}U and ^{232}Th .

Within each of the groups given by Goldschmidt and briefly described here, there is other classification that discriminates elements based on their condensation temperature. The elements with a high value of this parameter are called *refractory* while the ones with a lower one are called *volatile*. Let us then picture the scenario of the formation of the Earth. About 4500 *Gyr* ago, the solar system and the Earth were being formed, the latter through planetary accretion, this means, a hot, undifferentiated mass of molten rock that is slowly cooling down, the consequences of this are, first of all, density differentiation, that causes the heavier elements to accumulate near the center of mass, thus, forming the nucleus. It also causes the elements to con-

dense, so the refractory ones, are quickly trapped into the Earth, making it possible to infer the initial abundances of these elements in the primitive Earth by looking at their abundances in different meteorites. On the other hand, volatile elements will not be trapped so easily or quickly, instead, there is a possibility that these elements have escaped the primitive Earth while it was hotter; the gases are extremely volatile, while there are other elements, like potassium (whose relevance will become obvious in the following sections) that are moderately volatile.

I have mentioned that the cooling down of the primitive Earth, lead to density differentiation in layers, causing the layered structure of the planet. The less-dense melt that ascended all the way up made up the crust, and is enriched in lithophile elements such as U , Th , K . Nevertheless, there is still a large amount of these elements in the mantle.

3.2 Relative Abundances

This section is devoted to discuss the different views on the distribution of the radioactive isotopes, in depth, quantitatively. There are two main discussions to be considered; first, the average composition of the Bulk Silicate Earth (BSE) and, second, the distribution of Heat Producing Elements (HPE) in the mantle.

3.2.1 Bulk Silicate Earth

In the first place, we shall define the BSE as the “average” Earth composition when excluding the nucleus, in other words, an average between mantle and crust. Three main ideas on the BSE composition have arisen through the years and each one of them relies on different, valid, arguments.

The first model is the *geochemical*, that as the name suggests, is based in geochemical arguments to give values to the abundances and abundance ratios of elements in the planet. McDonough [8] analyzed peridotite samples in order to infer the BSE average compositions. Reference [8] reports values of $A_{Th}^{BSE} = 79.5 \text{ ppb}$, $A_K^{BSE} = 240 \text{ ppm}$ and $A_U^{BSE} = 20.3 \text{ ppb}$, while reference [12] reports $A_{Th}^{BSE} = 80 \pm 13 \text{ ppb}$, $A_K^{BSE} = 280 \pm 60 \text{ ppm}$ and $A_U^{BSE} = 20 \pm 4 \text{ ppb}$ for this BSE model, note that both references are in agreement. A_X^Y stands for the abundance of element X in reservoir Y .

The second model is the *geodynamical*. This model is based on the current measurements of the Earth’s heat flow and assumes that the fraction of this heat due to radioactive decay is higher than the fraction due to secular cooling of the planet. In reference [12], the reported values of the abundances

are $A_{Th}^{BSE} = 140 \pm 14 \text{ ppb}$, $A_K^{BSE} = 350 \pm 35 \text{ ppm}$ and $A_U^{BSE} = 35 \pm 4 \text{ ppb}$. This model has particularly high values of A_X^Y due to the assumption that the majority of the flux is product of radioactivity.

The third and final model is the *cosmochemical*. It relies on the analysis of enstatite chondrites to infer the BSE composition, under the argument that this abundances specially the iron one, will easily explain the presence of the core, as explained in reference [7]. Reference [12] reports the following values relevant to this project: $A_{Th}^{BSE} = 43 \pm 4 \text{ ppb}$, $A_K^{BSE} = 146 \pm 29 \text{ ppm}$ and $A_U^{BSE} = 12 \pm 2 \text{ ppb}$.

In this study, I will compare the results for geoneutrino measurements for the different BSE models presented above.

3.2.2 Crust

The relative abundance of HPE in the crust has been the object of various studies ([5, 9, 11, 13, 14]) among which, for this project, I chose reference [5]. In this article, the authors develop a reference Earth model for the HPE, based on geochemical analysis performed on different rocks and following the CRUST 2.0 model for crustal thickness and properties. The values given are $A_U^C = 1.31_{-0.25}^{+0.29} \text{ ppb}$ and $A_{Th}^C = 5.61_{-0.89}^{+1.56} \text{ ppb}$.

3.2.3 Mantle

The abundance of isotope X in the mantle, A_X^M , is calculated from the values reported above for A_X^{BSE} and A_X^C and the mass balance [12]:

$$m_{BSE} A_X^{BSE} = m_C A_X^C + m_M A_X^M, \quad (3.1)$$

that is, the total mass of isotope X is the sum of the mass in the crust and the mantle. The mass of reservoir Y , m_Y , is calculated through integration of the mass density of the Earth given by the PREM model in reference [1], the details of the method will be given in chapter 7.

Additionally, the mantle structure is currently debated. Two of the most accepted models are the uniform mantle and the two-layer (EL-DL) mantle. The former consists in uniformly distributed HPE over the mantle, thus, the abundance in every point will be given by equation 3.1. The latter, in which EL and DL stand for enriched and depleted layer, respectively, consists in two layers of slightly different HPE abundances. The motivation for this is that the difference between the types of basalts (OIB, MORB, etc.) is thought to be caused by the differentiation of the mantle into two chemically distinct layers, an enriched layer and a depleted one, respect to lithophile elements. The enriched layer is said to constitute the lowermost

10% in mass of the bulk mantle [12]. The abundance in these two layers is also calculated using a relation similar to equation 3.1.

4

Radioactivity,

4.1 Overview

Radioactivity is the phenomenon in which a parent isotope turns into a daughter isotope, with different characteristics, through the emission of a particle.

The history of radioactivity goes back to 1896, when H. Becquerel discovered that a uranium sample emitted some kind of penetrating radiation similar to the X rays (discovered earlier that year). In the following years, three different types of emissions were identified: alpha (α), beta (β) and gamma (γ). The nature of each of these emissions was identified through the years, concluding that α -particles correspond to ${}^4\text{He}$ nuclei, β^\pm -particles correspond to e^\pm (electrons or positrons) and γ -particles are nothing but high energy photons.

Rutherford discovered that radioactive phenomena was linked directly with the nucleus (size $\sim 1\text{ fm}$) of a given isotope, thus, it is an entirely quantum phenomenon.

The overall radioactive phenomena are described by a rather simple mathematical approach given the statistical behavior of the phenomena (which is a result, and evidence of how “quantum” it is). Given N parent isotopes in a sample at a given time t and assuming no more are added, the rate of decay (dN/dt) is proportional to N [6]

$$\frac{dN}{dt} = -\lambda N \quad (4.1)$$

That, upon integration becomes

$$N(t) = N_0 \exp(-\lambda t) \quad (4.2)$$

Where N_0 corresponds to the number of parent isotopes at time $t = 0$ and λ is called the decay constant and is unique for every isotope.

This simple model has been key to the use of the radioactive isotopes to, for

example, date rocks or organisms [6].

4.2 Types of radioactive decay

Let us now go deeper into the theory of the radioactivity, that is, briefly describe the different types of radioactive decays since not all of them produce antineutrinos, and these are the particles that concern this text.

4.2.1 Alpha decay

The theory of the alpha decay is rather simple. The α -particles are confined in a finite potential well (the nucleus X) and will, as expected, have a certain probability of tunneling said barrier. When this happens, the α -particle will escape and leave a “new” (daughter) nucleus X' [6]. The process may be written as

$${}^A_Z X_N \rightarrow {}^{A-4}_{Z-2} X'_{N-2} + \alpha \quad (4.3)$$

Where A is the atomic mass number, Z is the atomic number and N is the charge.

This decay involves both strong (nuclear) interactions and electromagnetic interactions since the potential well is given by confinement due to strong interaction between nucleons and the barrier has a decaying-exponential side, given by the Coulomb potential.

4.2.2 Gamma decay

Gamma decay is similar to the common electromagnetic emission due to atomic transitions, in fact, it is produced when a metastable state of an isotope decays into a more stable state through the emission of high energy photons. It should be noted that, in this case, no change in A or N is produced. These metastable states are common daughter isotopes to α and β decays.

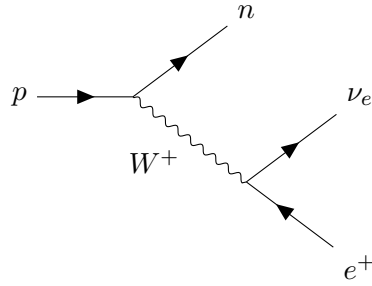
Lifetimes for this kind of process is generally short, taking only fractions of a second to decay while some α -decaying isotopes may have half-lives of the order of 10^3 *yr*.

4.2.3 Beta decay

This type of decay can be considered to be more complex than the other two, as it responds to more complicated underlying physics: the weak interaction. It consists in a set of semileptonic processes (that involve both leptons and hadrons) that will be described below.

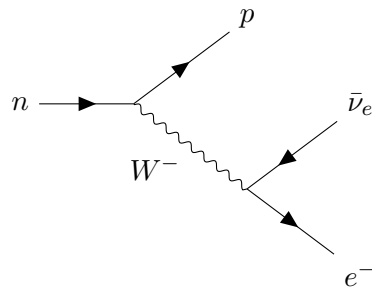
- **Positive Beta decay:**

$$p \rightarrow n + e^+ + \nu_e \quad (4.4)$$

Figure 4.1: Feynman diagram for β^+ process.

- **Negative Beta decay:**

$$n \rightarrow p + e^- + \bar{\nu}_e \quad (4.5)$$

Figure 4.2: Feynman diagram for β^- process.

- **Electron capture:**

$$p + e^- \rightarrow n + \nu_e \quad (4.6)$$

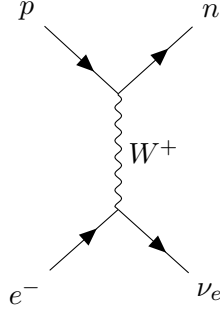


Figure 4.3: Feynman diagram for electron capture.

Note that in all these processes new particles are created in the nucleus, which is a consequence of the weak interaction that rules them.

The particles labeled with ν_e are electron neutrinos and were proposed by Pauli in order to solve the conundrum concerning the apparently continuous energy spectra of β -decay which is against the core of quantum mechanics, the inclusion of these particles in the energy spectrum of the decay made it discrete, thus, giving a solution to the problem [3].

For this project, we are strictly interested in the β^- -decay since it is the one that produces antineutrinos.

4.3 Decay Chains

Different isotopes in the Earth have decay chains that involve β^- -decay at some point, but there are a few ones that dominate the antineutrino production. The reader may have noticed that in chapter 3, I have said that the production of heat is dominated by ^{238}U , ^{232}Th and ^{40}K . Each isotope has a decay chain that includes some beta decays. These are summarized in the following equations.

$$^{238}\text{U} \rightarrow ^{206}\text{Pb} + 8\alpha + 8e^- + 6\bar{\nu}_e + 51.7\text{MeV}, \quad (4.7)$$

$$^{235}\text{U} \rightarrow ^{207}\text{Pb} + 7\alpha + 4e^- + 4\bar{\nu}_e + 46.4\text{MeV}, \quad (4.8)$$

$$^{232}\text{Th} \rightarrow ^{208}\text{Pb} + 6\alpha + 4e^- + 4\bar{\nu}_e + 42.7\text{MeV}, \quad (4.9)$$

$$^{40}\text{K} \rightarrow ^{40}\text{Ca} + e^- + \bar{\nu}_e + 1.31\text{MeV}, \quad (4.10)$$

$$^{40}\text{K} + e^- \rightarrow ^{40}\text{Ar} + \bar{\nu}_e + 1.505\text{MeV}. \quad (4.11)$$

Note that the processes shown in equations 4.10 and 4.11 have the same parent isotope, the former happens with a probability of 89.3% and the latter with the remaining 10.7%.

In equations 4.7 through 4.11, the particles $\bar{\nu}_e$ are electron antineutrinos, also called geoneutrinos.

4.4 (Anti)Neutrino emission

As seen in the previous section, naturally produced antineutrinos from beta decay are called geoneutrinos. These are produced following a given spectrum [2]. These spectra are shown in figure 4.4.

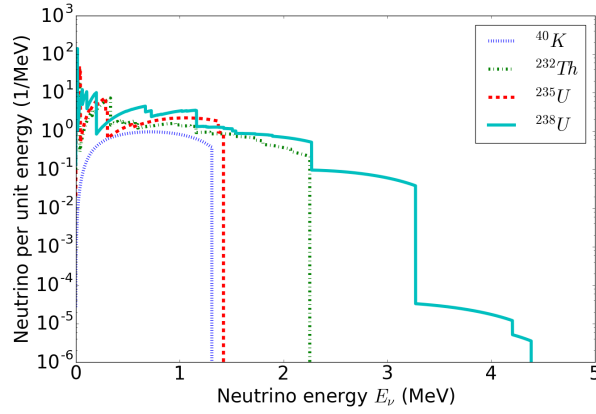


Figure 4.4: Geoneutrino energy spectrum from [2].

5

Neutrino Physics

This chapter is only an overview of the special physics that surrounds these particles, for a thorough explanation you should see reference [3] and for details on time evolution, reference [10].

Neutrinos are neutral leptons that interact weakly and are always produced in a given flavor eigenstate $|\nu_\alpha\rangle$ with $\alpha = e, \mu, \tau$. In equations 4.4 to 4.6, all neutrinos are labeled ν_e that stands for electron neutrino, this is because they are all associated with the “ordinary” matter which is composed by electrons, not tauons (τ) or muons (μ), which are other charged leptons, this is, electrically charged particles that lack color charge.

I just mentioned neutrinos are produced in flavor states, but these are not mass (energy) eigenstates and, since the Hamiltonian entirely controls the time evolution of a quantum system, the neutrinos will not evolve (travel) in these states but will rather oscillate between the real mass eigenstates which, as the name suggests, do have a definite (and unknown) mass, while flavor states are just a linear combination of them. This may be written as

$$|\nu_\alpha\rangle = \sum_{a=1}^3 U_{\alpha a}^* |\nu_a\rangle \quad (5.1)$$

Here, we have labeled $|\nu_a\rangle$ the mass eigenstates, with $a = 1, 2, 3$. $U_{\alpha a}^*$ is the matrix element of the transformation matrix U^* [10].

The matrix U^* is the complex conjugate of matrix U , called Pontecorvo-Maki-Nakagawa-Sakata (PMNS) matrix, that rules the mixing of the neutrinos and can be parametrized as shown in [10]

$$U = \begin{bmatrix} C_2 C_3 & S_3 C_2 & S_2 \\ -S_3 C_1 - S_1 S_2 C_3 & C_1 C_3 - S_1 S_2 S_3 & S_1 C_2 \\ S_1 S_3 - S_2 C_1 C_3 & -S_1 C_3 - S_2 S_3 C_1 & C_1 C_2 \end{bmatrix} \quad (5.2)$$

Where $S_i \equiv \sin(\theta_i)$ and $C_i \equiv \cos(\theta_i)$. Parameters θ_i are the vacuum mixing angles. Here we have assumed there is no charge-parity (CP) phase, thus, $U = U^*$.

5.1 Time evolution

According to the principles of quantum mechanics, the Hamiltonian of a system described by the ket $|\psi\rangle$ does define its time evolution according to

$$|\psi(t)\rangle = \exp(-i\hat{H}(t - t_0))|\psi(t_0)\rangle \quad (5.3)$$

In the case of the neutrinos, we know that they do not travel in flavor state but in mass states, then, we can write

$$|\nu_a(t)\rangle = \exp(-i\hat{H}_m(t - t_0))|\nu_a(t_0)\rangle \quad (5.4)$$

In this base, the Hamiltonian will be diagonal

$$\hat{H}_{m,unperturbed} = \begin{bmatrix} E_1 & 0 & 0 \\ 0 & E_2 & 0 \\ 0 & 0 & E_3 \end{bmatrix} \quad (5.5)$$

Inside matter, the Hamiltonian of the system is perturbed in the following way

$$\hat{V}_f = \mathcal{A} \begin{bmatrix} 1 & 0 & 0 \\ 0 & 0 & 0 \\ 0 & 0 & 0 \end{bmatrix} \quad (5.6)$$

Where \mathcal{A} depends on the density of the matter the neutrino is going through, and is defined as

$$\mathcal{A}(r) \approx \pm\sqrt{2}G_F \frac{\rho(r)}{m_N} \quad (5.7)$$

Here, G_F is the Fermi coupling constant, $\rho(r)$ is the matter density and m_N is the nucleon mass.

The perturbation V_f (f stands for flavor base) only affects electron-flavored neutrinos because matter, as we see it, is electronic. Taking this into account, the perturbed Hamiltonian, in mass basis is

$$\hat{H}_m = \begin{bmatrix} E_1 & 0 & 0 \\ 0 & E_2 & 0 \\ 0 & 0 & E_3 \end{bmatrix} + U^{-1}V_fU \quad (5.8)$$

Introducing equation 5.8 into 5.4 (this involves some complex algebra described in [10]) we can get the time evolution operator \hat{U}_f . Then, the transition probability from an initial flavor α state to a β one will be

$$P_{\alpha \rightarrow \beta} = |\langle \beta | \hat{U}_f | \alpha \rangle|^2 \quad (5.9)$$

This will be crucial on this project, as the neutrino oscillation phenomena will affect the detection, and the path that the neutrino follows defines how much it will be affected, therefore, the effect of the phenomena is not trivial at all.

6

Antineutrino detection

7

The Model,

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References

- [1] Adam M. Dziewonski and Don L. Anderson. “Preliminary reference Earth model”. In: *Physics of the Earth and Planetary Interiors* 25.4 (1981), pp. 297–356. ISSN: 00319201. DOI: 10.1016/0031-9201(81)90046-7.
- [2] S. Enomoto. “Neutrino Geophysics and Observation of Geo-Neutrinos at KamLAND”. PhD thesis. 2005, p. 233.
- [3] Carlo Giunti and Chung W. Kim. “Fundamentals of Neutrino Physics and Astrophysics”. In: *Fundamentals of Neutrino Physics and Astrophysics* (2010), pp. 1–728. ISSN: 0717-6163. DOI: 10.1093/acprof:oso/9780198508717.001.0001. arXiv: 9809069v1 [arXiv:gr-qc].
- [4] Victor Goldschmidt. *Geochemistry*. 1st ed. The Clarendon Press, 1958.
- [5] Yu Huang et al. “A reference Earth model for the heat-producing elements and associated geoneutrino flux”. In: *Geochemistry, Geophysics, Geosystems* 14.6 (2013), pp. 2003–2029. ISSN: 15252027. DOI: 10.1002/ggge.20129. arXiv: 1301.0365.
- [6] Kenneth S. Krane. *Introductory Nuclear Physics*. Wiley, 1988. ISBN: 9780471805533. DOI: 9780471805533.
- [7] L. Ludhova and S. Zavatarelli. “Studying the earth with geoneutrinos”. In: *Advances in High Energy Physics* 2013 (2013). ISSN: 16877357. DOI: 10.1155/2013/425693. arXiv: 1310.3961.
- [8] Shen-Su McDonough, William F; Sun. “The Composition of The Earth”. In: *Chemical Geology* 120 (1995), pp. 223–253. ISSN: 00092541. DOI: 10.1016/0009-2541(94)00140-4.
- [9] Scott M. McLennan. “Relationships between the trace element composition of sedimentary rocks and upper continental crust”. In: *Geochemistry, Geophysics, Geosystems* 2.4 (Apr. 2001), n/a–n/a. ISSN: 15252027. DOI: 10.1029/2000GC000109. URL: <http://doi.wiley.com/10.1029/2000GC000109>.

- [10] Tommy Ohlsson and Hakan Snellman. “Neutrino oscillations with three flavors in matter of varying density”. In: *The European Physical Journal C* 20.3 (2001), pp. 507–515. ISSN: 1434-6044. DOI: 10.1007/s100520100687. arXiv: 0103252 [hep-ph]. URL: <http://arxiv.org/abs/hep-ph/0103252> <http://www.springerlink.com/index/10.1007/s100520100687>.
- [11] R.L. Rudnick and S. Gao. “3.01 Composition of the Continental Crust”. In: *Treatise on Geochemistry*. 2003, pp. 1–64. ISBN: 9780080437514. DOI: 10.1016/B0-08-043751-6/03016-4. URL: <http://www.sciencedirect.com/science/article/pii/B0080437516030164>.
- [12] Ondrej Sramek et al. “Geophysical and geochemical constraints on geoneutrino fluxes from Earth’s mantle”. In: *Earth and Planetary Science Letters* 361 (2013), pp. 356–366. ISSN: 0012821X. DOI: 10.1016/j.epsl.2012.11.001. arXiv: arXiv:1207.0853v2.
- [13] Stuart Ross Taylor and Scott M. McLennan. “The geochemical evolution of the continental crust”. In: *Reviews of Geophysics* 33.2 (May 1995), p. 241. ISSN: 8755-1209. DOI: 10.1029/95RG00262. URL: <http://doi.wiley.com/10.1029/95RG00262>.
- [14] K Hans Wedepohl. “The composition of the continental crust”. In: *Geochimica et Cosmochimica Acta* 59.7 (Apr. 1995), pp. 1217–1232. ISSN: 00167037. DOI: 10.1016/0016-7037(95)00038-2. URL: <http://www.sciencedirect.com/science/article/pii/0016703795000382>.

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