

Supporting Information:

Synthesis, Purification, and Analysis of the

Weakly-Coordinating Anion $\text{BAr}_4^{\text{F}-}$

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1 General Information

2 Materials and Methods

2.1 Synthesis of $\text{NaBAr}_4^{\text{F}}$

To a 500 mL three-necked round-bottom flask was added 1.17 g (48.1 mmol) of magnesium turnings, 0.813 g (7.40 mmol) NaBF_4 , and a stir bar. A reflux condenser, an addition funnel, and a rubber septum secured with copper wire were fit on each neck. The apparatus was attached to the Schlenk line, and was vacuumed and refilled with nitrogen three times. Using an airtight syringe, 75 mL of diethyl ether was added to the flask, followed by 0.57 mL (6.7 mmol) of 1,2-dibromoethane added similarly. The mixture was then heated on a heating mantle attached to a 40 V Variac until the solution began to bubble softly and turn a cloudy gray, signaling the initiation of the magnesium turnings.

Upon initiation, the heating mantle was removed from contact with the system, and the solution was allowed to continue stirring until the reflux ceased. While the flask cooled, 7.15 mL (41.5 mmol) of $\text{BrC}_6\text{H}_3\text{-3,5-(CF}_3)_2$ was mixed with 50 mL of diethyl ether and injected into the separate, closed addition funnel. While the solution inside the flask was warm, but no longer refluxing, the addition funnel's stopcock was turned slightly to allow the dropwise addition of the $\text{BrC}_6\text{H}_3\text{-3,5-(CF}_3)_2$ solution over the course of 30 minutes. The flask solution gradually darkened into a caramel brown, indicative of successful $\text{BAr}_4^{\text{F-}}$ synthesis.

Upon complete addition of the $\text{BrC}_6\text{H}_3\text{-3,5-(CF}_3)_2$ solution, the reaction mixture was then refluxed again for 30 minutes on a heating mantle attached to a 40 V Variac. Afterwards, the solution was allowed to cool to room temperature, and the experimental apparatus was disassembled quickly to minimize exposure of the $\text{NaBAr}_4^{\text{F}}$ solution to ambient conditions. The round-bottom flask containing the newly-synthesized $\text{NaBAr}_4^{\text{F}}$ solution was sealed with glass stoppers, and stored at room temperature for a week.

2.2 Purification of $\text{NaBAr}_4^{\text{F}}$

The $\text{NaBAr}_4^{\text{F}}$ solution was transferred into a 1 L separatory funnel, and washed with 50 mL of diethyl ether and 100 mL of brine, followed by a second washing with 100 mL of DI water. The organic $\text{NaBAr}_4^{\text{F}}$ solution was collected into a 500 mL Erlenmeyer flask, was dried with 27 g of sodium sulfate, and was de-colored with 2 g of activated carbon charcoal. The solution was mixed vigorously for five minutes, and was filtered through a Buchner funnel into a 500 mL three-necked round-bottom flask. The side necks were sealed with rubber septa and copper wire, and the $\text{NaBAr}_4^{\text{F}}$ solution was dried with a rotary evaporator to yield a thick, dark-brown oil.

A standard Dean-Stark apparatus was then assembled with the $\text{NaBAr}_4^{\text{F}}$ flask, and was connected to the Schlenk line under an inert nitrogen atmosphere. 100 mL of benzene was injected into the flask, and the solution was heated until 75 mL of organic and aqueous solvent had been removed. The $\text{NaBAr}_4^{\text{F}}$ solution was dried again using a rotary evaporator, unfortunately yielding a thick, dark-brown $\text{NaBAr}_4^{\text{F}}$ slurry which failed to dry under 10 minutes of vacuum on the Schlenk line.

The “purified” $\text{NaBAr}_4^{\text{F}}$ sample was then weighed, and a small sample was saved for NMR analysis.

3 Analysis of $\text{NaBAr}_4^{\text{F}}$

Unfortunately, the NMR instrument had a severe software malfunction near the end of the experimental period, and NMR data was unable to be obtained for the $\text{NaBAr}_4^{\text{F}}$ sample.