MA348 Numerical Analysis, Thermodynamics

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Introduction

The goal of this lab is to use various algorithms to estimate the root of a function. In this case, the function was van der Waal's equation of state, $(P + \frac{a}{v^2})(v - b) = RT$, where P is the Pressure in atmospheres, R is the Gas Constant for oxygen in atmospheres per mole-Kelvin, T is the temperature in Kelvin, v is the modal volume and $v = \frac{V}{n}$, V is the total volume, n is the number of moles of gas present, a is the measure of the average attraction between particles, and b is the volume excluded by a mole of particles. In many chemical engineering models, a very accurate modal volume of an atom or molecule, in this case oxygen, is required in order to properly construct containment apparatuses for these gases. van der Waal's equation of state is an expansion upon the classic Ideal Gas Law formula, $PV = nRT$. Using the relationship $v = \frac{V}{n}$, the Gas Law formula used for this lab is $Pv = RT$. In this lab, the values for R, a, and b are constant and known, T and P are not constant but known, and v changes relative to the previous variables based on van der Waal's equation.

Theory-Analysis

The function for this lab is van der Waal's equation, $(P+\frac{a}{v^2})(v-b) = RT$, and the objective is to estimate roots for this function. v is the changing variable, essentially the 'x' value, so the function must be solved in terms of v:

$$Pv^3 - (bP + RT)v^2 + av - ab = 0$$

This function serves as the main '$f(v)$' function for the remainder of this report. The Ideal Gas Law formula solved for v is:

$$v = \frac{RT}{P}$$

The only assumptions in this report are the oxygen Gas Constant values for R, a, and b. For this report, $R \approx 0.082054$, $a \approx 1.360$, and $b \approx 0.03183$. Additionally, the Fortran installation used to compute root values is only capable of representing 16 decimal places and is ineffective at representing small numbers.

Numerical Solution

This lab was solved using Fortran code to estimate the roots of the function and gnuplot to plot and tabulate the values. Two methods were used to compute the root values for the function- the Bisection Method and the False Position Method.

The Bisection Method involves determining a range in which there should be a root ($[a, b]$), then finding the value at the midpoint of that range (referred to as x_m and calculated with the function $\{x_m = \frac{a+b}{2}\}$). Each iteration then involves comparing x_m with the function values for $f(a)$ and $f(b)$. If $\{f(a) \times f(x_m) < 0\}$ then the b value is replaced by x_m. If $\{f(a) \times f(x_m) > 0\}$ then the a value is replaced by x_m. This method iteratively moves each edge of the range closer and closer towards the true root value of the function, halving the maximum error each iteration.

The False Position Method is very similar to the Bisection Method, except x_m is calculated with the formula $\frac{[f(b)\times a]-[f(a)\times b]}{f(b)-f(a)}$. This formula is the equivalent of the secant line from a to b, with x_m the point where the secant line crosses the x-axis. In some scenarios this method can work much faster than the Bisection Method.

Results and Discussion

Figure 1 is the plot of the function over the range $[0.5, 2.5]$ and shows that the solution is definitively between 2.0 and 2.5.

Figure 2 in Appendix A plots the error during each iteration of the algorithm function, where each line represents one of the above-mentioned estimation methods, the x-axis is the number of iterations and the y-axis is the error. In this graph scenario $P = 10atm$ and $T = 300K$. This graph seems to indicate that the False Positive Roots Estimation Method both starts with a smaller error and converges towards the desired value faster. The Bisection Method converged to a solution within 52 iterations however the False Positive Method converged in 13 iterations, as shown in the table below.

n	Bisection Error	False Positive Error
1	2.25	2.090412
2	0.125	2.74E-02
3	6.25E-02	1.37E-03
4	3.13E-02	6.75E-05
5	1.56E-02	3.32E-06
6	7.81E-03	1.63E-07
7	3.91E-03	8.04E-09
8	1.95E-03	3.96E-10
9	9.77E-04	1.95E-11
10	4.88E-04	9.57E-13
11	2.44E-04	4.71E-14
12	1.22E-04	2.66E-15
13	6.10E-05	0
14	3.05E-05	
15	1.53E-05	
16	7.63E-06	
17	3.81E-06	
18	1.91E-06	
19	9.54E-07	
20	4.77E-07	
...	...	...
45	1.42E-14	
46	7.11E-15	
47	3.55E-15	
48	1.78E-15	
49	8.88E-16	
50	4.44E-16	
51	4.44E-16	
52	0	

Table 1: Table of values for the error on each estimation method

Conclusions

The False Positive Method appears to be much more efficient without sacrificing any error, in the future that may not always be the case though. If this were to be done by hand the most accurate way would be the False Positive Method as the Bisection Method is much slower to converge. There are other ways to optimize this function. The Bisection and False Positive Methods of finding roots are just a two from a plethora of various solvers and algorithms that accomplish the same this.

Appendix A

Important Graphs

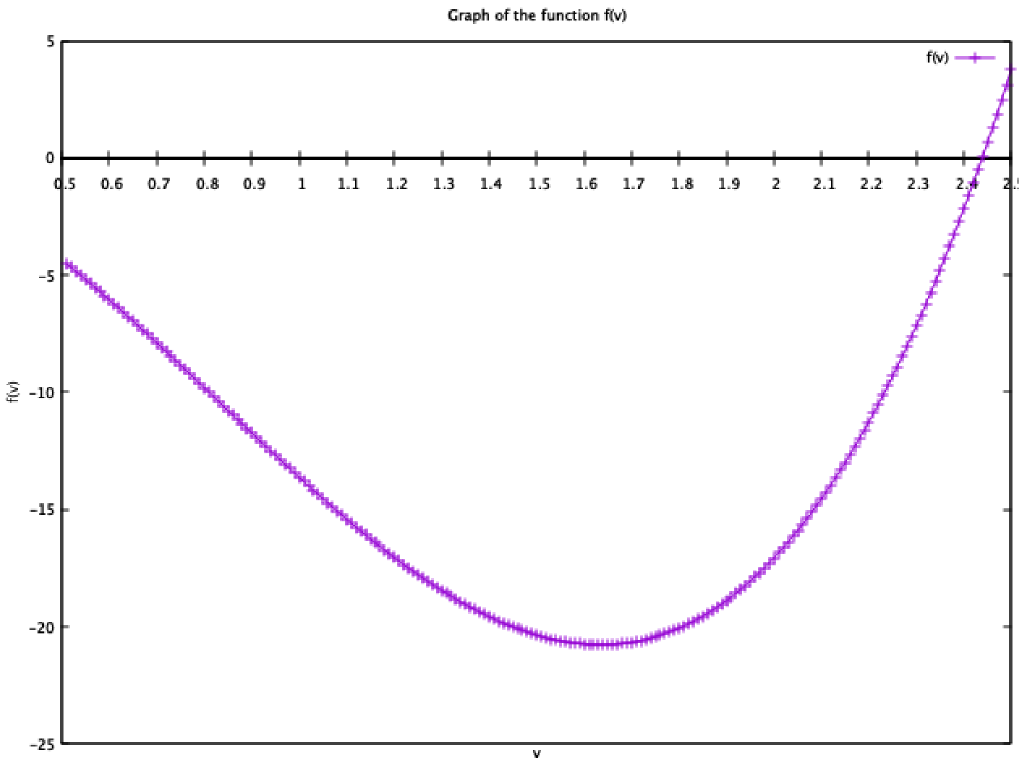

Figure 1: Graph of the $f(v)$ function

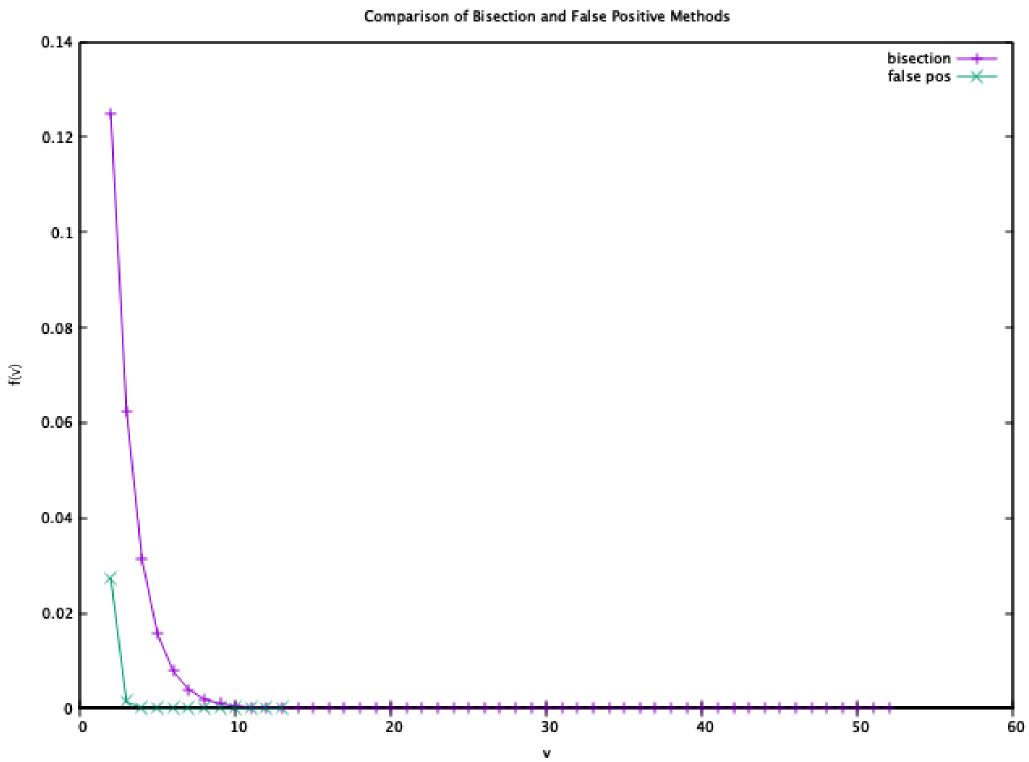

Figure 2: Bisection and False Position Errors when $P = 10atm$ and $T = 300K$

Appendix B

Graphs of each combination of Pressure/Temperature. The green line represents the graph of the Ideal Gas Law on the interval while the purple line represents the False Position Estimation function on the interval. v is the modal volume estimated by the False Position Method.

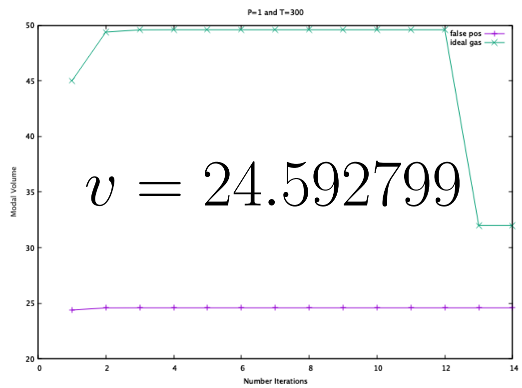

Figure 3: $P = 1atm$, $T = 300K$

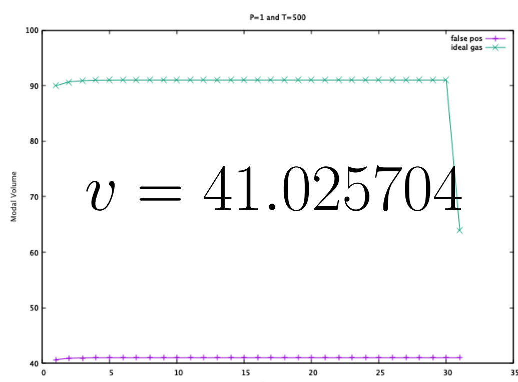

Figure 4: $P = 1atm$, $T = 500K$

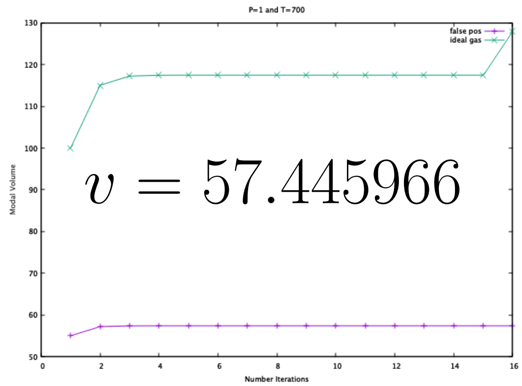

Figure 5: $P = 1atm$, $T = 700K$

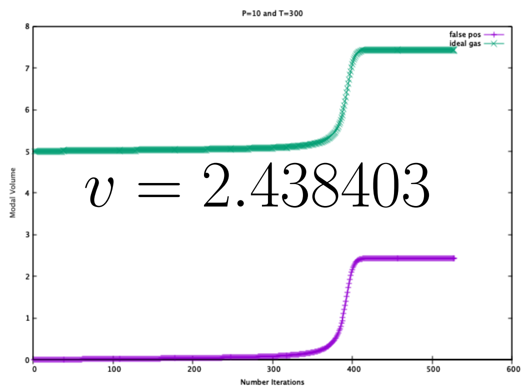

Figure 6: $P = 10atm$, $T = 300K$

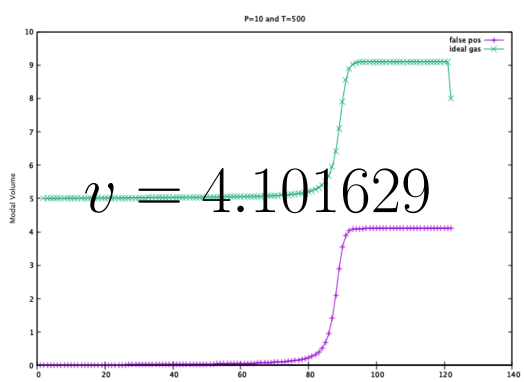

Figure 7: $P = 10atm$, $T = 500K$

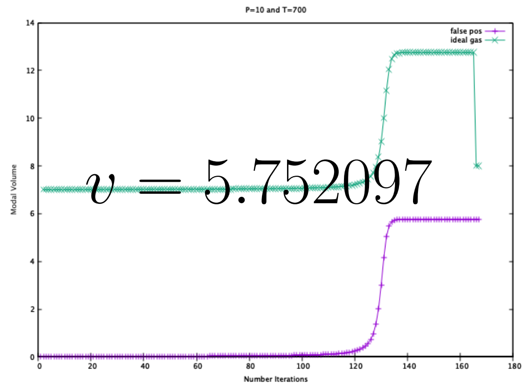

Figure 8: $P = 10atm$, $T = 700K$

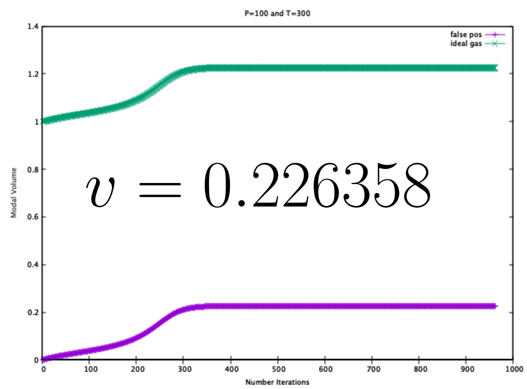

Figure 9: $P = 100atm$, $T = 300K$

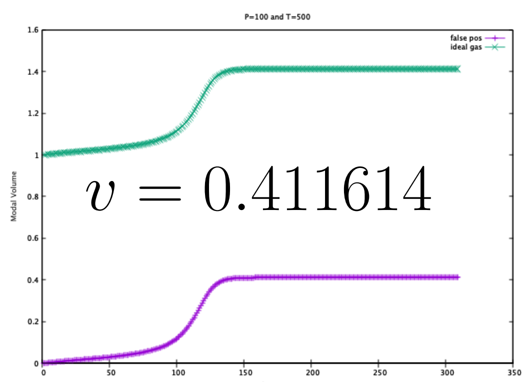

Figure 10: $P = 100atm$, $T = 500K$

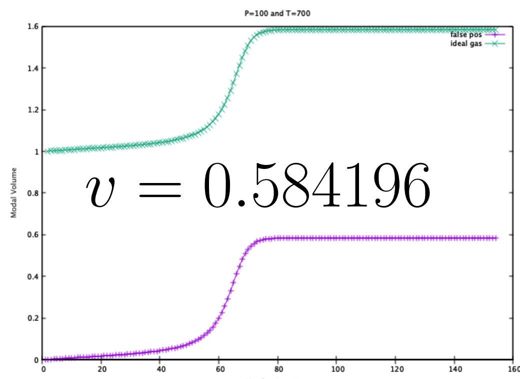

Figure 11: $P = 100atm$, $T = 700K$

Appendix C