

Franck-Hertz Experiment; Measuring Transition Energies of Neon and Mercury

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Abstract

We recreated the canonical Franck-Hertz experiment and measured the energy required to excite neon and mercury atoms. Hey! What happened to the key discovery F-H made, and why they received the Nobel prize? That is, you didn't just measure the energy required (which could equally well describe an experiment on a classical as well as a quantum system), you more importantly demonstrated that the energy levels in atoms are QUANTIZED. That should be in your lead sentence!

Be more specific in the next sentence (some suggestions in magenta): Using an applied voltage varying/ranging/spanning from 0 - 100? V, up to 100? V, or similar language we accelerated electrons through a gas of either neon or mercury atoms you did both, so consider rephrasing to something like 'first neon, then mercury' or 'a gas of the element under study' or similar. OK, but why? you try to address that later, of course, but adding a short little phrase like the following (not particularly artfully worded) example to the end of the previous sentence would clarify your PURPOSE and METHOD right away instead of guessing how the second sentence in your abstract is related to the first sentence: '...gases of various elements so as to study when collisions between the electrons and the gas molecules would result in energy transfers to the atoms' or perhaps 'how much kinetic energy an electron needed to acquire before it had enough energy to promote an atom to its next highest energy level' or 'your wording here'

This next paragraph is a bit confusing. Phrases like 'the number of electrons... decrease' wrongly suggests that electrons are disappearing, instead of just slowing down when they collide inelastically with gas atoms (thereby transferring part of their kinetic energy to the atoms). A drop in current doesn't mean a decrease in the number of electrons in the system, it means a decrease in their speed (because of a decrease in kinetic energy). We measured the electron current reaching the anode. As we increased the magnitude of the accelerating voltage we saw the number of electrons reaching the anode decrease to a minimum at evenly spaced intervals. The drops in the electron current were caused by electron obtaining enough energy, from the accelerating voltage, to cause the neon or mercury atoms to reach an excited state. After the colliding inelasticity with an atom the electrons did not have enough energy to reach the anode resulting the drop in electron current. By measuring the difference between the minima in the electron current we were able determine the excitation energy of neon. isn't this true of Hg as well?

Mercury has an optimal temperature were the number of minima is maximized and their am-

plitude is still discernible. After determining the optimum temperature range with direct measurements of electron current, we used a lock-in detector to get more precise measurements of the minima.

We plotted the difference in energy against the number of the minimum what does ‘the number of the minimum’ mean? and used fit lines to extrapolate to a minimum number of .5. huh? An abstract needs to “stand alone.” That is, it needs to make sense on its own, even if you don’t read the rest of the paper. No one reading the abstract will understand why you would do this particular extrapolation or what a minimum number of 0.5 means. Since what is really important is that this is method of using this information (your measurements) to infer the energy of the first excited state in Hg (or perhaps the first two???) after correcting for something involving mean free path — you get to chose the more specific wording to replace ‘something’ — why not say that? Averaging all of our data we got excitation energy for neon of $18.5 \pm .4$ V and for mercury of $4.58 \pm .4$ V. Volt is not a unit of energy. eV is a unit of energy. ‘electron volts’ are not volts, they are energies, just as $(1.609 \cdot 10^{-19}C)(18.5V) = 29.8 \cdot 10^{-19}J$ is a unit of energy (Joules). The difference between eV and J is that in eV units, the charge on an electron is ‘1 e’ rather than $1.609 \cdot 10^{-19}$ Coulomb.

Remember that the abstract should stand alone. Therefore, don’t leave the reader in suspense, give the actual values so they can compare theory and experiment without having to read the rest of the paper: Our results for both mercury and neon agree with the expected values of ? eV for the excitation energies of neon and of ? eV and ? eV for the first two excited states in mercury. However neither of our values were precise enough to correspond to a single excitation energy because of the resolution of our instruments. is it possible you are seeing more than one excitation energy? could you be exciting atoms into the 1st and the 2nd excited states from the ground state? or to the first excited state from neutral atoms and also from ionized atoms?

I. INTRODUCTION

In 1913, Niels Bohr postulated that electrons within an atom can only occupy certain discrete energy levels in his famous Bohr model for atoms. This theory also requires any change in an electron's energy to be discontinuous as well, equaling to a difference between two of the energy levels. It was not until the next year, 1914, that James Franck and Gustav Hertz provided experimental evidences to Bohr's energy level postulation. In a vacuum chamber they designed to study the dynamics of the electrons passing through a vapor of mercury, Franck and Hertz observed that a collision between an electron and a mercury atom could result in the electron losing a specific amount of its kinetic energy. The electrons with exactly this amount of kinetic energy were completely decelerated and stopped during the collisions with the atoms, while the faster electrons kept moving forward after the collisions, but lost the same amount of kinetic energy. Slower electrons may collide with the mercury atoms as well, but they did not lose an appreciable amount of kinetic energy. They also discovered that an electron could lose a multiple of this amount of energy when it passed through the vacuum chamber, provided that it is fast enough to endure this loss. Franck and Hertz's discovery was consistent with the Bohr model because the collision between an electron (unbound) and a mercury atom could transfer the kinetic energy of the unbound electron to the atom and excite a bound electron inside the atom from the lowest energy level to a higher one. The energy difference between these two energy levels was equal to the kinetic energy loss that occurs to the fast electrons. One should understand that although the Bohr model is regarded as a simplistic and insufficient interpretation of quantum mechanics, its proposal of the bound electron's energy quantization nonetheless provided insights into the atom's internal structures. Later on, Franck and Hertz published another paper about the emissions of light from the atoms that acquired energy from the collider electrons— the excited electron inside the atom decayed into the lowest energy level, emitting a photon with frequency $\nu = E/h$, where E is the energy difference between the two levels, and h is the Planck constant. The kinetic energy absorption and the photon emission jointly demonstrate the quantization of bound electron's energy levels.

Franck and Hertz's apparatus used a transparent vacuum glass tube to contain the experiment environment and enable visibility of the light emission patterns. A heated filament (cathode) was used to emit free electrons via thermionic emissions. The vacuum tube was

heated, so the mercury inside turned into vapor and filled the entire vacuum tube. There was a control grid, held at a distance from the filament, that was used to create an accelerating voltage. The electric field resulted from this accelerating voltage pulled the electrons towards the grid, and then further to an anode behind the grid to be collected. The anode was held at a voltage slightly negative to the accelerating grid to discourage the electrons that lost their kinetic energy from reaching the anode. Franck and Hertz changed the accelerating voltage and measured the current due to the electrons that reached the anode. They discovered that the current experienced several sharp drops at certain accelerating voltages. The spacings between the voltages of two adjacent drops were always approximately equal—4.9 V, meaning that an electron lost 4.9 eV of energy every time it collided with a mercury atom and excited a bound electron inside the atom from the ground level to a higher level. Franck and Hertz also discovered that the light emitted from the apparatus had wavelength 254nm, corresponding to exactly 4.9 eV of energy by $\nu = E/h$.

We attempt to reproduce the Franck-Hertz experiment in the lab. We conduct the experiment with both neon gas and mercury vapor. With the modern understanding of the atom’s internal energy structure, we would be able to find the exact energy level transition involved in the experiment. We use apparatuses similar to that of the original Franck-Hertz experiment, but with some modifications to improve signal-to-noise ratio that will be introduced in the next section. For both neon and mercury, we measure the currents based on varying accelerating voltages. In addition to the original Franck-Hertz experiment, we add temperature control to the apparatus for the mercury atom and study the temperature dependence of the internal transition of the mercury atom. However, in the mercury experiment, we encounter large uncertainties in our results, and to improve our measurements, we implement a lock-in technique for precision measurement of the voltage values where we have sharp drops in current.

II. METHODS

We perform the Franck-Hertz experiment with two types of atoms— neon and mercury. In each setup, we ought to bombard the atoms using electrons whose kinetic energy is controllable. We then can collect the electrons after they pass through the bombardment area, and measure the current due to this collection. The electrons are accelerated by a

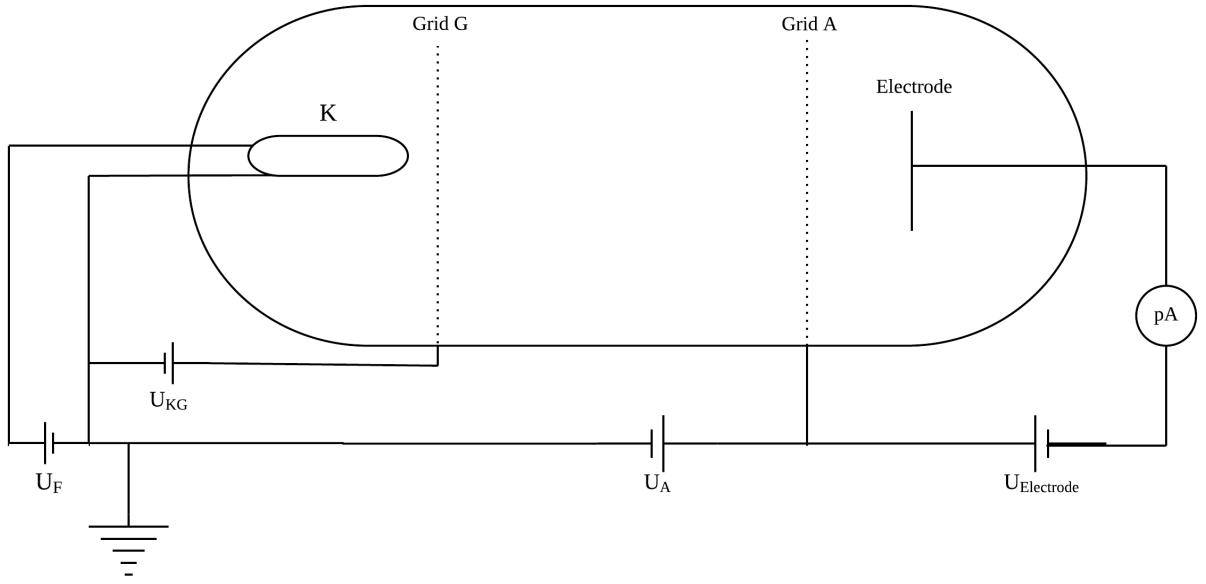


FIG. 1. Experiment setup for the neon sample.

voltage we can control using a power supply. If we measure the current due to electron collection under different accelerating voltage values, we can find the electrons' final kinetic energy's (after bombardment) dependence on their initial kinetic energy.

A. Neon Gas

We use a glass tube filled with neon gas at room temperature to perform the Franck-Hertz experiment. Fig 1 shows an illustrative diagram of our setup. K is a filament that emits electrons when heated, and it serves as the cathode. The current through the filament is controlled by the power source U_F . We ought to hold this current invariant through out the experiment because it controls the rate of electron emissions, thus any variance in this current could introduce a shift in the final collected current. The current we use to hold constant and optimize the electron emission rate is 149 mA.

The emitted electrons from the filament may roam freely near the cathode, so they we install a grid G in the vicinity of the cathode and applied an initial accelerating voltage to it so that the electrons were drawn towards the grid. The voltage on grid G is held at 2.6 V through a power source U_{KG} . This initial preparation makes the emitted electrons be effectively pulled by the field created by the the accelerating voltage. Grid A is where

the accelerating voltage is applied, by another source U_A , to make electrons acquire kinetic energy in an amount we can control. The accelerating voltage U_A can be set as any value between 0 and 80 V, manually or by ramp control. Ramp control sweeps through the 0-80 V range 60 times a second. It is worth noticing that the voltage applied to grid G contributes to the acceleration of the electrons as well, so the total accelerating voltage in the tube is the sum of the two voltages. Behind grid A, there is an electrode serving to collect the electrons flying through grid A. This collector electrode has to be held at a voltage relatively negative to grid A. This way the electrons that lose all of their kinetic energy during collisions are unable to reach the collector, because we do not want those electrons to contribute to the current. We have one multimeter measuring the accelerator voltage and another one measuring the current through the collector electrode. We ought to control the power supply U_A to increase the accelerating voltage. Initially we attempt to utilize ramp control to better resolve the change the current due to the change in the accelerating voltage, but have to give up since the multimeters we use are not capable of effectively responding that fast. Instead, we manually control the accelerating voltage and try to increase it as slowly as possible. The result current is plotted versus the accelerating voltage.

B. Mercury Vapor

Performing the Franck-Hertz experiment on mercury requires slight modification to the apparatus. We use a similar glass tube to bombard the mercury atoms with accelerated electrons, but this time the mercury has to be vaporized first. Under room temperature, the mercury inside the tube is in its liquid state. To get enough mercury vapor, we need to heat the glass tube up to a relatively high temperature. We do not have to reach mercury's boiling point as all we need is a sufficient amount of mercury vapor to be bombarded by the electrons. This requirement in the mercury experiment prompts us to seek the optimal temperature to measure the current's dependence on voltage, so we will perform measurements under different temperatures. We will use an oven enclosing the whole glass tube to control the temperature.

As seen in Fig 2, the initial accelerating voltage is not present in this experiment. A voltage of 6.5 V is applied to the electron-emitting filament and is held constant throughout the experiment. We still use an accelerator grid A to increase the electrons' kinetic energy.

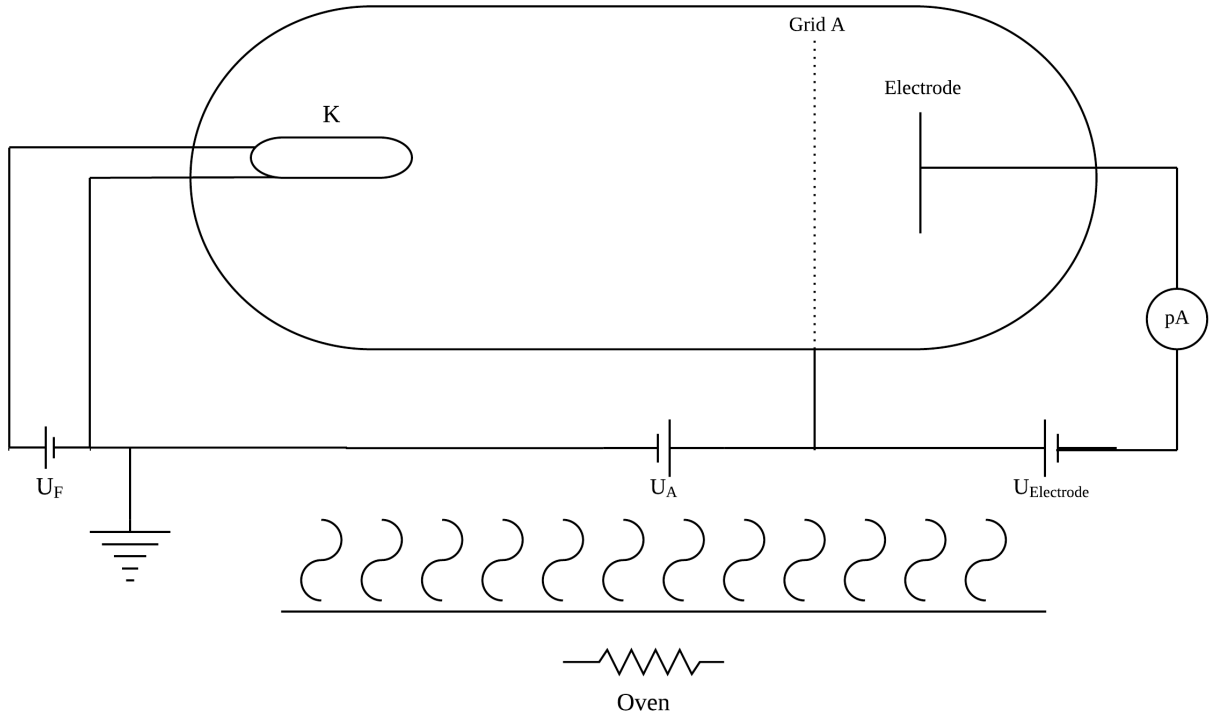


FIG. 2. Experiment setup for the mercury sample.

This time we can tune between 0 V and 70 V for the accelerating voltage. A retarding voltage of 1.5 V is applied to the collector electrode to serve the same purpose as in the neon example. We will again manually control the accelerating voltage to increase it as smoothly as possible. The results of the two experiments are presented in the next section.

III. RESULTS

A. Neon

We did four data runs using neon. For each run we recorded the accelerating voltage (x data) and the electron current measured by the anode, this current was recorded as a voltage measured across an internal resistor (y data). Each of our runs showed three discernible minima in the voltage corresponding to electron current. These dips are the voltages just before the electrons have enough energy (from the accelerating voltage) to reach the anode even after an inelastic collision with a neon atom. In an attempt to be concise we are not showing all of our data, Figure 3 is representative of all the data we collected. The apparent

double minima, see the second and third dips in Figure 3, is most likely caused by energy levels with very similar excitation energies.¹ Also of interest is the voltage corresponding to the steepest negative slope which is when the majority of electrons have enough energy to excite the neon atoms. However the location of the steepest negative slope was difficult to determine from our data as can be seen in Figure 3.

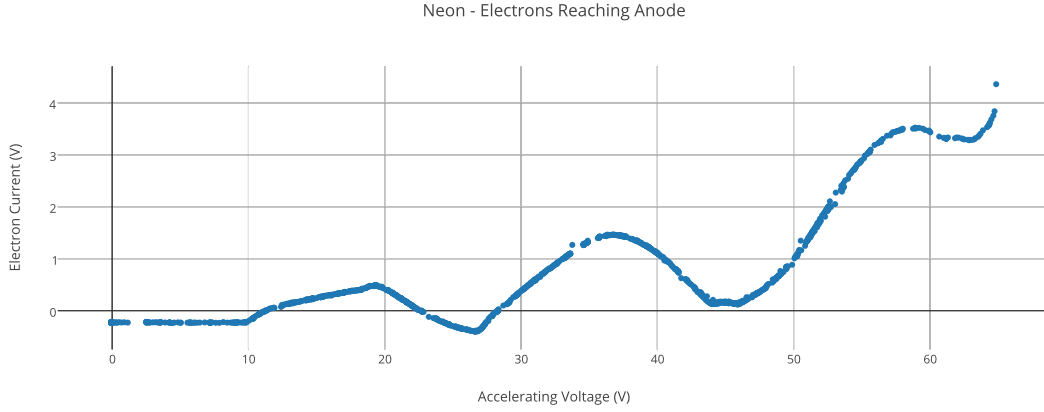


FIG. 3. Neon data with electron current as a voltage plotted against accelerating voltage. The three minima correspond to each electron exciting three neon atoms before reaching the anode. The double minima in the second and third dips is due to multiple excitation levels with similar energies.

B. Mercury

Since the number of discernible dips for mercury depends on temperature, we collected data in 10°C increments starting at 150°C and ranging to 210°C . From this data, see Figure 4, it can be seen that there is an optimum temperature at around 200°C at which the most minima can be observed.

Using the information that we obtained about the optimum range of temperature, we took a more data using a lock-in detector. The lock-in blocked the upward slope in electron current, see particularly 150°C in Figure 4. With the lock-in we took data from 185°C to 215°C in increments of 5°C . The output from the lock-in, see Figure 5, is the derivative of the direct output without the lock-in. Therefore, the minima in the lock-in output correspond to the steepest negative slope of the direct output and the places the lock-in data passes

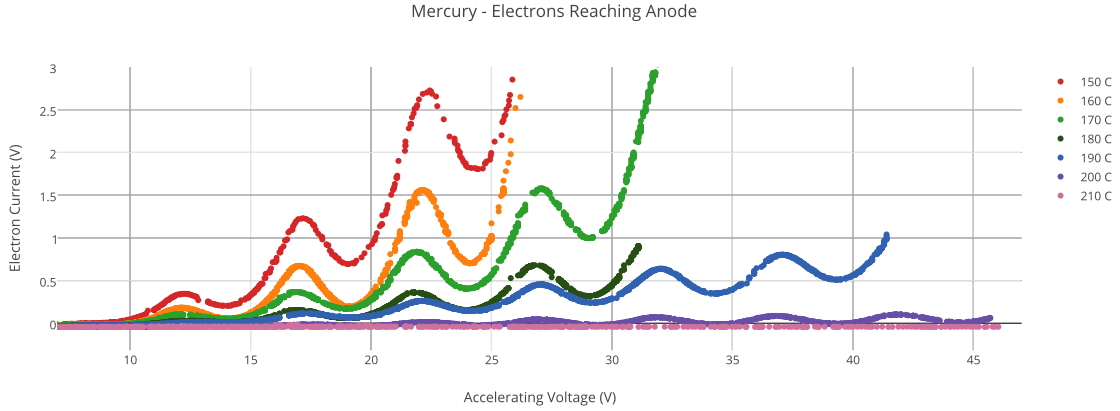


FIG. 4. Mercury data for temperatures ranging from 150°C to 210°C. The range of temperatures shows the optimum temperature to be approximately 190 - 200°C. For higher temperatures, such as 210°C, the minima in electron current are not discernible. For lower temperatures, for example 150°C and 160°C, there were fewer minima before the mercury atoms are ionized.

through zero correspond to the minima and maxima of the direct output. The lock-in detector was highly sensitive to ionization of the mercury atoms. Our 185°C data already showed significant reduction in the number of minima because of ionization. Because of space considerations we are not showing all of our lock-in data. Figure 5 is the data collected for 205°C and is a representative sample of the data collected.

IV. ANALYSIS

A. Neon

To determine the energy needed to excite the neon atoms we needed to find the spacing of the minima in electron current (which we measured as a voltage.) We visually determined the locations of the minima and the steepest negative slopes with uncertainty for each of our data runs, see sample data in Figure 3. We then calculated the change in voltage between adjacent minima and between adjacent steepest slopes. These values can be seen in Table I.

We then averaged the first column and got $17.9 \pm .3$ eV as the average difference between the minima, where the uncertainty is the standard deviation of the mean. From the difference between the steepest slopes we got an average value of $19.19 \pm .19$ eV. We then averaged

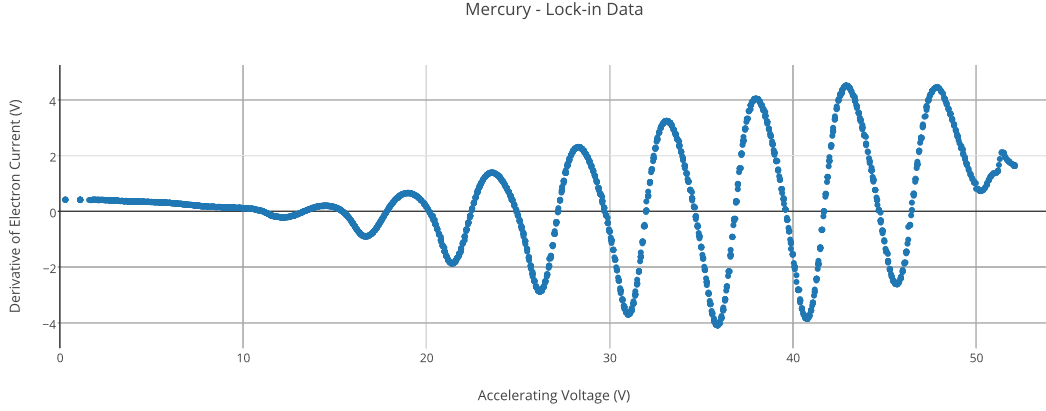


FIG. 5. Sample mercury data at 205°C using the lock-in detector. The lock-in output is the derivative of the direct output in Figure 4. Minima in the voltage correspond to the steepest negative slope of the electron current and the lock-in x intercepts with positive slopes correspond to the minima in the direct output.

TABLE I. The difference in voltage for adjacent minima and adjacent steepest negative slopes obtained from all of data runs. Our best value is the average of these values. The large errors are from the difficulty in determining the location of minima and steepest slopes.

ΔV Minima	Error ΔV Minima	ΔV Steepest Slope	Error ΔV Steepest Slope
19.0	2.0	19.2	1.5
16.4	2.0	19.5	1.5
18.5	1.3	20.0	2.5
17.6	1.6	18.5	2.5
18.4	1.2	19.5	2.0
17.0	1.5	18.8	2.0
18.5	1.5	18.5	2.5
17.5	2.5	19.5	2.0

those two values, to get a final value of $18.5 \pm .4$ eV where uncertainty is again the standard deviation of the mean. This value agrees with our expectations because neon has a large number of transitions with energies ranging from 18.3 eV to 18.9 eV.

B. Mercury

Because of the temperature dependence of the mercury data, the process by which we found the energy level was more involved. We began our analysis in a similar way to our analysis of neon by visually determining the location of the minima and steepest negative slope. However, when we calculated the difference in voltage between minima this method gave us an uncertainty on the order of 10%. We then plotted the distance between minima versus the minimum number as suggested by Rapior, Sengstock and Baev¹ and fit linear lines to the data for each temperature. Rapior, Sengstock and Baev found that the fit lines from each temperature converged toward the minima number .5. We found that the fit lines to our data did not converge at dip number .5, furthermore the uncertainty for our minima was large enough to make our results imprecise and unsatisfying.

To improve our results, we used that data collected using the lock-in detector. Because the lock-in output is the derivative of the direct output it is the x intercept with a positive slope that corresponds to the minima in the direct data and the minima of the lock-in output corresponds to the steepest negative slopes of the direct output. Figure 6 shows the data from the direct output steepest slope and minima and the corresponding lock-in outputs (minima and positive slope x-intercept) for 200°C. The data with the least uncertainty was that data from the lock-in using the positive x intercepts to calculate the differences in voltage. From this data we found the values seen in Table II.

The results of Table II are plotted in Figure 7. Following the analysis outlined by Rapior, Sengstock and Baev, we used our fit lines to extrapolate to the voltages corresponding to a minima number of .5. Averaging these values gave us $4.46 \pm .12$ V where the uncertainty is the standard deviation of the mean. We repeated this analysis for all lock-in and direct data. The results obtained from all of our mercury data are in Table III, note the much larger uncertainties for the direct output data than the lock-in data.

Averaging the best values from each set of data, from Table III, we got a final value of $4.56 \pm .4$ V. This value agrees with the known value of the lowest excitation 6^1S_0 to 6^3P_0 (4.67 eV) of mercury and also with the second excitation 6^1S_0 to 6^3P_1 (4.89 eV.)¹



FIG. 6. All the 200°C data shown with the relative errors. The blue data is the steepest negative slope from the direct data and the corresponding lock-in minima data. The lock-in data has less than half the error of the direct data. The red data is the direct output minima and the lock-in positive x-intercepts. The lock-in positive x-intercepts had the least uncertainty.

TABLE II. The difference in voltage for consecutive positive slope x-intercepts in from the lock-in output. The first dip was indiscernible for the 215°C data and the higher minima of the 195°C and 185°C data were lost due to ionization. The error for these values is 1.3%

Dips	ΔV 215°C	ΔV 210°C	ΔV 205°C	ΔV 200°C	ΔV 195°C	ΔV 195°C	ΔV 185°C
1 - 2		4.67	4.53	4.42	4.36	4.58	4.51
2 - 3	4.63	4.57	4.63	4.61	4.63	4.78	4.82
3 - 4	4.67	4.69	4.63	4.73	4.84	4.87	4.90
4 - 5	4.74	4.79	4.89	4.83	4.87	4.93	
5 - 6	4.76	4.80	4.87	4.88	4.95	4.99	
6 - 7	4.76	4.82	4.84	4.91		5.05	
7 - 8	4.76	4.73	4.77	4.85		4.90	

V. DISCUSSION

All of our data was consistent with our expectations based on theory and past experiments. We were able to see the drops in electron current at regular intervals as electron gained enough energy to excite the neon or mercury atoms. For mercury, we saw that there

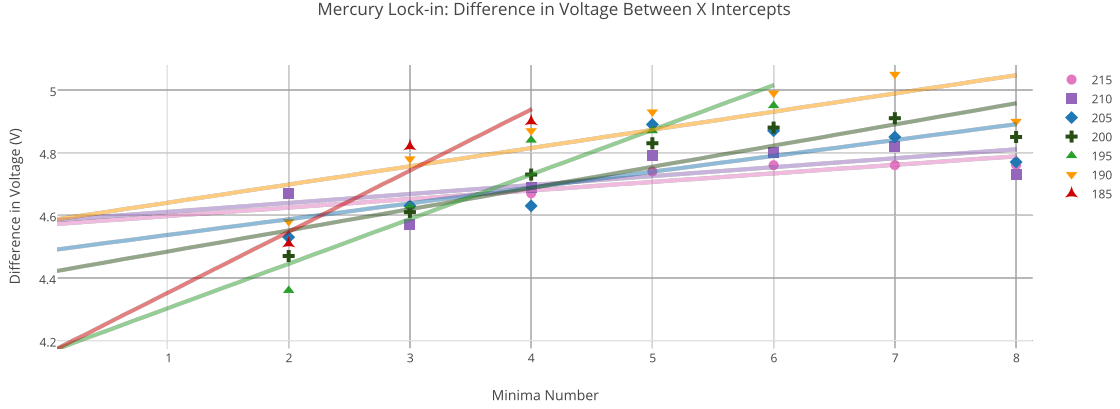


FIG. 7. The plot of the difference in voltage between minima versus the number of the minimum. All values have an uncertainty of 1.3%. The lines are linear fits to the data for each temperature. The two fits with steeper slopes correspond to the 195°C and 185°C data sets which were truncated due to ionization.

TABLE III. A summary of the best values and uncertainties from each of our data sets. All of these values, except the lock-in positive x-intercept value, agree with the expected energy value for mercury transitions.

Direct Steepest Slope	Lock-in Minima	Direct Minima	Lock-in Positive Slope X-Intercept
4.6 ± 0.7	4.58 ± 0.24 V	4.6 ± 0.4	4.46 ± 0.12 V

was an optimal temperature range when drops in electron current were maximized in number and still discernible in amplitude. As suggested by Rapior, Sengstock, and Baev, the difference in voltage between our minima tended to increase as the number of the minimum increased.

Our final values for the energy needed to excite mercury show a stark trade off of accuracy versus precision. Our direct data values have a very large uncertainty (10 - 15%), see Table III, so they are not very precise, but they are accurate. Whereas our lock-in x-intercept data has a much smaller uncertainty (2.7%) a much higher precision, but it is no longer accurate because it is not in agreement with the known value of the excitation energy.

One of the biggest limitations in our experiment was our uncertainty in the precise voltages corresponding to the minima and steepest slopes. Particularly for the lock-in data,

but also for the direct data, our uncertainty was largely based on the resolution of our graphs. The direct output data used the power supply that came with the apparatus, it had a resolution of 0.1 V. For the lock-in data we used an independent power supply that also increased the accelerating voltage in increments of 0.1 V. These increments of 0.1 V gave us an uncertainty on that order of magnitude and set a lower limit on our uncertainty. If we had increased the accelerating voltage in increments of 0.01 V, we would have had better resolution of our data and less uncertainty in our analysis.

VI. CONCLUSION

18.5 % .4 V, the value we got for the neon excitation energy, is in agreement with a large number of excitation energies for neon between 18.3 - 18.9 V. We were not able to determine the energy transitions more precisely from the data that we collected. With our apparatus we were only able to obtain three minima before the neon ionized. Using different instrumentation we might be able to obtain more data and calculate the transition more precisely.

Our final value for mercury energy transitions was $4.56 \pm .4$ V which agrees with the known values for both the lowest energy transition, 4.67 eV, and the next lowest energy transition, 4.89 eV. We would hope for a more precise answer that would correspond with just one of the energy transitions. However, the resolution of our data was restricted by the apparatus we used.

It is possible that a most sensitive measurement of electron current we could have seen more minima below the first minima that we could detect. Having more minima would increase the number of data points used to fit lines and extrapolate the transition energy. Since some of our fit lines were made with only a couple of data points, increasing the number of points would help to average out outliers that are undiscernable with only a few points.

Despite our difficulty in precisely determining the excitation energy of mercury, the overall results of this experiment are still as fascinating as they were for Franck and Hertz. This experiment clearly shows that electrons must have a certain quantified amount of energy before they can excite atoms. This result supports the premise of quantum mechanics and

makes an abstract theory more tangible.

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¹ Gerald Rapior, Klaus Sengstock, and Valery Baev, "New Features of the Franck-Hertz Experiment," Am. J. Phys. **74**, 423–428 (2006).

² Adrian C. Melissanos and Jim Napolitano, *Experiments in Modern Physics* 2nd edition (Academic Press, Boston, 2003).

³ Hyper-Physics: Franck-Hertz Experiment Website <<http://hyperphysics.phy-astr.gsu.edu/hbase/frhz.html>>