

Experiment 37 Stoichiometry Answers

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03 Stoichiometry with answers - PDF Free Download

Stoichiometry. Observations. Step 6 Data Table. Mass of Copper (II) Sulfate Mass of Iron Mass of Dry Copper Data/Observations. 1. Which of the observations you made support that a chemical reaction has occurred in this experiment? Analysis and Conclusions. 1. Which reactant was limiting? 2. What observations support this? 3.

Experiment 37 - Henry County Schools / Overview

Experiment 37 . Stoichiometry . Problem. ... From this information, you will determine which equation applies to the reaction you run in this experiment. Prelaboratory Assignment. Read the entire experiment before you begin. Answer the Prelaboratory Questions. 1.

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forming the question, or need help seeing how the lab relates to stoichiometry; performing the stoichiometry; special care should be spent making sure students are using the acetic acid mass, not the mass of the vinegar. To save time I have made this Stoichiometry lab answer key so I can quickly check student work. creating a step-by-step procedure

Stoichiometry lab answer key - BetterLesson

Chemistry: Stoichiometry and Baking Soda (NaHCO_3) Purposes: 1. Calculate theoretical mass of NaCl based on a known mass of NaHCO_3 . 2. Experimentally determine the actual mass of NaCl produced. 3. Calculate the percent yield for your experiment. Reaction Equation: $\text{NaHCO}_3 (\text{s}) + \text{HCl}(\text{aq}) \rightarrow \text{NaCl}(\text{s}) + \text{CO}_2 (\text{g}) + \text{H}_2\text{O}(\text{l})$

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