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Analyzing Oxidation-Reduction Reactions When chemical bonds form, electrons are either lost, gained or shared. Metals lose electrons. This is what happens when iron rusts. When the iron, a metal, combines with oxygen, a non metal, to form rust, it loses electrons. This process is called oxidation even when the nonmetal is not oxygen. Nonmetals

Analyzing Oxidation-Reduction Reactions - Evan's Regents ...

Writing Half Reactions During a redox reaction electrons are both lost and gained. The metal loses and the non metal gains. An equation showing either the gain or the loss of electrons but not both

is called a half reaction. Consider the reaction below: $2\text{Mg(s)} + \text{O}_2\text{(g)} \rightarrow 2\text{MgO(s)}$ Magnesium loses electrons while oxygen gains.

Writing Half Reactions - Evan's Regents Chemistry Corner

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Look up the oxidation state of each element or ion and reduce to lowest terms. Then cross over the oxidation states without the sign to find the subscripts as shown in the diagram to the right. Using the Cross -Over Rule, determine the formula for compounds of the elements and polyatomic ions below, and write your answer in the answer space ...

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Study the rules for assigning oxidation numbers and examine the sample problem below. Then determine the unknown oxidation state in each example. RULES FOR ASSIGNING OXIDATION NUMBERS 1. Oxidation numbers for atoms that are free elements are always zero 2. The oxidation numbers of ions are the same as the charge on the ion 3.

Regents Chemistry Date: ----- ~ 1

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