

## *Quantum Numbers Yahoo Answers*

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### Quantum Numbers Yahoo Answers

where each of the above obeys the rules for quantum numbers. If you wish to limit this only to the valence electrons then we can be more explicit 4s electrons  $[4, 0, 0, \{-\frac{1}{2}, \frac{1}{2}\}]$  - to a maximum of 2 electrons 4p electrons  $[4, 1, \{-1, 0, 1\}, \{-\frac{1}{2}, \frac{1}{2}\}]$  - to a maximum of 6 possible electron quantum numbers

### Quantum numbers? | Yahoo Answers

Yahoo UK & Ireland Answers ... Pretty sure of these.  $n$  is the principle quantum number and corresponds to electronic configuration,  $l$  is angular momentum quantum number and indicates subshell ( $0=s$ ,  $1=p$ ,  $2=d$ ,  $3=f$  orbitals),  $m_l$  is the magnetic quantum number and can equal from  $l$  to  $-l$ . And  $m_s$  is the spin quantum number and can only be  $+1/2$  or  $-1/2$  ...

### quantum numbers? | Yahoo Answers

Best Answer: There are four quantum numbers: Principal quantum number Angular quantum number Magnetic quantum number and Spin quantum number The most widely studied set of quantum numbers is that for a single electron in an atom: not only because it is useful in chemistry, being the basic notion behind the ...

### Quantum numbers? | Yahoo Answers

An electron has the  $n$ ,  $l$ ,  $m_l$ , and  $m_s$  quantum numbers 5, 4, 1,  $1/2$ . Choose the set of quantum numbers below that corresponds to an electron that is in the same orbital and paired with this electron. a.) 5, 4, 1,  $-1/2$  b.) 5, 4,  $-1$ ,  $1/2$  c.) 5, 4, 1,  $1/2$  d.) 5,  $-4$ , 1,  $1/2$

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Best Answer: Each electron has a set of four numbers, called quantum numbers, that specify it completely; no two electrons in the same atom can have the same four. That's a more precise statement of the Pauli exclusion principle. Primary quantum number, which is given the symbol  $n$ , tells you which of the "main" energy levels you're in.

### What are quantum numbers? And how many are there? - Yahoo

Yahoo Malaysia Answers ... Identify allowable combinations of quantum numbers for an electron.?  $n = 3$ ,  $l = 2$ ,  $m_l = 2$ ,  $m_s = 1/2$  ... ( $m_s$ ) have strict rules. Identify allowable combinations of quantum numbers for an electron? Identify allowable combinations of quantum numbers for an electron. Select all that apply? Answer Questions.

### Identify allowable combinations of quantum numbers ... - Yahoo

QUANTUM NUMBERS WORKSHEET 1. State the four quantum numbers, then explain the possible values they may have and what they actually represent.  $n$  - Principal Quantum Number: represents the energy level the electron is in, linked to the

### QUANTUM NUMBERS WORKSHEET answers

Chemistry Question: Quantum numbers? Answer Questions C<sub>3</sub>H<sub>4</sub>O<sub>2</sub>.A 0.472-g sample is combusted to produce 0.267 g of water and 0.446 g of carbon dioxide.Find theoretical yield of C and H in the acid?

### Chemistry quantum numbers help? | Yahoo Answers

Identify which sets of quantum numbers are valid for an electron. Each set is ordered ( $n, l, m_l, m_s$ )?

### Identify which sets of quantum numbers are valid ... - Yahoo

Yahoo India Answers ... Answers. Best Answer: The question of how many quantum numbers are needed to describe any given system has no universal answer, hence for each system, one must find the answer for a full analysis of the system. A quantized system requires at least one quantum number.

### which of the following combinations of quantum numbers are ...

Best Answer: N can be any whole number value greater than 0, which knocks out the 11th choice..  
L can be any whole number value between 0 and n-1 for a given n. This eliminates choices 2, 6, 8  
and 12. Ml can be any whole number between +l and -l.

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