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out how to get past. The rest of the answers might be posted by me again or contributed by  
someone else. \*Note: You may use the skip signs whenever you fell like it because this guide tells  
you at least how to get to question 66\*

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(b) The molecular formula, CH<sub>4</sub>, gives the number of each kind of atom. Follow steps 1-3 as given  
in the student textbook. Step 1 Carbon has the lower electronegativity in the molecule and will be  
the central atom. Step 2 Find the total number of valence electrons. (1 atom C × 4e<sup>-</sup>/C) + (4 atoms  
H × 1e<sup>-</sup>/H) = 8e<sup>-</sup> Determine the total number of electrons required for a noble gas

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