

## *Saturated Unsaturated Supersaturated Graph Answers*

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**Saturated Unsaturated Supersaturated Graph Answers**

Unsaturated, Saturated, and Supersaturated Examples. Check Your Answer The liquid inside reusable hand warmers will crystalize over time. However, they can be reheated in the microwave and go back to the liquid form. Explain why the crystals form and why reheating the hand warmers "removes" the crystals. Check Your Answer.

**Types of Solutions: Saturated, Supersaturated, or Unsaturated**

A graph can illustrate what solution is saturated and unsaturated. If the point is on the line, then the solution is saturated, while if it is below the line, the solution is unsaturated. share with ...

**What is saturated and unsaturated solutions - answers.com**

Best Answer: The solution will be saturated when the grams of solute dissolved = the solubility of the salt. In other word the solution will be saturated at the line on the graph. In other word the solution will be saturated at the line on the graph.

**Chemistry: How can you tell if the solution is saturated ...**

Identify the following solutions as unsaturated, saturated, or supersaturated. Some solubilities are given below. The solubility chart: Solute: - Solubility in water: Sucrose (Table sugar)-2000g/L at 25 C NaCl (Salt) - 359 g/L at 25 C Atenolol (drug) - 26.5 mg/ml at 37... show more Identify the following solutions as unsaturated, saturated, or supersaturated.

**Chemistry question: unsaturated, saturated, or ...**

A graph can illustrate what solution is saturated and unsaturated. If the point is on the line, then the solution is saturated, while if it is below the line, the solution is unsaturated.

**How does a graph show a saturated solution ... - answers.com**

That intersection of 19 g at 10 C is when the solution is saturated (at that temp). Anything on the 10 C line below 19 is unsaturated. Anything above on the 10 C line is supersaturated.

**Chemistry Q ( Look at graph provided ... - answers.yahoo.com**

In a nutshell, unsaturated solutions have points below the line on the graph, saturated is on the line and supersaturated is above the line on the graph.

**Solubility Curves - Saturated, Unsaturated, Supersaturated Solutions**

To answer this, find where the curve of KCl crosses 30°C. I have marked this point A on the graph. You then read across and determine the grams of solute which is 37 g. Solubility Curve Practice Problems Worksheet. For most substances, solubility increases as temperature increases.

**Solubility Curve Practice Problems Worksheet 1**

Saturated Unsaturated Supersaturated Graph Answers Learn the definition of saturated solution, a term is used in chemistry, plus see examples of saturated solutions.

**Saturated And Unsaturated Solutions Pogil Answer Key**

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**2 POGIL Saturated and Unsaturated Solutions and Solubility ...**

approximately 26 more grams [the solubility of potassium chlorate is `approximately 30g at 65 Celcius] 10. A solution of potassium iodide at 70°C contains 200 g of dissolved solute in 100 cm<sup>3</sup> water. The solution is allowed to cool.

**KEY Name, Date, Hr/Per Solubility Curves**

On the line=saturated (full can not hold anymore solute) Below the line=unsaturated (can hold more solute) Above the line= supersaturated (holding more solute than it should, unstable condition) ... Answer--> The graph show the decrease in solubility as temperature increases: 1/04.

### **Solubility Curves - AP Chemistry**

Use the provided solubility graph to answer the following questions: For questions 1 - 4 an amount of solute is given, and a temperature is stated. If all of the solute could be dissolved in 100 g of water at the given temperature, would the resulting solution be unsaturated, saturated, or supersaturated? 1. 60 g KCl at 70 °C \_\_\_\_ 2. 10 g KClO

### **Worksheet: Solubility Graphs Name**

Classify the following solutions as saturated, unsaturated, or supersaturated, based on your graph. Note that c. and d. require you to convert from the concentrations give to g/100 g H<sub>2</sub>O before you refer to the graph.

### **Classify the following solutions as saturated, unsaturated ...**

I'm taking an online chemistry course, and it doesn't explain this concept very well. I am confused on how to identify whether a solution on the graph is unsaturated, saturated, or supersaturated. I added the picture of the graph that goes along with the questions. Please explain how you got your answer, I do not want just the answers.

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