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Physical Science If8767 Worksheets - Leary Kids

Balance the equation: $\text{KMnO}_4 + \text{HCl} = \text{KCl} \dots$ To have the same amount of each element on each side of the eq... $2\text{KMnO}_4 + 16\text{HCl} = 2\text{KCl} + 2\text{MnCl}_2 + 8\text{H}_2\text{O} + 5\text{Cl}_2$ What is the purpose of balancing chemic... To have the same amount of each element on each side of the eq... What law does balancing chemical equati... The law of conservation...

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CLASSIFYING CHEMICAL REACTIONS Name. Classify the following reactions as synthesis, decomposition, single replacement or double replacement. $2\text{KCl} + 3\text{O}_2 \rightarrow 2\text{KClO}_3$ 2. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ 3. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ 5. $2\text{Al} + 3\text{NiBr}_2 \rightarrow 2\text{AlBr}_3 + 3\text{Ni}$ 6. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ 7. $2\text{NaCl} + 2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$ 9. $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$ -Physical Science IF8767-+3ZnCl₂ + 2K₃P₀₄ 63 @Instructional Fair, Inc.

CLASSIFYING CHEMICAL REACTIONS Name

You have too many H's to start with, the equation is unbalanced. $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$ can not give you 3 H's

out! without having more than one molecule! you could get NaOH +H though which is sodium Hydroxide and a single hydrogen.

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Balancing Chemical Equations Activity Worksheet Answers ...

Balancing Chemical Equations -Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$

Balancing Chemical Equations Answer Key

chemical property is observed, the original substance is changed into a different substance. For example, the ability of iron to rust is a chemical property- The iron has reacted with oxygen and the original iron metal is gone. It is now iron oxide, a new substance. All chemical changes include physical changes.

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