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Section Review 2-1 1. protons; neutrons 2. electrons 3. neutrons 4. electrons 5. ionic 6. The two main types of chemical bonding are ionic and covalent bonding. Ionic bonds are formed when a transfer of electrons takes place from one atom to another, and covalent bonds are formed when electrons are shared between atoms. 7.

Ch. 2 Answer Key - lawndalehs.org

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Chapter 14

The goal of chemistry is to accumulate knowledge. ____13. Biochemistry is the study of the chemistry of living organisms. 10 9 8 7 6 5 4 3 2 1 • chemistry • organic chemistry • inorganic chemistry • analytical chemistry • physical chemistry • biochemistry CHEMISTRY SECTION REVIEW 1.1

1.1 Chemistry Section Review - LPS

Section 19.1 I. H_2SO_4 and H_3O^+ are proton donors and H_2O - 0.35 - 4.1×10^{-2} 0.35 - 0.041 0.309 (4.1×10^{-2}) (4.1×10^{-2}) 0.309 5.4×10^{-3} Section Review 19.4 Part A Completion and HSO_4^- and H_2O are the proton acceptors. The conjugate acid—base pairs are $\text{H}_2\text{SO}_4/\text{HSO}_4^-$ and $\text{H}_2\text{O}/\text{H}_3\text{O}^+$. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ H_3PO_4 is a triprotic acid able to ionize three hydrogens.

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84 Guided Reading and Study Workbook SECTION 9.4 NAMING AND WRITING FORMULAS FOR ACIDS AND BASES (pages 271–273) This section explains the three rules for naming acids and shows how these rules can also be used to write the formulas for acids.

Name Date Class CHEMICAL NAMES AND FORMULAS 9

11.2 Types of Chemical Reactions> 30 A halogen can also replace another halogen from a compound. •The activity of halogens decreases as you go down Group 7A of the periodic table—fluorine, chlorine, bromine, and iodine. •Bromine is more active than iodine, so this reaction occurs: $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

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