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Chapter Objectives The major goals of this chapter are to enable you to: Use Charles' law to determine thermal expansion. Apply Boyle's law to calculate volume changes of gases. Relate gas density to pressure and temperature.

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14.1 Properties of Gases > Factors Affecting Gas Pressure Collisions of particles with the inside walls of the raft result in the pressure that is exerted by the

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understand why gases behave the way they do. Remember first that 2. gas particles are very in relation to the distances between 3. them. Therefore, the volumes of the individual particles are . 4. This assumption helps to explain gas . 5. The second property of gas particles assumed by the kinetic 6.

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gas law represents, the unit in which it is measured and the abbreviation of the unit. 8. When would you use the ideal gas law instead of the combined gas law? Ideal Gases and Real Gases(pages 428-429) 9. Circle the letter of each sentence that is true about ideal gases and real gases. a.

SECTION 14.1 PROPERTIES OF GASES(pages 413-417)

Key Equation. Ideal gas law: P(V = n ... Chapter 14 The Behavior of Gases 351 Name Date ... Answer the following in the space provided. 19. Ideal Gas Law Worksheet <math>PV = nRT - Poudre School District ...

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Prentice Hall Physical Science. Chapter 2 Properties of Matter. 2.2 Physical Properties. A physical property is one that can be observed or measured without changing the composition of the material. A. Examples of Physical Properties. Viscosity – the tendency of a liquid to keep from flowing (resistance to flow)

Prentice Hall Physical Science - Farmington High School

Gases are easily , or squeezed into a smaller volume 1. because of the between particles in a gas. The four variables 2. used to describe a gas are pressure, (P), (V), (T), 3. and number of (n). 4. You can use theory to predict and explain how gases 5. will respond to a change in conditions. Doubling the amount of 6.

THE PROPERTIES OF GASES 14 - flemingchem.weebly.com

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