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Holt Chemistry Gases Problem Solving Instructor: Elizabeth (Nikki) Wyman. Nikki has a master's degree in teaching chemistry and has taught high school chemistry, biology and astronomy. Calculating Boiling-Point Elevation of a Solution - Video ... To understand real gas behavior we use the van der Waals equation. This allows us to account for

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Holt ChemFile: Problem-Solving Workbook 184 Stoichiometry of Gases. Name Class Date. Problem Solving continued. Mass of substance A — solid, liquid, or gas Convert using the mole ratio between A and B. The volume ratio is the same as the mole ratio if the two gases are under the same set of conditions.

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Holt ChemFile: Problem-Solving Workbook 55 Mole Concept. Name Class Date. Problem Solving continued. CONVERTING THE AMOUNT OF AN ELEMENT IN MOLES TO THE NUMBER OF ATOMS In Sample Problem 3, you were asked to determine the number of moles in 1.204 1024 atoms of lithium.

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Holt ChemFile: Problem-Solving Workbook 3 Conversions Name Class Date Problem Solving continued The correct conversion factor is the one that when multiplied by the given quantity causes the units to cancel. COMPUTE 45.5 g calcium chloride 100 1 0 g mg 45 500 mg calcium chloride EVALUATE Are the units correct? Yes; milligrams are the desired units.

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Holt ChemFile: Problem-Solving Workbook 175 The Ideal Gas Law. Name Class Date. Problem Solving continued. APPLICATIONS OF THE IDEAL GAS LAW You have seen that you can use the ideal gas law to calculate the moles of gas, n, in a sample when you know the pressure, volume, and temperature of the sample.

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Holt ChemFile: Problem-Solving Workbook 130 Percentage Yield Name Class Date Problem Solving continued Additional Problems 1. Ethyl acetate is a sweet-smelling solvent used in varnishes and fingernail-polish remover. It is produced industrially by heating acetic acid and ethanol together in the presence of sulfuric acid, which is added to speed ...

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Holt ChemFile: Problem-Solving Workbook 97 Stoichiometry Stoichiometry So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of

moles and that they relate moles of a substance to mass by using the molar mass. In addition, you have learned to

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Holt ChemFile: Problem-Solving Workbook 191 Stoichiometry of Gases Name Class Date Problem Solving continued Rearrange the ideal-gas-law equation to solve for the unknown quantity, V. PV nRT V nR P T COMPUTE EVALUATE Are the units correct? Yes; units canceled to give liters of SO 2. Is the number of significant figures correct?

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Ideal Gas Law Answers Holt Chemfile - pdfsdocuments2.com

CHEMFILE MINI-GUIDE TO PROBLEM SOLVING CHAPTER 13 Stoichiometry of Gases Now that you have worked with relationships among moles, mass, and volumes of gases, you can easily put these to work in stoichiometry calculations. Many reactions have gaseous reactants, gaseous products, or both. Reactants and products that are not gases are usually ...

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