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170 Guided Reading and Study Workbook 4. Look at Figure 16.8 on page 481. Circle the letter of the best procedure for making a 0.50-molar (0.50M) solution in a 1.0-L volumetric flask. a. Add distilled water exactly to the 1.0-L mark, add 0.50 mol of solute, and

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### Chapter 14

11.2 Types of Chemical Reactions> 30 A halogen can also replace another halogen from a compound. •The activity of halogens decreases as you go down Group 7A of the periodic table—fluorine, chlorine, bromine, and iodine. •Bromine is more active than iodine, so this reaction occurs:  $\text{Br}_2(\text{aq}) + 2\text{NaI}(\text{aq}) \rightarrow 2\text{NaBr}(\text{aq}) + \text{I}_2(\text{aq})$

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138 Core Teaching Resources SECTION 6.2 CLASSIFYING THE ELEMENTS 1. Use the periodic table to write the electron configuration for silicon. Explain your thinking. 2. Use the periodic table to write the electron configuration for iodine.

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