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Holt ChemFile: Problem-Solving Workbook 97 Stoichiometry Stoichiometry So far in your chemistry course, you have learned that chemists count quantities of elements and compounds in terms of moles and that they relate moles of a substance to mass by using the molar mass. In addition, you have learned to

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Holt ChemFile: Problem-Solving Workbook 58 Mole Concept Name Class Date Problem Solving continued CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4 uses the progression of steps 1 → 2 → 3 to convert from the mass of an element to the number of atoms.

Holt ChemFile Problem Solving Workbook 58 Mole Concept ...

Holt ChemFile: Problem-Solving Workbook 191 Stoichiometry of Gases Name Class Date Problem Solving continued Rearrange the ideal-gas-law equation to solve for the unknown quantity, V. PV nRT V nR P T COMPUTE EVALUATE Are the units correct? Yes; units canceled to give liters of SO₂. Is the number of significant figures correct?

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Problem Solving continued Sample Problem 1 Ammonia is made industrially by reacting nitrogen

and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The ... Holt ChemFile: Problem-Solving Workbook 102 Stoichiometry Name Class Date Problem Solving continued

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Holt ChemFile: Problem-Solving Workbook 318 Electrochemistry Name Class Date Problem Solving continued Sample Problem 2 A cell is constructed in which the following two half-reactions can occur in either direction. $\text{Zn}^{2+} + 2\text{e}^{-} \rightleftharpoons \text{Zn}$ $\text{Br}_2 + 2\text{e}^{-} \rightleftharpoons 2\text{Br}^{-}$ Write the full ionic equation for the cell in the spontaneous direction,

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