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a. stronger acid. c. weaker acid. b. stronger base. d. weaker base	_ 6. Proton-transfer reactions
favor production of the a. stronger acid and stronger base. b. weaker a	icid and weaker base. c.
stronger acid and weaker base. d. weaker acid and stronger base	_ 7. Aqueous solutions of
most bases contain a. hydroxide ions and cations. c	

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4 is considered acid 1, identify the other three terms as acid 2, base 1, and base 2 to indicate the conjugate acid-base pairs. base 2 CO 3 2 acid 2 HCO 3 base 1 NH 3 True b. A proton has been transferred from acid 1 to base 2 in the above reaction. True or False? 3. Consider the neutralization reaction described by the equation: HCO 3(aq) OH ...

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What is the correct acid name for an aqueous solution of HCIO 4? a. hypochlorous acid b. chlorous acid c. chloric acid d. perchloric acid _____ 9. A substance that increases the concentration of OH ions in an aqueous solution is known as a(n) a. Arrhenius acid. b. Arrhenius base. c. Lewis acid. d.

Lewis base. 10. An Arrhenius acid in an

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Compare the Brønsted-Lowry definitions of acids and bases with the Arrhenius definitions of acids and bases. Identify conjugate acid-base pairs. Write chemical equations that show how an amphoteric species can behave as either an acid or a base. Acids Vinegar is acidic. So are the juices of the fruits that you see in Figure 1.

CHAPTER

Holt Chemistry 60 Acids and Bases Section: Equilibria of Weak Acids and Bases In the space provided, write the letter of the term or phrase that best answers the question. _____ 1. In an acid-base reaction, the conjugate base of the acid is the _____ base. a. stronger, stronger, b. weaker, stronger c. weaker, weaker d. none of the above

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