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$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ 2. $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$ 3. $\text{Al}(\text{NO}_3)_3 + 3\text{NaOH} \rightarrow \text{Al}(\text{OH})_3 + 3\text{NaNO}_3$
chemical equation Reagents products coefficients balance Conservation equal created destroyed
coefficients formulas / subscripts

Worksheet: Balancing Equations Name KEY I. Fill in the ...

Problems: Show all work and circle your final answer. 1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required? $4.00 \text{ M} = \frac{\text{moles of solute}}{12.0 \text{ L}}$
 $\text{moles of solute} = 48.0 \text{ mol}$ 2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M? $? \text{ L} = 250 \text{ mL} \times 1 \text{ L}$

Molarity: Molarity = 1. 2. - Central Bucks School District

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