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1. d It might be easiest to balance the equation with mostly whole numbers: $2 \text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow 2 \text{NO} + 3 \text{H}_2\text{O}$. The question asks about the amount of oxygen reacting with ONE mole of ammonia, thus cut the $\frac{7}{2}$ (3.5) of oxygen in half to 1.75

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formula units KClO_3 decompose to form two formula units KCl and three molecules O_2 . b Chapter 12 stoichiometry test b answer key. Four molecules NH_3 react with six molecules NO to form five mol-Chapter 12 Stoichiometry Test B Answer Key - fullexams.com a Chapter 12 stoichiometry test b answer key. 2 moles of oxygen and 1 mole of hydrogen in the

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Chapter 11 Small-Scale Lab Section 11.3 Precipitation Reactions: Formation of Solids, page 345 Analysis 1. a. $\text{Na}_2\text{CO}_3 + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + 3\text{Ag}_2\text{CO}_3(\text{s})$ b. $2\text{Na}_3\text{PO}_4 + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 6\text{NaNO}_3 + \text{Pb}_3(\text{PO}_4)_2(\text{s})$

Chapter 11 Small-Scale Lab

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Stoichiometry CHAPTER 12 What You'll Learn You will write mole ratios from balanced chemical equations. You will calculate the number of moles and the mass of a reactant or product when given the number of moles or the mass of another reactant or product. You will identify the limiting reactant in a chemical reaction. You will determine the

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