

## *Periodic Trends Section Review Answers*

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6.3 Periodic Trends. Prentice Hall Chemistry, 2005. STUDY. PLAY. Atomic radii generally \_\_\_\_\_ as you move from left to right in a period. (energy levels, increases, charge, decreases) decreases.

### 6.3 Periodic Trends Flashcards | Quizlet

Answers for Comparing Atomic Sizes Here are answers for the questions above. a. Li, C, F All are in the same period and thus have the same number of energy levels. Therefore, the important factor is the nuclear charge. Li is the largest because it has the smallest nuclear charge and pulls the ... Microsoft Word - Periodic Trends Worksheet Author:

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- Interpret group trends in atomic radii, ionic radii, ionization energies, and ... Answer the following in the space provided. 20. For the following pairs of atoms, tell which one of each pair has the largest ionic radius. a. ... 14.2 Periodic Trends Section Review

### 14.2 Periodic Trends Section Review - LPS

Concept Review: Trends in the Periodic Table 1. ionization energy 2. bond radius 3. electron affinity 4. electronegativity 5. increases, decreasing 6. d 7. a 8. c 9. The electron cloud model is based on the probability of finding an electron at a specific location. As you move far-ther out from the nucleus, the proba-bility of finding electrons ...

### Skills Worksheet Concept Review

Review of Periodic Trends You will need to periodic table to complete this activity. Choose the correct answer for each question. Show all questions <= => Generally speaking, the group of elements with the highest first ionization energy is: ? Group 1 ? Group 16 ? ...

### Review of Periodic Trends - sciencegeek.net

174 Chapter 6 • The Periodic Table and Periodic Law Section 66.1.1 Development of the Modern Periodic Table MAIN Idea The periodic table evolved over time as scientists discovered more useful ways to compare and organize the elements.

### Chapter 6: The Periodic Table and Periodic Law

Periodic(Trends(Worksheet(((Answer Key 1. Circle)the)element)withthe)largest)atomic(radius)andpu t)a)square)aroundthe)element)withthe)smallest) atomic(radius:))) Cu) K) Ni) Br))) Largest - K Smallest - Br) Explainwhy)you)madethesechoices.) Atomic radius decreases as you go left to right across a period. Potassium is in the far

### Periodic Trends Worksheet Answers

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### **6.3 Periodic Trends Worksheet Answers - FREE Printable ...**

Section 6.3 Periodic Trends 3 sessions 11/ 2 blocks P LS 1. Trace the development and identify key features of the periodic table. 2. Explain why elements in the same group have similar properties. 3. Identify the four blocks of the periodic table based on electron configuration. 4. Compare period and group trends of several properties. 5.

### **Chapter 6: The Periodic Table and Periodic Law**

Periodic Trends > Trends in Ionization Energy The energy required to remove an electron from an atom is called ionization energy. • The energy required to remove the first electron from an atom is called the first ionization energy. • The energy required to remove an electron from an ion with a 1+ charge is called the second ionization energy. 6.3

### **6.3 Periodic Trends - Mr. Walk**

SECTION 6.3 PERIODIC TRENDS . 1. Explain why a magnesium atom is smaller than atoms of both sodium and calcium. 2. Predict the size of the astatine (At) atom compared to that of tellurium (Te). Explain your prediction. 3. Would you expect a  $\text{Cl}^-$  ion to be larger or smaller than an  $\text{Mg}^{2+}$  ion? Explain. 4.

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