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An oxidation-reduction (redox) reaction is a type of chemical reaction that involves a transfer of electrons between two species. An oxidation-reduction reaction is any chemical reaction in which the oxidation number of a molecule, atom, or ion changes by gaining or losing an electron.

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Oxidation Reduction Reaction Prentice Hall Answers

DOWNLOAD PRENTICE HALL OXIDATION AND REDUCTION ANSWER KEY prentice hall oxidation and pdf Etymology "Redox" is a portmanteau of the words "reduction" and "oxidation". The word oxidation originally implied reaction with oxygen to form an oxide, since dioxygen ($O_2(g)$) was historically the first recognized

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Chapter 20 Oxidation–Reduction Reactions221 SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION (pages 631–638) This section explains oxidation and reduction in terms of the loss or gain of electrons, and describes the characteristics of a redox reaction. It also explains how to identify oxidizing and reducing agents.

SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION (pages ...

The element undergoing oxidation is the one which assumes a more positive oxidation state when going from reactants and products. The element undergoing reduction is the one which assumes a more negative oxidation state when going from reactants and products.

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"Redox" is a portmanteau of the words "reduction" and "oxidation". The word oxidation originally implied reaction with oxygen to form an oxide, since dioxygen ($O_2(g)$) was historically the first recognized oxidizing agent. Later, the term was expanded to encompass oxygen-like substances that accomplished parallel chemical reactions.

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