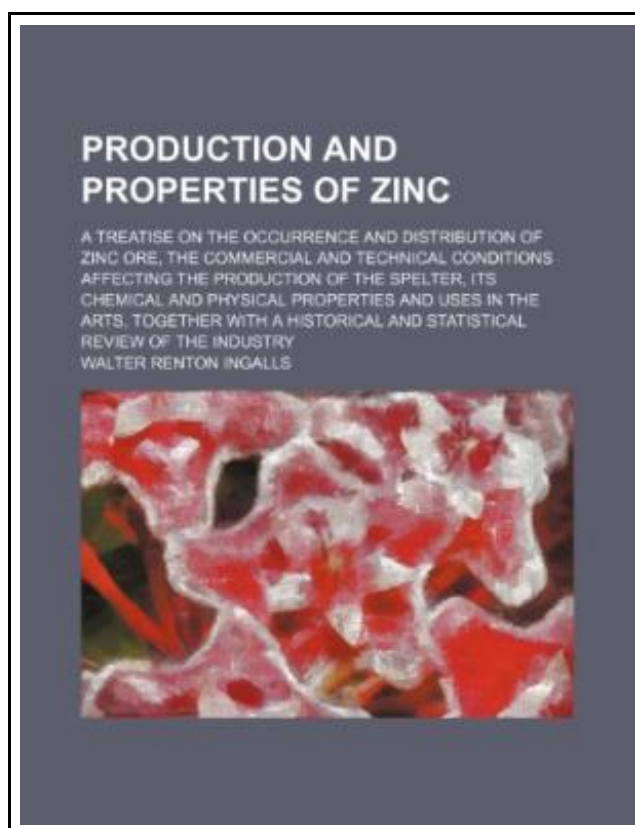


# Production and Properties of Zinc; A Treatise on the Occurrence and Distribution of Zinc Ore, the Commercial and Technical Conditions Affecting the Production of the Spelter, Its Chemical and Physical Properties and Uses



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## ***Reviews***

*The publication is easy in read through preferable to fully grasp. It is writter in simple phrases instead of hard to understand. You will not sense monotony at at any moment of your respective time (that's what catalogs are for concerning if you request me).*

*(Kevin Bergstrom Sr.)*

## **PRODUCTION AND PROPERTIES OF ZINC; A TREATISE ON THE OCCURRENCE AND DISTRIBUTION OF ZINC ORE, THE COMMERCIAL AND TECHNICAL CONDITIONS AFFECTING THE PRODUCTION OF THE SPELTER, ITS CHEMICAL AND PHYSICAL PROPERTIES AND USES**



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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book \*\*\*\*\* Print on Demand \*\*\*\*\*.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1902 Excerpt: . disengagement of carbonic acid and the precipitate formed is a basic hydrocarbonate. This hydrocarbonate is soluble in an aqueous solution of carbonic acid; upon exposure to the air a granular powder deposits from the solution, which has been taken for neutral carbonate of zinc by some chemists, but by others has been held to be a basic carbonate. According to Thorpe<sup>2</sup> neutral zinc carbonate,  $\text{ZnCO}_3$ , can be produced artificially by adding acid carbonate of sodium to a solution of zinc sulphate, while Roscoe and Sehormer state that it is produced by precipitation from  $\text{ZnSO}_4$  with an excess of hydrogen potassium car 1 In a paper on Sulphide Ore Treatment, Metallurgy, London, June 19, 1901. read before the Institution of Mining and 2 Dictionary of Applied Chemistry, p. 1057. bonate ( $\text{KHCO}_3$ ), the neutral carbonate of sodium ( $\text{Na}_2\text{CO}_3$ ) throwing down a basic carbonate of zinc, which is the more basic the higher the temperature of precipitation and the more dilute the solution. Zinc carbonate,  $\text{ZnCO}_3$ , is produced when the solution of a zinc salt is precipitated by an alkaline carbonate under pressure. As obtained by heating a solution of  $\text{ZnCl}_2$  to 150-160 C. and precipitating with calcium carbonate or sodium bicarbonate it is a white microcrystalline powder. The neutral hydrated carbonate ( $2\text{ZnCO}_3, \text{H}_2\text{O}$ ) is an amorphous powder obtained by digesting a basic carbonate of zinc with ammonium bicarbonate. The anhydrous carbonate of zinc is soluble in carbonic acid water. A solution of  $\text{CO}_2$  at five atmospheres, taking...



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