Organic Chemistry Concepts LOKT.09.051

 π -bond reactivity

π – bonds

$$c = c$$
 $-c = c$

- C C
- C = C
- C ≡ C
- C O
- C = O
- C − N
- C = N
- C = N

- 83 85 kcal/mol
- 146 151 kcal/mol
- 199 200 kcal/mol
 - 85 91 kcal/mol
- 173 181
 - 69 75 kcal/mol
 - 143 kcal/mol
 - 204 kcal/mol

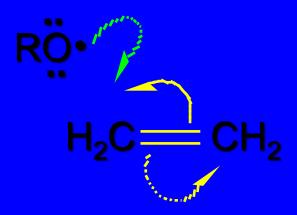
Radical reaction mechanism

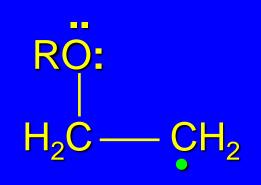
Radical stability determines the reaction center

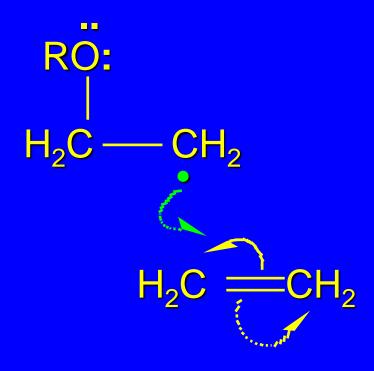
Ethene polymerization

$$H_2C = CH_2$$

Polyethene or polyethylene







RÖ:
$$H_2C - CH_2$$

$$H_2C - CH_2$$

RO:
$$H_2C - CH_2$$

$$H_2C - CH_2$$

$$H_2C - CH_2$$

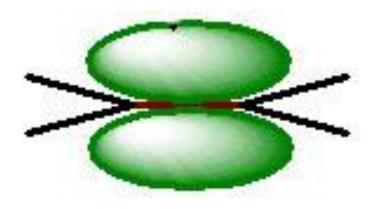
RÖ:
$$H_{2}C - CH_{2}$$

$$H_{2}C - CH_{2}$$

$$H_{2}C - CH_{2}$$

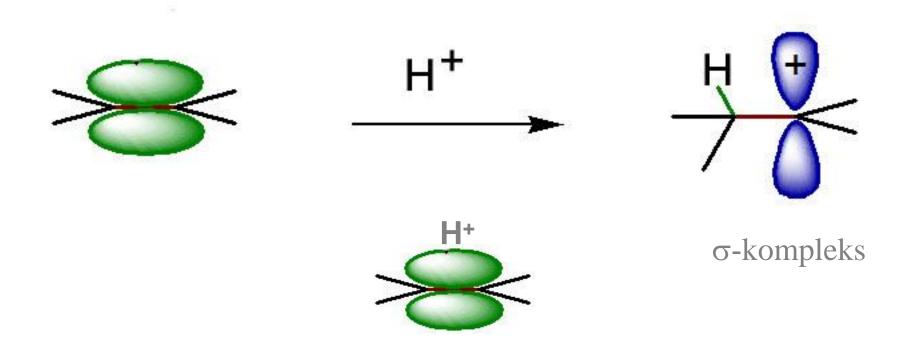
$$H_{2}C - CH_{2}$$

Ionic reaction mechanism

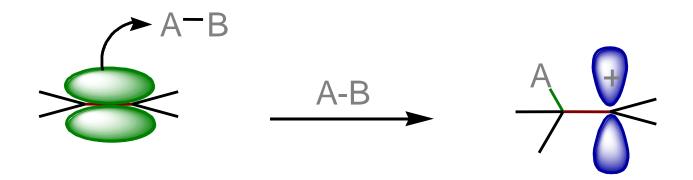


$$C = C$$

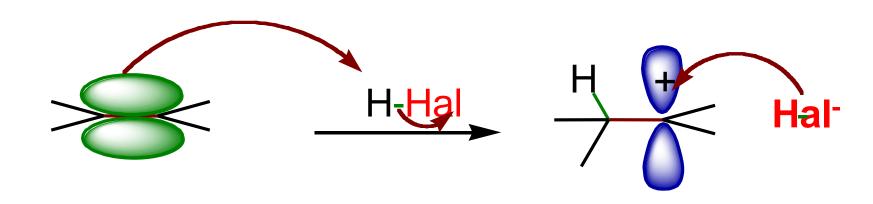
π - bond as base



π - bond as nucleophile



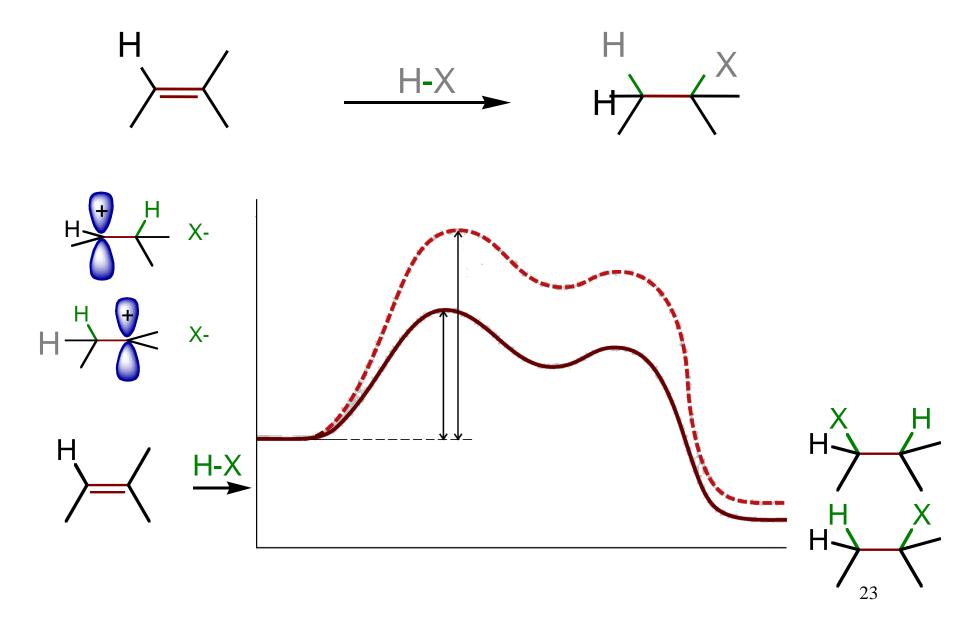
H-Hal addition



Markovnikov Rule

Markovnikov Rule

Main product Side product



Conjugated double bonds

Triple bonds

$$RC \stackrel{\longleftarrow}{=} CH \xrightarrow{\stackrel{\longleftarrow}{H} \stackrel{\frown}{=} Br} \begin{bmatrix} R - \overset{\longleftarrow}{C} = C \\ H \end{bmatrix} \xrightarrow{\stackrel{:Br:}{:Br:}} R = C = C \\ H$$

Water addition in acidic medium

$$CH_3 \longrightarrow CH_2 + \mathbf{HOH} \xrightarrow{\mathbf{H}_3\mathbf{O}^+} CH_3 \longrightarrow CH_2 - \mathbf{H}$$

$$CH_3 \longrightarrow CH_2 \longrightarrow CH_2 - \mathbf{H}$$

$$H_{3}C$$
 $C = C$
 $H_{2}SO_{4}$
 CH_{3}
 CH_{3}
 $CH_{2}CH_{3}$
 CH_{3}
 $CH_{2}CH_{3}$
 CH_{3}
 CH_{3}
 $CH_{2}CH_{3}$
 CH_{3}
 CH_{3}
 $CH_{2}CH_{3}$
 CH_{3}
 CH_{3}
 $CH_{2}CH_{3}$
 CH_{3}
 CH_{3}
 CH_{3}
 CH_{4}
 $CH_{5}CH_{5}$
 $CH_{5}CH_{5}$
 $CH_{5}CH_{5}$
 $CH_{5}CH_{5}$
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 $CH_{5}CH_{5}$
 $CH_{5}CH_{5}CH_{5}$
 $CH_{5}CH_{5}CH_{5}CH_{5}$
 $CH_{5}CH_{5}CH_{5}CH_{5}$
 $CH_{5}CH_{5}CH_{5}CH_{5}$
 $CH_{5}CH_{5}CH_{5}$

Polymerization

Halogen addition

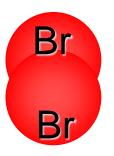
$$C = C + X_2 \longrightarrow X - C - C - X$$

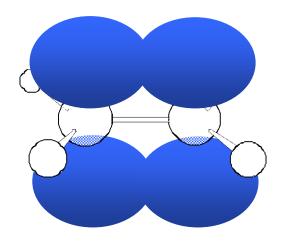
$$CH_{3}CH = CHCH(CH_{3})_{2} \xrightarrow{Br_{2}} CH_{3}CHCHCH(CH_{3})_{2}$$

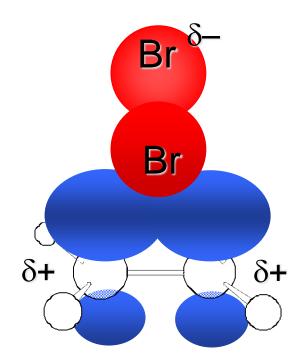
$$CHCI_{3} \qquad | \qquad |$$

$$0^{\circ} C \qquad Br Br$$

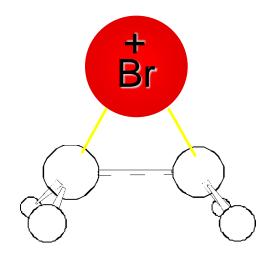
$$(100\%)$$



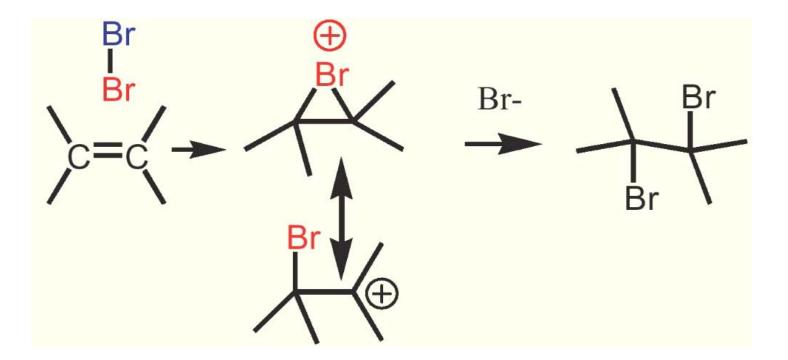








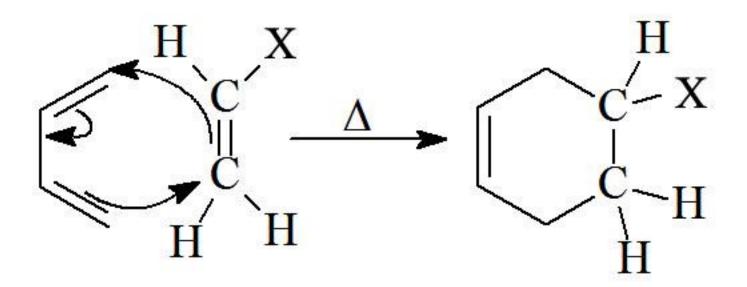
Bromonium ion



$$(CH_3)_2C=CH_2$$
 $\xrightarrow{Br_2}$ $(CH_3)_2CCH_2Br$ OH

Diels-Alderi reaktsioon

- Otto Diels, Kurt Alder; Nobel prize, 1950
- Dieen + alkeen (alküün)



Otto Paul Hermann Diels 1876-1954



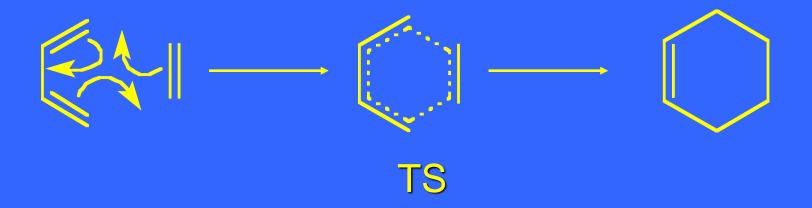
Kurt Alder 1902-1958





Dien Alkene, Dienophile

Cyclohexene



Cyclic reaction

