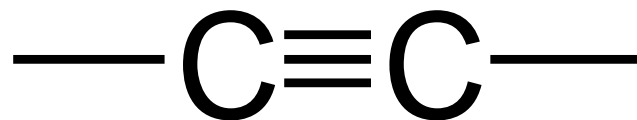
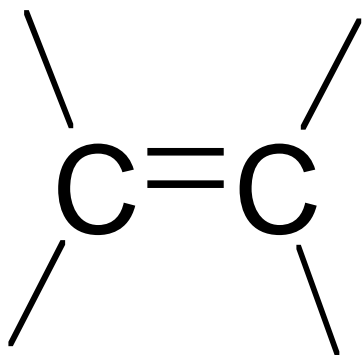


Organic Chemistry Concepts

LOKT.09.051

π -bond reactivity

π – bonds

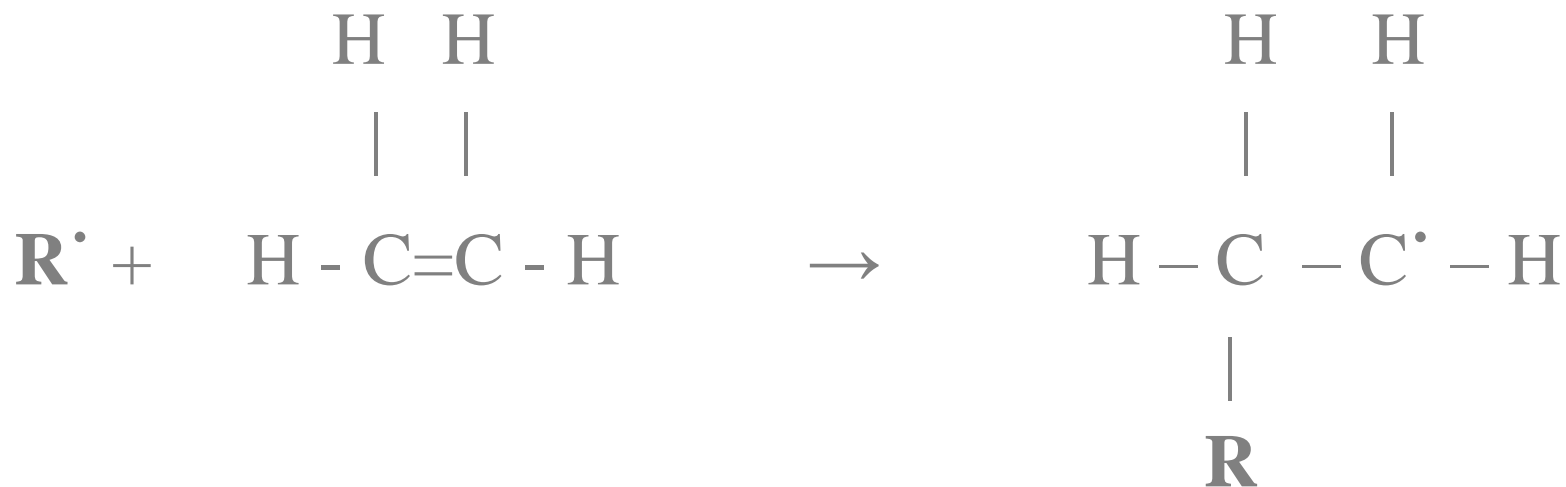


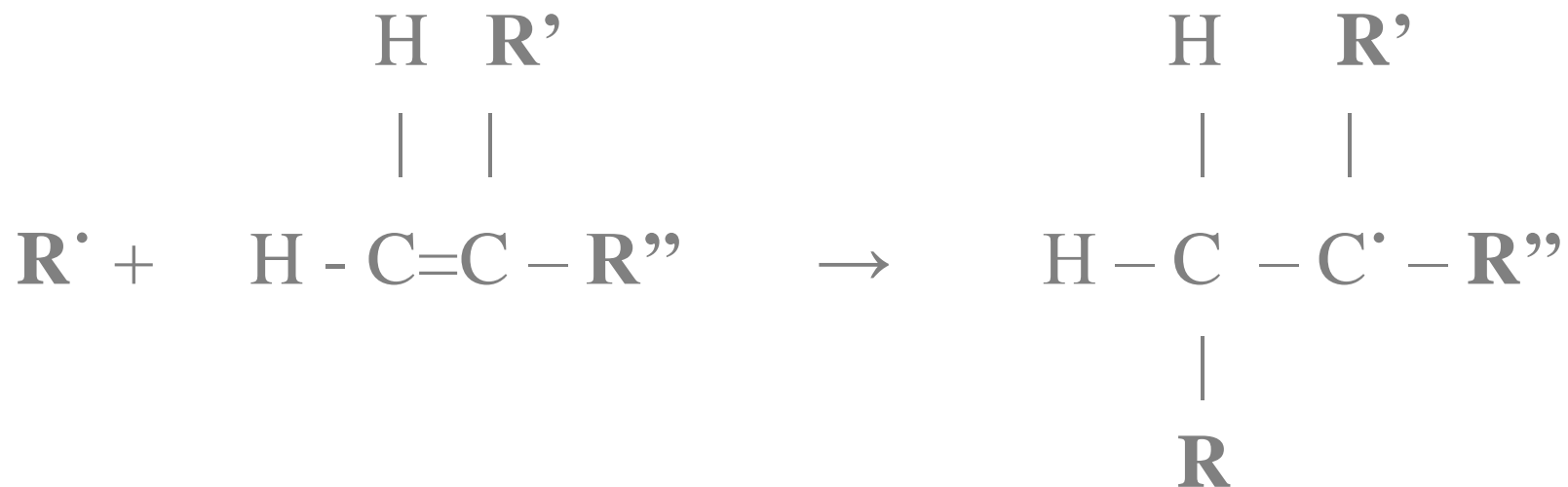
- C – C 83 - 85 kcal/mol
- C = C 146 – 151 kcal/mol
- C \equiv C 199 – 200 kcal/mol

- C – O 85 – 91 kcal/mol
- C = O 173 – 181

- C – N 69 – 75 kcal/mol
- C = N 143 kcal/mol
- C \equiv N 204 kcal/mol

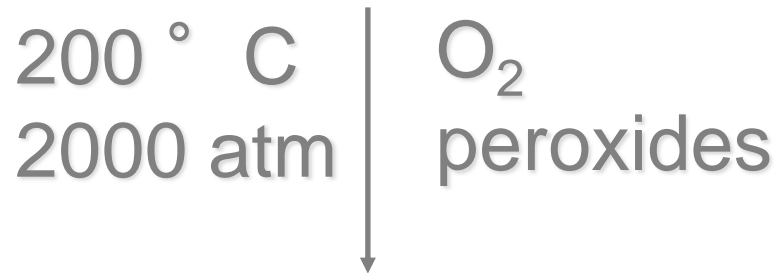
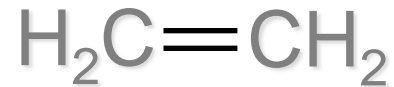
Radical reaction mechanism



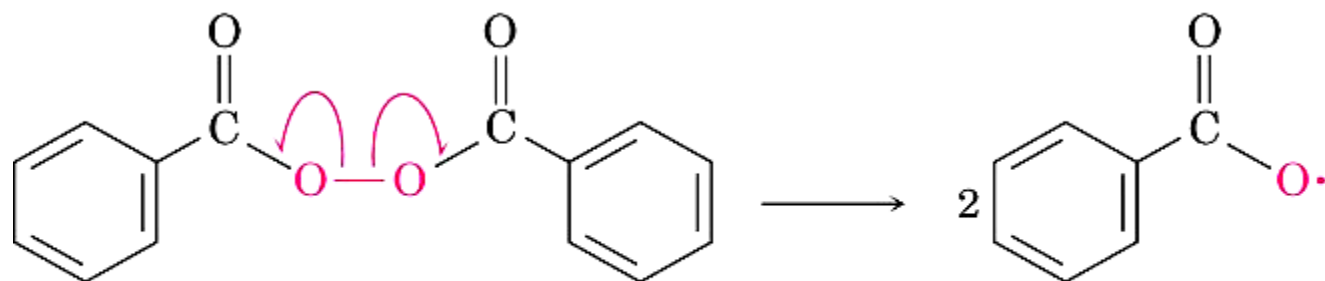
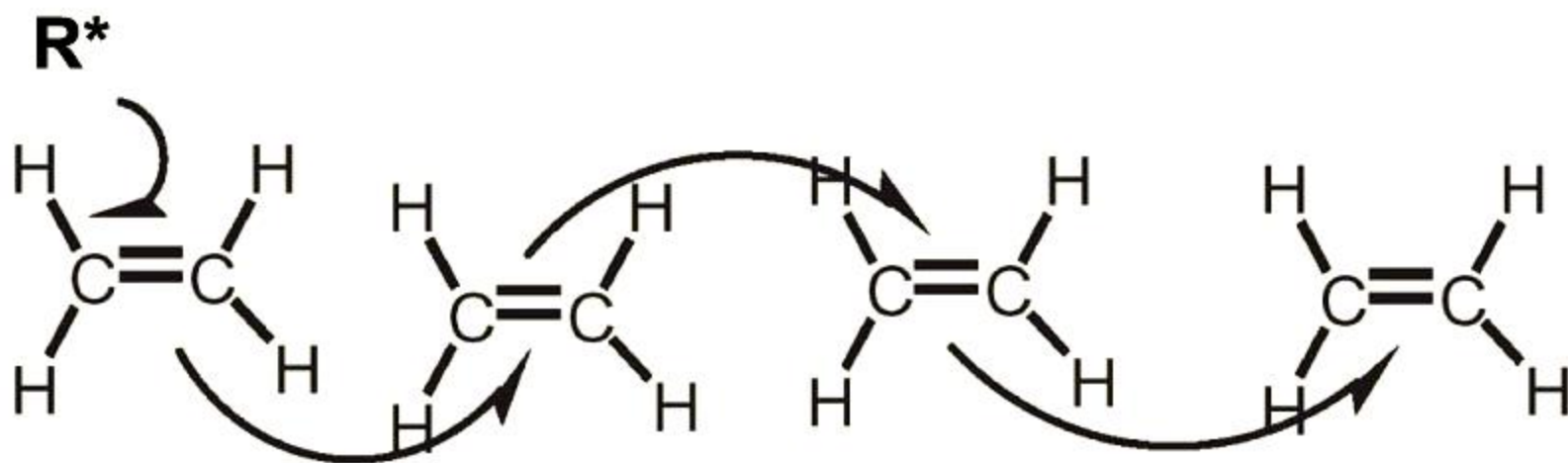


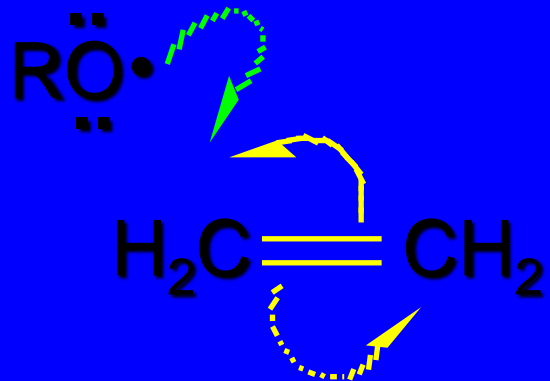
Radical stability determines the reaction center

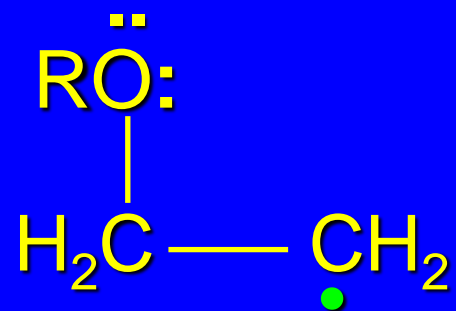
Ethene polymerization

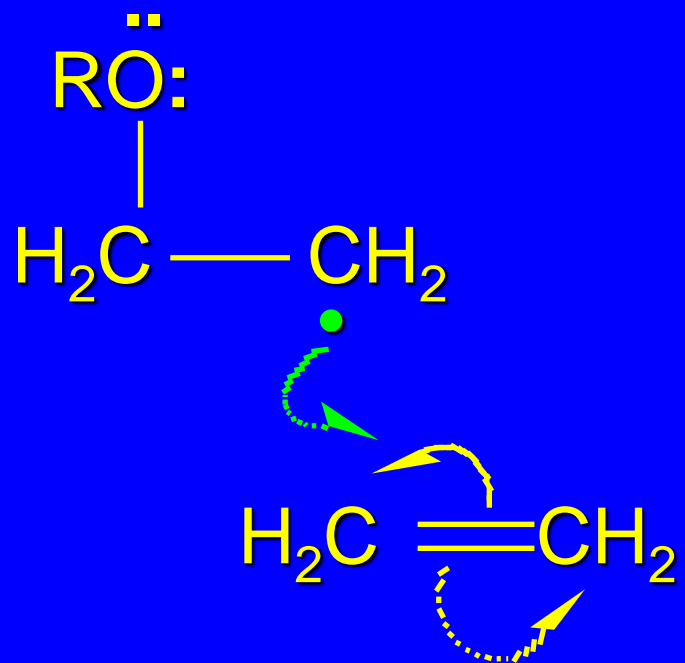


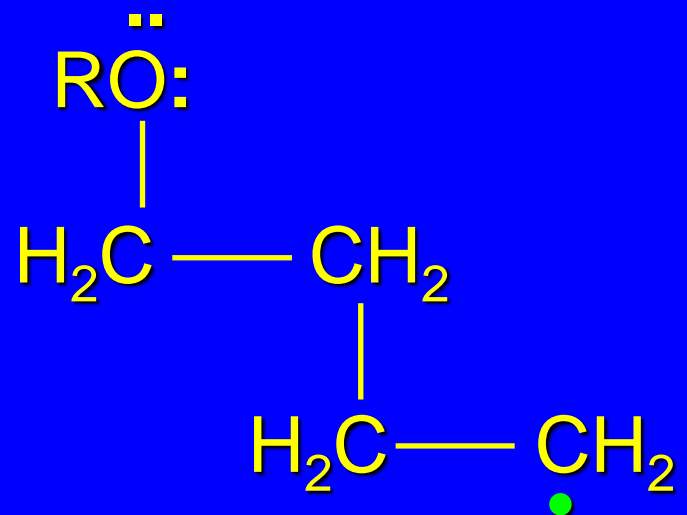
Polyethene or polyethylene

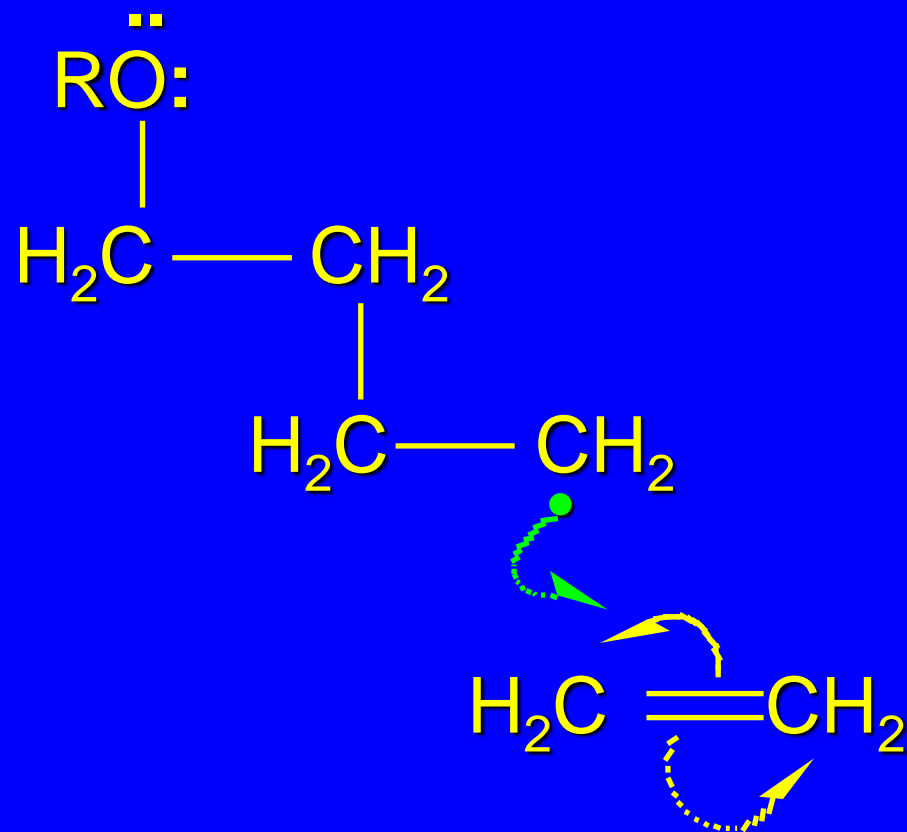


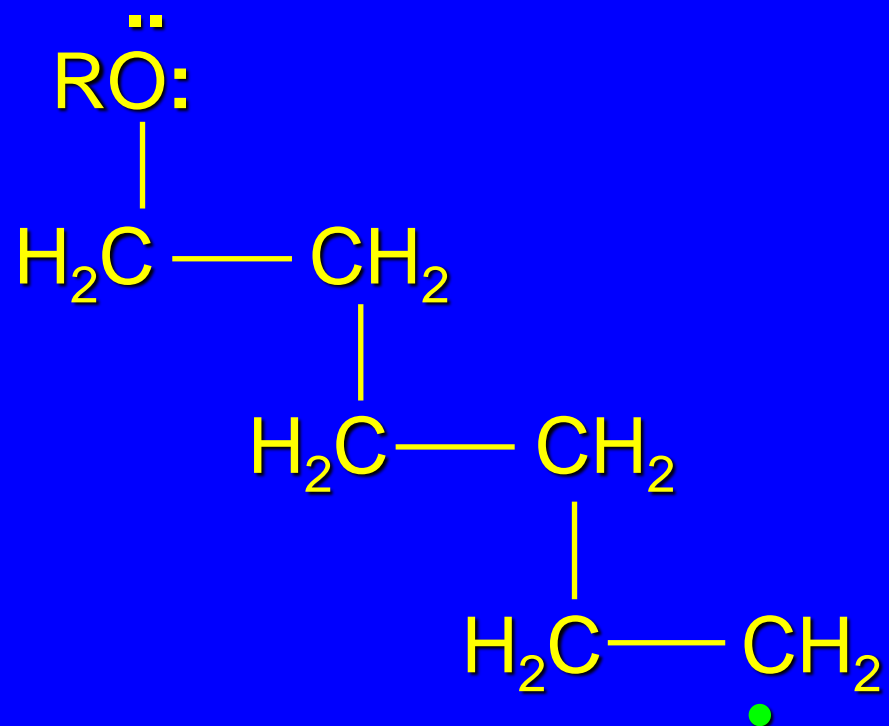


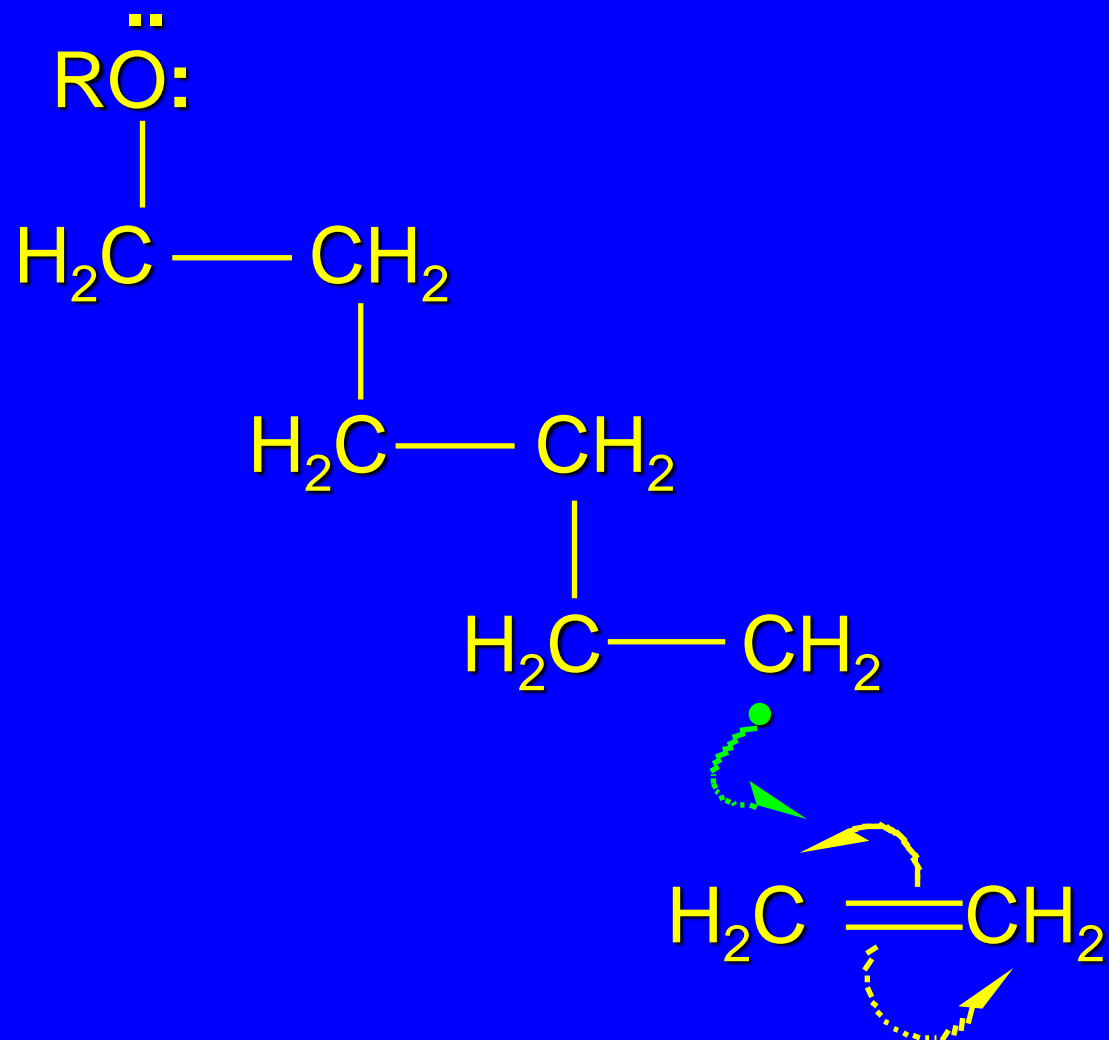




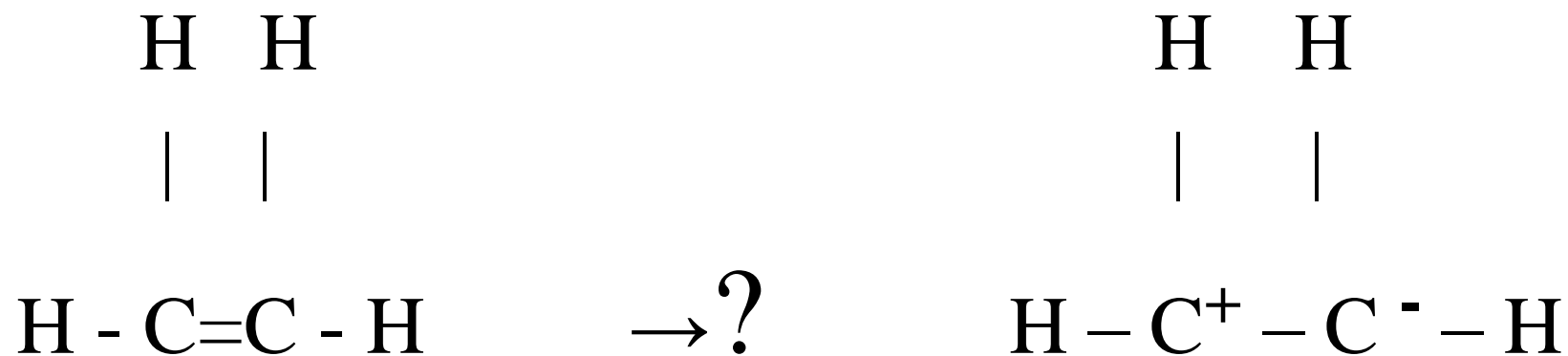


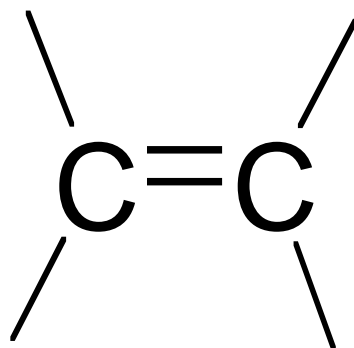
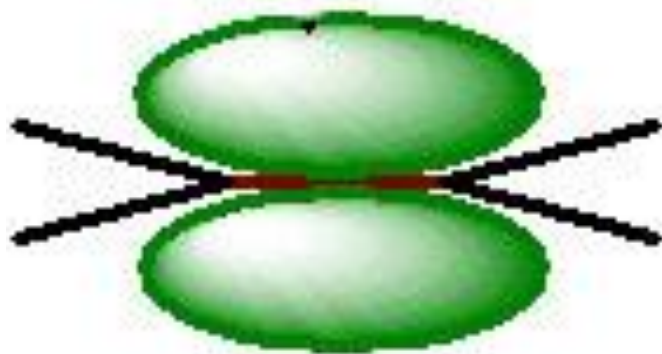




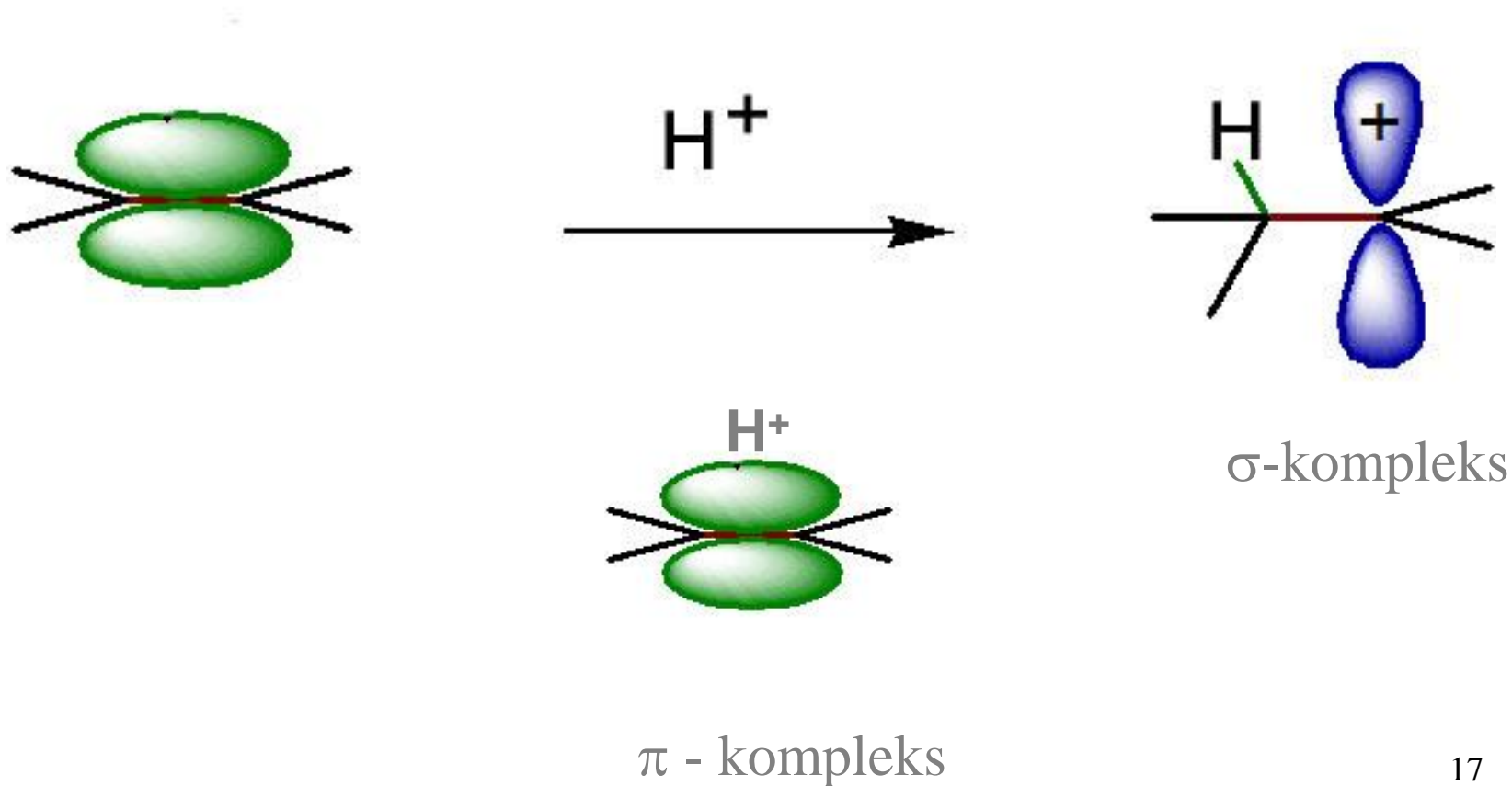


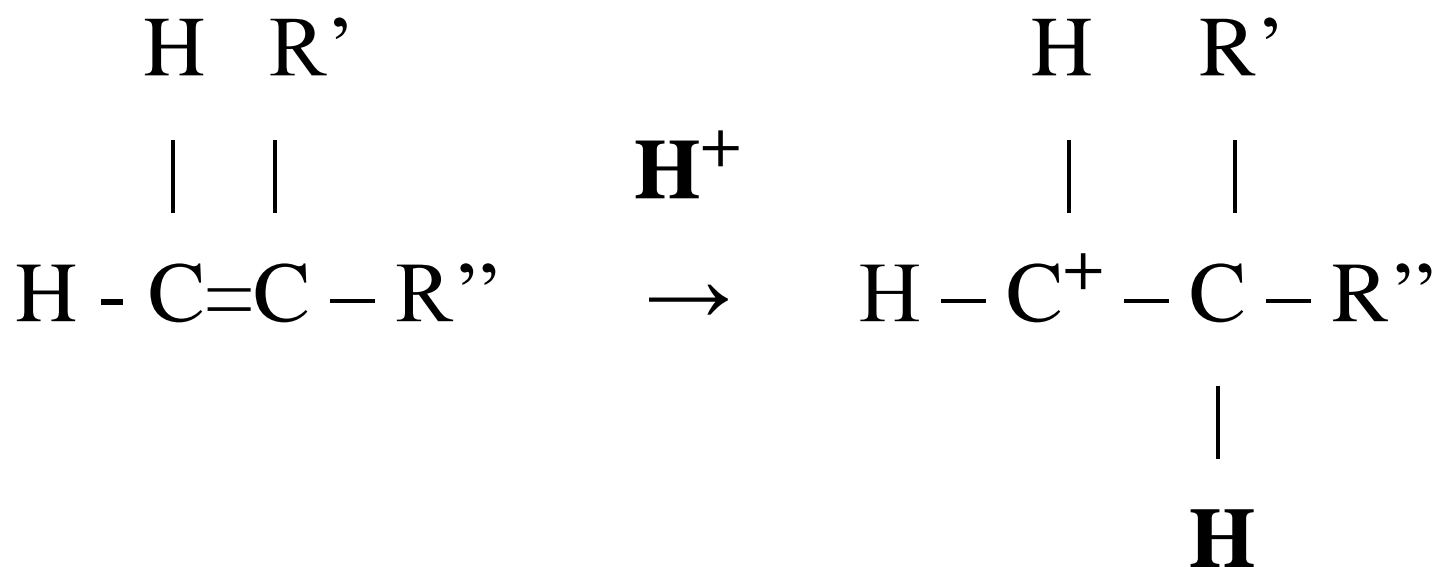
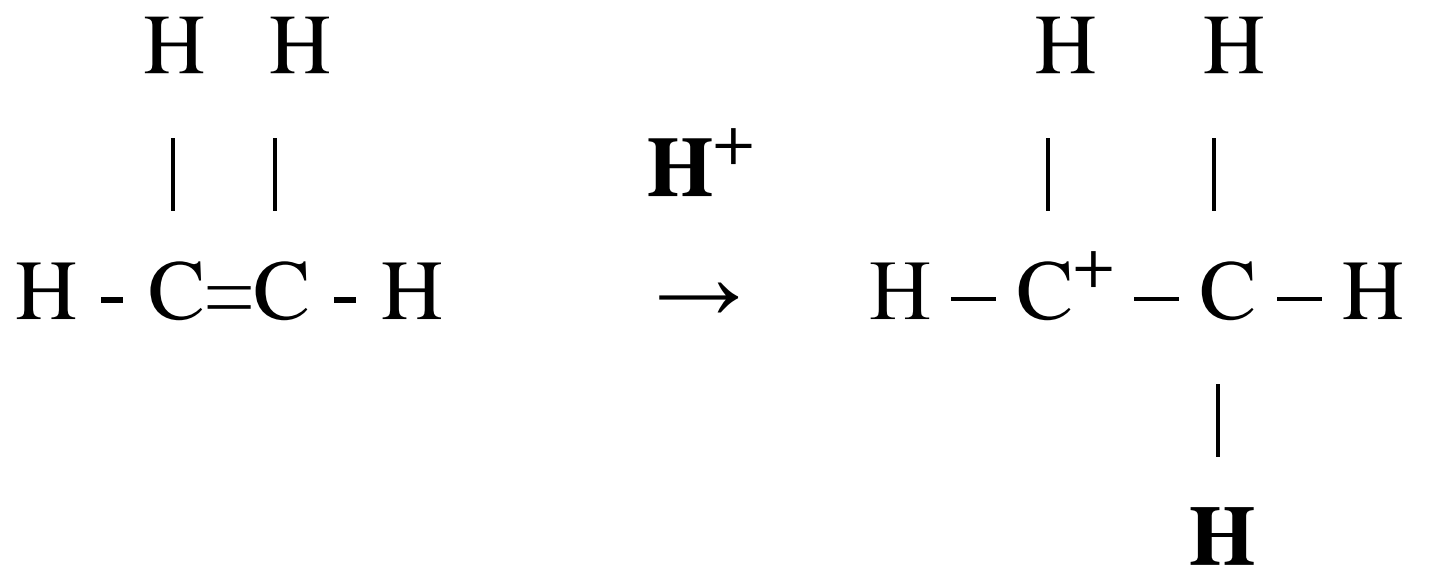
Ionic reaction mechanism



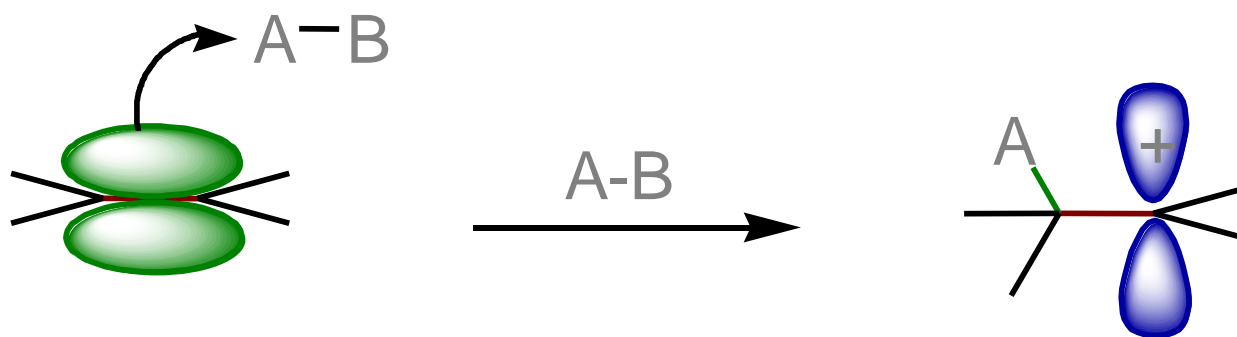


π - bond as base

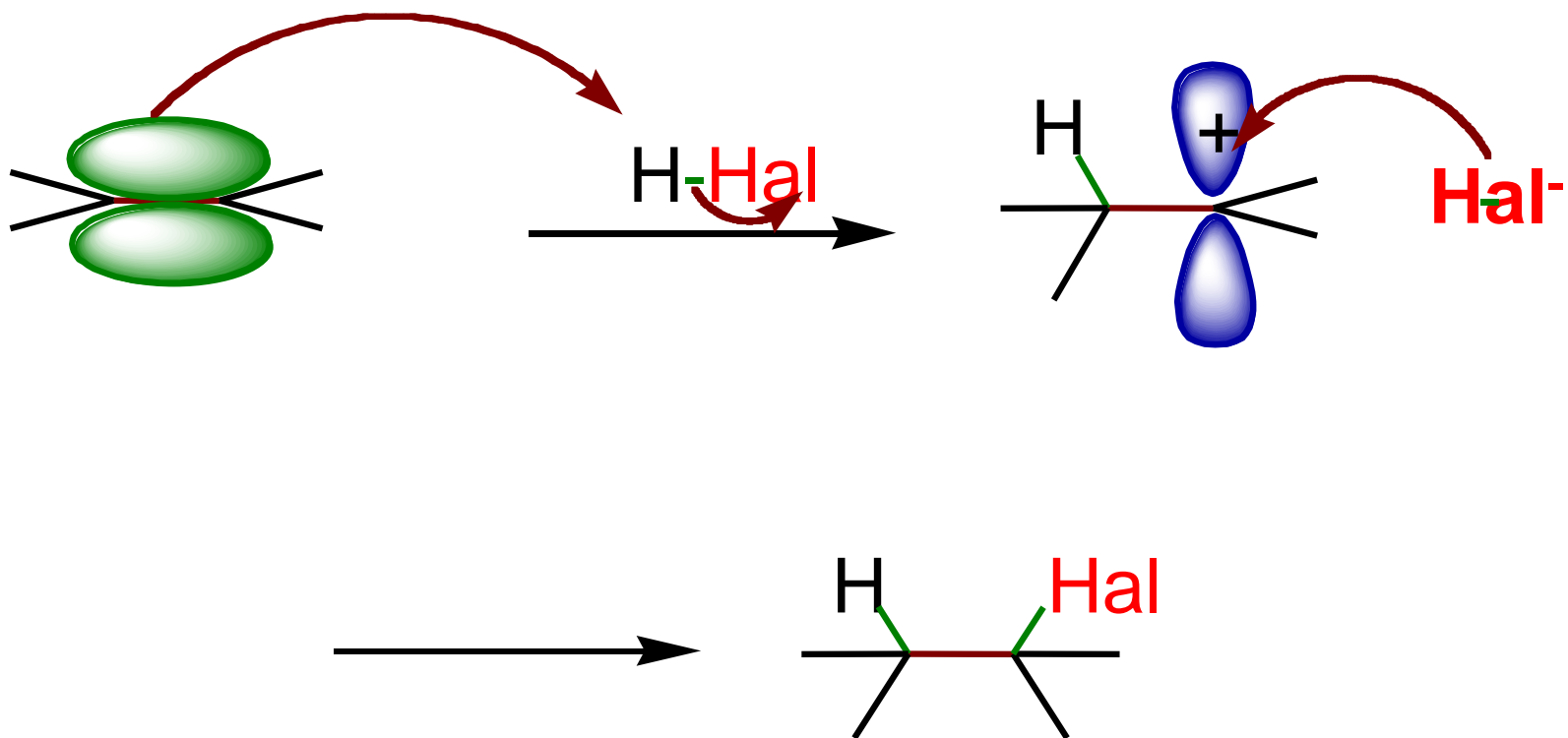




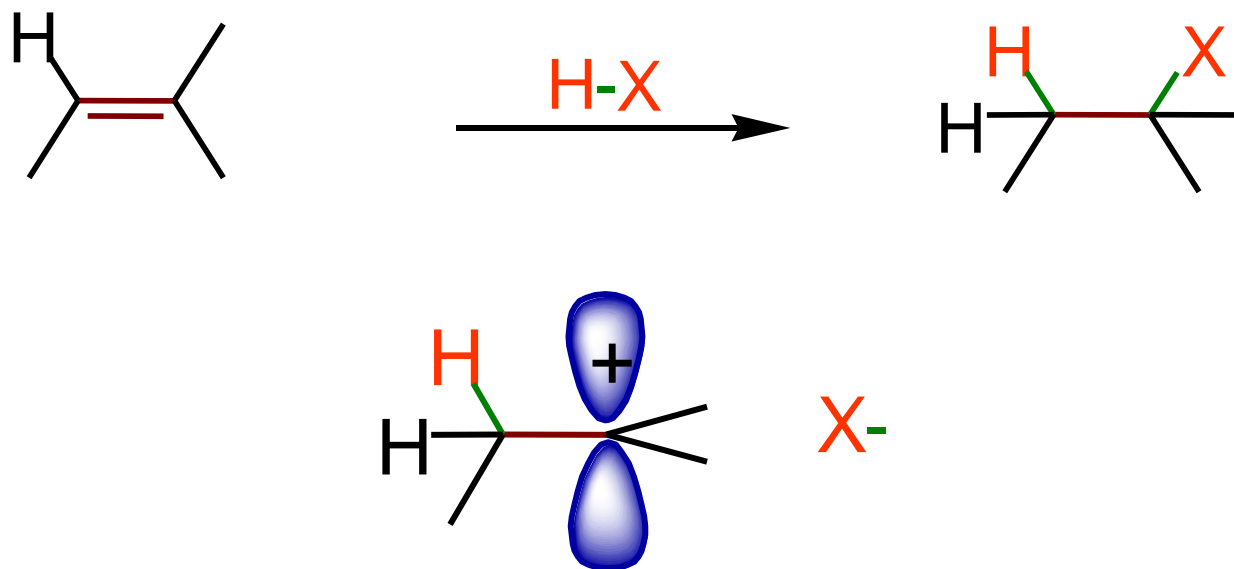
π - bond as nucleophile



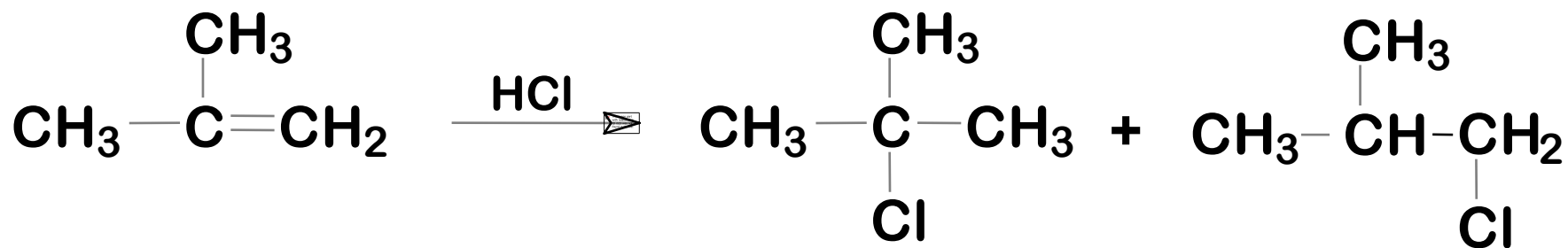
H-Hal addition



Markovnikov Rule

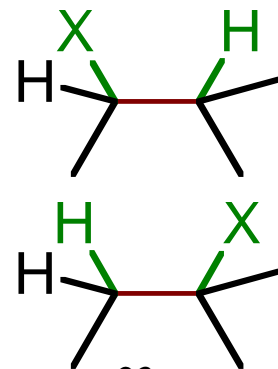
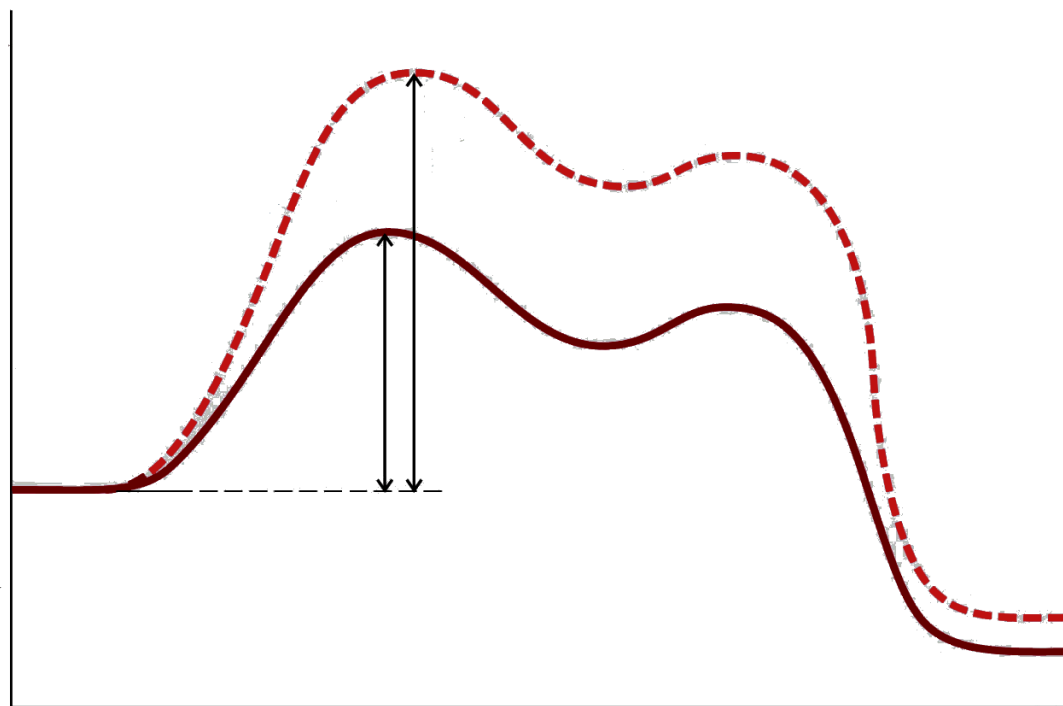
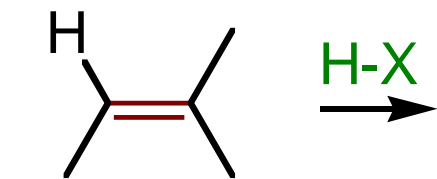
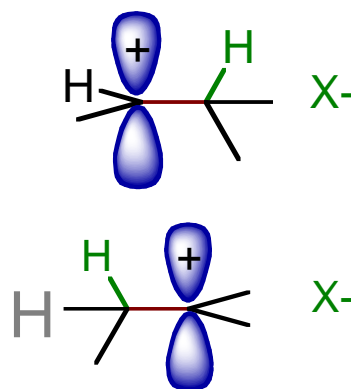
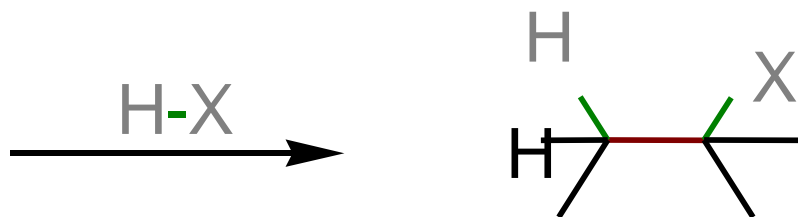
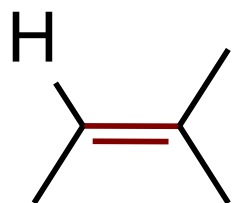


Markovnikov Rule

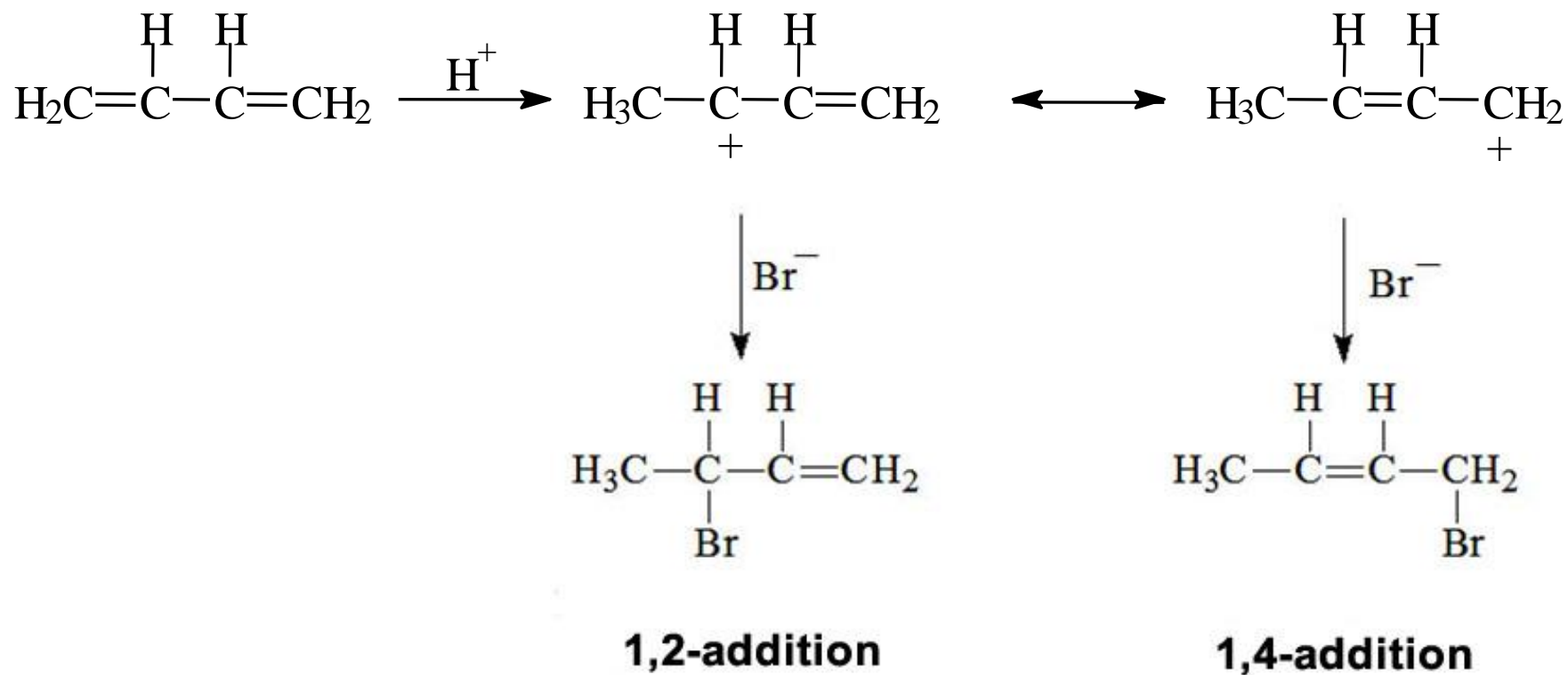


Main product

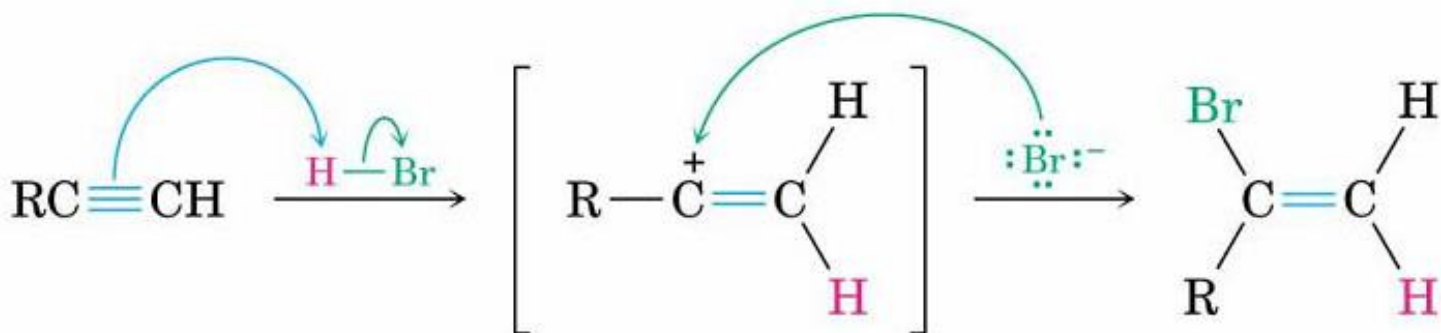
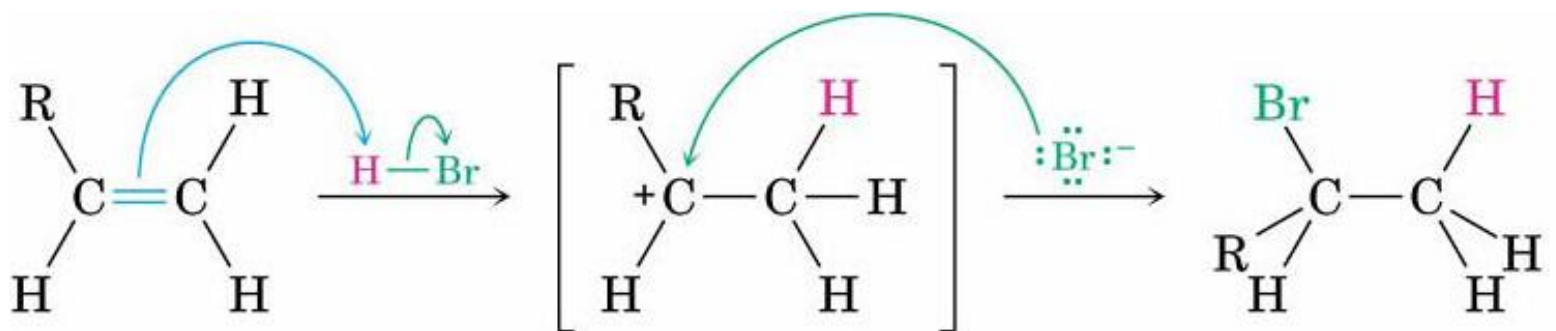
Side product



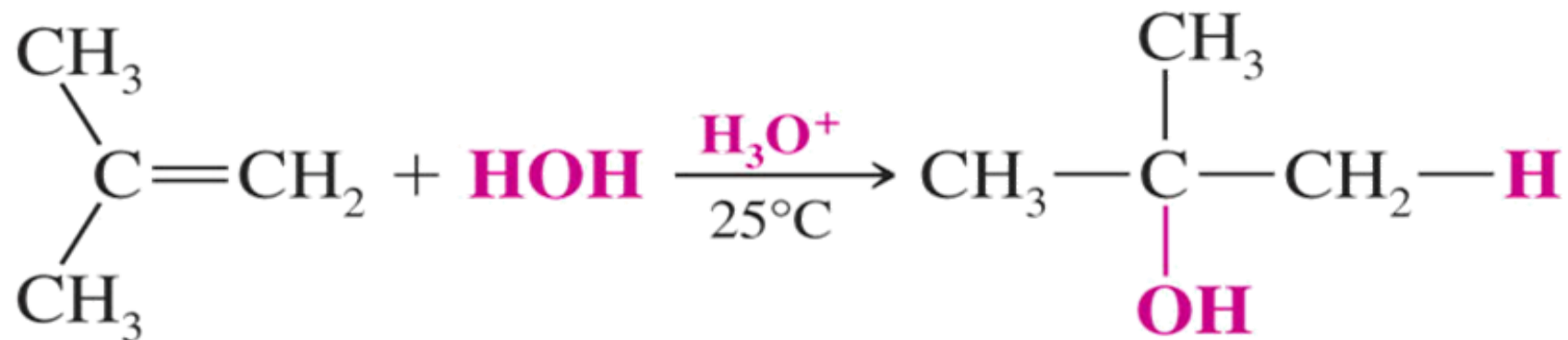
Conjugated double bonds

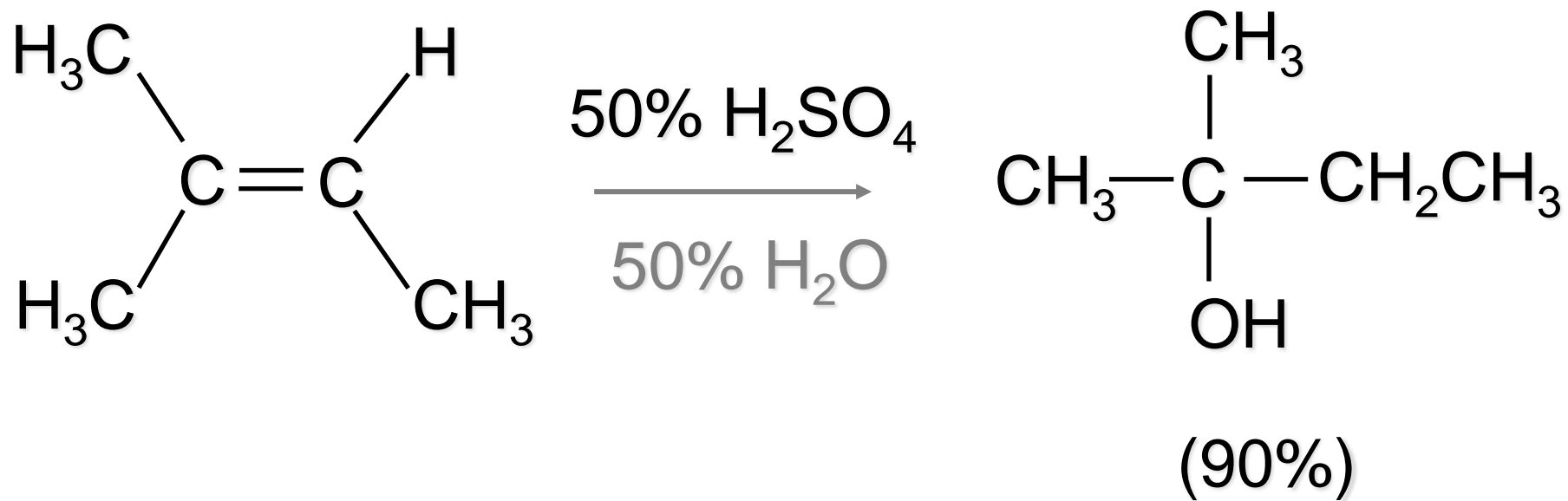


Triple bonds

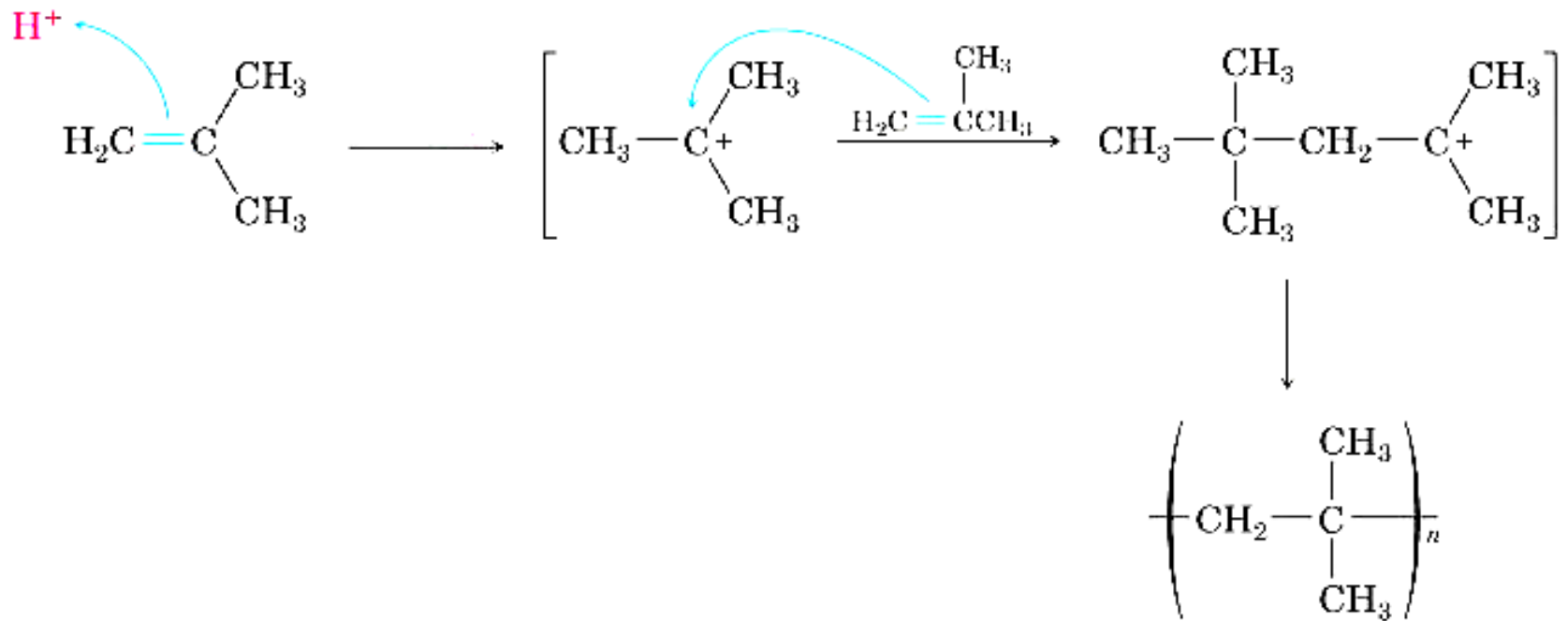


Water addition in acidic medium

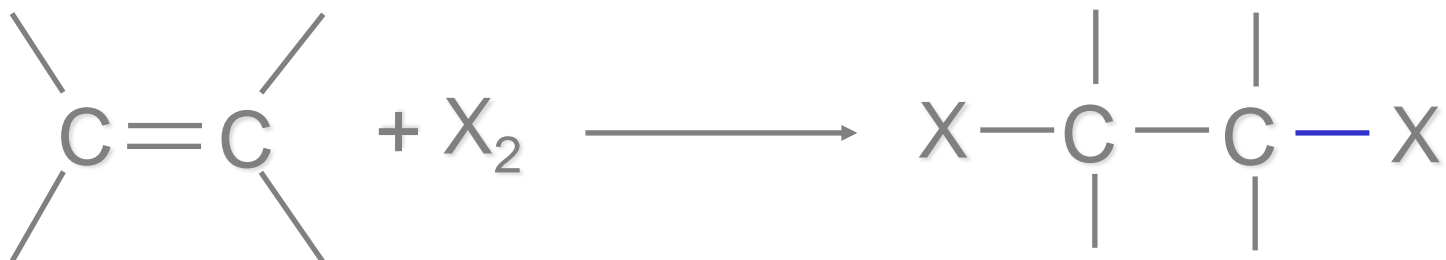


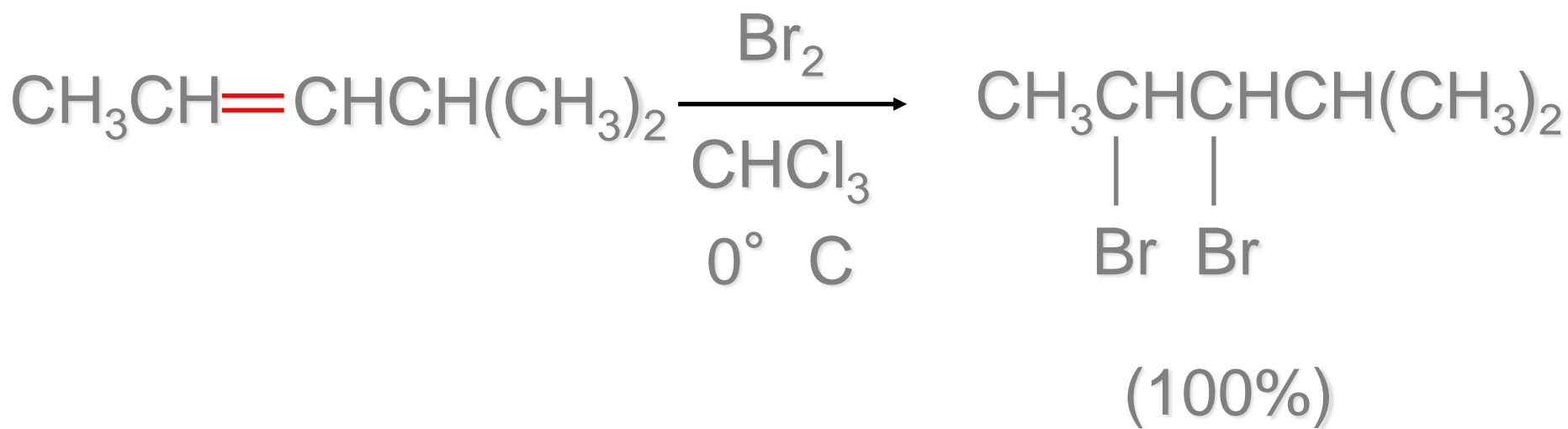


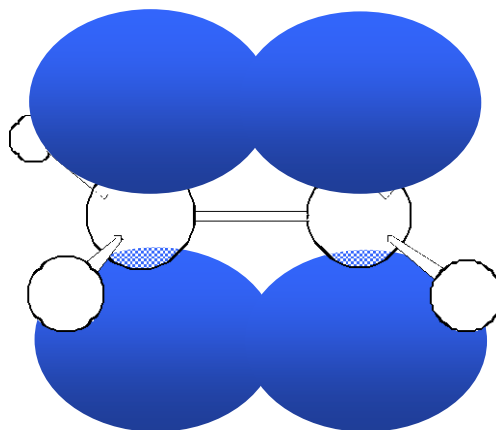
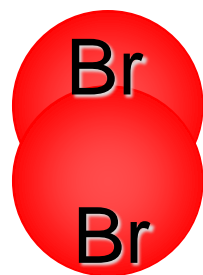
Polymerization

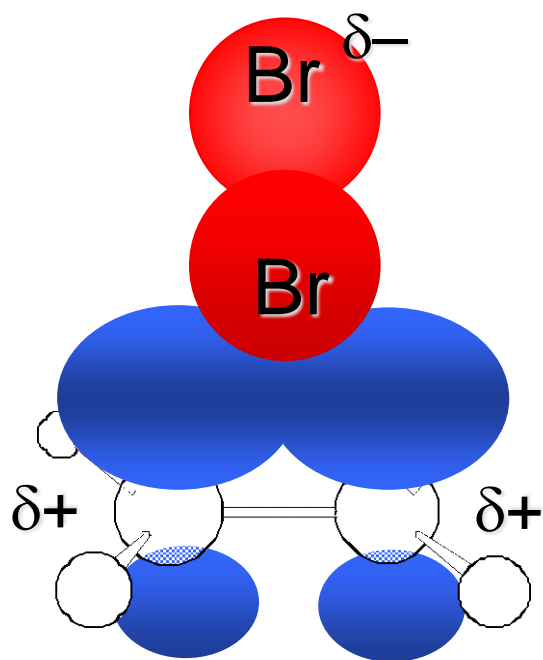


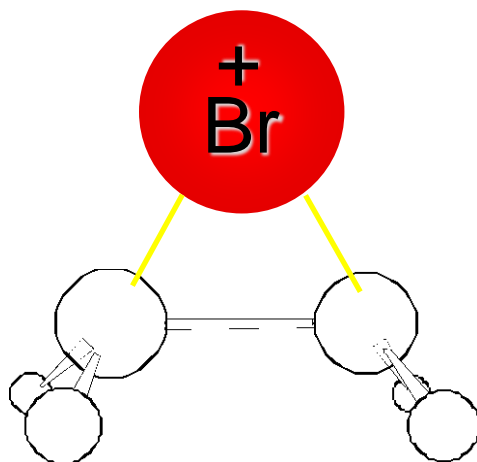
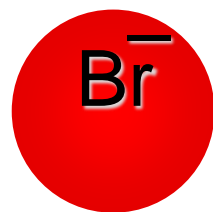
Halogen addition



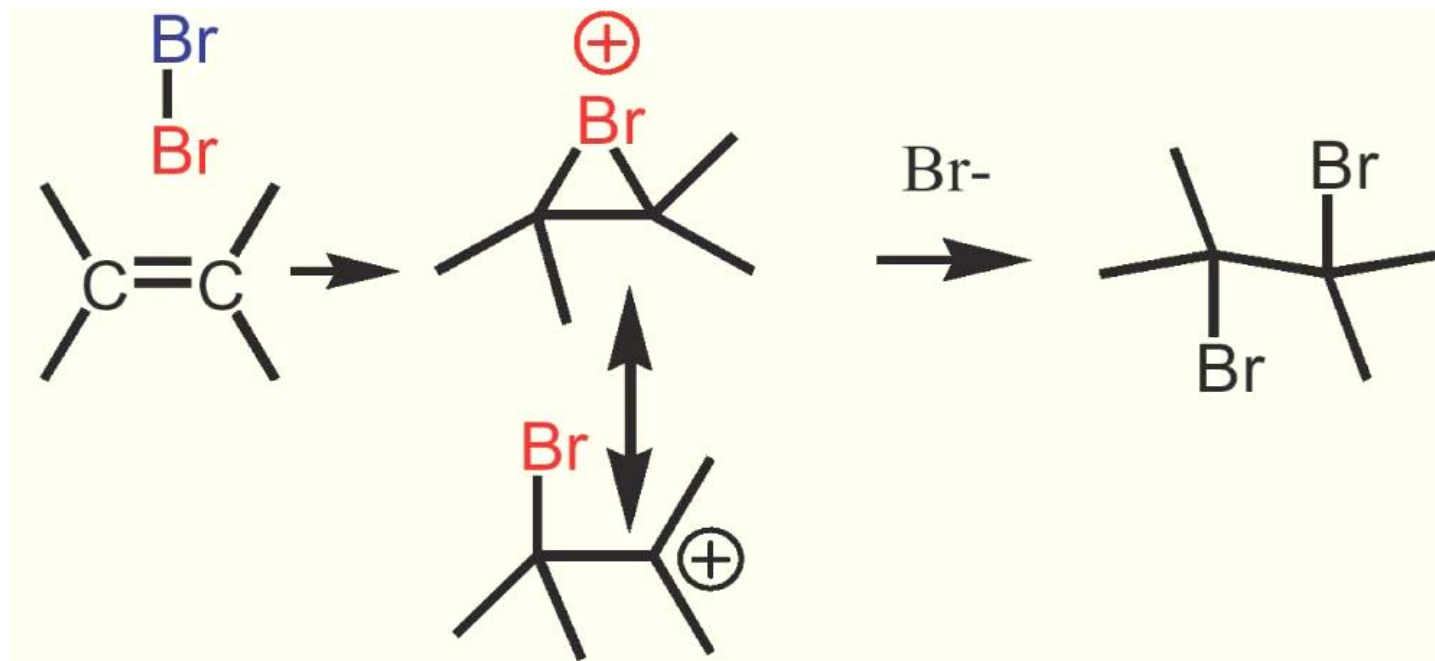


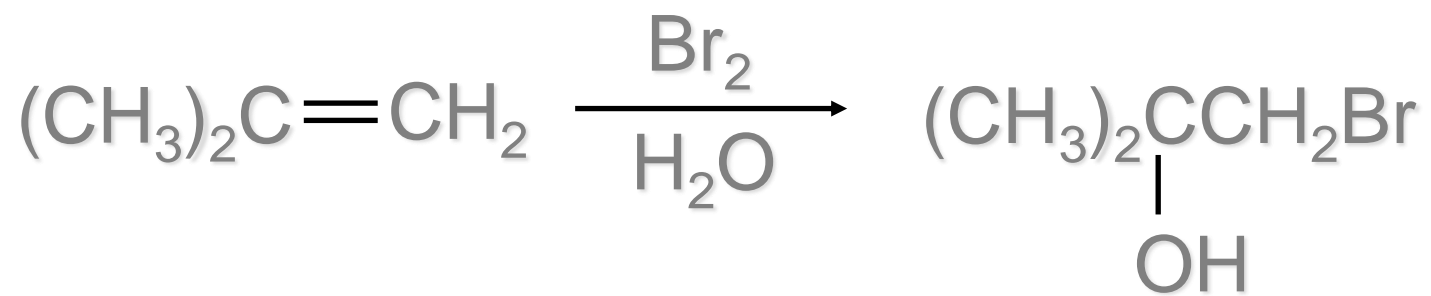






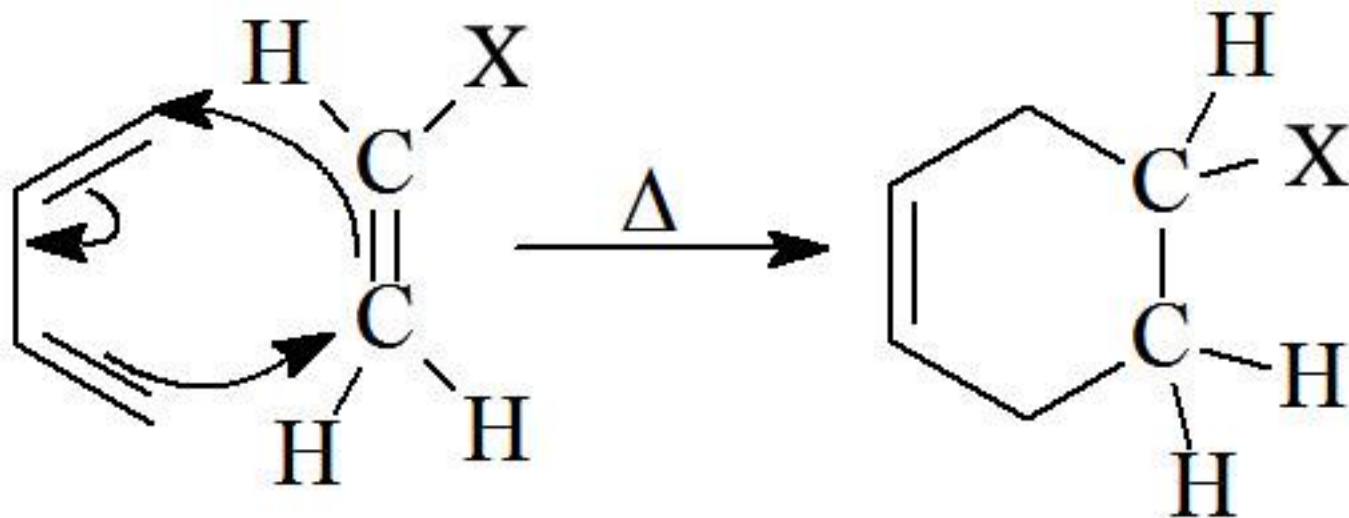
Bromonium ion





Diels-Alderi reaktio

- Otto Diels, Kurt Alder; Nobel prize, 1950
- Dieen + alkeen (alküün)



Otto Paul Hermann Diels

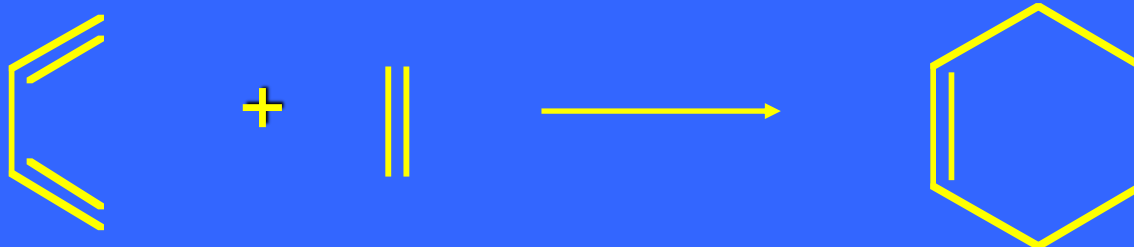
1876-1954



Kurt Alder

1902-1958

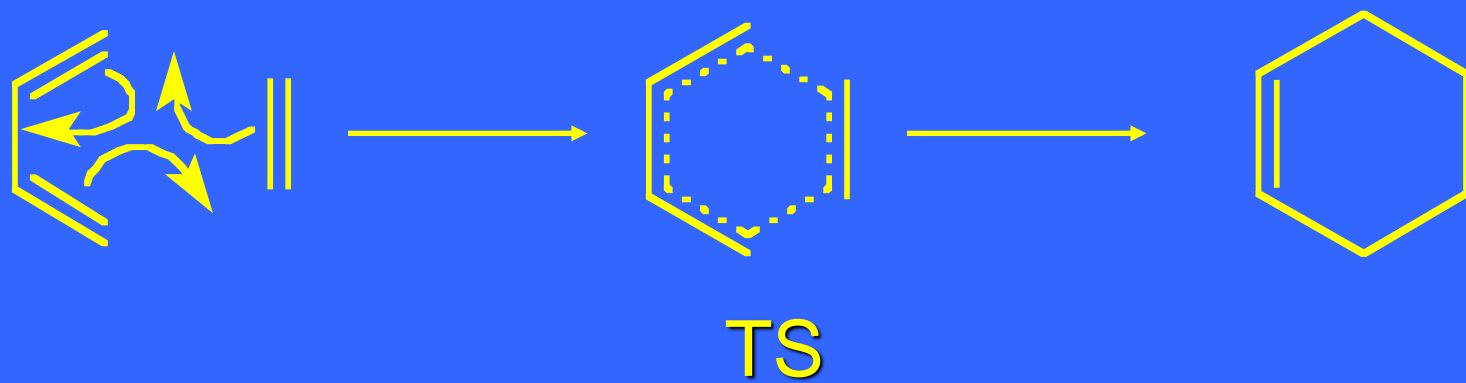




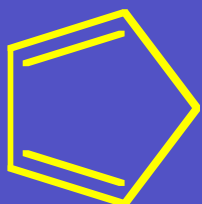
Dien

Alkene,
Dienophile

Cyclohexene



Cyclic reaction



+

