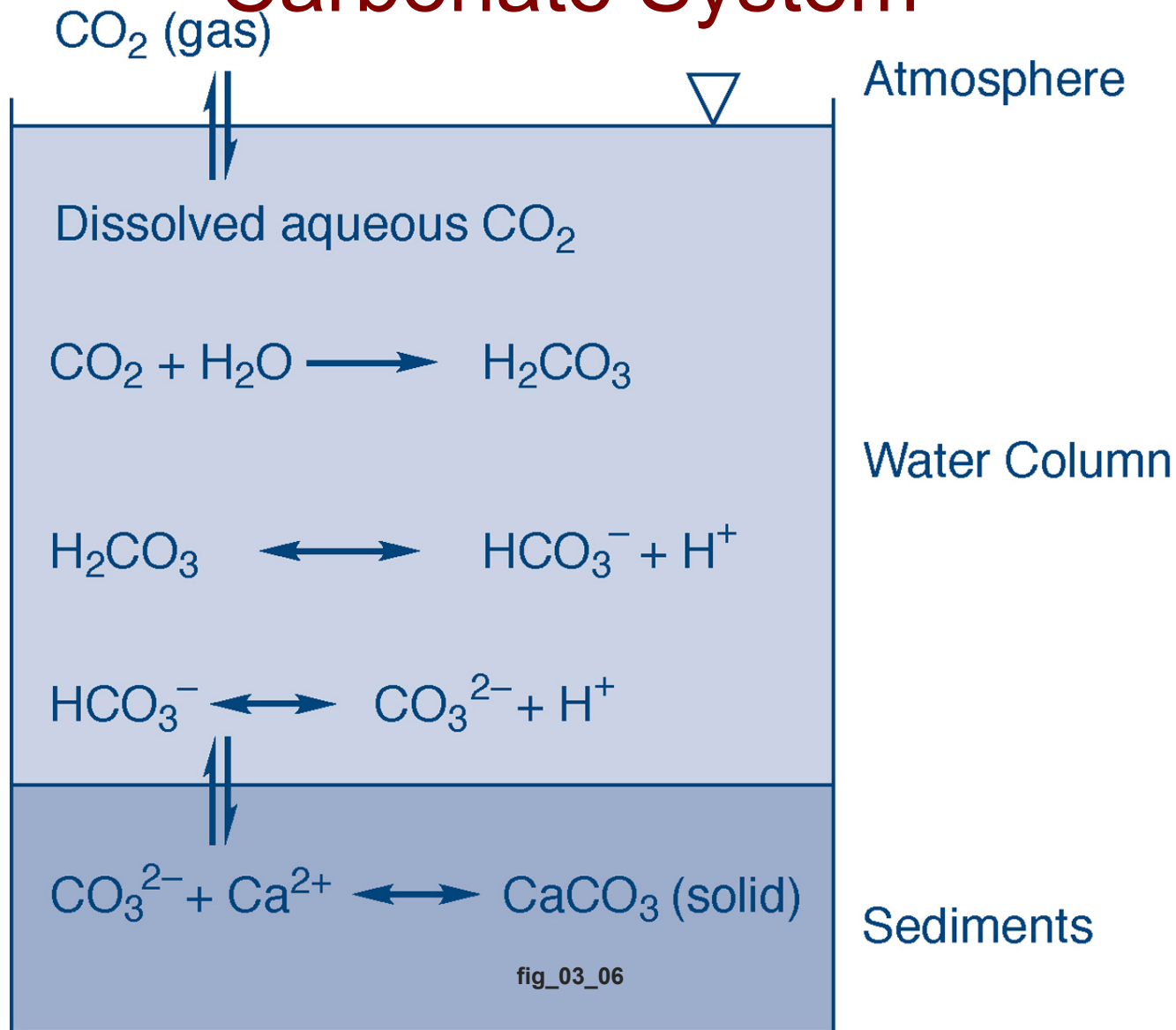


Announcements

Monday 9/16

- Lab tomorrow in Dana 325—Model fitting, Kinetics & Discussion of *Big Necessity*.
- Problem Set 2 due Wed 9/18 in class.
- Halfway through responding to blog responses
- PSet1 Graded
- Quizz Prep Extra Credit (should be graded by tWed)
- Trying to figure out a time to have Dr. Heavner's visit. Afternoon/evening in place of class on Oct. 7

Acid-Base Carbonate System



Alkalinity

$$\text{Alkalinity (eq/L)} = [\text{HCO}_3^-] + 2[\text{CO}_3^{2-}] + [\text{OH}^-] - [\text{H}^+],$$

where [] refer to concentrations in moles/L

- Alkalinity is the ability of a water to neutralize acid.
- In most waters, only carbonate species contribute significantly
- Between pH 6 to 8, primarily $[\text{HCO}_3^-]$