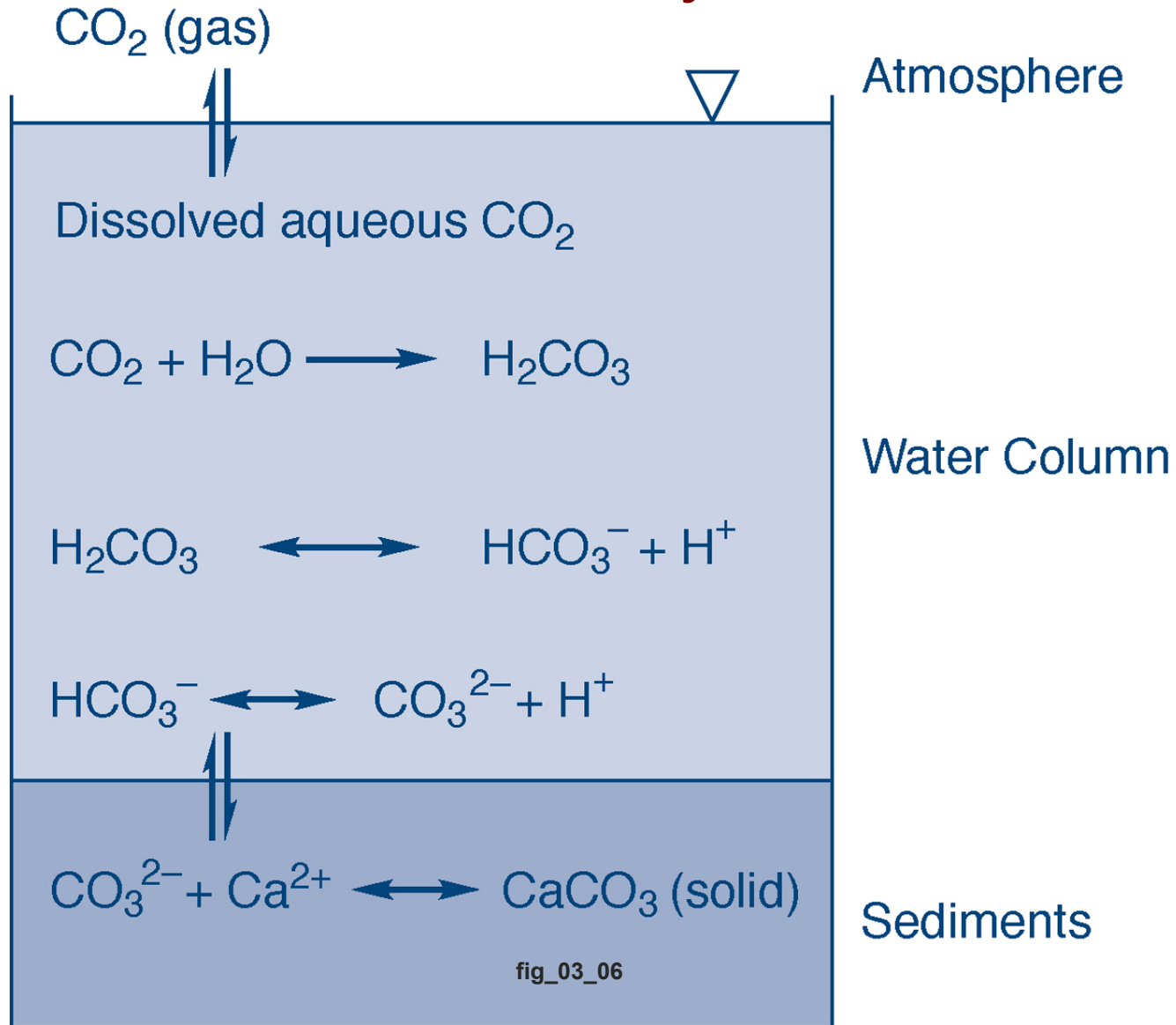


Announcements

Wednesday 9/18

- Problem Set 3 posted; due Wed 9/25 in class.
- Graded PSets in mailboxes
- Will announce time(s) for meeting with Dr. Heavner on Friday. **Please fill out Doodle Poll** if you haven't.
- Midterm course evaluations on Monday

Acid-Base Carbonate System



Alkalinity

$$\text{Alkalinity (eq/L)} = [\text{HCO}_3^-] + 2[\text{CO}_3^{2-}] + [\text{OH}^-] - [\text{H}^+],$$

where [] refer to concentrations in moles/L

- Alkalinity is the ability of a water to neutralize acid.
- In most waters, only carbonate species contribute significantly
- Between pH 6 to 8, primarily $[\text{HCO}_3^-]$