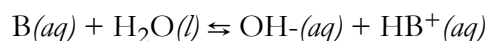


Acid/Base Properties of a Pharmaceutical Compound

Many pharmaceutically interesting compounds are weak acids or weak bases, a fact of some importance when preparing them for use by the public. For example, if the active ingredient in a nasal spray is too acidic, then the preparation might include an additional ingredient to neutralize some of the acid; after all, no one wants to spray something as acidic as, say, lemon juice up his or her nose!

Pseudoephedrine is a central nervous system stimulant used in many cold and allergy tablets. In its molecular form it is a weak base (it is an amine), which, for convenience, we may represent as B. In water, the following equilibrium reaction exists



The Merck Index reports that a 0.030 M solution of pseudoephedrine has an equilibrium pH of 11.44. What is the value of K_b for this compound?

Answer. With a pH of 11.44, the pOH is 2.56 and the equilibrium concentration of OH^- is 2.75×10^{-3} M. This value goes in the change column for OH^- and determines the change for B and HB^+

	B(aq)	+	H ₂ O	↔	OH ⁻ (aq)	+	HB ⁺ (aq)
I	0.030 M		—		0		0
C	-2.75×10^{-3} M		—		2.75×10^{-3} M		2.75×10^{-3} M
E	0.02725		—		2.75×10^{-3} M		2.75×10^{-3} M

Now that we have the equilibrium concentrations, we substitute into the K_b expression

$$K_b = \frac{[\text{HB}^+][\text{OH}^-]}{[\text{B}]} = \frac{(2.75 \times 10^{-3})^2}{0.02725} = 2.77 \times 10^{-4}$$

Because its base form is only slightly soluble in water, pseudoephedrine typically is dispensed in its weak acid form, HB^+ . Pseudoephedrine hydrochloride, therefore, is an ionic compound consisting of protonated pseudoephedrine, HB^+ , and Cl^- as a counter ion. The formula for this often is written as $\text{B} \cdot \text{HCl}$; don't confuse this with the strong acid HCl. A solution of pseudoephedrine hydrochloride, therefore, is acidic due to the presence of HB^+ . Write the K_a reaction responsible for making the solution acidic and report the value for K_a .

Answer. The reaction is $\text{BH}^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{B}$, for which the equilibrium constant is

$$K_a = K_w/K_b = (1.00 \times 10^{-14})/(2.77 \times 10^{-4}) = 3.61 \times 10^{-11}$$

Suppose you dissolve three tablets of Sudafed[®], each containing 30.0 mg of pseudoephedrine hydrochloride, in 200.0 mL of water. What is the pH of the resulting solution? The molar mass for pseudoephedrine hydrochloride is 201.7 g/mol.

Answer. Begin by finding the concentration of pseudoephedrine hydrochloride, which is

$$\frac{3 \text{ tablets} \times \frac{0.030 \text{ g}}{\text{tablet}} \times \frac{1 \text{ mole}}{201.7 \text{ g}}}{0.200 \text{ L}} = 2.23 \times 10^{-3} \text{ M}$$

Taking this as the initial concentration of HB^+ , we solve the problem

	$\text{HB}^+(aq)$	+	$\text{H}_2\text{O}(l)$	\rightleftharpoons	$\text{H}_3\text{O}^+(aq)$	+	$\text{B}(aq)$
I	$2.23 \times 10^{-3} \text{ M}$		—		0		0
C	-X		—		+X		+X
E	$2.23 \times 10^{-3} - X$		—		X		X

Substituting into the K_a expression and solving gives

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{B}]}{[\text{HB}^+]} = \frac{X^2}{2.23 \times 10^{-3} - X} = 3.61 \times 10^{-11}$$

Assuming that $2.23 \times 10^{-3} - X \approx 2.23 \times 10^{-3}$ simplifies this to

$$K_a = \frac{X^2}{2.23 \times 10^{-3}} = 3.61 \times 10^{-11}$$

which gives

$$X = 2.84 \times 10^{-7}$$

Checking our assumption, we find that this produces an error of

$$\% \text{error} = \frac{2.23 \times 10^{-3} - (2.23 \times 10^{-3} - 2.84 \times 10^{-7})}{2.23 \times 10^{-3}} \times 100 = 0.01\%$$

which is acceptable. The concentration of H_3O^+ is X, thus

$$[\text{H}_3\text{O}^+] = 2.84 \times 10^{-7} \quad \text{pH} = 6.55$$