Ka for acetic acid =	1.78E-05	
molarity acetic acid as supplied =	0.7494	M
molarity diluted acetic acid =	5.00E-02	M
initial [H3O+] =	9.43E-04	M
initial pH =	**3.03**	
molarity of NaOH =	0.1038	M
equivalence point volume =	14.44	mL
molarity acetate at equiv. point =	3.37E-02	M
Kb for acetate =	5.62E-10	
[OH-] at equiv. point =	4.35E-06	M
pOH at equiv. point =	5.36	
pH at equiv. point =	**8.64**	

before equiv. point solution is a buffer of acetic acid and acetate
moles acetate = moles OH- added
moles acetic acid = initial moles acetic acid - moles OH- added

volume of titrant added (mL)	4.00	6.00	8.00	10.00	12.00
moles OH- added =	4.15E-04	6.23E-04	8.30E-04	1.04E-03	1.25E-03
moles acetate =	4.15E-04	6.23E-04	8.30E-04	1.04E-03	1.25E-03
moles acetic acid =	1.08E-03	8.76E-04	6.68E-04	4.61E-04	2.53E-04
pH = pKa - log(mol acetate/mol acetic acid) =	**4.33**	**4.60**	**4.84**	**5.10**	**5.44**

after equiv. point pH is controlled by excess NaOH
moles excess OH- = moles titrant added - initial moles acetic acid

volume titrant added (mL)	16.00	18.00	20.00	22.00	24.00
moles excess OH- =	1.62E-04	3.70E-04	5.77E-04	7.85E-04	9.92E-04
[OH-] =	3.52E-03	7.70E-03	1.15E-02	1.51E-02	1.84E-02
pOH =	2.45	2.11	1.94	1.82	1.74
pH =	**11.55**	**11.89**	**12.06**	**12.18**	**12.26**

data to plot	V titrant	pH
	0.00	3.03
	4.00	4.33
	6.00	4.60
	8.00	4.84
	10.00	5.10
	12.00	5.44
	14.44	8.64
	16.00	11.55
	18.00	11.89
	20.00	12.06
	22.00	12.18
	24.00	12.26