

Antiperovskite Ionic Conductor Layer for Stabilizing the Interface of NASICON Solid Electrolyte Against Li Metal in All-Solid-State Batteries**

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The instability of the interface between solid-state electrolytes (SSEs) and Li metal anodes is a critical factor restricting the application of NASICON-type ionic conductors in all-solid-state lithium batteries (ASSLBs). Here, the lithium-rich antiperovskite (LiRAP) Li_2OHBr ionic conductor is introduced as a protective layer for the NASICON-type solid electrolyte $\text{Li}_{1.5}\text{Al}_{0.5}\text{Ge}_{1.5}(\text{PO}_4)_3$ (LAGP) by a simple way of rapid solidification of molten Li_2OHBr , to prevent the adverse reaction caused by direct contact

between LAGP and Li metal. The Li/Li symmetric cells with Li_2OHBr protective layer exhibit long-term cycle stability compared with the bare LAGP, and the mechanisms are analyzed by XPS measurements for the interface between SSEs and Li metal. Also, the $\text{LiFePO}_4/\text{Li}$ ASSLBs have been assembled to further demonstrate this advantage. This work shows that LiRAP Li_2OHBr can improve the cycling stability of ASSLBs from the perspective of Li metal anode protection.

1. Introduction

Li-ion batteries (LIBs) have been widely applied in consumer electronics, electric vehicles, and stationary energy storage systems owing to their excellent electrochemical performances, and the demand for safer Li secondary batteries with higher energy density is still increasing simultaneously.^[1] Whereas, the safety of current LIBs is challenged due to the flammability of the liquid electrolytes in thermal runaway behavior,^[2] and the application of Li metal anode is limited in LIBs with liquid electrolytes due to the formation of dendrites,^[3] which leads to the restriction on the increase in energy density.^[4] To overcome these obstacles, inorganic solid-state electrolytes (SSEs) have been introduced into LIBs for their merits of nonflammability,^[5] good electrochemical stability,^[6] and applicability for the next generation of high-safety and high energy density all-solid-state lithium batteries (ASSLBs).^[7] Consequently, a variety of

SSEs with high ionic conductivity have been discovered and investigated in recent years.^[8]

Among these candidate SSEs, the NASICON-type ionic conductors ($\text{Li}_{1-x}\text{Al}_x\text{M}_{2-x}(\text{PO}_4)_3$, $\text{M}=\text{Zr, Ti, Ge, etc.}$) are considered to be one of the most promising SSEs for commercial ASSLBs, due to their high ionic conductivity, stability to air and moisture, and easiness for synthesis and sintering.^[9] However, there is still a key challenge limiting the development of ASSLBs based on the NASICON-type SSEs, that is, the instability of interface between the NASICON-type SSEs and Li metal anode.^[10] Representatively, recent studies on the chemical reaction between $\text{Li}_{1.5}\text{Al}_{0.5}\text{Ge}_{1.5}(\text{PO}_4)_3$ (LAGP) and Li metal showed that the chemically formed interphase strongly affects the mechanical integrity of LAGP pellet.^[11] Rapid capacity loss due to the formation of mixed conducting interphase at the LAGP/lithium interface has also been reported.^[12] Besides, the morphological evolution of the interphase during cycling in symmetric Li/LAGP/Li cells is dependent on the current density, which has been proven by *in situ* transmission electron microscopy (TEM).^[13] Therefore, stabilizing the Li/LAGP interface has become the prerequisite for LAGP to be used as a practical candidate SSE for all-solid-state Li metal batteries.^[10b,14]

A variety of strategies have been proposed to improve the stability of the interface between the NASICON-type SSEs and Li metal. For example, polymer coatings were reported to improve electrochemical stability and compatibility of LAGP to Li metal.^[12b,15] However, whether the polymer can prevent the penetration of Li dendrites has been the focus of debate among researchers. The Al_2O_3 -coated LATP [$\text{Li}_{1.3}\text{Al}_{0.3}\text{Ti}_{1.7}(\text{PO}_4)_3$] by atomic layer deposition (ALD) exhibited a stable cycling behavior and small resistance, which suggested that ALD is effective in improving SSE/electrode interface stability.^[10c] Whereas, in fact, Al_2O_3 is not an ionic conductor, which is unfavorable for Li-ions to shuttle between LATP and Li metal, and similar limitation exists in the recently reported work of Ge

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[**] NASICON: Sodium (Na) Super Ionic CONductor

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metal coating on LAGP.^[11c] Also, the amorphous LAGP layer was formed on the ceramic LAGP via a spin-coating co-sintering method to insulate the internal electrolyte from the external highly reducing environment,^[14a] but the existence of Ge^{4+} in the amorphous LAGP layer may lead to the inevitable reduction reaction. To address these challenges, we propose to use lithium-rich antiperovskite (LiRAP)^[16] Li_2OHBr ionic conductor as a protective layer to prevent the adverse reaction caused by direct contact between LAGP and Li metal,^[11c,17] and the reason that LiRAP Li_2OHBr is specifically chosen in this work is as follows. Firstly, the LiRAP family has been proven to be stable to metallic lithium in symmetric $\text{Li}/\text{LiRAP}/\text{Li}$ cells^[18] and ASSLBs.^[19] In this case, the interface between LiRAP and Li metal would be stable during cycling. Secondly, Li_2OHBr possesses high ionic conductivity ($>10^{-6}$ S/cm) at room temperature (RT),^[19–20] which could support the fast shuttle of lithium ions between LAGP and Li metal. Thirdly, since the melting point of Li_2OHBr is only $\sim 254^\circ\text{C}$,^[21] it is easy for the direct solidification of molten Li_2OHBr into desired shape and forming a protective layer on LAGP,^[18,20,22] and does not need a special customized mold to provide mechanical pressure to avoid imperfect physical contact between SSEs and protective layer.^[23] Fourthly, the mechanical property and flame retardancy of inorganic Li_2OHBr are better than those of polymer composite solid electrolyte membranes,^[12b,15,24] and it could protect LAGP from Li anode more effectively in ASSLBs.

Herein, the protective layer on the surface of LAGP pellet was formed by a simple way of rapid solidification of molten Li_2OHBr , and the dense and smooth Li_2OHBr layer was observed with scanning electron microscopy (SEM). The Li/Li symmetric

cells with Li_2OHBr protective layer exhibit long-term cycle stability, compared with those using bare LAGP. XPS analysis was conducted to reveal the possible reasons for the stability of the interface between Li_2OHBr -protected LAGP and Li metal. Besides, $\text{LiFePO}_4/\text{SSEs}/\text{Li}$ ASSLBs with protective layer of Li_2OHBr were also assembled and showed good cycle and rate performances.

2. Results and Discussion

2.1. Characterizations of SSEs

The XRD pattern of LAGP powder and sintered LAGP pellet at room temperature are shown in Figure 1a. All the main peaks in the XRD pattern match well with the NASICON-type structure of $\text{LiGe}_2(\text{PO}_4)_3$ (LGP, PDF#80-1922, $R\bar{3}c$ (167) space group), and details of the crystal can be obtained from previous work.^[25] Due to the annealing at 950°C for 10 h, the crystallinity of LAGP pellet becomes better after sintering, and thus exhibits a sharper peak shape than LAGP powder. The XRD patterns of Li_2OHBr powder and layer on the LAGP pellet are shown in Figure 1b. The XRD pattern of Li_2OHBr powder exhibits single-phase characteristics and matches well with the standard PDF card (#84-0101, $Pm\bar{3}m$ (221) space group), except for the bulge at the low angle, which is caused by the Kapton film. The OH^- ions are coordinated by Li^+ ions, forming $<\text{Li}_6\text{OH}>$ octahedra with one-third of the Li vacancies for charge balance, and the Br^- ions are located at the 12-fold coordinated cavity of the octahedral network. Interestingly, the Li_2OHBr layer solidified

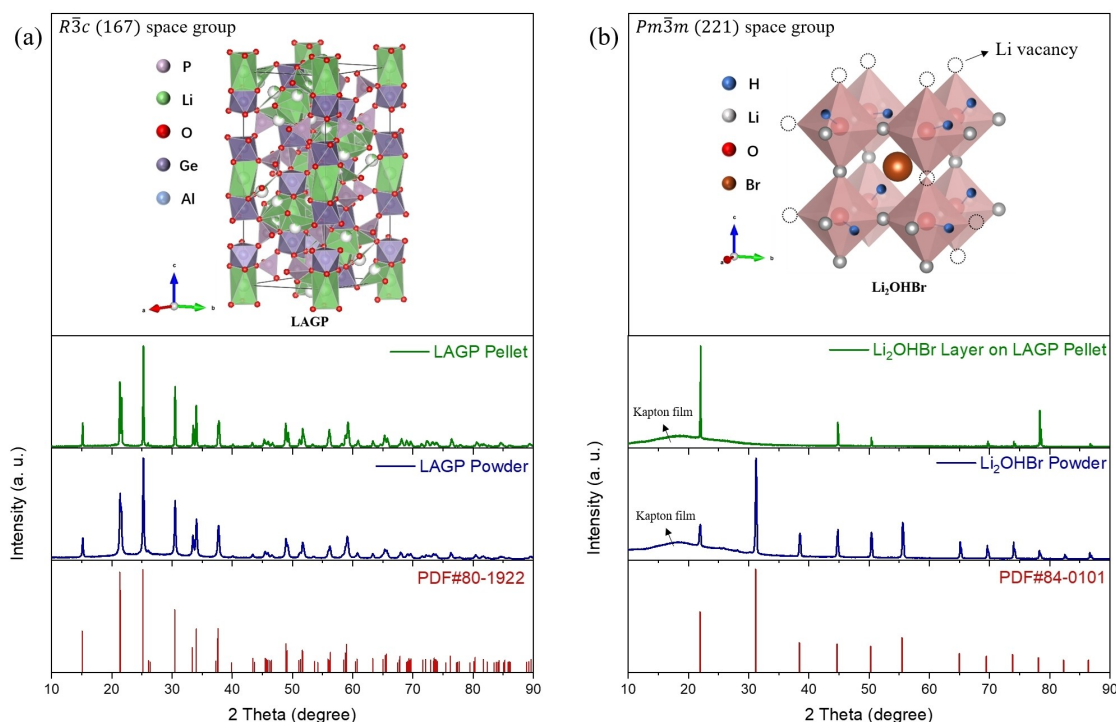


Figure 1. The schematic crystal structures and XRD patterns of a) LAGP and b) Li_2OHBr .

on the surface of LAGP after melting and rapid cooling shows obvious orientation (e.g. the Miller index (100) direction), which has also been observed in our other experiments involving the rapid cooling of molten Li_2OHBr . This orientation issue will be studied more systematically in our future works.

Before the preparation of Li_2OHBr -protected LAGP pellets, the melting temperature of Li_2OHBr was determined by DSC measurement. As shown in Figure 2S, the endothermic and exothermic peaks represent the melting and crystallization temperature of Li_2OHBr , respectively. The melting point of Li_2OHBr is about 257°C , which is higher than the reported value of 254°C ,^[21] and the slight difference in melting points may be related to the different synthesis conditions of the samples. The reported Li_2OHBr was synthesized by ball milling at room temperature, which may help reduce the crystallinity of the samples and lead to a lower melting point.^[26]

Figure 2a and Figure 2b show the Nyquist plots of Li_2OHBr and LAGP measured in the temperature range between 25 and 100°C , respectively. Taking the Nyquist curve of LAGP at 30°C as a representative, as shown in Figure 2e, it exhibits a typical semicircle at high frequency representing the bulk and grain boundary resistance in series, and the linear part at low frequency representing ion blocking electrode. Fitting the Nyquist curve by the equivalent circuit leads to a total resistance R of $484\ \Omega$, including the bulk resistance R_b of $200\ \Omega$ and the grain boundary resistance R_{gb} of $284\ \Omega$. Further, we calculate the ionic conductivity is $2.1 \times 10^{-4}\ \text{S/cm}$ based on the formula of $\sigma = L/(RS)$, where L ($\sim 1.2\ \text{mm}$), R ($\sim 484\ \Omega$), S ($\sim 1.13\ \text{cm}^2$) are the thickness, total resistance and area of LAGP pellet, respectively. Figure 2c and Figure 2d are the Arrhenius plots of Li_2OHBr and LAGP illustrating the relationship between ionic conductivity and temperature in the range between 25

and 100°C , and the ionic conductivities of Li_2OHBr and LAGP are $1.03 \times 10^{-6}\ \text{S/cm}$ and $1.7 \times 10^{-4}\ \text{S/cm}$ at RT (25°C), respectively. The activation energies E_a of Li_2OHBr and LAGP are $0.597\ \text{eV}$ and $0.360\ \text{eV}$ respectively, which are fitted according to the slope of the plots following the Arrhenius law: $\sigma = A \exp(-E_a/k_B T)$. In which σ is the conductivity, A is the preexponential parameter, E_a is the activation energy, T is absolute temperature, and k_B is the Boltzmann constant. In addition, the Nyquist curve of the Li_2OHBr -protected LAGP electrolyte at RT (25°C) is shown in Figure 2f, which is fitted by the equivalent circuit leads to a total resistance R of $8.24\ \text{k}\Omega$, including the bulk resistance R_b of $1.2\ \text{k}\Omega$, the grain boundary resistance R_{gb} of $5.52\ \text{k}\Omega$, and the interface resistance R_i of $1.52\ \text{k}\Omega$. Further, the ionic conductivity is calculated to be $1.4 \times 10^{-5}\ \text{S/cm}$ based on the formula of $\sigma = L/(RS)$, where L ($\sim 1.3\ \text{mm}$), R ($\sim 8.24\ \text{k}\Omega$), S ($\sim 1.13\ \text{cm}^2$) are the thickness, total resistance and area of the composite electrolyte, respectively.

Figure 3a shows the surface morphology of the sintered bare LAGP pellet, and Figure 3b shows the Li_2OHBr protective layer on LAGP surface observed by SEM. For the LAGP pellet, distinct graininess could be observed, even after sintering and densification at 950°C for 10 h. In contrast, the Li_2OHBr protective layer on the surface of LAGP presents a smooth and dense morphology after Li_2OHBr melting, rapid cooling, and solidification. These features may be beneficial to the uniformity of the current density during lithium ions intercalation/deintercalation and prevent lithium dendrite from penetrating the LAGP pellet through the pores and causing a short-circuit of the batteries. Figure 3c shows the cross-sectional morphology of the Li_2OHBr -protected LAGP pellet, and the thickness of Li_2OHBr layer is about $20\ \mu\text{m}$. Meanwhile, good adhesion can be observed in the transition region between LAGP pellet and

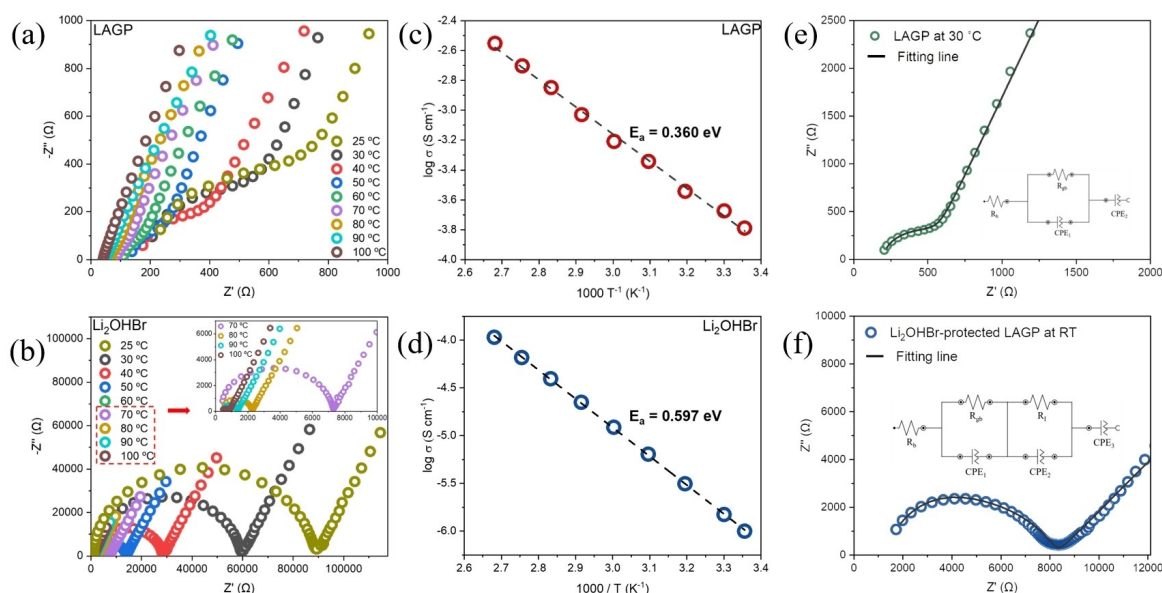


Figure 2. a, b) Nyquist plots of Li_2OHBr and LAGP measured in the temperature range from 25 to 100°C . c, d) The Arrhenius plots of Li_2OHBr and LAGP illustrating the relationship between ionic conductivity and temperature in the range from 25 to 100°C . e) Fitting curve with the equivalent circuit for Nyquist plot of LAGP at 30°C . f) Fitting curve with the equivalent circuit for Nyquist plot of Li_2OHBr -protected LAGP at RT (25°C).

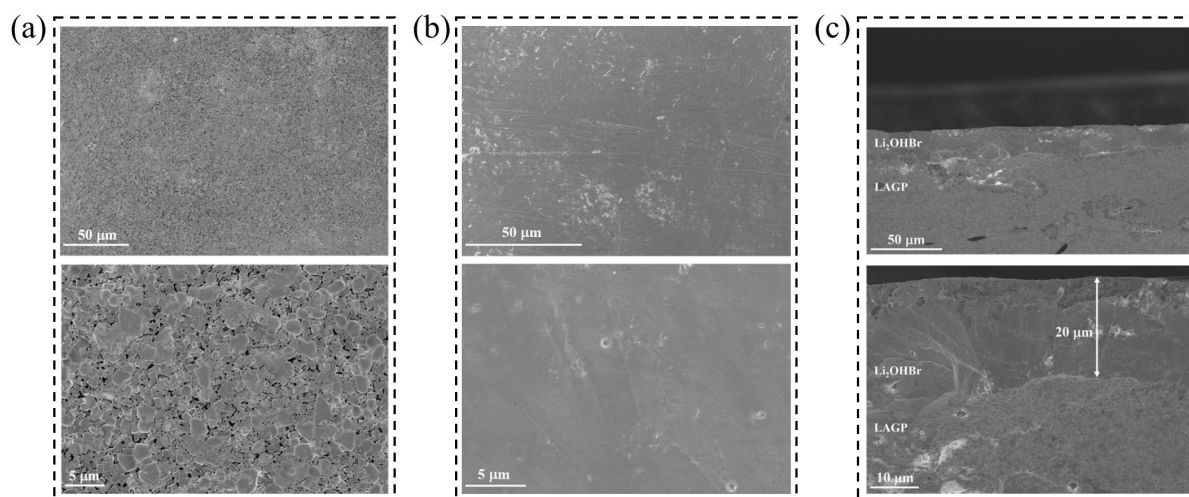


Figure 3. a) Surface morphology of the sintered bare LAGP pellet. b) Surface morphology of Li_2OHBr protective layer on LAGP pellet. c) Cross-sectional morphology of the Li_2OHBr -protected LAGP pellet, and the thickness of Li_2OHBr layer is about 20 μm .

Li_2OHBr layer, which may be due to the fact that molten Li_2OHBr infiltrated into the pores of LAGP pellet.

2.2. Electrochemical Performance of Li/Li Symmetric Cells

Before the Li/Li symmetric cell tests, cyclic voltammetry (CV) measurement is conducted to investigate whether Li_2OHBr

could prevent LAGP from being reduced with Li metal or not. In the CV tests, Li metal was used as the counter electrode and Au was used as the working electrode. As shown in Figure 3S, the Li_2OHBr -protected LAGP is highly stable in the range of -0.5 – 5.0 V versus Li/Li^+ , demonstrating no side reactions happen in the interphase.

The electrochemical performance of Li/Li symmetric cells is shown in Figure 4. With the galvanostatic cycling of Li/LAGP/Li

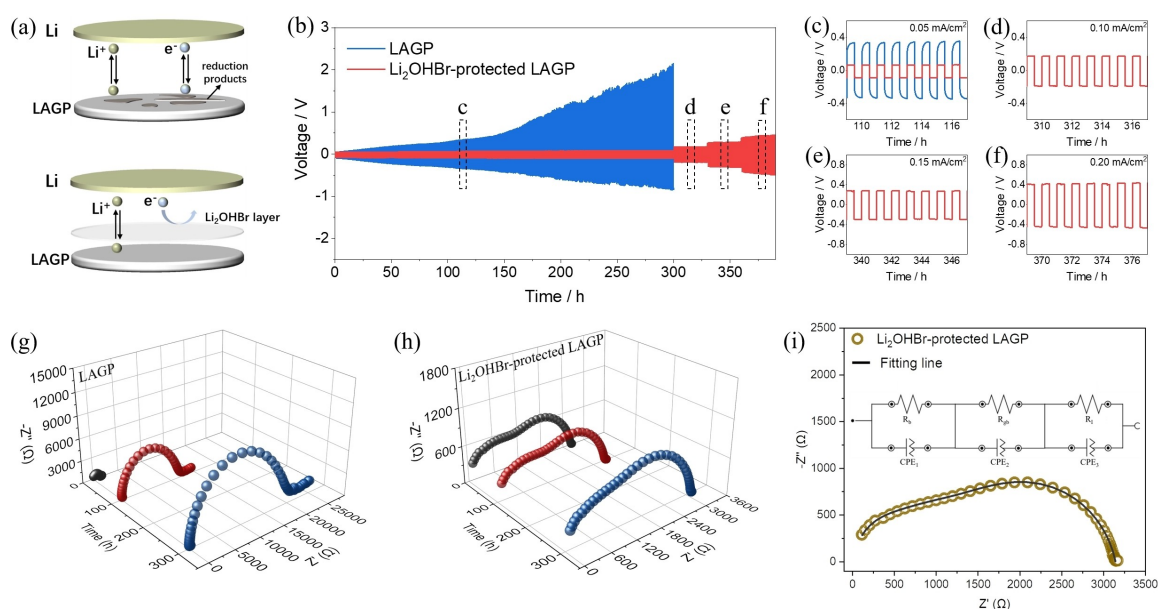


Figure 4. a) Schematic diagram of the evolution of the interface between LAGP and Li metal, direct contact of bare LAGP pellet with Li metal leads to the formation of reduction products (upper part), the interface between LAGP and Li metal is stabilized by the protection of Li_2OHBr layer (lower part). b) Galvanostatic cycling of Li/LAGP/Li symmetric cell at current density of $0.05\text{ mA}/\text{cm}^2$ (300 h) and Li/ Li_2OHBr /LAGP/ Li_2OHBr /Li symmetric cell at current density of $0.05\text{ mA}/\text{cm}^2$ (300 h), $0.10\text{ mA}/\text{cm}^2$ (30 h), $0.15\text{ mA}/\text{cm}^2$ (30 h) and $0.20\text{ mA}/\text{cm}^2$ (30 h). c), d), e), f) Insets of the specific period of the galvanostatic cycling. g), h) EIS of Li/LAGP/Li symmetric cell and Li/ Li_2OHBr /LAGP/ Li_2OHBr /Li symmetric cell at different cycle times. i) Fitting curve with the equivalent circuit for Nyquist plot of Li/Li symmetric cell of Li_2OHBr -protected LAGP at 80°C .

symmetric cell (Figure 4b), the obvious polarization was observed along with the increase of voltage at the current density of 0.05 mA/cm^2 . In contrast, $\text{Li/Li}_2\text{OHBr/LAGP/Li}_2\text{OHBr/Li}$ symmetric cell remained stable after a long-term galvanostatic cycling (Figure 4b-f; 0.05 mA/cm^2 -300 h, 0.10 mA/cm^2 -30 h, 0.15 mA/cm^2 -30 h, and 0.20 mA/cm^2 -30 h). Meanwhile, the EIS plots of Li/Li symmetric cells (Figure 4g and 4 h) were analyzed to verify whether the Li_2OHBr layer is beneficial to alleviate polarization. Taking the EIS plot of $\text{Li/Li}_2\text{OHBr/LAGP/Li}_2\text{OHBr/Li}$ symmetric cell after 300 h cycle as a representative, as shown in Figure 4i, fitting it by the equivalent circuit leads to a total resistance R of $3.14 \text{ k}\Omega$, including the bulk resistance R_b of $0.26 \text{ k}\Omega$, the grain boundary resistance R_{gb} of $1.76 \text{ k}\Omega$, and the interface resistance R_i of $1.12 \text{ k}\Omega$. Before the galvanostatic cycling of the symmetric cells, the $\text{Li/Li}_2\text{OHBr/LAGP/Li}_2\text{OHBr/Li}$ cell ($2.63 \text{ k}\Omega$) shows a higher impedance than that of Li/LAGP/Li cell ($1.46 \text{ k}\Omega$) due to the introduction of the Li_2OHBr layer between LAGP and Li metal. However, with the galvanostatic cycling tests, the impedance of Li/LAGP/Li cell changed from $1.46 \text{ k}\Omega$ to $20.82 \text{ k}\Omega$ (after 0.05 mA/cm^2 -300 h). Comparatively, $\text{Li/Li}_2\text{OHBr/LAGP/Li}_2\text{OHBr/Li}$ cell only changed from $2.63 \text{ k}\Omega$ to $3.14 \text{ k}\Omega$ (after 0.05 mA/cm^2 -300 h), which is due to the existence of Li_2OHBr layer to stabilize the interface of LAGP against Li metal. Noteworthily, the resistance of $2.63 \text{ k}\Omega$ is slightly high for Li/Li symmetric cell, which is unfavorable for operation at high current densities. The method of vacuum evaporation or spin-coating could be considered to reduce the thickness of Li_2OHBr layer and the resistance of Li/Li cells in future experiments.

To further demonstrate the role of Li_2OHBr layer in stabilizing LAGP against Li metal, we disassembled the Li/Li symmetric cells to observe the surface and cross-sectional morphology by SEM after galvanostatic cycling. For LAGP, as shown in Figure 5a (left), the morphology of surface deteriorated obviously, accompanied by surface cracks and reaction by-products (Figure 4S). In contrast, the surface morphology of Li_2OHBr protective layer on LAGP still maintained flat (Figure 5a (right)), and there is no obvious difference from that before the Li/Li symmetric cell tests, except that the holes are found on the Li_2OHBr layer, which may be caused by the continuous Li ions shuttling between the Li metal and the Li_2OHBr layer after a long galvanostatic cycling of Li/Li symmetric cell. Meanwhile, the good adhesion can still be observed in the transition region between Li_2OHBr layer and Li metal, as well as the transition region between Li_2OHBr layer and LAGP (Figure 5b), with correlating energy-dispersive X-ray spectroscopy (EDS) mapping of oxygen (existing in Li_2OHBr , LAGP and Li metal) in green, bromine (existing in Li_2OHBr) in orange, and germanium (existing in LAGP) in yellow. Noteworthily, the existence of oxygen in Li metal may be attributed to its susceptibility to oxidation.

Furthermore, the stability of the interface region was analyzed by XPS. For LAGP, as shown in Figure 5c, the binding energy of Ge^{4+} for the 3d orbit is 33.0 eV and the binding energy corresponding to the Li 1s orbit is 55.7 eV before the galvanostatic cycling. After electrochemical cycles, in addition to the Ge^{4+} peak, the peaks of Ge metal (29.5 eV) and Ge^{3+} (31.2 eV) appeared in the Ge 3d data. Also, the Li 1s peak showed an obvious shift, which included the original peak of

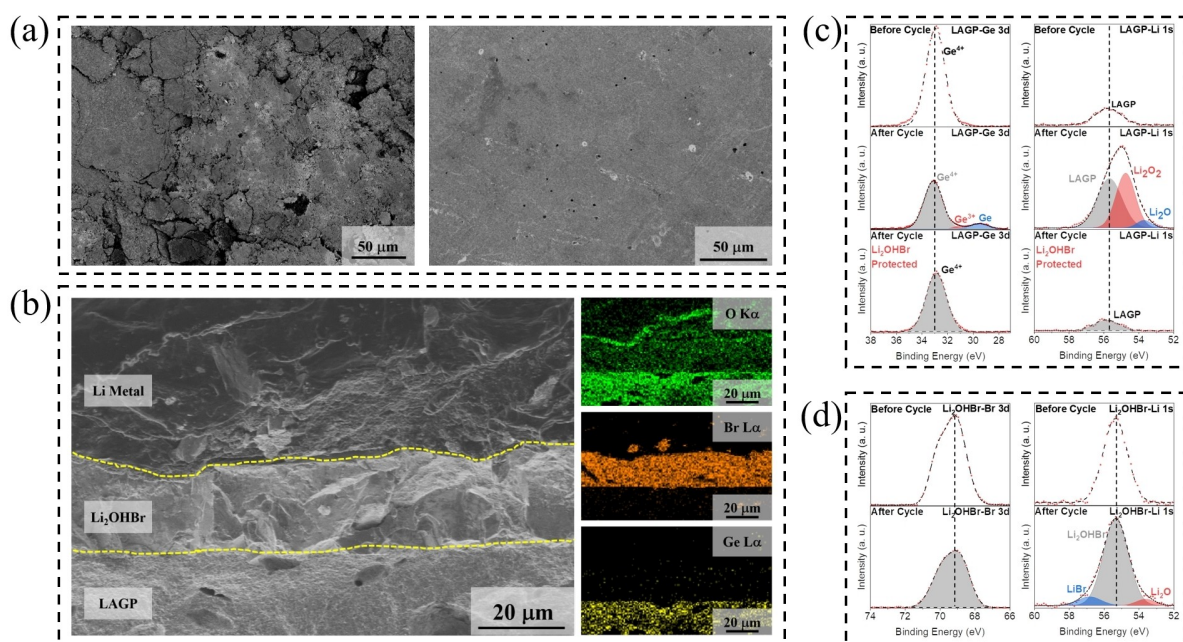
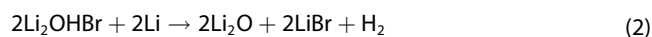
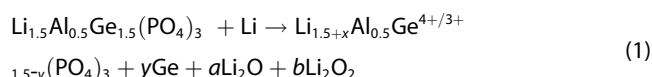


Figure 5. a) The morphology of LAGP surface (left) and Li_2OHBr layer surface (right) after galvanostatic cycling of Li/Li symmetric cells. b) Cross-sectional morphology of the Li_2OHBr -protected LAGP pellet after galvanostatic cycling of Li/Li symmetric cells, with correlating EDS mapping of oxygen in green, bromine in orange, and germanium in yellow. c) XPS of LAGP, data of Ge 3d and Li 1s before cycles (top), data of Ge 3d and Li 1s after cycles without Li_2OHBr protecting (middle), data of Ge 3d and Li 1s after cycles with Li_2OHBr protecting (bottom). d) XPS of Li_2OHBr layers, data of Br 3d and Li 1s before cycles (top), data of Br 3d and Li 1s after cycles (bottom).

LAGP (55.7 eV), the peak of Li_2O_2 (54.8 eV), and the peak of Li_2O (53.7 eV). These results indicated that LAGP was reduced by Li metal after cycling in Li/LAGP/Li cell, and the possible reaction mechanism [Equation (1)] is suggested. In particular, Ge metal, the decomposition product of LAGP, has electronic conductivity, which may lead to the continuous reactions between LAGP and Li metal, resulting in the increase of polarization and interface impedance of Li/LAGP/Li cells. Comparatively, the LAGP in Li/Li₂OHBr/LAGP/Li₂OHBr/Li cell did not show obvious changes after cycling, and the peaks of Ge 3d orbit and Li 1s orbit are consistent with those before cycling. For Li₂OHBr layer, as shown in Figure 5d, the peaks of Br 3d and Li 1s after cycling remained similar to the main peaks before cycling, which indicates that Li₂OHBr is stable to Li metal during cycling. This result is due to the excellent electronic insulation (as shown in Figure 5S, $\sim 1.1 \times 10^{-8} \text{ S/cm}$ at 80°C) and ionic conductivity of Li₂OHBr. As a protective layer between LAGP and Li metal, Li₂OHBr not only acts as an ion-transport layer but also prevents electron conduction between LAGP and Li metal, thereby effectively preventing the reaction between LAGP and Li metal (Figure 4a). Interestingly, in the XPS data of Li 1s, we noticed that there may be traces of LiBr and Li₂O in Li₂OHBr by fitting the spectra. This result is similar to that of Li₂OHCl previously reported,^[18] which infers Li₂OHCl may react with Li metal to form LiCl and Li₂O in the interface region during the high-temperature cycling at 195°C , and this bulk solid electrolyte interphase (SEI, a mixture of LiCl and Li₂O) forms an interconnected network that protects the Li₂OHCl from further reacting with Li metal. Therefore, we speculate that a small amount of above-mentioned bulk SEI (a mixture of LiBr and Li₂O) may be formed at the interface between Li₂OHBr and Li metal during cycling of Li/Li₂OHBr/LAGP/Li₂OHBr/Li cell at 80°C ,

which prevents a further reaction between Li₂OHBr and Li metal [as shown in Equation (2)].



2.3. Electrochemical Performance of LiFePO₄/Li ASSLBs

After verifying that the Li₂OHBr layer can be used to stabilize the interface of LAGP against the Li metal, LFP/SSEs/Li cells were assembled to further examine whether the Li₂OHBr-protected LAGP electrolytes can be applied in all-solid-state batteries. As shown in Figure 6a, despite the LFP/Li ASSLB without Li₂OHBr layer exhibits the initial specific capacity of 113.7 mAh g^{-1} , it decays to less than 96 mAh g^{-1} after 20th cycle at 0.1 C, only possessing poor capacity retention of 25.7 mAh g^{-1} after 40th cycle. Such fast capacity decay is stimulated by the deteriorating LAGP/Li interface and rapidly increasing overpotential (Figure 6d). In contrast, the LFP/Li ASSLB with Li₂OHBr layer exhibits the higher initial specific capacity of 119.9 mAh g^{-1} , and remains the capacity of 110.1 mAh g^{-1} after 20th cycle and the capacity of 96.3 mAh g^{-1} after 40th cycle at 0.1 C, respectively (Figure 6b). Meanwhile, the increase in overpotential has been well alleviated compared to LFP/Li ASSLB without Li₂OHBr layer (Figure 6e), which is due to the Li₂OHBr layer effectively avoiding the reactions between LAGP and Li metal anode. Besides, rate performance was also

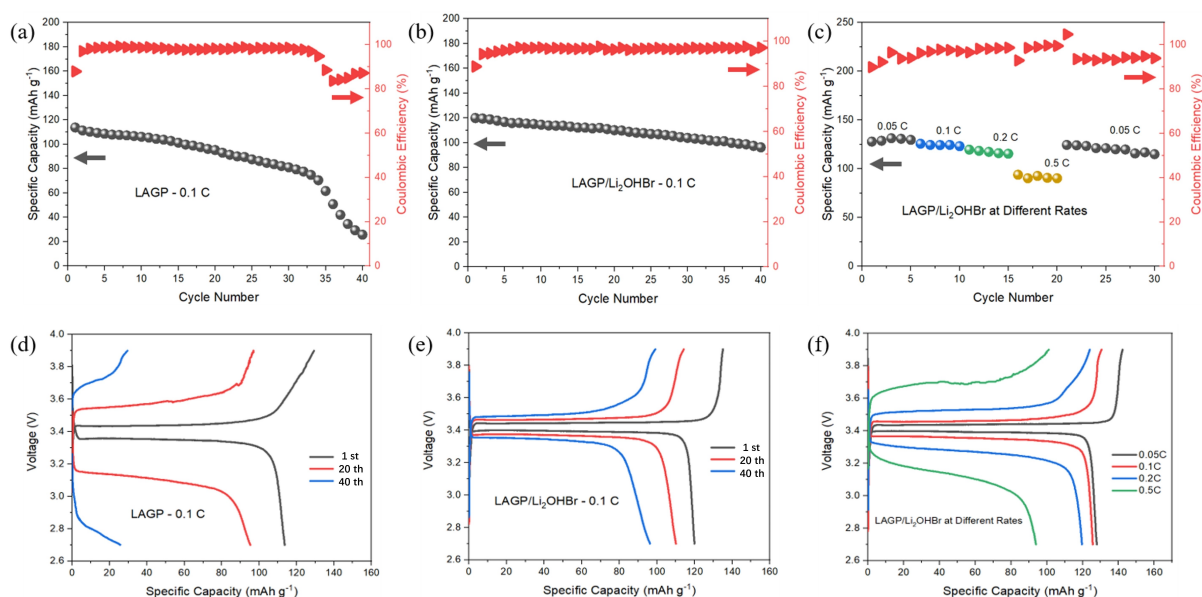


Figure 6. a) Cycling performance of LFP/Li ASSLB without Li₂OHBr layer at 0.1 C. b) Cycling performance of LFP/Li ASSLB with Li₂OHBr layer at 0.1 C. c) Cycling performance of LFP/Li ASSLB with Li₂OHBr layer at different rates. d) Voltage profiles of LFP/Li ASSLB without Li₂OHBr layer at 0.1 C. e) Voltage profiles of LFP/Li ASSLB with Li₂OHBr layer at 0.1 C. f) Voltage profiles of LFP/Li ASSLB with Li₂OHBr layer at different rates. All the above measurements were conducted on the battery tester at 80°C .

tested, and the specific capacity remains 124.5 mAhg^{-1} at 21th cycle after switching the rate from 0.5 C to 0.05 C (Figure 6c), which indicates that the LFP/Li ASSLB with Li_2OHBr layer can withstand moderate rate tests. Overall, according to the above cycle performances of the LFP/Li ASSLBs, it can be concluded that the introduction of Li_2OHBr layer in the interface between LAGP and Li metal can improve the cycling stability of LFP/Li ASSLBs from the perspective of Li metal anode protection. However, the discharge polarization is still inevitable after a long-term cycle of the ASSLBs due to the low ionic conductivity of the SSEs and the complex interface mechanism, pending further study and resolution.

3. Conclusions

In summary, taking the advantage of low melting point ($\sim 258^\circ\text{C}$) of the antiperovskite Li_2OHBr , the protective layer on the surface of LAGP pellet was formed by a simple method of rapid solidification of molten Li_2OHBr , and the dense and smooth Li_2OHBr layer was observed with SEM. The Li/Li symmetric cells with Li_2OHBr protective layer exhibit long-term cycle stability, compared with those using bare LAGP pellet. XPS analysis revealed the possible mechanism of the reaction between LAGP and Li metal, that is, Ge^{4+} is reduced to Ge^{3+} and Ge metal. Also, the stability of Li_2OHBr -protected LAGP against Li metal was verified by XPS analysis, which suggests that the ionic conductivity and electronic insulation of Li_2OHBr layer between LAGP and Li metal not only facilitate the shuttle of Li-ions, but also can effectively avoid the reduction of LAGP by Li metal. Furthermore, the assembled LFP/Li ASSLBs with Li_2OHBr layer remained the specific capacity of 96.3 mAhg^{-1} after 40th cycle at 0.1 C and 124.5 mAhg^{-1} at 21th cycle after switching the rate from 0.5 C to 0.05 C. This work shows that LiRAP Li_2OHBr possesses a considerable effect on Li-metal anode protection, and offers a straightforward strategy to protect SSEs that are unstable against Li metal.

Experimental Section

Preparation of Li_2OHBr -protected LAGP Pellet

The LAGP pellets were prepared by pressing the commercial LAGP powder (mean particle size: $0.3 \mu\text{m}$; purity: 99.9%) in the mold and sintering at 950°C for 10 h. The LiRAP- Li_2OHBr was prepared by simply grinding stoichiometric amounts of LiOH ($>99\%$) and LiBr ($>99\%$) and one-step sintering at 300°C for 1 h. Interestingly, it has been recently reported that Li_2OHBr was prepared by ball milling at room temperature, which indicates that the LiRAP family has a significant advantage in preparation compared to other electrolytes such as garnet-type and perovskite-type (more than 1000°C). Then, the prepared Li_2OHBr was heated to molten-state in a nickel crucible, and the sintered LAGP pellet was put into the molten liquid of Li_2OHBr and clamped out with tweezers (Figure 1S). Naturally, the Li_2OHBr -protected LAGP pellet was formed with the rapid solidification of molten- Li_2OHBr on LAGP surface at room temperature. Further, the Li_2OHBr -protected LAGP pellet was annealed at 240°C for 10 h to densify the Li_2OHBr layer and

optimize the interface between Li_2OHBr and LAGP. All the above preparations were performed under Ar atmosphere.

Material Characterizations

X-ray diffraction (XRD) measurements were conducted on a Rigaku SmartLab 9 kW diffractometer with Cu $\text{K}\alpha 1$ radiation ($\lambda = 1.5406 \text{ \AA}$) to identify the phase of the powder and pellet samples. Before measurements, the Li_2OHBr powder and Li_2OHBr -protected LAGP pellet were sealed with Kapton film under the inert atmosphere for the hygroscopicity of LiRAP. Differential Scanning Calorimetry (DSC) analysis was conducted on a NETZSCH STA 449 instrument, and samples were sealed in aluminum crucible under N_2 atmosphere with a heating and cooling rate of $5^\circ\text{C}/\text{min}$ from 30 to 280°C . The field emission scanning electron microscopy (SEM, Hitachi SU8230) was used to observe the surface and cross-sectional morphology of the solid electrolyte. The stability of the interface between SSEs and Li metal was analyzed by the X-ray photoelectron spectroscopy (XPS, PHI 5000 VersaProbe II) with Al $\text{K}\alpha$ X-rays. All the preparations of samples and measurements were performed under Ar atmosphere.

Electrode Preparation and Cell Assembly

The cathodes for the ASSLBs were prepared by mixing the commercial LiFePO_4 (LFP), super P, PVDF, and LITFSI (60:12:20:8 by weight) in NMP to form a uniform slurry. Then the slurry was cast on a carbon-coated aluminum foil to produce an electrode film with an active material loading of around $1.29 \text{ mg}/\text{cm}^2$, which was dried at 80°C for 12 h under Ar atmosphere. The cathode films were cut into 10 mm diameter discs, and 2032 coin-type cells were fabricated in an Ar-filled glove box with Li foil as the anode. It should be noted that the composite membrane (CM) was introduced in the interface to realize intimate interfacial contact between cathode and SSEs, which was prepared by mixing the ZIF-67, PEO, and LITFSI (4:8:1 by weight) in Acetonitrile and casting on the side (no Li_2OHBr layer) of the SSEs with enough drying time ($\sim 12 \text{ h}$) at 60°C . Here, ZIF-67 was chosen as the component of the composite membrane in order to increase the mobility of PEO chains, and ZIF-67 possesses a high surface area, rich porosity, and uniform channels, which can effectively benefit the Li^+ plating/stripping between the interface.^[27]

Electrochemical Measurements

Ionic conductivities were obtained from electrochemical impedance spectroscopy (EIS) in the temperature range of $25\text{--}100^\circ\text{C}$, using an electrochemical workstation analyzer (AUTOLAB M204) at frequencies ranging from 1 Hz to 1 MHz. During the EIS measurements, both sides of the sintered LAGP pellets were coated with sputtered Au films as the current collectors. Li_2OHBr powders were hot-pressed into pellets (diameter = 10 mm) in an insulative die with an annealing temperature of 120°C for 10 h, and two stainless-steel rods were clamped on both sides of Li_2OHBr as the current collectors. Galvanostatic cycling of symmetric batteries (Li/ Li_2OHBr -LAGP/Li) were sealed in 2032 coin-type cells and carried out at different current densities (0.05, 0.10, 0.15, $0.20 \text{ mA}/\text{cm}^2$) after heating at 130°C for 10 h. When testing the impedance of Li/Li symmetric cells, the battery tester will be automatically suspended at the time set by the computer program, and the cells were always placed in an 80°C oven during the measurement. Galvanostatic charge-discharge cycling of all-solid-state $\text{LiFePO}_4/\text{Li}$ batteries was carried out between 2.7 and 3.9 V at a rate of 0.1 C. For the rate performance, ASSLBs were tested at different rates of 0.05, 0.1, 0.2, and 0.5 C. All the above measurements were

conducted on the Wuhan Land battery tester at 80 °C, and all cells were laid aside at this temperature for 6 h before testing.

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Conflict of Interest

The authors declare no conflict of interest.

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