

## Misconception:

- The Concept between Strong & Concentrate:
  - | [ **Concentrate Acid** <> **Strong Acid** ]
  - “ **Concentrated** → **Molarity** ”
  - ◆ NOT HAVE THE CONCEPT ABOUT [ **Strong / Weak** ]
  - “ **Strong / Weak** → Level of **Ionization / Dissociation** ”
  - ◆ Example: “ Which of the following acid is Strong Acid ? ”  
[ 1M of  $\text{CH}_3\text{COOH}$  ] [ 12M of  $\text{CH}_3\text{COOH}$  ] [ **0.001M of HCl** ] [ 15M of Sulphurous Acid ]
- [ True & False ]: All Acid can't react with copper.
  - False: The **Concentrated** Acid which
  - ◆ have **Strong Oxidizing properties** |  $\text{H}_2\text{SO}_{4(l)}$  &  $\text{HNO}_{3(l)}$  | → Can React with Cu

