

We also need to balance redox reactions.



- ① Pull out half reactions.
(the elements are the same)



- ② Balance w/ regards to mass...

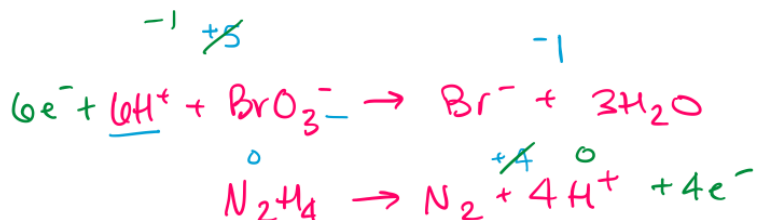
A Balance weird

B Balance O w/ H_2O

C Balance H w/ H^+



- ③ Balance w/ respect to charge. Add e^- .

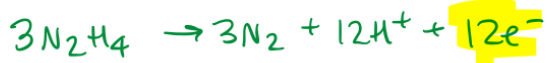
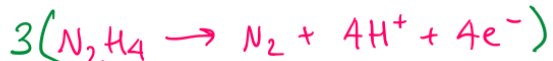


★ Electrons must go on opposite sides!

★ Red gains e^- (reactant)

★ Ox loses e^- (product)

- ④ Now balance electrons between 2 half reactions.



- ⑤ Recombine, cancelling where you can.

- ⑥ Check balance w/ mass & charge.