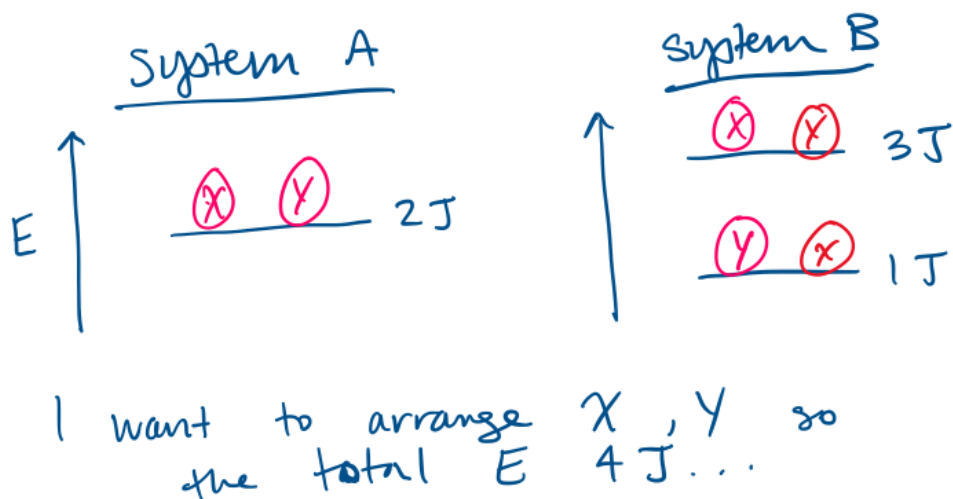


CHEM 1032 – Week 4 Questions

1. What is the boiling point of the salt solution? $K_b = 0.512\text{ }^{\circ}\text{C/m}$
2. If a blood cell is placed in pure water...what would happen?
 - a. Swell and burst
 - b. Shivel up
 - c. Nothing
3. What are the units of Entropy?
4. Which system has a higher entropy?



5. Predict the sign of ΔS_{rxn} for the equation below....
 $\text{CaCO}_3 (\text{s}) \rightarrow \text{CaO} (\text{s}) + \text{CO}_2 (\text{g})$
6. Let's say a reaction releases 10 J of heat. The reaction is performed at $-40\text{ }^{\circ}\text{C}$, $0\text{ }^{\circ}\text{C}$, and $40\text{ }^{\circ}\text{C}$. Which reaction will see the smallest increase in entropy of the surroundings?
7. Which compound has the highest standard entropy, S° ?
 - a. $\text{H}_2 (\text{g})$
 - b. $\text{NH}_3 (\text{g})$
 - c. $\text{Br}_2 (\text{l})$
 - d. $\text{NaCl} (\text{s})$