Exp. No. 5	Experiment/Subject	Date 10/10/2023	
Name Emmeankailash Ramesh	Lab Partner	Locker/ Desk No. 24	Course & Section No. 213 203

Empirical formula of zinc iodide.

Reference: "An Experiment in thinking of Scientifically", D.J Sardella Chem Ed. 69.933 (1992).

Purpose: "to determined the Empirical formula of Zinc iodide, Binary Compound form product that been produced by react between Zinc and iodide.

Materials - Electronic balance, ~~stand~~, volumic flask, Breaker, hood, zinc, idonic Electra heater.

Produre.

1. Record the mass of dry (PED) directly on notebook.
2. Weight the 2.00 grams of zinc and add PED. Weight PED with 2.00 grams.
3. Calculet the differnt mass of PED before and after.
4. Weight Vail Containing the idoine crystal (Put the cap on)
5. Tranfer the idoine Crystal into (PED) that haue zinic
6. Weight the (PED) with zine an idoine.
7. Determined the zine crystal.
8. take (PED) with iodide to fume hood and observee the physical changes.
9. Add 5mL distilled water to (PED) that contioned iclorde and swizde

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obsere it.

10. In the fanel hood Add 6-Maeetric solation to iodine and water

11. Obsere an chemical change (it take 10 minutes)

12. decco clean (PED) with zinc idode three time with 1mL prption distalled water with two or three drop of 6·M acide solution

13 drg the PED by and place it on hot plate (if drg oven is full)

14. drg zinc and rest cintill it reach room temporture

15. Weight (PED) with drg and unreacted zinc

16. deterime the mass of unreacted zinic

17. deterime the mass of zinc that react to iodide.

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Mass of Evaporating Dish (g)	41.786 g
Mass of Evaporating Dish + Zinc (g)	42.787 g
Mass of Zinc (g)	1.001 g
Mass of Viral filled with iodine (g)	6.765
Mass of Viral Empty (g)	5.760
Mass of iodine (g)	1.005
Mass of Evaporating Dish + Unreacted zinc	42.522 g
Mass of unreacted zinc (g)	0.736 g

Calcation

Mass of zinc

41.786 - 42.787 = 1.001

Mass of iodine

6.765 - 5.760 = 1.005

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