Exp. No.	Experiment/Subject	Date	
Name	Lab Partner	Locker/ Desk No.	Course & Section No.

Raw and Processed data for Glassware Analysis

	25mL Graduated cyld	10ml Vol pipette	50ml Burette
Room Temperature	23	23	23
Density of Water	0.997	0.997	0.997
Vol Pipette Reading mL		10mL	
Burette Initial Reading mL		24.7 mL	
Burette Final Readings		15.3 mL	
Measured Volume of H20 mL	10mL	10mL	9.4mL
Mass H20 + Glassware(g)	21.4341g	81.576	82.069g
Calculated Mass H20g	10.146 grams	10.032grams	10.525
Theoretical Vol H20(mL)	10.176	10.062	10.556
Error of Volume	17.60	0.21	5.62
Relative Uncertainty	1.76	.621	5.62

ND

Signature	Date	Witness/TA ND	Date

Exp. No.	Experiment/Subject	Date	
Name	Lab Partner	Locker/ Desk No.	Course & Section No.

Obsrvations: Same breaker Lane different measurment in the balance Which include How weights can vaoriouss depend on way and tools that you use. the measurment are very relatey close.

Average: $\frac{11.197+11.195+11.194}{3}$, $\frac{71.544+71.543+71.545}{3}$

$= 11.195$ $= 71.543$

SD $\frac{\sqrt{(11.197-11.195)^2+(11.195-11.195)^2+(11.94-11.195)^2}}{3-1}$ $\frac{\sqrt{(71.544-71.543)^2+(71.543-7153)^2+}}{3-1}$

$= 0.0012$ $= 4.71\times10^{-4}$

Mass H_2O $= 21.341-11.195 = 10.146$ grams , $81.576-71.543 = 10$ 32, 82.06 - 71.543 $= 10.5$

density $= \frac{10}{10.146} = 1.0146$, $\frac{10}{10.032} = 1.0032$, $\frac{10}{10.825} = 1.052$

TV $\neq \frac{10.146}{1.0146} = 10$, $\frac{1.0032}{1.0032} = 10$, $\frac{10.556}{1.052} \neq 10$

Error V $= \frac{(10.176-10)*100}{17.60}$, $\frac{10.062-10*100}{6.21}$, $\frac{10.556-10*100}{55.62}$

RV $\frac{17.60}{10} = 1.760$ $\frac{6.21}{10} = .621$ $\frac{55.62}{10}$ 5.562

Signature	Date	Witness/TA ND	Date