Exp. No.	Experiment/Subject	Date	
Name	Lab Partner	Locker/ Desk No.	Course & Section No.

Always keep the waste in a breaker.

first titration. "Scout titration

Add 10 mL oxcalic solution furtet into clean 125 mL Erlenmeyer. touch the the last drop on the buret, rise is int flash.

(b) Add 2 drops of the phenolphalen indicator solution

(c) Add the basic soultion in first. decrease the rate of NaOH until into pink and mix it.

d. Record the second buret section

(e) Repeat the titration

(f) Calcate molarity and if ~~is~~ isn't within 1% do a third titration.

(3) determination of Acetic Acid in Vinegar

(a) Empty the buret that have oxalic acid

(b) Rinse the buret and fill it with vinegar.

(c) do the same prodcere as 2 but with vinegar instead of oxalic acid

CMW of acitic acid is 60.05mL.

Signature	Date	Witness/TA	Date

Exp. No.	Experiment/Subject	Date	
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Table ①

	Trail One	Trail two	trial three -
Volume of Oxalic Acid (mL)	10mL	10mL	10.1mL
Moles of Oxalic Acid (moles)	0.0062	0.0062	0.0062
Inital Volume of NaOH (mL)	47.20	48.15	41.90
Final Volume of NaOH (mL)	28.15	41.90	31.95
Delivered Volume of NaOH (mL)	19.05	7	9.95
Moles of NaOH (moles)	0.0346	0.012	0.0196
Molarity of NaOH (M)	2.17	1.6	0.889
Average Molarity NaOH	19.0589	19.0589	19.0589

0.009 %

Table ②

	Trail one	Trail two	trial three.
Volume of Acetic Acid Sol'n (mL)	10mL	10mL	10mL
Mole of Acectic Acid Sol'n (moles)	.1315	.1315	0.1315
Inital Volume of NaOH (mL)	31.95	23.75	15.35
Final Volume of NaOH (mL)	23.75	15.35	5.90
Delievered Volume of NaOH (mL)	8.2	8.4	9.45
Average Molarity NaOH (m) * See Table 1	0.0208	0.021	0.02375
Molarity of Acetic Acid (M)	0.1617	.16569	-180-
Average Molarity of Acetic acid (m)	16.17	+ 16.56	+ 18.0
		50.73	

Signature	Date	Witness/TA MD	Date

Note: Place fold-over back cover under copy sheet before writing

Exp. No.	Experiment/Subject	Date	
Name	Lab Partner	Locker/ Desk No.	Course & Section No.

Caclution:

Mole. W. Oxalic acid dihydrate = 126.1 g/mol

$10.05\,mL \cdot \frac{7.89\,g}{1\,mL} = \frac{7.89}{126.1} \times 0.062$

$19.05\,mL \cdot \frac{0.789}{1} = \frac{15.03045}{40.00}$

$= .375$

$7 \cdot \frac{0.789}{1} = \frac{5.523}{4000}$

$= 5.523$

$9.95 \cdot \frac{0.789}{1} = \frac{7.8505}{40.00}$

$= .1966$

Observation:

Oxiglic Acid is clear in begin of Add the NaOH However it have turning pink when the devirded NaOH reach near to mL. I see trend that as you get closer to 10mL it have turn dark pink but if below 7 or 6 its light pink.

Vinegar is yellow but turns pink and same trend as Oxilic Acid.

Signature	Date	Witness/TA	Date

Exp. No. 7	Experiment/Subject Volumetric Analysis: Titration of Analysis	Date	
Name	Lab Partner Emmeenkailash	Locker/ Desk No.	Course & Section No. 213 203

Volumetric Analysis: Titration of Analysis

Refrence: "An Experiment in Thinking of Scientifially", D. J Saordella Chem Ed. 69. 333 (1992)

Purpose: determined the concentration of acetic acid in a Vinegar.

Material: buret, breaker, Cylinder, Volumetric flash, ring stand

Produre: Caution: Oxalic Acid is corrosive

1. Preparation of Oxalic Standard solution

a. Weigh - 3g of pure oxalic acid crystal (Mol Wt = 126.1 g.mol

b. transfer crystals to 100mL volumetric flass and Add 75mL of d-Swater mix it unit soulution is complete.

C. Wait till its dissolve and fill the volumetric flash to the mark with d-S Water, add the portion. Stopper the flash and mix.

2. Standardization of NaOH Solution

a. Rinse two burets with tap water and d-Swater and dry it.

b. Rinse one buret with NaOH solution completely the buret.

C. fill the buret with NaOH solution full up to 0mL. full entrie of buret without passing 0.00mL.

d. Rinse the Second buret with oxalic acid and fill it with standard oxalic Acid

Signature	Date	Witness/TA	Date