

Result: The experiment was able to achieve its purposes as to determine the contraction of acetic in the Vinger. The experiment verificate the legal requirement that 5 percent of acetic acid should be present in the Vinger. The indicator show how does it turns pink when it reaches amount of NaOH.

Discussion: Equivalence Point or stoichiometry point is a point in titration where the amount titration to completely neutralize the analyte solution. It is recognized by the change of color during the experiment. The weight percent acetic acid in vinegar solution of the experiment have been closer it closer but its precise percent as the legal percent. This experiment is a systematic deviation as the measurements and environments can change solely depending on the ways of recording or process during the experiment. The possible error on the amount of the control you have in the measurements of each instrument and the opening and closing of the Burette as it drops can change the results.

Conclusion: The purpose of the experiment is to determine the concentration of acetic acid in the Vinger through the process of the titration of the Vinger with reaction of the NaOH. It was anticipated to be in the 5% as it is the legal standard percent of the acetic acid in Vinger from the government whoever in the experiment is it various in the different trails of the titrations. It various may be because of the number of changes for trail or the error may have been acquired during the experiments. In the Future it may better if the precise of the measurements or calculation during the experiments.