The Equation of the hypothetical equation is y= 0.017(m) + 1.00 . The Slope and intercept have been a little different as there may be a slit different value recorded in the experiment, which can result in diverse types of slopes and intercepts. The Real Solution may have a different concentration, which can result in changes in the density that may cause a different slope and intercept also the experiment has water with no concentration which can two intercepts in the x values. Sources of Error in this experiment are the amount of the solution that has been measured as it may be different for each experiment as it is harder to control the amount of the solution that has been poured and the mass of equipment may be different for each can result in different values that can cause an error in the experiments.

Conclusion: The purpose of this experiment is to calculate a concentration with an unknown concentration, which can be calculated by having a linear equation between density and concentration that different values of data help precede an unknown solution concentration. In the experiment, it has been expected that density and concentration have a rate of exchange, which shows the trend of increase of density can result in increased concentration. It expects outcome and actual are related same it can be changed if an outsources can change a concentration without change in density. The sources of error can occur in the recording process or different values of the mass of the equipment as data can apply value as it changes for each data set. In future experiments we will try to measure as precisely as possible to the expected values so that can help to remove any types of error for values.