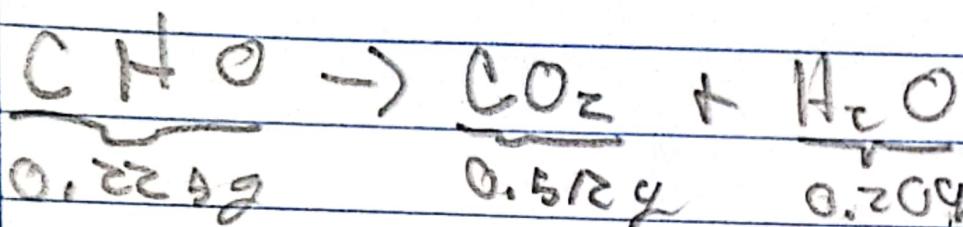


ΔH_{req}



$$H^? = 6.209 \left(\frac{\text{mol H}_2\text{O}}{18.02} \right) \left(\frac{\text{mol H}}{\text{mol H}_2\text{O}} \right) \left(\frac{1.015 \text{ H}}{1 \text{ mol H}} \right) = 6.02342 \text{ H}$$

$$C^? = 0.5122 \left(\frac{\text{mol CO}_2}{44.01} \right) \left(\frac{\text{mol C}}{\text{mol CO}_2} \right) \left(\frac{12 \text{ g C}}{\text{mol C}} \right) = 0.1862 \text{ C}$$
$$= \underline{0.2262}$$
$$\underline{0.015620}$$

$$\text{H} = \frac{0.0234}{1.01} = 0.0232 \text{ mol H} / 975 \times 10^{-6} = 29$$

$$\text{C} = \frac{0.186}{12.01} = 0.0155 \text{ mol C} / 975 \times 10^{-6} = 16 \rightarrow \text{C}_{16}\text{H}_{24}\text{O}_4$$

$$\text{O} = \frac{0.0156}{16} = 975 \times 10^{-6} \text{ mol O} / 925 \times 10^{-6} = 1$$

116 g/mol

$$29 * 1.01 = 29.29 \quad \underline{116} = 1.03 \quad 21 \Rightarrow \text{C}_{16}\text{H}_{24}\text{O}_4$$

$$16 * 12.01 = 72.06 \quad \underline{112.3}$$

$$1 * 16 = \underline{16} \\ \underline{112.3}$$