Mn_2O_7

Practiquemos formula empirica y formular:

- $C = 38.7g / 12.01 g = \boxed{3.22 \text{ mol } C}$

$H = 9.7g / 9.01g = 9.6 \text{ mol } H$

$O = 51.7g / 16g = 3.23 \text{ mol } O$

masa molar real $= 62.1 g/mol$

Formula empirica

$C = \dfrac{3.22}{3.22} = 1$; $H = \dfrac{9.6}{3.22} = 3$; $O = \dfrac{3.23}{3.22} = 1$; CH_3O

$C = 1 \times 12.01 = 12.01$
$H = 3 \times 1.01 = 3.03$
$O = 1 \times 16 = 16$

$\overline{\quad\quad\quad}$
31.04

$n = \dfrac{62.1}{31.04} = 2$; $(CH_3O)_2 = C_2H_6O_2$