Tarea

$\underbrace{CHO}_{0.222g} \rightarrow \underbrace{CO_2}_{0.512g} + \underbrace{H_2O}_{0.204g}$

$H? = 0.209 \left(\frac{1 mol H_2O}{18.02}\right)\left(\frac{2 mol H}{1 mol H_2O}\right)\left(\frac{1.01g H}{1 mol H}\right) = 0.0234g H$

$C? = 0.512g \left(\frac{1 mol CO_2}{44.01}\right)\left(\frac{1 mol C}{1 mol CO_2}\right)\left(\frac{12g C}{mol C}\right) = 0.1162g C$

$0.222g - 0.1162g = 0.0156g O$

H = 0.0234 / 1.01 = 0.0232 mol H / 975×10⁻⁶ = 24

C = 0.186 / 12.01 = 0.0155 mol C / 975×10⁻⁶ = 16 → $C_{16}H_{24}O_1$

O = 0.0156 / 16 = 975×10⁻⁶ mol O / 925×10⁻⁶ = 1

116 g/mol

24 * 1.01 = 24.24
16 * 12.01 = 72.06
1 * 16 = 16
112.3

116 / 112.3 = 1.033 1 ⇒ $C_{16}H_{24}O_1$