ADDITIONAL MULTIPLE CHOICE QUESTIONS

1. **Non-metals are essential**
   (a) for the maintenance (b) for the existence
   (c) for the safety of life (d) all of these
2. **It has the highest percentage in earth's crust and oceans.**
   (a) calcium (b) carbon (c) oxygen (d) nitrogen
3. **The electronegative and non-metallic character from top to bottom**
   (a) increases (b) decreases (c) remain same (d) stable
4. **Chlorine only reacts with methane in**
   (a) darkness (b) sunlight (c) yellow light (d) screened light
5. **Chlorine has colour:**
   (a) pale yellow (b) greenish yellow (c) reddish brown (d) purple black
6. **These are highly oxidizing agents**
   (a) alkali metals (b) alkaline earth metals
   (c) transition metals (d) halogens
7. **The highest electronegative element in periodic table is**
   (a) fluorine (b) chlorine (c) bromine (d) iodine
8. **The electronic configuration of halogens is**
   (a) $ns^2\ np^5$ (b) $ns^2\ np^3$ (c) $ns^3\ np^5$ (d) $ns^2\ np^2$
9. **Non-metals are electronegative in nature and form oxides.**
   (a) acidic (b) basic (c) neutral (d) suboxide
10. **It is used to make jewelry items because of its unique characteristics like colour beauty, strength, flexibility and resistance to tarnish.**
    (a) gold (b) platinum (c) silver (d) copper
11. **White gold is an alloy of**
    (a) gold + palladium + zinc (b) palladium + silver + nickel
    (c) gold + silver + copper (d) palladium + nickel + zinc
12. **The alkaline earth metals are smaller in size and have more**
    (a) ionization energy (b) electron affinity (c) nuclear change (d) electropositive
13. **All alkali metals have the largest size and the lowest in their respective periods.**
    (a) electro negativity (b) ionization energy (c) electron affinity (d) electro positivity
14. **Metals have the tendency to lose their valence electron. This property of a metal is termed as**
    (a) electro negativity (b) electro positivity (c) electron affinity (d) ionization power
15. **Copper, mercury, silver and gold are the examples of metals**
    (a) very reactive (b) moderately (c) least reactive (d) none of these
16. **Cation in formed, when an element – electron to its outermost shell**
    (a) loses (b) gains (c) donates (d) shares
17. **Metals which are easily oxidized are said**
    (a) negative metals (b) state metals (c) reactive metals (d) non-reactive metals
18. **A metal in a compound always exists in which oxidation sate**
    (a) negative (b) positive (c) neutral (d) zero
19. **Ionization energy of sodium is less than**
    (a) aluminum (b) magnesium (c) copper (d) all of these

20. All metals are solids except
(a) sodium (b) magnesium (c) mercury (d) gold

21. The most valuable metal among the following is
(a) Gold (b) Uranium
(c) Osmium (d) Rubidium

22. The heaviest metal is
(a) Uranium (b) Gold
(c) Osmium (d) Calcium

23. Lithium has density
(a) $0.53gcm^{-3}$ (b) $1.53gcm^{-3}$
(c) $15.3gcm^{-3}$ (d) $3.5gcm^{3}$

24. Mark which one is non-metal?
(a) Sodium (b) Calcium
(c) Nitrogen (d) Mercury

25. Which one of the following will not react with HCl?
(a) Carbon (b) Silver
(c) Zinc (d) Copper

26. Sodium does not react with
(a) Carbon (b) Nitrogen
(c) Hydrogen (d) Both a and b

27. Which metal burns with golden yellow flame
(a) Calcium (b) Barium
(c) Sodium (d) Potassium

28. Which metal is used in Thermite process to ignite aluminum powder
(a) Na (b) Mg
(c) Ca (d) Be

29. Silver is get tarnished by
(a) Atmosphere oxygen (b) Nitrogen
(c) Hydrogen sulphide (d) Carbon dioxide

30. Which metal is used for making mirrors?
(a) Lead (b) Iron
(c) Silver (d) Lithium

31. Which one of the following metal has yellow colour
(a) Lead (b) Gold
(c) Iron (d) Potassium

32. Gold is always alloyed with one, among the following metals
(a) Sodium (b) Mercury
(c) Copper (d) Calcium

33. Sodium is used for the refining of which metal
(a) U (b) Zn
(c) Ti (d) Zr

34. How many series of d-block elements are there in the periodic
(a) Three (b) Four
(c) Five (d) Two

35. Three transition elements Cu, Ag and Au constitute group number
(a) 9 (b) 10
(c) 11 (d) 12

36. The least electronegative element among the halogens is
(a) Fluorine (b) Chlorine (c) Bromine (d) Iodine

37. Which of the following halogen has pale yellow colour?

(a) F_2 (b) Cl_2 (c) Br_2 (d) I_2

38. Which one of the following makes covalent bond with halogens

(a) Na (b) K (c) O (d) Mg

39. Identify among the following which one is semimetal

(a) Lead (b) Zinc (c) Silicon (d) Galium

40. The general electronic valence shell configuration of alkali metals is

(a) ns^2 (b) ns^1 (c) ns^2, np^1 (d) ns^2, np^2

41. Hydrogen is released when water reacts with

(a) Na (b) Mg (c) K (d) All of them

42. Which metal releases electron from its outermost shell the most easily

(a) Na (b) K (c) Mg (d) Ca

43. Which one among the halogens has least affinity with hydrogen?

(a) F_2 (b) Cl_2 (c) Br_2 (d) I_2

44. In diffused sun lights chlorine reacts with methane to form

(a) CH_3Cl (b) CH_2Cl_2 (c) CCL_4 (d) All of above

45. In the earth crust the highest %age of oxygen is

(a) 47% (b) 86% (c) 90% (d) 24%

46. The %age of oxygen in oceans is:

(a) 47% (b) 86% (c) 90% (d) 24%

47. The %age of oxygens in air is:

(a) 21% (b) 24% (c) 26% (d) 30%

48. Which one of the following will form amphoteric oxides?

(a) Na (b) N (c) Si (d) S

49. Which halogen reacts with water in dark and cold state

(a) F_2 (b) Cl_2 (c) Br_2 (d) I_2

50. Indicate the least reactive metal among the following

(a) Gold (b) Zinc (c) Iron (d) Tin

ANSWER KEY

1	d	11	d	21	b	31	b	41	d
2	c	12	b	22	c	32	c	42	b
3	b	13	a	23	a	33	c	43	d
4	b	14	b	24	c	34	a	44	d
5	b	15	c	25	a	35	c	45	a
6	d	16	a	26	d	36	d	46	b
7	a	17	c	27	c	37	a	47	a
8	c	18	b	28	b	38	c	48	c
9	a	19	d	29	c	39	d	49	a
10	b	20	c	30	c	40	b	50	a