

## ADDITIONAL MULTIPLE CHOICE QUESTIONS

- Symbol of hydronium ion is:**  
(a)  $H^+$  (b)  $OH^-$  (c)  $H_3O^+$  (d) none of these
- An example of non electrolyte is:**  
(a) Glucose (b) aq NaCl (c) HCl (d)  $H_2SO_4$
- An example of weak electrolyte is :**  
(a)  $HNO_3$  (b) HCl (c)  $H_2SO_4$  (d)  $H_2CO_3$
- During electrolysis \_\_\_\_\_ takes place at anode.**  
(a) Catenation (b) oxidation (c) reduction (d) addition
- Which is not the characteristic of electrolyte?**  
(a) Cheap (b) conductor (c) easily oxidized (d) soluble in water
- In dry cell \_\_\_\_\_ acts as cathode:**  
(a) Zn cup (b) graphite rod (c) paste (d) steel rod
- In Zn-Cu galvanic cell, Zn is dipped in:**  
(a)  $ZnSO_4$  (b)  $Zn(NO_3)_2$  (c)  $CuSO_4$  (d) both a & b
- In Zn-Cu galvanic cell, Zn is used as:**  
(a) Cathode (b) anode (c) electrode (d) all of above
- In  $Zn + Cu^{+2} \longrightarrow Zn^{+2} + Cu$ , Zn is:**  
(a) Oxidized (b) reduced (c) redoxed (d) decomposed
- Anions are \_\_\_\_\_ ions.**  
(a) Positive (b) neutral (c) negative (d) amphoteric
- Cations are \_\_\_\_\_ ions.**  
(a) Negative (b) positive (c) neutral (d) none of these
- Electrolysis of NaCl is done in the cell:**  
(a) Electrolytic (b) voltaic (c) down's (d) faradays'
- Which one is not produced during electrolysis of aqueous NaCl?**  
(a)  $Na^+$  (b)  $OH^-$  (c)  $H_2O$  (d)  $Cl^-$
- In pure water, out of \_\_\_\_\_ only one molecule ionizes.**  
(a)  $6 \times 10^6$  (b)  $6 \times 10^8$  (c)  $6 \times 10^{16}$  (d)  $6 \times 10^{12}$
- Which one is weak electrolyte?**  
(a) Citric acid (b) carbonic acid (c) tartaric acid (d) all of these
- Metallic, ammonium and hydrogen ions carry \_\_\_\_\_ charges.**  
(a) Negative (b) neutral (c) positive (d) partial positive
- Electrode through which electrons enter the external circuit:**  
(a) Anode (b) Cathode (c) electrode (d) none of these
- Electrode through which electrons leave the external circuit:**  
(a) Anode (b) Cathode (c) graphite (d) electrolyte
- Which one is conductor?**  
(a) Naphthalene (b) paraffin wax (c) plastic (d) HCl

20. Which solution is not a conductor:  
(a) NaCl (b) KI (c)  $\text{CO}(\text{NH}_2)_2$  (d)  $\text{CuCl}_2$
21. Rods through which electric current enters or leaves the cell:  
(a) Protons (b) electrons (c) electrodes (d) all of these
22. Spontaneous redox reaction produce current in:  
(a) Voltaic cell (b) electrolytic cell (c) galvanic cell (d) Both a & b
23. In an oxidation reaction electrons are:  
(a) Absorbed (b) lost (c) moved (d) increased
24. In reduction reaction electrons are:  
(a) Lost (b) absorbed (c) kept constant (d) all of these
25. Which of the following is a good electrolyte?  
(a) NaCl (b)  $\text{H}_2\text{SO}_4$  (c) NaOH (d) All of these
26. Which ionizes in small extent in water  
(a)  $\text{Ca}(\text{OH})_2$  (b) NaCl (c) NaOH (d)  $\text{H}_2\text{SO}_4$
27. Oxidation always takes place at  
(a) Anode (b) Cathode (c) Both of them (d) None of them
28. Who invented first electrolytic cell?  
(a) Berzelius (b) Volta (c) J. Dalton (d) Newton
29. In galvanic cell, cathode carries.  
(a) Positive charge (b) Negative charge (c) No charge (d) Neutral charge
30.  $\text{Cl}_2$  gas is formed, when  $\text{Cl}^-$  ions are  
(a) Reduced (b) Oxidized (c) Removed (d) Reacted with metals
31. Which ion is not formed during is of aqueous sodium chloride?  
(a)  $\text{H}^+$  (b)  $\text{H}_2$  (c)  $\text{Cl}_2$  (d) NaOH
32. Which medium accelerates the process of rusting?  
(a) Acidic (b) Basic (c) Buffer (d) Neutral
33. Stainless steel contains  
(a) Nickel (b) Iron (c) Chromium (d) All of these
34. Which of the followings is a corrosion resistant metal?  
(a) Fe (b) Zn (c) Sn (d) Sr
35. The electrolytic cell is made up of  
(a) Cement (b) Glass (c) Wood (d) All of these
36. Which of the following is a common example of silver plating?  
(a) Wares (b) Cutlery (c) Jewellery (d) All of these
37. Which metal has a great tendency to corrosion?  
(a) Silver (b) Copper (c) Aluminum (d) All of these
38. Chemical formula of sodium thio-sulphate is  
(a)  $\text{Na}_2\text{SO}_4$  (b)  $\text{Na}_2\text{SO}_3$  (c)  $\text{Na}_2\text{S}_2\text{O}_3$  (d)  $\text{NaSO}_4$
39. Sterling silver is an alloy of silver and  
(a) Iron (b) Copper (c) Chromium (d) Aluminum
40. In HCl, oxidation number of H is  
(a) -1 (b) +1 (c) +2 (d) -2



41. The oxidation number of all elements in free state is  
(a) One (b) Two (c) Three (d) Zero
42. The oxidation number of Group-I elements is  
(a) +1 (b) +2 (c) +3 (d) +4
43. The oxidation number of hydrogen in metal hydrides is  
(a) +1 (b) -1 (c) +2 (d) -2
44. The oxidation number of oxygen is +2 in  
(a)  $H_2O$  (b)  $OF_2$  (c)  $Na_2O$  (d)  $HNO_3$
45. The oxidation number of sulphur in  $H_2SO_4$  is  
(a) +1 (b) +4 (c) +6 (d) +8

**ANSWER KEY**

1	c	11	b	21	c	31	a	41	d
2	a	12	c	22	d	32	a	42	a
3	d	13	c	23	b	33	d	43	b
4	b	14	b	24	b	34	c	44	b
5	c	15	d	25	d	35	d	45	c
6	b	16	c	26	a	36	d		
7	a	17	a	27	a	37	c		
8	b	18	b	28	b	38	c		
9	a	19	d	29	a	39	b		
10	c	20	c	30	b	40	b		

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