

**ADDITIONAL MULTIPLE CHOICE QUESTIONS**

- The emulsions are the type of solution:**  
(a) Solid in solid (b) Liquid in solid (c) Gas in liquid (d) None of these
- In true solution, the particle are the type of solution:**  
(a)  $0.1\mu$  (b)  $10.0\mu$  (c)  $0.5\mu$  (d)  $1.0\mu$
- Which property is not for suspensions?**  
(a) Homogeneous (b) Cannot be seen by naked eye  
(c) Not a true solution (d) Solid in liquid
- The concentration of the solute in solution, when it is in equilibrium with solid substance, at a particular temperature is called:**  
(a) Molarity (b) Dilution  
(c) Colloidal solution (d) Supersaturated solution
- A solution containing relatively higher concentration of solute is called:**  
(a) Dilute solution (b) Saturated solution (c) Concentrated solution (d) Suspension
- Fog is an example of solution:**  
(a) Gas in liquid (b) Liquid in gas (c) Liquid in gas (d) Solid in liquid
- Emulsions are the colloidal dispersion of liquid in:**  
(a) Solid (b) Gas (c) liquid (d) Water
- The suspension particles have appearance:**  
(a) Uniform (b) Transparent (c) Opaque (d) Both 'A' and 'B'
- Starch, glue are the examples of:**  
(a) Colloidal solution (b) Solution (c) Colloids (d) Crystalloids
- Solution which can dissolve further amount of a solute at particular temperature is called:**  
(a) Saturated solution (b) Unsaturated solution  
(c) Colloidal solution (d) Supersaturated solution
- The solution which can easily pass through parchment membrane is considered as:**  
(a) Colloidal solution (b) True solution (c) Suspensions (d) Crystalloids
- Change of temperature can change the \_\_\_\_\_ of a solute in a solvent.**  
(a) Dilution (b) Solubility (c) Molarity (d) Dipole moment
- The ionic and polar compounds like NaCl and HCl are more soluble in water than non-polar covalent compounds like.**  
(a)  $\text{CCl}_4$  (b) Benzene (c)  $\text{CS}_2$  (d) All of these
- Which one produced colloidal solution**  
(a) Blood (b) Cooper Sulphate solution  
(c) Silver nitrate solution (d) None of these
- Concentration is most often expressed as the ratio of the amount of \_\_\_\_\_ to the amount of solution.**  
(a) Solute (b) solvent (c) brine (d) salt
- Number of moles of solute per  $\text{dm}^3$  of the solution is called:**  
(a) Molarity (b) molality (c) normality (d) density
- Sodium amalgam which two metals as present**  
(a) Na and Mg (b) Na and  $\text{M}\mu$  (c) Na and Hg (d) Zn and Co
- The concentrated solution of NaCl is called:**  
(a) Fluid (b) Brass (c) Brine (d) Plasma
- Brass is a familiar alloy, made of:**  
(a) zinc + carbon (b) Zinc + copper (c) copper + iron (d) zinc + lead
- The homogeneous mixture of two or more compounds is called:**  
(a) solute (b) solvent (c) solution (d) fluid

21. Which of the following solutions contains least amount of water?  
(a) 2M (b) 1M  
(c) 0.5 M (d) 0.25M
22. Opal is an example of solution  
(a) Liquid in solid (b) Solid in gas  
(c) Solid in solid (d) Gas in liquid
23. 0.1M solution is diluted to ten time, its new molarity is:  
(a) 0.01M (b) 0.9M  
(c) 0.2M (d) 0.1M
24. Which one of the following is true solution?  
(a) Paints (b) Alcohol in water  
(c) Milk (d) Milk of magnesia
25. Heat is absorbed on dissolving which one of the following salts:  
(a) NaCl (b)  $\text{Ce}_2(\text{SO}_4)_3$   
(c)  $\text{NaNO}_3$  (d)  $\text{Li}_2\text{SO}_4$
26. Smoke is an example of solution  
(a) Solid in gas (b) Gas in liquid  
(c) Liquid in solid (d) Gas in solid
27. KCl is soluble in  
(a) Water (b) Benzene  
(c) Ether (d) Kerosene oil
28. Naphthalene is soluble in  
(a) Water (b) Ether  
(c) Carbon tetrachloride (d) Both (b) and (c)
29. Blood is an example of  
(a) True solution (b) Suspension  
(c) Colloid (d) None of all
30. 20 g of NaOH has been dissolved in  $0.5\text{dm}^3$  of the solution its molarity is:  
(a) 0.1M (b) 1.0M  
(c) 0.5M (d) 1.5M
31. What mass of NaOH is required to prepare 1M  $500\text{cm}^3$  of the solution:  
(a) 10g (b) 20g  
(c) 30g (d) 40g
32. What volume of 2M solution of  $\text{H}_2\text{SO}_4$  is required to prepare  $500\text{ cm}^3$  0.1M solution?  
(a)  $10\text{cm}^3$  (b)  $15\text{cm}^3$   
(c)  $20\text{cm}^3$  (d)  $25\text{cm}^3$
33. 2g of NaCl has been dissolved in 20.0g of  $\text{H}_2\text{O}$ , its % m/m is  
(a) 9.09% (b) 19.09%  
(c) 1.9% (d) 10.0%
34. Identify among the following which one is universal solvent?  
(a)  $\text{CH}_3\text{COOH}$  (b) Liquid ammonia  
(c) Water (d) Ethanol
35. Which one of the following salts gives out heat on dissolving in water?  
(a) NaCl (b)  $\text{Ce}_2(\text{SO}_4)_3$   
(c)  $\text{KNO}_3$  (d) KCl
36. Mist is an example of solution  
(a) Liquid in gas (b) Gas in liquid  
(c) Solid in gas (d) Gas in solid
37. Which one of the following is a liquid in solid solution?  
(a) Sugar in water (b) Butter



- (c) Solvent to solution (d) Fog
38. Which one of the following solutions contains more water?  
(a) 2 M (b) 1 M  
(c) 0.5 M (d) 0.25 M
39. Tynadall effect is shown by  
(a) Sugar solution (b) Jelly  
(c) Paints (d) Chalk solution
40. Tyndall effect is due to  
(a) Blockage of beam of light (b) Non-scattering of beam of light  
(c) Scattering of beam of light (d) Passing through beam of light
41. If  $10\text{cm}^3$  of alcohol is dissolved in 100g of water, it is called  
(a) % m/m (b) % m/v  
(c) % v/m (d) % v/v
42. Sea water is a source of  
(a) Fog (b) Opal  
(c) Soft drink (d) Cheese
43. Which one of the following solutions contains maximum amount of water?  
(a) 2M (b) 1M  
(c) 0.5 M (d) 0.25 M
44. Brass is a solid solution of  
(a) Zn and Cd (b) Zn and Cu  
(c) Zn and Hg (d) Zn and Mg
45.  $10\text{cm}^3$  of 0.01 molar  $\text{KMnO}_4$  solution has been diluted to  $100\text{ cm}^3$  its molarity will become  
(a) 0.001 M (b) 0.1 M  
(c) 0.01M (d) 1 M

**ANSWER KEY**

1	b	11	b	21	a	31	b	41	c
2	a	12	b	22	c	32	d	42	b
3	b	13	d	23	a	33	a	43	d
4	d	14	a	24	b	34	c	44	b
5	c	15	a	25	c	35	b	45	a
6	b	16	b	26	a	36	a		
7	a	17	c	27	a	37	d		
8	c	18	c	28	d	38	d		
9	c	19	b	29	c	39	b		
10	b	20	c	30	b	40	c		