



## Smart Guess & Test Papers

Student Name \_\_\_\_\_ Father Name \_\_\_\_\_ Roll Number \_\_\_\_\_

Class: 1st /year - Chemistry Marks : 42 Exam Format : Chapter Wise MCQs

Time : notespk.com\_Nauman Sadaf | Date \_\_\_\_\_ Examiner Sig \_\_\_\_\_ Chapter#: 1

MCQ's		S/Q		L/Q		Total	
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### Objective Type

#### 1. Encircle the Correct Option. (1 x 32 = 32)

1. درست جواب کے گرد دائرہ لگائیں۔

1) Isotopes differ in

a) Properties Which depends upon mass	b) Arrangement of electrons in orbitals	c) Chemical properties	d) The extent to which they may be effected in electromagnetic field
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2) 27 g of Al will react completely with how much mass of O<sub>2</sub> to produce Al<sub>2</sub>O<sub>3</sub>.

a) 8 g of oxygen	b) 16 g of oxygen	c) 32 g of oxygen	d) 24 g of oxygen
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3) A limiting reactant is the one which

a) Is taken in lesser quantity in grams as compared to other reactants.	b) Is taken in lesser quantity in volume as compared to the other reactants.	c) Gives the maximum amount of the product which is required.	d) Gives the minimum amount of the product under consideration.
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4) Isotopes differ in

a) Properties which depend upon mass	b) Arrangement of electrons in orbitals	c) Chemical properties	d) The extent to which they may be affected in electromagnetic field
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5) The number of Isotopes of Gold ( Au ) is

a) 1	b) 3	c) 7	d) 11
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6) The number of natural Isotopes is

a) 280	b) 150	c) 300	d) 400
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7) Isotopes differ in

a) Number of atoms	b) Number of Ptoion	c) Number of electron	d) Number of neutron
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8) Which of the following statement is not true.

a) Isotopes with even atomic masses are comparatively abundaat	b) Isotopes with odd atomic masses are comparatively abundant	c) Isotopes with even atomic masses and even atomic numbers are comparatively abundant	d) Isotopes with even atomic masses and oss atomic numbers are comparatively abundant
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9) The number of isotopes of calcium is

a) Six	b) Seven	c) Five	d) Nine
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10) Silver has \_\_\_\_\_ natural isotopes.

a) Nine	b) Two	c) Eleven	d) Sixteen
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11) Silver has \_\_\_\_\_ isotopes.

a) Nine	b) Ten	c) Eleven	d) Sixteen
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12) What is the maximum mass of chromium that can be extracted from 76g of Cr<sub>2</sub>O<sub>3</sub> ?

a) 48g	b) 52g	c) 104g	d) 152g
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13) Which of the following has highest mass at S.T.P ?

a) $6.02 \times 10^{22}$ molecules of CO <sub>2</sub>	b) 2 moles of CH <sub>4</sub>	c) 0.1 g mole of N <sub>2</sub> O	d) 22.4 dm <sup>3</sup> of CO <sub>2</sub> at . S.T.P.
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14) The number of atoms present in 0.5 moles of Na is

a) $1.0 \times 10^{23}$	b) $6.02 \times 10^{23}$	c) $2.04 \times 10^{23}$	d) $3.01 \times 10^{23}$
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15) The volume occupied by 16g of CH<sub>4</sub> at S.T.P.

a) 224.14 dm <sup>3</sup>	b) 22.414 dm <sup>3</sup>	c) 1.12 dm <sup>3</sup>	d) 2.24 dm <sup>3</sup>
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16) The volume occupied by 1.6 g of O<sub>2</sub> at S.T.P.

a) 24.4 dm <sup>3</sup>	b) 2.24 dm <sup>3</sup>	c) 1.12 dm <sup>3</sup>	d) 112 dm <sup>3</sup>
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17) The volume occupied by 1.4g of N<sub>2</sub> at S.T.P is

a) 24.4 dm <sup>3</sup>	b) 22.4 dm <sup>3</sup>	c) 1.12 dm <sup>3</sup>	d) 112 dm <sup>3</sup>
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18) What are the number of moles of hydrogen atom in 3.2g of CH<sub>4</sub> ? ( Relative atomic mass C = 12 )

a) 0.2	b) 0.4	c) 0.6	d) 0.8
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19) One mole of water contains .

a) 81 g water	b) $6.02 \times 10^{23}$ atoms	c) $6.02 \times 10^{23}$ ions	d) $6.02 \times 10^{23}$ molecules
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20) What is the ratio of volume of 2g of  $H_2$  to the volume of 16g  $CH_4$  both volumes are at STP

a) 1:8	b) 1:2	c) 1:1	d) 2:1
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21) Number of moles in 100g of  $KClO_3$  .

a) 0.76	b) 0.56	c) 0.014	d) 0.816
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22) Stoichiometry is a branch of \_\_\_\_\_

a) Physics	b) Chemistry	c) Biology	d) English
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23) Stoichiometry tells us the \_\_\_\_\_ relationship between reactants and products .

a) Qualitative	b) Quantitative	c) Both A & B	d) None of these
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24) In mass mole relationship if we are given with mass of substance the we can calculate \_\_\_\_\_ of other substance .

a) Mass	b) No. of electron	c) Moles	d) None of these
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25) In Mass-volume relationship if we are given with mass of one substance we can calculate \_\_\_\_\_ of other substance .

a) Mole	b) Volume	c) Both A & B	d) None of these
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26) In stoichiometric calculate all the reactants must converted into

a) Products	b) Reactants	c) Both A & B	d) Molar mass
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27) The reactant which consumes earlier and gives least quantity of product is called.

a) Reactant	b) Limiting reactant	c) Stoichiometry	d) Stoichiometric amount
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28) A limiting reactant is the one which

a) Is taken is lesser quantity to the amount in gram required	b) Is taken is lesser quantity in volume as per its required	c) Gives the maximum amount of the product required	d) Gives the minimum amount of the required product
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29) The amount of products obtained in a chemical reaction is called \_\_\_\_\_ yield of that reaction .

a) Required	b) Low	c) High	d) Actual
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30) The amount of products calculated from balanced chemical equation represents .

a) Actual yield	b) Theoretical yield	c) Both A & B	d) None of these
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31) In most of the reactions the actual yield is \_\_\_\_\_ than theoretical yield .

a) More	b) Less	c) None	d) Expected yield
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32) Many elements have fractional atomic masses . This is because .

a) The mass of the atom is itself fractional	b) Atomic masses are average masses of isobars	c) Atomic masses are average masses of isotopes	d) Atomic masses are average masses of isotopes proportional to their relative abundance
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2. Write "T" for a true statement and "F" for a false statement (1 x 4 = 4)

2. درست جواب کے سامنے ( ✓ ) نشان لگائیں اور غلط کے سامنے (X) کا نشان لگائیں۔

33) Neon has three isotopes and the fourth one with atomic mass 20.18 amu.

☐ True ☐ False

34) The number of atoms in 1.79 g of gold and 0.023 g of sodium are equal.

☐ True ☐ False

35) The number of electrons in the molecules of CO and  $N_2$  are 14 each, so 1 g of each gas will have same number of electrons.

☐ True ☐ False

36) Actual yield of a chemical reaction may be greater than the theoretical yield.

☐ True ☐ False

3. Fill in the blanks. (1 x 6 = 6)

3. خالی جگہ پُر کریں۔

37) The unit of relative atomic mass is \_\_\_\_\_.

38) The exact masses of isotopes can be determined by \_\_\_\_\_ spectrograph.

39) The phenomenon of isotopy was first discovered by \_\_\_\_\_.

40) A limiting reagent is that which controls the quantities of \_\_\_\_\_.

41) 4g of  $CH_4$  at  $0^\circ C$  and 1 atm pressure has \_\_\_\_\_ molecules of  $CH_4$

42) Stoichiometric calculations can be performed only when \_\_\_\_\_ is obeyed.