Platinum

Please note that the value in green is indicative. Values by Azaroual et al. (2001) come from a thorough literature analysis combined with new experiments. Values for $Pt(OH)_2$ are in line with data by Wood (1991), which is the reference considered by Brown and Ekberg (2016). Final differences come from the value of redox potential used to convert this constant to the hydrolysis constant in the two studies.

Equilibrium reactions	lg*K* at infinite dilution and *T* = 298 K	
	Azaroual et al., 2001	Brown and Ekberg, 2016
$Pt^{2+} + H_2O \rightleftharpoons Pt(OH)^+ + H^+$	10.91	
$Pt^{2+} + 2\ H_2O \rightleftharpoons Pt(OH)_2 + 2\ H^+$	0.81	−7.3 ± 1.0

M. Azaroual, B. Romand, P. Freyssinet, J.R. Disnar. Solubility of platinum in aqueous solutions at 25°C and pHs 4 to 10 under oxidizing conditions. Geochim. Cosmochim. Acta 65, 4453–4466 (2001).

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 739–740.

S.A. Wood, Experimental determination of the hydrolysis constants of Pt^{2+} and Pd^{2+} at 25°C from the solubility of Pt and Pd in aqueous hydroxide solutions. Geochim. Cosmochim. Acta 55, 1759–1767 (1991).

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