# Iridium

| Equilibrium reactions | lg*K* at infinite dilution and *T* = 298 K |
|---|---|
| | Brown and Ekberg, 2016 |
| $Ir^{3+} + H_2O \rightleftharpoons Ir(OH)^{2+} + H^+$ | −3.77 ± 0.10 |
| $Ir^{3+} + 2\ H_2O \rightleftharpoons Ir(OH)_2^{+} + 2\ H^+$ | −8.46 ± 0.20 |
| $Ir(OH)_3(s) + 3\ H^+ \rightleftharpoons Ir^{3+} + 3\ H_2O$ | 8.88 ± 0.20 |

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 736–739.

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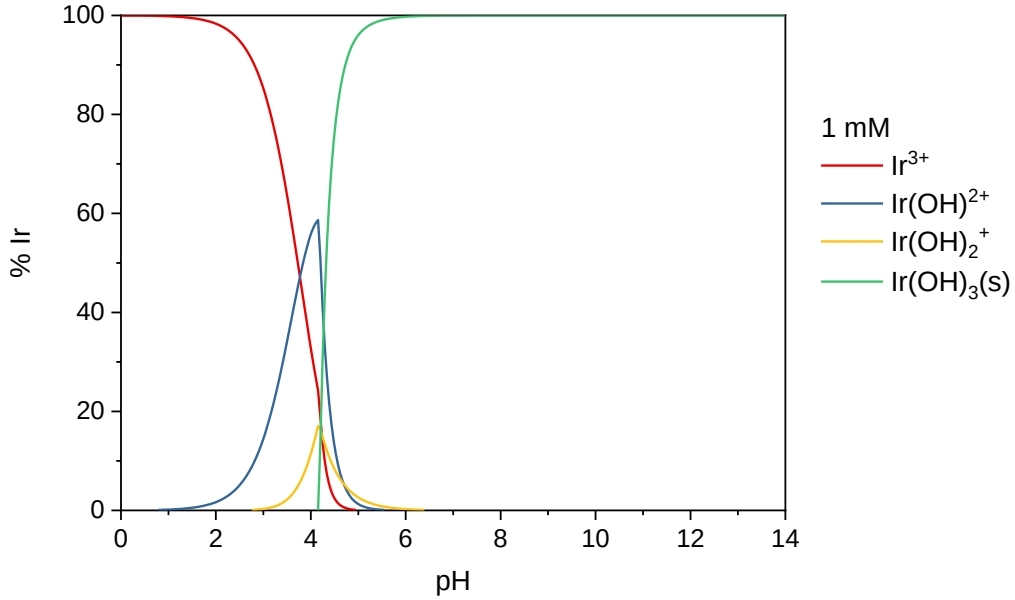
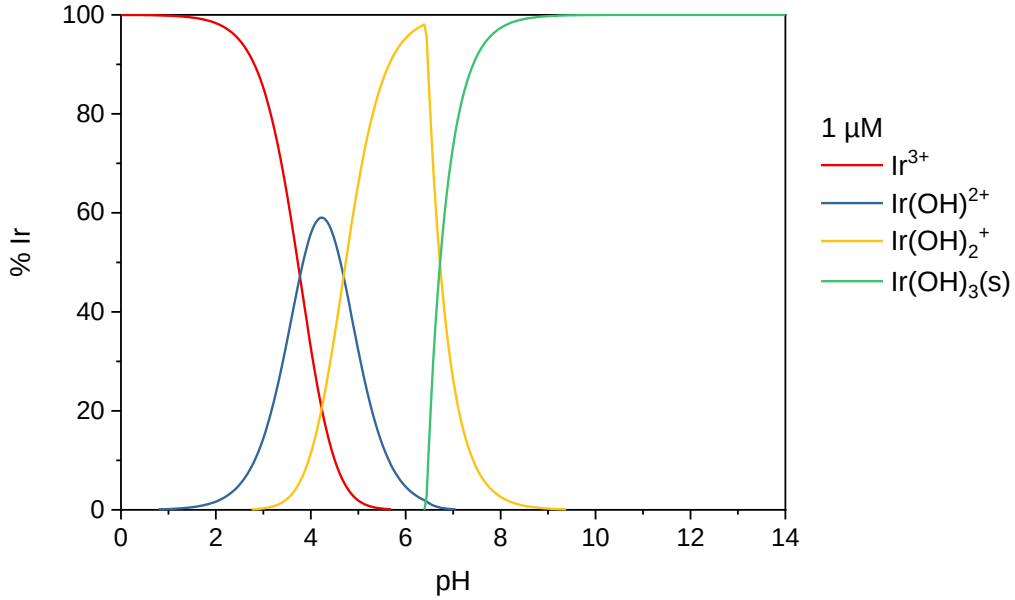
# Distribution diagrams

These diagrams have been computed at two Ir concentrations (1 mM = $1x10^{-3}$ mol $L^{-1}$ and 1 µM = $1x10^{-6}$ mol $L^{-1}$) with the 'best' equilibrium constants above. Calculations assume *T* = 298 K for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).

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