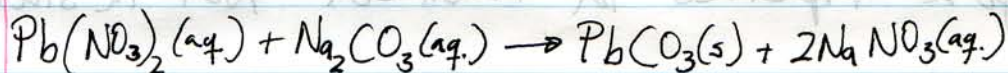


# Writing Chemical Equations

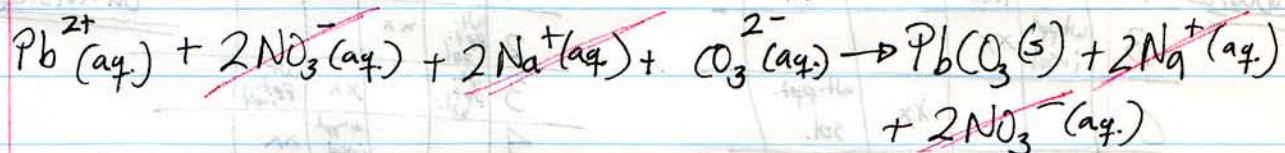
## REMEMBER:

• Write Conclusions and show calculations. Fill up all parts of the lab report.

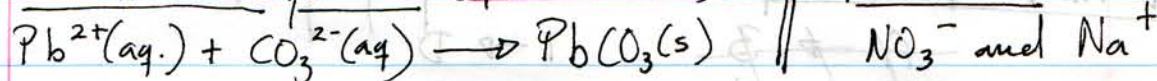
## Chemical or Molecular Equation



## Complete Ionic Equation



## Net Ionic Equation (spectator ions)



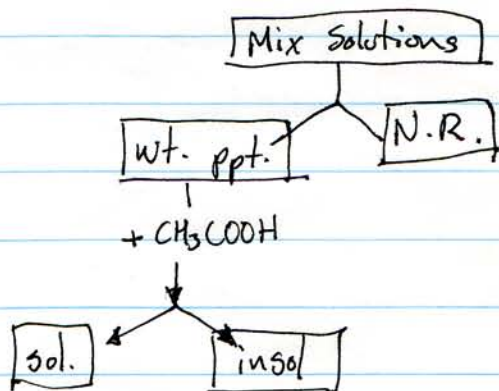
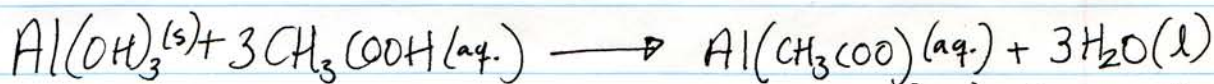
Spectator Ions

$\text{NO}_3^{-}$  and  $\text{Na}^{+}$

## Solubility Rules

• Memorize table for final, it's a hierarchical table.

## Solubility in $\text{CH}_3\text{COOH}$

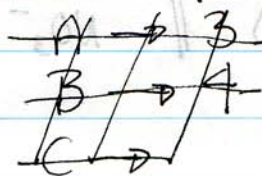


	$\text{Al}_2\text{SO}_4$	$\text{Na}_2\text{SO}_3$	$\text{Zn}(\text{NO}_3)_2$	$\text{NaCl}$	$\text{ZnCl}_2$	$\text{NaIO}_3$
$\text{Al}_2\text{SO}_4$	○	○	○	○	○	○
$\text{Na}_2\text{SO}_3$	○	○	○	○	○	○
$\text{Zn}(\text{NO}_3)_2$	○	○	○	○	○	○
$\text{NaCl}$	○	○	○	○	○	○

- 1/2 group knows first, other 1/2 unknowns then switch
- No more than 2ml. stock solns.
- If more unknown - 5pts.
- Label tubes & pipettes
- Dispose Pipettes in trash-can NOT in sink.

knowns

	A	B	C	D
A	xx	wt ppt insol		wt ppt sol
B	wt ppt insol	xx		
C			xx	wt ppt sol.
D	wt ppt sol		wt ppt sol.	xx



	1	2	3	4
1	xx	wt ppt sol		
2	wt ppt sol.	xx		
3	wt ppt sol.		xx	wt ppt insol
4			wt ppt insol.	xx

unknowns

