

Example: Quantum Mechanics



Even Marius Nordhagen

University of Oslo

evenmn@fys.uio.no

March 4, 2020

2020-03-04

- This is an example presentation about quantum mechanics
- The front frame is generated using *frontframe*
- Note also that the notes can be turned on and off in the first line of this file



2020-03-04

- ## Outline

Here is the outline of the presentation

Outline

- ▶ The Schrödinger Equation
- ▶ The Probability distribution
- ▶ References



2020-03-04

Outline

Outline

- ▶ The Schrödinger Equation
- ▶ The Probability distribution
- ▶ References

Additional note for the last bullet point. Note that the progression wheel is not moving when elements are added to a slide.

2020-03-04

*"The general theory of quantum mechanics is now almost complete...
...The underlying physical laws necessary for the mathematical theory of a large part of physics and the whole of chemistry are thus completely known, and the difficulty is only that the exact application of these laws leads to equations much too complicated to be soluble.*

Paul M. Dirac, Quantum Mechanics of Many-electron Systems²

“ The general theory of quantum mechanics is now almost complete...
...The underlying physical laws necessary for the mathematical theory of a large
part of physics and the whole of chemistry are thus completely known, and the
difficulty is only that the exact application of these laws leads to equations much
too complicated to be soluble.
Paul M. Dirac, Quantum Mechanics of Many-electron Systems¹ ”

Frame without title or subtitle

The Schrödinger Equation

The time-independent Schrödinger equation

The time-independent Schrödinger equation is given by

$$\hat{\mathcal{H}}\Psi_n = \varepsilon_n \Psi_n,$$

with $\hat{\mathcal{H}}$ as the Hamiltonian, Ψ_n as the wave function and ε_n as the corresponding energy².



2020-03-04

The Schrödinger Equation

Here, we have a basic slide with subtitle

The time-independent Schrödinger equation is given by

$$\hat{\mathcal{H}}\Psi_n = \varepsilon_n \Psi_n,$$

with $\hat{\mathcal{H}}$ as the Hamiltonian, Ψ_n as the wave function and ε_n as the corresponding energy².

The Probability Distribution

The probability distribution in quantum mechanics is given by

$$P(\boldsymbol{r}) = \frac{\Psi_n(\boldsymbol{r})^* \Psi_n(\boldsymbol{r})}{\int d\boldsymbol{r} \Psi_n(\boldsymbol{r})^* \Psi_n(\boldsymbol{r})},$$

where \boldsymbol{r} is a set of spatial and spin coordinates³.

The Probability Distribution

Here is a basic slide without a subtitle.

$$P(\boldsymbol{r}) = \frac{\Psi_n(\boldsymbol{r})^* \Psi_n(\boldsymbol{r})}{\int d\boldsymbol{r} \Psi_n(\boldsymbol{r})^* \Psi_n(\boldsymbol{r})}$$

The Probability Distribution

The probability distribution in quantum mechanics is given by

$$P(\mathbf{r}) = \frac{\Psi_n(\mathbf{r})^* \Psi_n(\mathbf{r})}{\int d\mathbf{r} \Psi_n(\mathbf{r})^* \Psi_n(\mathbf{r})},$$

where \mathbf{r} is a set of spatial and spin coordinates³. However, often the wave function is assumed to be normalized, and the equation is simply written as

$$P(\mathbf{r}) = \Psi_n(\mathbf{r})^* \Psi_n(\mathbf{r}),$$

2020-03-04

The Probability Distribution

The pause function can be used to add more elements to a slide

The Probability Distribution

The probability distribution in quantum mechanics is given by

$$P(\mathbf{r}) = \frac{\Psi_n(\mathbf{r})^* \Psi_n(\mathbf{r})}{\int d\mathbf{r} \Psi_n(\mathbf{r})^* \Psi_n(\mathbf{r})}$$

where \mathbf{r} is a set of spatial and spin coordinates³. However, often the wave function is assumed to be normalized, and the equation is simply written as

$$P(\mathbf{r}) = \Psi_n(\mathbf{r})^* \Psi_n(\mathbf{r}).$$

Thank you!

The title frame contains just a large centered text (should not be confused with frontframe)

References

1. Dirac, P. A. M. & Fowler, R. H. Quantum mechanics of many-electron systems. *Proceedings of the Royal Society A* **123**, 714 (1929).
2. Schrödinger, E. An Undulatory Theory of the Mechanics of Atoms and Molecules. *Physical Review* **28**, 1049 (1926).
3. Born, M. Zur Quantenmechanik der Stoßvorgänge. *Zeitschrift für Physik* **37**, 863 (1926).



References

1. Dirac, P. A. M. & Fowler, R. H. Quantum mechanics of many electron systems. *Proceedings of the Royal Society A* **123**, 714 (1929).
2. Schrödinger, E. An Undulatory Theory of the Mechanics of Atoms and Molecules. *Physical Review* **28**, 1049 (1926).
3. Born, M. Zur Quantenmechanik der Stoßvorgänge. *Zeitschrift für Physik* **37**, 863 (1926).