

Tabla periódica de los elementos

grupo 1

periodo 1

2

3

4

5

6

7

1	2											18							
H Hidrógeno 1s ¹												He Helio 1s ²							
3	4											10							
Li Litio 1s ² 2s ¹	Be Berilio 1s ² 2s ²											Ne Neón 1s ² 2s ² 2p ⁶							
11	12											18							
Na Sodio [Ne] 3s ¹	Mg Magnesio [Ne] 3s ²											Ar Argón [Ne] 3s ² 3p ⁶							
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36		
K Potasio [Ar] 4s ¹	Ca Calcio [Ar] 4s ²	Sc Escandio [Ar] 3d ¹ 4s ²	Ti Titanio [Ar] 3d ² 4s ²	V Vanadio [Ar] 3d ³ 4s ²	Cr Cromo [Ar] 3d ⁵ 4s ¹	Mn Manganeso [Ar] 3d ⁵ 4s ²	Fe Hierro [Ar] 3d ⁶ 4s ²	Co Cobalto [Ar] 3d ⁷ 4s ²	Ni Níquel [Ar] 3d ⁸ 4s ²	Cu Cobre [Ar] 3d ¹⁰ 4s ¹	Zn Zinc [Ar] 3d ¹⁰ 4s ²	Ga Gallio [Ar] 3d ¹⁰ 4s ² 4p ¹	Ge Germanio [Ar] 3d ¹⁰ 4s ² 4p ²	As Arsénico [Ar] 3d ¹⁰ 4s ² 4p ³	Se Selenio [Ar] 3d ¹⁰ 4s ² 4p ⁴	Br Bromo [Ar] 3d ¹⁰ 4s ² 4p ⁵	Kr Kriptón [Ar] 3d ¹⁰ 4s ² 4p ⁶		
37	38	39	40	41	42	(98)	43	44	45	46	47	48	49	50	51	52	53	54	
Rb Rubidio [Kr] 5s ¹	Sr Estroncio [Kr] 5s ²	Y Itrio [Kr] 4d ¹ 5s ²	Zr Zirconio [Kr] 4d ² 5s ²	Nb Niobio [Kr] 4d ⁴ 5s ¹	Mo Molibdeno [Kr] 4d ⁵ 5s ¹		Tc Tecnecio [Kr] 4d ⁵ 5s ²	Ru Rutenio [Kr] 4d ⁷ 5s ¹	Rh Rodio [Kr] 4d ⁸ 5s ¹	Pd Paladio [Kr] 4d ¹⁰	Ag Plata [Kr] 4d ¹⁰ 5s ¹	Cd Cadmio [Kr] 4d ¹⁰ 5s ²	In Indio [Kr] 4d ¹⁰ 5s ² 5p ¹	Sn Estañ [Kr] 4d ¹⁰ 5s ² 5p ²	Sb Antimonio [Kr] 4d ¹⁰ 5s ² 5p ³	Te Telurio [Kr] 4d ¹⁰ 5s ² 5p ⁴	I Yodo [Kr] 4d ¹⁰ 5s ² 5p ⁵	Xe Xenón [Kr] 4d ¹⁰ 5s ² 5p ⁶	
55	56	71	72	73	74	75	76	77	78	79	80	81	82	83	(210)	84	(210)	85	(220)
Cs Cesio [Xe] 6s ¹	Ba Bario [Xe] 6s ²	Lu Lutecio [Xe] 4f ¹⁴ 5d ¹ 6s ²	Hf Hafnio [Xe] 4f ¹⁴ 5d ² 6s ²	Ta Tantalio [Xe] 4f ¹⁴ 5d ³ 6s ²	W Wolframio [Xe] 4f ¹⁴ 5d ⁴ 6s ²	Re Renio [Xe] 4f ¹⁴ 5d ⁵ 6s ²	Os Osmio [Xe] 4f ¹⁴ 5d ⁶ 6s ²	Ir Iridio [Xe] 4f ¹⁴ 5d ⁷ 6s ²	Pt Platino [Xe] 4f ¹⁴ 5d ⁹ 6s ¹	Au Oro [Xe] 4f ¹⁴ 5d ¹⁰ 6s ¹	Hg Mercurio [Xe] 4f ¹⁴ 5d ¹⁰ 6s ²	Tl Talio [Xe] 4f ¹⁴ 5d ¹⁰ 6s ² 6p ¹	Pb Plomo [Xe] 4f ¹⁴ 5d ¹⁰ 6s ² 6p ²	Bi Bismuto [Xe] 4f ¹⁴ 5d ¹⁰ 6s ² 6p ³	(210)	Po Polonio [Xe] 4f ¹⁴ 5d ¹⁰ 6s ² 6p ⁴	(210)	At Astatio [Xe] 4f ¹⁴ 5d ¹⁰ 6s ² 6p ⁵	(220)
87	88	103	104	105	106	107	108	109	110	111	112	113	114	115	(292)	116	(294)	117	(294)
Fr Francio [Rn] 7s ¹	Ra Radio [Rn] 7s ²	Lr Laurencio [Rn] 5f ¹⁴ 6d ¹ 7p ¹	Rf Rutherfordio [Rn] 5f ¹⁴ 6d ² 7s ²	Db Dubnio [Rn] 5f ¹⁴ 6d ³ 7s ²	Sg Seaborgio [Rn] 5f ¹⁴ 6d ⁴ 7s ²	Bh Bohrio [Rn] 5f ¹⁴ 6d ⁵ 7s ²	Hs Hassio [Rn] 5f ¹⁴ 6d ⁶ 7s ²	Mt Meitnerio [Rn] 5f ¹⁴ 6d ⁷ 7s ²	Ds Darmstatio [Rn] 5f ¹⁴ 6d ⁸ 7s<										

- * 1 kJ/mol \approx 96.485 eV.
- * Todos los elementos tienen un estado de oxidación implícito cero.
- * Los estados de oxidación de los elementos 109,110, 111,112,113,114,115,116,117 y 118 son predicciones.
- * Las configuraciones electrónicas de los elementos 105, 106,107,108,109,110,111,112,113,114,115,116,117 y 118 son predicciones.



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