

Aluno:	Turma:	Data
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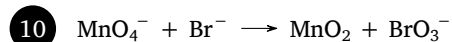
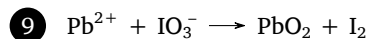
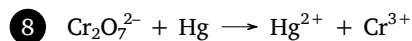
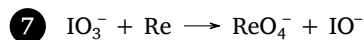
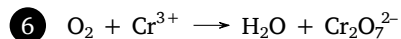
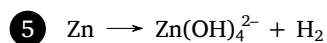
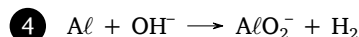
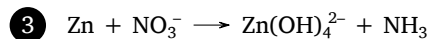
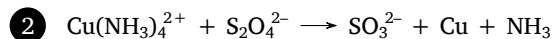
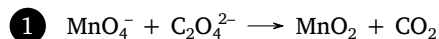
1 Realize o balanceamento das reações redox

- 1 $\text{As}_2\text{O}_3 + \text{HNO}_3 + \text{H}_2\text{O} \longrightarrow \text{H}_3\text{AsO}_4 + \text{NO}$
- 2 $\text{KI} + \text{KNO}_2 + \text{H}_2\text{SO}_4 \longrightarrow \text{I}_2 + \text{NO} + \text{K}_2\text{SO}_4 + \text{H}_2\text{O}$
- 3 $\text{KI} + \text{H}_2\text{SO}_4 \longrightarrow \text{K}_2\text{SO}_4 + \text{I}_2 + \text{H}_2\text{S} + \text{H}_2\text{O}$
- 4 $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$
- 5 $\text{Cu} + \text{H}_2\text{SO}_4 \longrightarrow \text{CuSO}_4 + \text{SO}_2 + \text{H}_2\text{O}$
- 6 $\text{KSCN} + \text{H}_2\text{O} + \text{I}_2 \longrightarrow \text{KHSO}_4 + \text{HI} + \text{ICN}$
- 7 $\text{AlBr}_3 + \text{KMnO}_4 + \text{H}_2\text{SO}_4 \longrightarrow \text{Al}_2(\text{SO}_4)_3 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{Br}_2 + \text{H}_2\text{O}$
- 8 $\text{FeSO}_4 + \text{KMnO}_4 + \text{H}_2\text{SO}_4 \longrightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O}$
- 9 $\text{Na}_2\text{C}_2\text{O}_4 + \text{KMnO}_4 + \text{H}_2\text{SO}_4 \longrightarrow \text{K}_2\text{SO}_4 + \text{Na}_2\text{SO}_4 + \text{MnSO}_4 + \text{CO}_2 + \text{H}_2\text{O}$
- 10 $\text{KMnO}_4 + \text{Na}_2\text{SO}_3 + \text{H}_2\text{SO}_4 \longrightarrow \text{MnSO}_4 + \text{Na}_2\text{SO}_4 + \text{K}_2\text{SO}_4 + \text{H}_2\text{O}$

2 Equilibre cada reação redox em solução ácida.

- 1 $\text{Mn}^{2+} + \text{BiO}^{3-} \longrightarrow \text{MnO}_4^- + \text{Bi}^{3+}$
- 2 $\text{MnO}_4^- + \text{S}_2\text{O}_3^{2-} \longrightarrow \text{S}_4\text{O}_6^{2-} + \text{Mn}^{2+}$
- 3 $\text{ClO}_3^- + \text{Cl}^- \longrightarrow \text{Cl}_2 + \text{ClO}_2$
- 4 $\text{P} + \text{Cu}^{2+} \longrightarrow \text{Cu} + \text{H}_2\text{PO}_4^-$
- 5 $\text{PH}_3 + \text{I}_2 \longrightarrow \text{H}_3\text{PO}_2^- + \text{I}^-$
- 6 $\text{NO}_2 \longrightarrow \text{NO}_3^- + \text{NO}$
- 7 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} \longrightarrow \text{O}_2 + \text{Cr}^{3+}$
- 8 $\text{PbO}_2 + \text{I}_2 \longrightarrow \text{Pb}^{2+} + \text{IO}_3^-$
- 9 $\text{ReO}_4^- + \text{IO}^- \longrightarrow \text{IO}_3^- + \text{Re}$
- 10 $\text{As} \longrightarrow \text{H}_2\text{AsO}_4^- + \text{AsH}_3$

3 Equilibre cada reação redox em solução básica.



Gabarito

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