



SECRETÁRIA DE EDUCAÇÃO DE MATO GROSSO
DO SUL

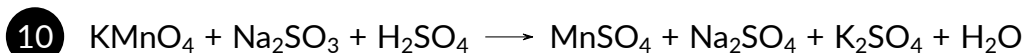
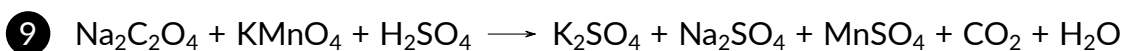
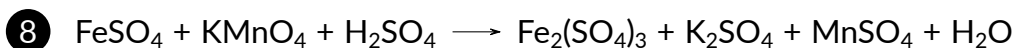
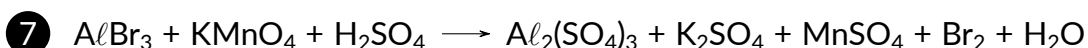
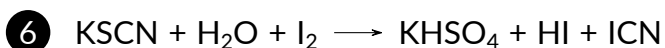
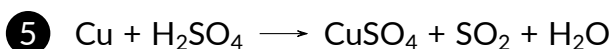
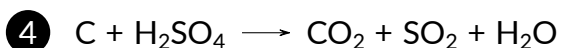
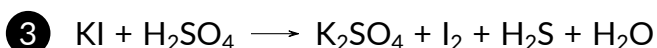
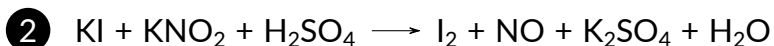
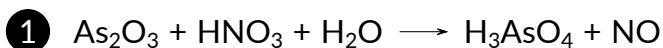
Escola _____

Prof: Fábio Lima

Disciplina: Química



Aluno:	Turma:	Data
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Solutions to the Exercises

- 1 $3 \text{As}_2\text{O}_3 + 4 \text{HNO}_3 + 7 \text{H}_2\text{O} \longrightarrow \text{H}_3\text{AsO}_4 + \text{NO}$
- 2 $2 \text{KI} + 2 \text{KNO}_2 + 2 \text{H}_2\text{SO}_4 \longrightarrow \text{I}_2 + 2 \text{NO} + 2 \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
- 3 $8 \text{KI} + 5 \text{H}_2\text{SO}_4 \longrightarrow \text{K}_2\text{SO}_4 + 4 \text{I}_2 + \text{H}_2\text{S} + 4 \text{H}_2\text{O}$
- 4 $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2$
- 5 $\text{Cu} + 2 \text{H}_2\text{SO}_4 \longrightarrow \text{CuSO}_4 + \text{SO}_2 + 2 \text{H}_2\text{O}$
- 6 $\text{KSCN} + 4 \text{H}_2\text{O} + 4 \text{I}_2 \longrightarrow \text{KHSO}_4 + 7 \text{HI} + \text{ICN}$
- 7 $10 \text{AlBr}_3 + 6 \text{KMnO}_4 + 24 \text{H}_2\text{SO}_4 \longrightarrow 5 \text{Al}_2(\text{SO}_4)_3 + 3 \text{K}_2\text{SO}_4 + 6 \text{MnSO}_4 + 15 \text{Br}_2 + 24 \text{H}_2\text{O}$
- 8 $10 \text{FeSO}_4 + 2 \text{KMnO}_4 + 8 \text{H}_2\text{SO}_4 \longrightarrow 5 \text{Fe}_2(\text{SO}_4)_3 + \text{K}_2\text{SO}_4 + 2 \text{MnSO}_4 + \text{H}_2\text{O}$
- 9 $5 \text{Na}_2\text{C}_2\text{O}_4 + 2 \text{KMnO}_4 + 8 \text{H}_2\text{SO}_4 \longrightarrow \text{K}_2\text{SO}_4 + 5 \text{Na}_2\text{SO}_4 + 2 \text{MnSO}_4 + 10 \text{CO}_2 + \text{H}_2\text{O}$
- 10 $2 \text{KMnO}_4 + 5 \text{Na}_2\text{SO}_3 + 3 \text{H}_2\text{SO}_4 \longrightarrow 2 \text{MnSO}_4 + 5 \text{Na}_2\text{SO}_4 + \text{K}_2\text{SO}_4 + 3 \text{H}_2\text{O}$