SECRETÁRIA DE EDUCAÇÃO DE MATO GROSSO DO SUL



Escola _____

Disciplina: Química



Aluno: Turma: Data

Balanceie as reações a seguir

$$\underbrace{ \text{H}_3 \text{PO}_4 + \underline{\qquad} \text{KOH} \longrightarrow \underline{\qquad} \text{K}_3 \text{PO}_4 + \underline{\qquad}$$

Prof: Fábio Lima

2 ____K + ___B₂O₃
$$\longrightarrow$$
 ____K₂O + ___B

6 ____Na + ____O
$$_2 \rightarrow$$
 ____Na $_2$ O

9 __NaOH + ___H
$$_2$$
CO $_3$ \longrightarrow ___Na $_2$ CO $_3$ +

11 ____Na + ____O₂
$$\longrightarrow$$
 ____Na₂O

$$13 \quad \underline{\hspace{1cm}} A\ell + \underline{\hspace{1cm}} S_8 \longrightarrow \underline{\hspace{1cm}} A\ell_2S_3$$

$$14 \quad \underline{\hspace{1cm}} Cs + \underline{\hspace{1cm}} N_2 \longrightarrow \underline{\hspace{1cm}} Cs_3N$$

$$18 \quad \underline{\hspace{1cm}} N_2 + \underline{\hspace{1cm}} H_2 \longrightarrow \underline{\hspace{1cm}} NH_3$$

23
$$A\ell$$
 Li + ____A ℓ C ℓ_3 \longrightarrow ___LiC ℓ +

$$26 \quad \underline{\hspace{1cm}} Rb + \underline{\hspace{1cm}} P \longrightarrow \underline{\hspace{1cm}} Rb_3P$$

29 ____Na + ____C
$$\ell_2 \longrightarrow$$
 ____NaC ℓ

30 ____Rb + ____S
$$_8 \longrightarrow$$
 ____Rb $_2$ S

31 ____H₃PO₄ + ____Ca(OH)₂
$$\longrightarrow$$
 ____Ca₃(PO₄)₂ + ____H₂O

32 ____NH₃ + ____HC
$$\ell \rightarrow$$
 ____NH₄C ℓ

$$39 \quad \underline{\hspace{1cm}} S_8 + \underline{\hspace{1cm}} NO_2 \longrightarrow \underline{\hspace{1cm}} SO_2 + \underline{\hspace{1cm}} N_2$$

$$S_8 + \underline{\hspace{1cm}} NO_3 \longrightarrow \underline{\hspace{1cm}} SO_2 + \underline{\hspace{1cm}}$$

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$$C_5H_{10} + O_2 \longrightarrow CH_2O$$