

Aluno:

## SECRETÁRIA DE EDUCAÇÃO DE MATO GROSSO DO SUL



Data

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Disciplina: Química

Turma:

PROF. QUÍMICA
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Balanceie as reações a seguir

1 
$$\underline{\hspace{1cm}} H_3PO_4 + \underline{\hspace{1cm}} KOH \longrightarrow \underline{\hspace{1cm}} K_3PO_4 + \underline{\hspace{1cm}} H_2O$$

2 \_\_\_\_K + \_\_\_B<sub>2</sub>O<sub>3</sub> 
$$\longrightarrow$$
 \_\_\_\_K<sub>2</sub>O + \_\_\_B

3 \_\_\_\_HC
$$\ell$$
 + \_\_\_\_NaOH  $\longrightarrow$  \_\_\_\_NaC $\ell$  + \_\_\_\_H<sub>2</sub>O

4 \_\_\_\_Na + \_\_\_\_NaNO<sub>3</sub> 
$$\rightarrow$$
 \_\_\_\_Na<sub>2</sub>O + \_\_\_\_N<sub>2</sub>

$$N_2 + _O_2 \rightarrow _N_2O_5$$

8 \_\_\_\_H<sub>3</sub>PO<sub>4</sub> + \_\_\_\_Mg(OH)<sub>2</sub> 
$$\longrightarrow$$
 \_\_\_\_Mg<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + \_\_\_\_H<sub>2</sub>O

9 \_\_\_NaOH + \_\_\_H
$$_2$$
CO $_3$   $\rightarrow$  \_\_\_Na $_2$ CO $_3$  + \_\_\_H $_2$ O

10 \_\_\_\_KOH + \_\_\_HBr 
$$\rightarrow$$
 \_\_\_KBr + \_\_\_H $_2$ O

11 \_\_\_\_Na + \_\_\_O<sub>2</sub> 
$$\longrightarrow$$
 \_\_\_\_Na<sub>2</sub>O

12 
$$A\ell(OH)_3 + H_2CO_3 \rightarrow A\ell_2(CO_3)_3 + H_2O$$

$$13 \quad \underline{\hspace{1cm}} A\ell + \underline{\hspace{1cm}} S_8 \longrightarrow \underline{\hspace{1cm}} A\ell_2S_3$$

$$15 \quad \underline{\hspace{1cm}} Mg + \underline{\hspace{1cm}} C\ell_2 \longrightarrow \underline{\hspace{1cm}} MgC\ell_2$$

16 \_\_\_\_Rb + \_\_\_\_RbNO<sub>3</sub> 
$$\longrightarrow$$
 \_\_\_\_Rb<sub>2</sub>O + \_\_\_\_N<sub>2</sub>

17 \_\_\_\_C<sub>6</sub>H<sub>6</sub> + \_\_\_\_O<sub>2</sub> 
$$\longrightarrow$$
 \_\_\_\_CO<sub>2</sub> + \_\_\_\_H<sub>2</sub>O

- $18 \quad \underline{\hspace{1cm}} N_2 + \underline{\hspace{1cm}} H_2 \longrightarrow \underline{\hspace{1cm}} NH_3$
- 19  $C_{10}H_{22}^{+}C_{02} \rightarrow CO_{2}^{+}H_{2}O$
- 20  $\underline{\hspace{1cm}}$   $A\ell(OH)_3 + \underline{\hspace{1cm}}$   $HBr \longrightarrow \underline{\hspace{1cm}}$   $A\ell Br_3 + \underline{\hspace{1cm}}$   $H_2O$
- 21 \_\_\_\_CH<sub>3</sub>CH<sub>2</sub>CH<sub>3</sub> + \_\_\_\_O<sub>2</sub>  $\longrightarrow$  \_\_\_\_CO<sub>2</sub> + \_\_\_\_H<sub>2</sub>O
- 22  $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$
- 23 \_\_\_\_Li + \_\_\_\_ $A\ell C\ell_3 \longrightarrow$  \_\_\_\_Li $C\ell$  + \_\_\_\_ $A\ell$
- 24  $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$
- 25 \_\_\_NH<sub>4</sub>OH + \_\_\_H<sub>3</sub>PO<sub>4</sub>  $\longrightarrow$  \_\_\_(NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> + \_\_\_H<sub>2</sub>O
- 27 \_\_\_\_CH<sub>4</sub> + \_\_\_\_O<sub>2</sub>  $\longrightarrow$  \_\_\_\_CO<sub>2</sub> + \_\_\_\_H<sub>2</sub>O
- 28 \_\_\_\_A $\ell$ (OH)<sub>3</sub> + \_\_\_\_H<sub>2</sub>SO<sub>4</sub>  $\longrightarrow$  \_\_\_\_A $\ell$ <sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + \_\_\_\_H<sub>2</sub>O
- 29 \_\_\_\_Na + \_\_\_\_ $C\ell_2 \longrightarrow$  \_\_\_\_Na $C\ell$
- $30 \quad \underline{\hspace{1cm}} Rb + \underline{\hspace{1cm}} S_8 \longrightarrow \underline{\hspace{1cm}} Rb_2S$
- 31 \_\_\_\_H<sub>3</sub>PO<sub>4</sub> + \_\_\_\_Ca(OH)<sub>2</sub>  $\longrightarrow$  \_\_\_\_Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> + \_\_\_\_H<sub>2</sub>O
- 32 \_\_\_NH<sub>3</sub> + \_\_\_HC $\ell \rightarrow$  \_\_\_NH<sub>4</sub>C $\ell$
- 33 \_\_\_\_Li + \_\_\_\_ $H_2O \longrightarrow$  \_\_\_\_LiOH + \_\_\_\_ $H_2$
- 35  $\_\_NH_3 + \_\_O_2 \longrightarrow \_\_N_2 + \_\_H_2O$
- 36 \_\_\_\_FeS<sub>2</sub> + \_\_\_\_O<sub>2</sub>  $\longrightarrow$  \_\_\_\_Fe<sub>2</sub>O<sub>3</sub> + \_\_\_\_SO<sub>2</sub>
- 37 \_\_\_\_C + \_\_\_\_SO<sub>2</sub>  $\longrightarrow$  \_\_\_\_CS<sub>2</sub> + \_\_\_\_<sup>+</sup>\_\_\_CO

- $\_\_S_8 + \_\_Br_2 \longrightarrow \_\_SBr_2$
- $\_\_S_8 + \_\_NO_2 \longrightarrow \_\_SO_2 + \_\_N_2$
- $\_\__S_8 + \_\__NO_3 \longrightarrow \_\__SO_2 + \_\__NO_3$
- $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$
- $C_7H_{14} + O_2 \rightarrow CO + H_2$
- $C_6H_6 + HNO_3 \rightarrow C_6H_5NO_2 + H_2O_3$
- $C_3H_4 + I_2 \longrightarrow C_3H_4I_2$
- $CO_2 + C\ell_2 \rightarrow C\ell_4 + O_2$
- $\_\_S_7 + \_\_P_2O_5 + \_\_O_2 \longrightarrow \_\_SO_3 + \_\_P_4$
- 47 \_\_\_\_N<sub>2</sub> + \_\_\_\_C<sub>2</sub>H<sub>6</sub>  $\longrightarrow$  \_\_\_\_N<sub>2</sub>H<sub>4</sub> + \_\_\_\_C<sub>2</sub>H<sub>2</sub>
- $C_5H_{10} + O_2 \rightarrow CH_2O$
- 50 \_\_\_\_NaOH + \_\_\_H $_2$ SO $_4$   $\longrightarrow$  \_\_\_\_Na $_2$ SO $_4$  + \_\_\_H $_2$ O
- 51 \_\_\_\_C $_6O_6Cr + ___C\ell_2 \longrightarrow __CrC\ell_3 + ___CO$
- $P_4 + HC\ell + O_2 \rightarrow PC\ell_3 + H_2O$
- $H_3PO_4 + C \rightarrow P_4 + CO + H_2O$
- NOC $\ell$  + WC<sub>6</sub>O<sub>6</sub>  $\longrightarrow$  WN<sub>2</sub>O<sub>2</sub>C $\ell$ <sub>2</sub> + CO
- 56 \_\_\_NH<sub>3</sub> + \_\_\_CO  $\longrightarrow$  \_\_\_CH<sub>4</sub> + \_\_\_N<sub>2</sub> + \_\_\_O<sub>2</sub>
- 57 \_\_\_\_PC $\ell_3$  + \_\_\_\_H $_2$ O  $\longrightarrow$  \_\_\_\_H $_3$ PO $_3$  + \_\_\_\_HC $\ell$