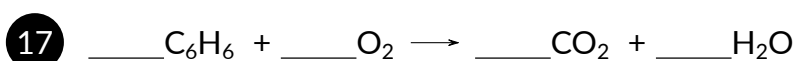
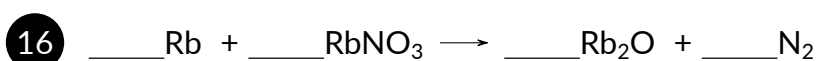
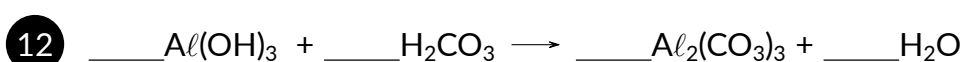
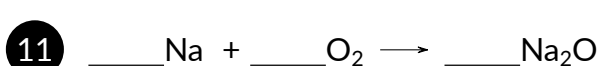
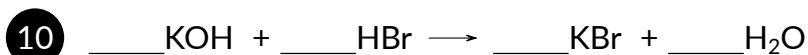
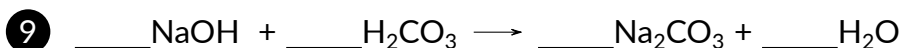
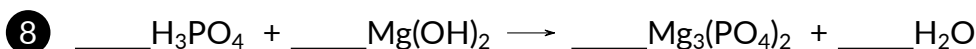
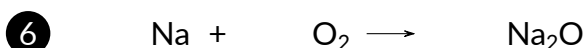
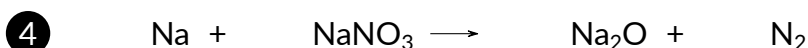
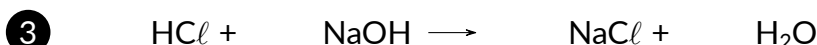
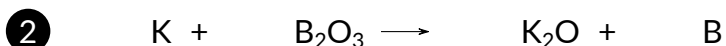
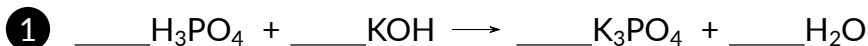


Aluno: _____

Turma: _____

Data _____

Balançeie as reações a seguir



- 18 $\text{N}_2 + \text{H}_2 \rightarrow \text{NH}_3$
- 19 $\text{C}_{10}\text{H}_{22} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 20 $\text{Al}(\text{OH})_3 + \text{HBr} \rightarrow \text{AlBr}_3 + \text{H}_2\text{O}$
- 21 $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 22 $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 23 $\text{Li} + \text{AlCl}_3 \rightarrow \text{LiCl} + \text{Al}$
- 24 $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 25 $\text{NH}_4\text{OH} + \text{H}_3\text{PO}_4 \rightarrow (\text{NH}_4)_3\text{PO}_4 + \text{H}_2\text{O}$
- 26 $\text{Rb} + \text{P} \rightarrow \text{Rb}_3\text{P}$
- 27 $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 28 $\text{Al}(\text{OH})_3 + \text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{H}_2\text{O}$
- 29 $\text{Na} + \text{Cl}_2 \rightarrow \text{NaCl}$
- 30 $\text{Rb} + \text{S}_8 \rightarrow \text{Rb}_2\text{S}$
- 31 $\text{H}_3\text{PO}_4 + \text{Ca}(\text{OH})_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + \text{H}_2\text{O}$
- 32 $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$
- 33 $\text{Li} + \text{H}_2\text{O} \rightarrow \text{LiOH} + \text{H}_2$
- 34 $\text{Ca}_3(\text{PO}_4)_2 + \text{SiO}_2 + \text{C} \rightarrow \text{CaSiO}_3 + \text{CO} + \text{P}$
- 35 $\text{NH}_3 + \text{O}_2 \rightarrow \text{N}_2 + \text{H}_2\text{O}$
- 36 $\text{FeS}_2 + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
- 37 $\text{C} + \text{SO}_2 \rightarrow \text{CS}_2 + \text{CO}$

- 38 $\text{S}_8 + \text{Br}_2 \rightarrow \text{SBr}_2$
- 39 $\text{S}_8 + \text{NO}_2 \rightarrow \text{SO}_2 + \text{N}_2$
- 40 $\text{S}_8 + \text{NO}_3 \rightarrow \text{SO}_2 + \text{NO}$
- 41 $\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
- 42 $\text{C}_7\text{H}_{14} + \text{O}_2 \rightarrow \text{CO} + \text{H}_2$
- 43 $\text{C}_6\text{H}_6 + \text{HNO}_3 \rightarrow \text{C}_6\text{H}_5\text{NO}_2 + \text{H}_2\text{O}$
- 44 $\text{C}_3\text{H}_4 + \text{I}_2 \rightarrow \text{C}_3\text{H}_4\text{I}_2$
- 45 $\text{CO}_2 + \text{Cl}_2 \rightarrow \text{CCl}_4 + \text{O}_2$
- 46 $\text{S}_7 + \text{P}_2\text{O}_5 + \text{O}_2 \rightarrow \text{SO}_3 + \text{P}_4$
- 47 $\text{N}_2 + \text{C}_2\text{H}_6 \rightarrow \text{N}_2\text{H}_4 + \text{C}_2\text{H}_2$
- 48 $\text{C}_5\text{H}_{10} + \text{O}_2 \rightarrow \text{CH}_2\text{O}$
- 49 $\text{C}_6\text{H}_{12}\text{O}_6 + \text{F}_2 \rightarrow \text{C}_6\text{H}_6\text{F}_6 + \text{H}_2\text{O} + \text{O}_2$
- 50 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
- 51 $\text{C}_6\text{O}_6\text{Cr} + \text{Cl}_2 \rightarrow \text{CrCl}_3 + \text{CO}$
- 52 $\text{P}_4 + \text{HCl} + \text{O}_2 \rightarrow \text{PCl}_3 + \text{H}_2\text{O}$
- 53 $\text{H}_3\text{PO}_4 + \text{C} \rightarrow \text{P}_4 + \text{CO} + \text{H}_2\text{O}$
- 54 $\text{Na} + \text{C}_2\text{Cl}_6 \rightarrow \text{NaCl} + \text{C}_2\text{Cl}_2$
- 55 $\text{NOCl} + \text{WC}_6\text{O}_6 \rightarrow \text{WN}_2\text{O}_2\text{Cl}_2 + \text{CO}$
- 56 $\text{NH}_3 + \text{CO} \rightarrow \text{CH}_4 + \text{N}_2 + \text{O}_2$
- 57 $\text{PCl}_3 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{PO}_3 + \text{HCl}$