

ASTR 792 HW 2

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Let

a = radius of dust

r_H = radius of hydrogen atom = Bohr radius = $5.3 \cdot 10^{-9} \text{ cm}$

n_H = the number density of hydrogen

T = the temperature of the hydrogen gas

$$a \gg r_H$$

$$m_{dust} \gg m_H$$

2a

$$\begin{aligned} \langle v_H \rangle &= \sqrt{8 \frac{k_b T}{\pi m_H}} \\ &= \sqrt{\frac{8 \cdot 10^{-22} \text{ J/K} \cdot T}{7\pi \cdot 1.67 \cdot 10^{-27} \text{ kg}}} \end{aligned}$$

2b

Let τ_m be the mean free path for a hydrogen atom

$$\begin{aligned} \tau_m &= \frac{1}{n_b \sigma_{AB}} \\ \sigma &= \pi(a + r_H)^2 \\ \rightarrow \tau_m &= \frac{1}{n_H \pi(a + r_H)^2} \end{aligned}$$

Letting t be the time for a grain to be hit by a single hydrogen atom, then we can find the time to encounter a hydrogen atom by taking the mean free path τ_m and dividing by the velocity of the hydrogen

$$\begin{aligned} t &= \frac{\tau_m}{\langle v_H \rangle} \\ &= \frac{1}{\langle v_H \rangle n_H \cdot \pi(a + r_H)^2} \end{aligned}$$

Since this is the mean free path for a single collision, we need to determine the number of collisions it would take to accumulate the mass of a dust grain. we can calculate the mass of the dust particle by

$$N_H = \frac{m_{dust}}{m_H}$$

$$\rightarrow t_{total} = \frac{m_{dust}}{m_H} \frac{1}{\langle v_H \rangle n_H \cdot \pi(a + r_H)^2}$$

$$\approx \frac{m_{dust}}{m_H} \frac{1}{\langle v_H \rangle n_H \cdot \pi(a)^2}$$

2c

$$a = 10^{-5} \text{ cm}$$

$$n_H = 30 \text{ cm}^{-3}$$

$$\rho_{dust} = 3 \text{ g cm}^{-3}$$

$$T = 10^2 \text{ K}$$

c1

$$\langle v_H \rangle = \sqrt{\frac{8 \cdot 10^{-22} \text{ J/K} \cdot 10^2 \text{ K}}{7\pi \cdot 1.67 \cdot 10^{-27} \text{ kg}}}$$

$$\langle v_H \rangle \approx 1.5 \cdot 10^3 \text{ m/s}$$

$$= 1.5 \cdot 10^5 \text{ cm/s}$$

c2

First, let's calculate the mass of a single dust grain

$$\rho \cdot V = m_{dust}$$

$$\rightarrow 3 \text{ g cm}^3 \cdot \frac{4\pi}{3} (10^{-5} \text{ cm})^3 = m_{dust}$$

$$\rightarrow m_{dust} = 4\pi \cdot 10^{-15} \text{ g}$$

Now we can divide this mass by the mass of a hydrogen atom to obtain the number N_H of atoms necessary to equal the mass of a single dust grain

$$\rightarrow N_H = \frac{m_{dust}}{m_H}$$

$$N_H = \frac{4\pi \cdot 10^{-15} \text{ g}}{1.67 \cdot 10^{-25} \text{ g}}$$

$$= 7.5 \cdot 10^9 \text{ H atoms per dust grain}$$

Now we can plug this and 2c1 into the result from 2b

$$\begin{aligned}
\rightarrow t_{total} &= \frac{m_{dust}}{m_H} \frac{1}{<v_H> n_H \cdot \pi(a)^2} \\
&= 7.5 \cdot 10^9 \frac{1}{(1.5 \cdot 10^5 \text{ cm/s})(30 \text{ cm}^3)\pi(10^{-5} \text{ cm})^2} \\
&= 5.3 \cdot 10^{12} \text{ s} \\
&\approx 10^5 \text{ years}
\end{aligned}$$