

$$\begin{aligned}
 \Delta H_{\text{rxn}} &= \sum \Delta H_{\text{comb}} (\text{reactant}) - \sum \Delta H_{\text{comb}} (\text{Products}) \\
 &= [(-1411) + (-286)] - [(-1560)] = -137 \text{ kJ/mol}
 \end{aligned}$$