

## REVISÃO 8



## RESUMO

$$\Delta G^0 = \Delta H^0 - T\Delta S^0 \quad (T = 298 \text{ K})$$

$$\Delta G = \Delta H^0 - T\Delta S^0 \quad (\text{Qualquer } T)$$

$$\Delta G = \Delta G^0 + RT \ln Q \quad (Q = Q_p)$$

$$0 = \Delta G^0 + RT \ln K \quad (K = K_p)$$

$$\Delta G^0 = -RT \ln K$$

$$-RT \ln K = \Delta H^0 - T\Delta S$$

$$\ln K = \frac{-\Delta H^0}{R} \left( \frac{1}{T} \right) + \frac{\Delta S^0}{R}$$

$$\ln \frac{K_2}{K_1} = \frac{-\Delta H^0}{R} \left( \frac{1}{T_2} - \frac{1}{T_1} \right)$$

$$\ln P_v = \frac{-\Delta H_v^0}{R} \left( \frac{1}{T} \right) + \frac{\Delta S^0}{R}$$

$$P_v(1) = X(1) \cdot P_v^0(1)$$

$$\frac{\Delta P}{P} = K_T \cdot W$$

$$K_T = \frac{M(\text{solvente})}{1000}$$

$$K_T(\text{H}_2\text{O}) = \frac{18 \text{ g/mol}}{1000} = 0,018 \text{ kg/mol}$$