



**RAJARATA UNIVERSITY OF SRI LANKA
FACULTY OF APPLIED SCIENCES**

B.Sc. (Joint Major) Degree in Chemistry & Physics
Fourth Year Semester II Examination- September/November 2014

CHE 4201 APPLIED MOLECULAR CHEMISTRY

Answer both Parts I and II

Time allowed: 2 hours

All symbols carry their standard meanings. Any standard symbol may be used without defining it.

PART I

Answer in the spaces provided. Return this section.

1. State the value of following quantities in atomic units:
(a) Mass of electron (2) Charge of electron (3) Planck constant

2. Write down the Slater determinant for He ($Z = 2$)

3. State variational principle. What is meant by the term "self consistence"?

4. How a potential energy surface for a molecule is resulted from Bon Oppenheimer approximation? How many dimensionalities do occur for H_2O in PES?
5. Why Hatree method is not accepted? Give brief answers.
6. Write the electronic Hamiltonian for H_2O .
7. What is the major difference between *ab initio* and semi-empirical methods of electronic structure calculations?
8. Write the operators for following observables:

Electron – electron interaction term

Electron – nuclear interaction term
9. Why molecular mechanics methods cannot be used to describe chemical reactivity?

10. A 6-32G basis set notation is given. Calculate the number of primitive Gaussians in 6-31G for Li ($Z=3$).

PART II

Answer in separate sheets. Retain this section.

1. (a) Discuss the advantages of Hartree-Fock theory in comparison to Hartree method.
 - (b) State Hartree Fock equation of a many electrons system. Clearly identify all terms used.
 - (c) Derive Roothan Hall equation using basis set terminology.
 - (d) A single point energy calculation was performed for formaldehyde with Gaussian09TM code. A section of the output file is attached. Answer following questions.
 - (i) The level of theory used in the calculation
 - (ii) State the wave functions of HOMO-LUMO orbitals
 - (iii) Prove that these wave functions are normalized.
2. (a) What is a force field of a given molecule? How they are used in MM codes?

(b) The atoms in a methylene molecule (CH₂) have the following cartesian coordinates (in Å):

Atom	x	y	z
C	0.000	0.000	0.000
H1	0.000	-0.785	0.901
H2	0.000	1.096	1.143

The molecule is described by a harmonic stretching and bending force field with the following equilibrium values for the bond lengths and bond angle, respectively, $r_{\text{CH,eq}} = 1.150 \text{ Å}$ and $\theta_{\text{HCH,eq}} = 103.3^\circ$.

When the molecular mechanics energy of the system is calculated using the MMFF force field at the geometry listed above, the **total energy** is 42.82 kcal/mol. The stretching force constant for the two C-H bonds is $k_{\text{s,CH}} = 406.0 \text{ kcal mol}^{-1} \text{ Å}^{-2}$. From this information, **determine the bending force constant, $k_{\text{b,HCH}}$.**

Express your answer in units of $\text{kcal mol}^{-1} \text{ deg}^{-2}$.

Sample printout

Gaussian 09: IA32W-G09RevB.01 12-Aug-2011
10-Oct-2014

#T RHF/6-31G(d) Pop=Reg Test

Formaldehyde Single Point

		(A1)	(A1)	(A1)	(B1)	(A1)	(B1)	(B2)		
Occupied		(A1)	(A1)	(A1)	(B1)	(A1)	(B1)	(B2)		
Virtual		(B1)	(A1)	(B1)	(A1)	(B1)	(B2)	(A1)	(A1)	(B2)
		(A1)	(B1)	(B2)	(A1)	(A2)	(B1)	(A1)	(A2)	(A1)
		(A1)	(B1)	(B2)	(A1)	(A1)				
Molecular Orbital Coefficients:										
		4	5	6	7	8				
		(A1)--O	(B2)--O	(A1)--O	(B1)--O	(B2)--O				
Eigenvalues --		-0.27257	-0.69706	-0.63916	-0.52290	-0.44043				
1 1 C 1S		-0.24261	0.00000	0.01957	0.00000	0.00000				
2 2S		0.33994	0.00000	-0.06111	0.00000	0.00000				
3 2PX		0.00000	0.00000	0.00000	0.32494	0.00000				
4 2PY		0.00000	0.42032	0.00000	0.00000	-0.13770				
5 2PZ		-0.18473	0.00000	-0.37609	0.00000	0.00000				
6 3S		0.31305	0.00000	0.03916	0.00000	0.00000				
7 3PX		0.00000	0.00000	0.00000	0.21194	0.00000				
8 3PY		0.00000	0.15769	0.00000	0.00000	-0.04436				
9 3PZ		-0.07978	0.00000	-0.08877	0.00000	0.00000				
10 4XX		-0.31333	0.00000	0.00548	0.00000	0.00000				
11 4YY		0.03021	0.00000	0.02730	0.00000	0.00000				
12 4ZZ		-0.00166	0.00000	-0.01928	0.00000	0.00000				
13 4XY		0.00000	0.00000	0.00000	0.00000	0.00000				
14 4XZ		0.00000	0.00000	0.00000	0.03560	0.00000				
15 4YZ		0.00000	-0.01398	0.00000	0.00000	0.06037				
16 2 O 1S		0.08887	0.00000	-0.06975	0.00000	0.00000				
17 2S		-0.20345	0.00000	0.15375	0.00000	0.00000				
18 2PX		0.00000	0.00000	0.00000	0.49070	0.00000				
19 2PY		0.00000	0.32078	0.00000	0.00000	0.56595				
20 2PZ		-0.14191	0.00000	0.50921	0.00000	0.00000				
21 3S		-0.27033	0.00000	0.32427	0.00000	0.00000				
22 3PX		0.00000	0.00000	0.00000	0.35378	0.00000				
23 3PY		0.00000	0.17949	0.00000	0.00000	0.44356				
24 3PZ		-0.06784	0.00000	0.28713	0.00000	0.00000				
25 4XX		-0.00041	0.00000	0.00484	0.00000	0.00000				
26 4YY		-0.00412	0.00000	0.00743	0.00000	0.00000				
27 4ZZ		0.00965	0.00000	-0.03499	0.00000	0.00000				
28 4XY		0.00000	0.00000	0.00000	0.00000	0.00000				
29 4XZ		0.00000	0.00000	0.00000	-0.04165	0.00000				
30 4YZ		0.00000	-0.02339	0.00000	0.00000	-0.01932				
31 3 H 1S		0.17907	0.19092	0.09090	0.00000	-0.18058				
32 2S		0.06482	0.12037	0.07400	0.00000	-0.22518				
33 4 H 1S		0.17907	-0.19092	0.09090	0.00000	0.18058				
34 2S		0.06482	-0.12037	0.07400	0.00000	0.22518				
		9	10	11	12	13				
		(B1)--V	(A1)--V	(B2)--V	(A1)--V	(B1)--V				
Eigenvalues --		0.13577	0.24838	0.33337	0.37335	0.73665				
1 1 C 1S		0.00000	-0.12210	0.00000	0.05347	0.00000				
2 2S		0.00000	0.14892	0.00000	0.00027	0.00000				
3 2PX		0.40261	0.00000	0.00000	0.00000	-1.07309				
4 2PY		0.00000	0.00000	-0.43254	0.00000	0.00000				
5 2PZ		0.00000	-0.21094	0.00000	-0.05318	0.00000				
6 3S		0.00000	1.58075	0.00000	-1.80268	0.00000				
7 3PX		0.71157	0.00000	0.00000	0.00000	1.10559				
8 3PY		0.00000	0.00000	-1.64055	0.00000	0.00000				
9 3PZ		0.00000	-0.74592	0.00000	-1.99713	0.00000				
10 4XX		0.00000	-0.00273	0.00000	-0.00207	0.00000				
11 4YY		0.00000	-0.01265	0.00000	-0.01784	0.00000				
12 4ZZ		0.00000	-0.00457	0.00000	-0.00327	0.00000				
13 4XY		0.00000	0.00000	0.00000	0.00000	0.00000				
14 4XZ		-0.02288	0.00000	0.00000	0.00000	-0.01572				
15 4YZ		0.00000	0.00000	0.00599	0.00000	0.00000				
16 2 O 1S		0.00000	-0.00101	0.00000	-0.10774	0.00000				
17 2S		0.00000	-0.01033	0.00000	0.02902	0.00000				
18 2PX		-0.38114	0.00000	0.00000	0.00000	-0.01006				
19 2PY		0.00000	0.00000	0.12594	0.00000	0.00000				
20 2PZ		0.00000	0.05330	0.00000	-0.15411	0.00000				
21 3S		0.00000	0.10056	0.00000	2.27053	0.00000				
22 3PX		-0.52769	0.00000	0.00000	0.00000	-0.09028				
23 3PY		0.00000	0.00000	0.39608	0.00000	0.00000				
24 3PZ		0.00000	0.05062	0.00000	-0.94543	0.00000				
25 4XX		0.00000	0.00129	0.00000	-0.08344	0.00000				
26 4YY		0.00000	-0.01009	0.00000	-0.08067	0.00000				
27 4ZZ		0.00000	-0.00050	0.00000	0.00148	0.00000				
28 4XY		0.00000	0.00000	0.00000	0.00000	0.00000				
29 4XZ		0.00355	0.00000	0.00000	0.00000	0.08787				
30 4YZ		0.00000	0.00000	0.00368	0.00000	0.00000				
31 2 H 1S		0.00000	-0.04750	-0.03426	-0.01673	0.00000				
32 2S		0.00000	-1.47347	1.83003	-0.04451	0.00000				
33 4 H 1S		0.00000	-0.04750	0.03426	-0.01673	0.00000				
34 2S		0.00000	-1.47347	-1.83003	-0.04451	0.00000				

Library
Faculty of Applied Science
Rajarata University of Sri Lanka
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