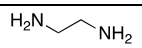
DATOS CERTAMEN 2

Datos	
en=Etilendiamina 	
K_w (H_2O, 25°C) = 10^{-14} = K_a K_b	
pH + pOH = pK_w=14	
$\%_{ionización} = \frac{[especie]_{ionizada}}{[especie]_{inicial}} x100$	
$pH = pK_a + \log\frac{[base\ conjugada]}{[ácido]}$	
$E^0{}_{celda} = E^0_{Red.\ cátodo} - E^0_{Red.\ ánodo}$	
$1J = 1C\ x\ 1V$	$1\text{F} = 96500\ \text{C}$
$E° = \frac{0.0257}{n} lnK$	
$E° = \frac{0.0592}{n} logK$	
$E = E° - \frac{0.0592}{n} logQ$	
$\Delta G° = -RTlnK$	
$\Delta G° = -nFE°$	
$R = 0.08206\ atmL/molK$	
$R = 8.314\ J/molK$	
$C = I \cdot t$	

Potenciales estándar de reducción a 25°C			
MnO_4^-(ac) + $8H^+$ + $5e^-$	→	Mn^{2+} (ac) + $4H_2O$(l)	+1.51
Au^{3+} (ac) + $3e^-$	→	Au (s)	+1.50
O_2(g) + $4H^+$(ac) + $4e^-$	→	$2H_2O$(l)	+1.23
Br_2(l) + $2e^-$	→	$2Br^-$(ac)	+1.07
NO_3^-(ac) + $4H^+$(ac) + $3e^-$	→	NO(g) + H_2O(l)	+0.92
Ag^+(ac) + e^-	→	Ag(s)	+0.80
Cu^{2+}(ac) + $2e^-$	→	Cu(s)	+0.34
SO_4^{2-}(ac) + $4H^+$(ac) + $2e^-$	→	$2H_2O$(l) + SO_2(ac)	+0.17
Cu^{2+}(ac) + e^-	→	Cu^+(ac)	+0.16
$2H^+$(ac) + $2e^-$	→	H_2(g)	0.00
Sn^{2+}(ac) + $2e^-$	→	Sn(s)	−0.13
Pb^{2+}(ac) + $2e^-$	→	Pb(s)	−0.13
Fe^{2+}(ac) + $2e^-$	→	Fe(s)	−0.44
Ni^{2+}(ac) + $2e^-$	→	Ni(s)	−0.25
Co^{2+}(ac) + 2e	→	Co(s)	−0.28
Cr^{3+}(ac) + $3e^-$	→	Cr(s)	−0.74
Zn^{2+}(ac) + $2e^-$	→	Zn(s)	−0.76
Al^{3+}(ac) + $3e^-$	→	Al(s)	−1.68
Mg^{2+}(ac) + $2e^-$	→	Mg(s)	−2.37