# The Influence of Temperature on Ozone Production under varying $NO_x$ Conditions – a modelling study

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**Abstract**

Surface ozone is a secondary air pollutant produced during the degradation of emitted volatile organic compounds (VOCs) in the presence of sunlight and nitrogen oxides ($NO_x$). Temperature directly influences ozone production through speeding up the rates of the chemical reactions producing ozone and increasing the emissions of VOCs, such as isoprene, from vegetation. In this study, we used a box model to reproduce the non-linear relationship of ozone with $NO_x$ and temperature from previous observational studies. An increase in ozone of up to 20 ppbv was due to faster reaction rates while increased isoprene emissions added a further 11 ppbv of ozone under high-$NO_x$ conditions. The increased rate of emitted VOC loss with temperature controlled the rate of $O_x$ production with temperature increasing net $O_x$ production by $\sim 1$ molecule of $O_x$ per loss of VOC. The rate of increase in ozone mixing ratios with temperature from our box model simulations was about half the rate of increase in ozone with temperature over central Europe compared to both observed values and WRF-Chem output. The missing sensitivity in our simulations compared to observations and 3D model output is related to not including stagnant atmospheric conditions in our experiment.

## 1 Introduction

Surface-level ozone ($O_3$) is a secondary air pollutant formed during the photochemical degradation of volatile organic compounds (VOCs) in the presence of nitrogen oxides ($NO_x \equiv NO + NO_2$). Due to the photochemical nature of ozone production it is strongly influenced by meteorological variables such as temperature (Jacob and Winner, 2009). Otero et al. (2016) showed that

temperature was a major meteorological driver for summertime ozone in many areas of central Europe.

Temperature primarily influences ozone production in two ways: speeding up the reaction rates of many chemical reactions leading to ozone production, and increasing emissions of VOCs from biogenic sources (BVOCs) (Sillman and Samson, 1995). While emissions of anthropogenic VOCs (AVOCs) are generally not dependent on temperature, evaporative emissions of AVOCs increase with temperature (Rubin et al., 2006). The review of Pusede et al. (2015) provides further details of the temperature-dependent processes impacting ozone production.

Regional modelling studies over the US (Sillman and Samson, 1995; Steiner et al., 2006; Dawson et al., 2007) examined the sensitivity of ozone production during a pollution episode to increased temperatures. These studies noted that increased temperatures (without changing VOC or $NO_x$-conditions) led to higher ozone levels, often exceeding local air quality guidelines. Sillman and Samson (1995) and Dawson et al. (2007) varied the temperature dependence of the PAN (peroxy acetyl nitrate) decomposition rate during simulations of the eastern US determining the sensitivity of ozone production with temperature to the PAN decomposition rate. In addition to noting the influence of PAN decomposition to ozone production, Steiner et al. (2006) correlated the increase in ozone mixing ratios with temperature over California to increased mixing ratios of formaldehyde, a secondary degradation production of many VOC and an important radical source. Steiner et al. (2006) also noted increased emissions of BVOCs at higher temperatures in urban areas with high $NO_x$ emissions also increased ozone levels with temperature.

Pusede et al. (2014) used an analytical model constrained by observations over the San Joaquin Valley, California to infer a non-linear relationship of ozone production with temperature and $NO_x$, similar to the well-known non-linear relationship of ozone production on $NO_x$ and VOC levels (Sillman, 1999). Morever, Pusede et al. (2014) showed that temperature can be used as a surrogate for VOC levels when considering the relationship of ozone under different $NO_x$ conditions.

Environmental chamber studies have also been used to analyse the relationship of ozone with temperature using a fixed mixture of VOCs. The chamber experiments of Carter et al. (1979) and Hatakeyama et al. (1991) showed increases in ozone from a VOC mix with temperature. Both studies compared the concentration time series of ozone and nitrogen-containing compounds ($NO_x$, PAN, $HNO_3$) at various temperatures and linked the maximum in ozone concentration to the decrease in PAN concentrations at temperatures greater than 303 K.

Despite many studies considering the effects of temperature on ozone production from an observational and chamber study perspective, modelling studies focusing on the detailed chemical processes of the influence of temperature on ozone production under different $NO_x$ conditions have not been performed (to our knowledge). The regional modelling studies described previously concentrated on reproducing ozone levels (using a single chemical mechanism) over regions with known meteorology and $NO_x$ conditions then varying the temperature. These modelling studies did not consider the relationship of ozone with $NO_x$ and temperature. The review of Pusede et al. (2015) also highlights a lack of modelling studies looking at the non-linear relationship of ozone on temperature under different $NO_x$ conditions.

Comparisons of different chemical mechanisms, such as Emmerson and Evans (2009) and Coates and Butler (2015), showed that different representations of tropospheric chemistry influenced ozone production. Neither of these studies examined the ozone-temperature relationship differences between chemical mechanisms. Furthermore, Rasmussen et al. (2013) acknowledged that the modelled ozone-temperature relationship may be sensitive to the choice of chemical mechanism and recommended investigating this sensitivity. Comparing the ozone-temperature relationship predicted by different chemical mechanisms is potentially important for modelling of future air quality due to the expected increase in heatwaves (Karl and Trenberth, 2003).

In this study, we use an idealised box model to determine how ozone levels vary with temperature under different $NO_x$ conditions. We determine whether faster chemical reaction rates or increased BVOC emissions have a greater influence on instantaneous ozone production with higher temperature under different $NO_x$ conditions. Furthermore, we compare the ozone-temperature relationship produced by different chemical mechanisms.

# 2 Methodology

## 2.1 Model Setup

We used the MECCA box model (Sander et al., 2005) to determine the important gas-phase chemical processes for ozone production under different temperatures and $NO_x$ conditions. The MECCA box model was set up as described in Coates and Butler (2015) and updated to include vertical mixing with the free troposphere using a diurnal cycle for the PBL height. The supplementary material includes further details of these updates.

Simulations broadly representative of urban conditions in central Europe with equinoctical

conditions were performed. The simulations started at 06:00 with a total run time of two days. Methane was fixed at 1.7 ppmv throughout the model run, carbon monoxide (CO) and ozone were initialised at 200 ppbv and 40 ppbv and then allowed to evolve freely throughout the simulation. All VOC emissions were held constant until noon of first day simulating a plume of freshly-emitted VOC.

Separate box model simulations were performed by systematically varying the temperature between 288 and 313 K (15 – 40 °C) in steps of 0.5 K. NO emissions were systematically varied between $5.0 \times 10^9$ and $1.5 \times 10^{12}$ molecules (NO) $cm^{-2}$ $s^{-1}$ at each temperature used in this study. At 20 °C, these NO emissions corresponded to peak $NO_x$ mixing ratios of 0.02 ppbv and 10 ppbv respectively, this range of $NO_x$ mixing ratios covers the $NO_x$ conditions found in pristine and urban conditions (von Schneidemesser et al., 2015).

All simulations were repeated using different chemical mechanisms to investigate whether the relationship between ozone, temperature and $NO_x$ changes using different representations of ozone production chemistry. The reference chemical mechanism was the near-explicit Master Chemical Mechanism, MCMv3.2, (Jenkin et al., 1997, 2003; Saunders et al., 2003; Rickard et al., 2015). The reduced chemical mechanisms in our study were Common Representative Intermediates, CRIv2 (Jenkin et al., 2008), Model for ozone and related chemical tracers, MOZART-4 (Emmons et al., 2010), Regional Acid Deposition Model, RADM2 (Stockwell et al., 1990) and the Carbon Bond Mechanism, CB05 (Yarwood et al., 2005). Coates and Butler (2015) described the implementation of these chemical mechanisms in MECCA. These reduced chemical mechanisms were chosen as they are commonly used by modelling groups in 3D regional and global models (Baklanov et al., 2014).

Model runs were repeated using a temperature-dependent and temperature-independent source of BVOC emissions to determine the relative importance of increased emissions of BVOC and faster reaction rates of chemical processes for the increase of ozone with temperature. MEGAN2.1 (Guenther et al., 2012) specified the temperature-dependent BVOC emissions of isoprene, Sect. 2.3 provides further details. As isoprene emissions are the most important on the global scale, we considered only isoprene emissions from vegetation (Guenther et al., 2006). Only isoprene emissions were dependent on temperature, all other emissions were constant in all simulations. In reality, many other BVOC are emitted from varying vegetation types (Guenther et al., 2006) and increased temperature can also increase AVOC emissions through increased evaporation (Rubin et al., 2006).

Table 1: Total AVOC emissions in 2011 in tonnes from each anthropogenic source category assigned from TNO-MACC_III emission inventory and temperature-independent BVOC emissions in tonnes from Benelux region assigned from EMEP. The allocation of these emissions to MCMv3.2, CRIv2, CB05, MOZART-4 and RADM2 species are found in the supplementary material.

| **Source Category** | **Total Emissions** | **Source Category** | **Total Emissions** |
|---|---|---|---|
| Public Power | 13755 | Road Transport: Diesel | 6727 |
| Residential Combustion | 21251 | Road Transport: Others | 1433 |
| Industry | 62648 | Road Transport: Evaporation | 2327 |
| Fossil Fuel | 15542 | Non-road Transport | 17158 |
| Solvent Use | 100826 | Waste | 1342 |
| Road Transport: Gasoline | 24921 | BVOC | 10702 |

## 2.2 VOC Emissions

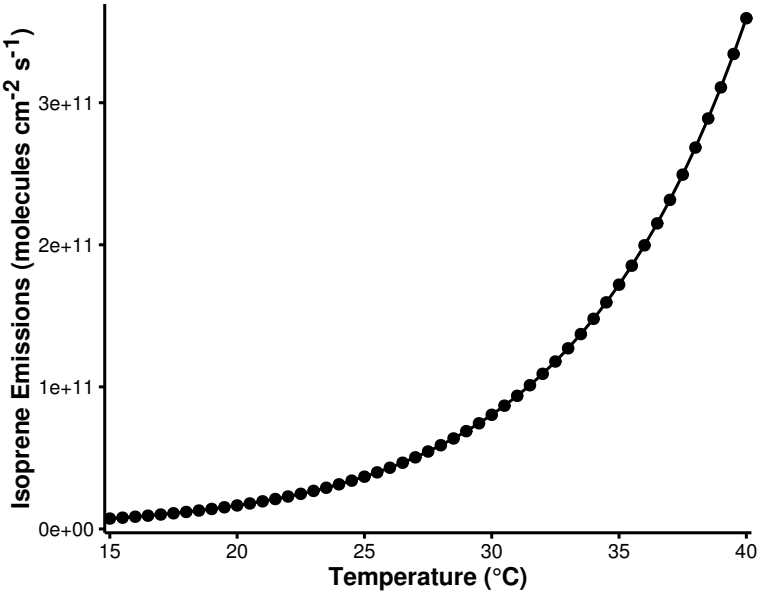
Emissions of urban AVOC over central Europe were taken from TNO-MACC_III emission inventory for the Benelux (Belgium, Netherlands and Luxembourg) region for the year 2011. TNO-MACC_III is the updated TNO-MACC_II emission inventory created using the same methodology as Kuenen et al. (2014) and based upon improvements to the existing emission inventory during AQMEII-2 (Pouliot et al., 2015).

Temperature-independent emissions of isoprene and monoterpenes from biogenic sources were calculated as a fraction of the total AVOC emissions from each country in the Benelux region. This data was obtained from the supplementary data available from the EMEP (European Monitoring and Evaluation Programme) model (Simpson et al., 2012). Temperature-dependent emissions of isoprene are detailed in Sect. 2.3.

Table 1 shows the quantity of VOC emissions from each source category and the temperature-independent BVOC emissions. These AVOC emissions were assigned to chemical species and groups based on the profiles provided by TNO. Most individual chemical species are represented by the MCMv3.2, otherwise the individual contributions of a group of VOC were further split into individual components using the detailed speciation of Passant (2002). For example, 'xylenes' are one of the component chemical groups in many source categories but the MCMv3.2 treats xylenes as the individual isomers (m-, o-, p-xylene) and the contributions of the individual isomers to a source category was provided by Passant (2002). This approach was also used in von Schneidemesser et al. (2016) to allocate AVOC emissions from different solvent sector speciations to MCMv3.2 species.

For simulations done with other chemical mechanisms, the VOC emissions represented by the MCMv3.2 were mapped to the mechanism species representing VOC emissions in each reduced

Figure 1: The estimated isoprene emissions (molecules isoprene $cm^{-2}$ $s^{-1}$) using MEGAN2.1 at each temperature used in the study.




chemical mechanism based on the recommendations of the source literature and Carter (2015). The VOC emissions in the reduced chemical mechanisms were weighted by the carbon numbers of the MCMv3.2 species and the emitted mechanism species, thus keeping the amount of emitted reactive carbon constant between simulations. The supplementary data outlines the primary VOC and calculated emissions with each chemical mechanism.

### 2.3 Temperature Dependent Isoprene Emissions

Temperature-dependent emissions of isoprene were estimated using the MEGAN2.1 algorithm for calculating the emissions of VOC from vegetation (Guenther et al., 2012). Emissions from nature are dependent on many variables including temperature, radiation and age but for the purpose of our study all variables except temperature were held constant.

The MEGAN2.1 parameters were chosen to give similar isoprene mixing ratios at 20 °C to the temperature-independent emissions of isoprene in order to compare the effects of increased isoprene emissions with temperature. The estimated emissions of isoprene with MEGAN2.1 using these assumptions are illustrated in Fig. 1 and show the expected exponential increase in isoprene emissions with temperature (Guenther et al., 2006).

The estimated emissions of isoprene at 20 °C lead to 0.07 ppbv of isoprene in our simulations while at 30 °C, the increased emissions of isoprene using MEGAN2.1 estimations lead to 0.35 ppbv of isoprene in the model. A measurement campaign over Essen, Germany (Wagner and Kuttler,

Table 2: Increase in mean ozone mixing ratio (ppbv) due to chemistry and temperature-dependent isoprene emissions from the reference temperature (20 °C) at 40 °C in the $NO_x$-regimes of Fig. 3.

| **Chemical Mechanism** | **Source of Difference** | **Increase in Ozone from 20 °C at 40 °C (ppbv)** | | |
|---|---|---|---|---|
| | | **Low-$NO_x$** | **Maximal-$O_3$** | **High-$NO_x$** |
| MCMv3.2 | Isoprene Emissions | 4.6 | 7.7 | 10.6 |
| | Chemistry | 6.8 | 12.5 | 15.2 |
| CRIv2 | Isoprene Emissions | 4.8 | 7.9 | 10.8 |
| | Chemistry | 6.0 | 11.1 | 13.7 |
| MOZART-4 | Isoprene Emissions | 4.1 | 6.7 | 10.0 |
| | Chemistry | 6.0 | 10.2 | 12.3 |
| CB05 | Isoprene Emissions | 4.6 | 7.4 | 9.8 |
| | Chemistry | 9.3 | 16.0 | 19.9 |
| RADM2 | Isoprene Emissions | 3.8 | 5.7 | 7.8 |
| | Chemistry | 8.6 | 14.1 | 17.3 |

2014) measured 0.1 ppbv of isoprene at temperature 20 °C and 0.3 ppbv of isoprene were measured at 30 °C. The similarity of the simulated and observed isoprene mixing ratios indicates that the MEGAN2.1 variables chosen for calculating the temperature-dependent emissions of isoprene were suitable for simulating urban conditions over central Europe.

# 3 Results and Discussion

## 3.1 Ozone as a Function of $NO_x$ and Temperature

Figure 2 depicts the peak mixing ratio of ozone of each simulation as a function of the total $NO_x$ emissions and temperature on the first day of simulations when using a temperature-independent and temperature-dependent source of isoprene emissions for each chemical mechanism. A non-linear relationship of ozone mixing ratios with $NO_x$ and temperature is reproduced by each chemical mechanism. This non-linear relationship is similar to that determined by Pusede et al. (2014) using an analytical model constrained to observational measurements over the San Joaquin Valley, California.

Higher ozone mixing ratios are produced when using a temperature-dependent source of isoprene emissions (Fig. 2). The highest mixing ratios of ozone are produced at high temperatures and moderate emissions of $NO_x$ regardless of the temperature dependence of isoprene emissions. Conversely, the least amount of ozone is produced with low emissions of $NO_x$ over the whole temperature range (15 – 40 °C) when using both a temperature-independent and temperature-dependent source of isoprene emissions.

The contours of ozone mixing ratios as a function of $NO_x$ and temperature can be split into three $NO_x$ regimes (Low-$NO_x$, Maximal-$O_3$ and High-$NO_x$), similar to the $NO_x$ regimes defined

Figure 2: Contours of peak ozone mixing ratios (ppbv) as a function of the total $NO_x$ emissions on the first day and temperature for each chemical mechanism using a temperature-dependent and temperature-independent source of isoprene emissions. The contours can be split into three separate regimes: High-$NO_x$, Maximal-$O_3$ and Low-$NO_x$ indicated in the figure.

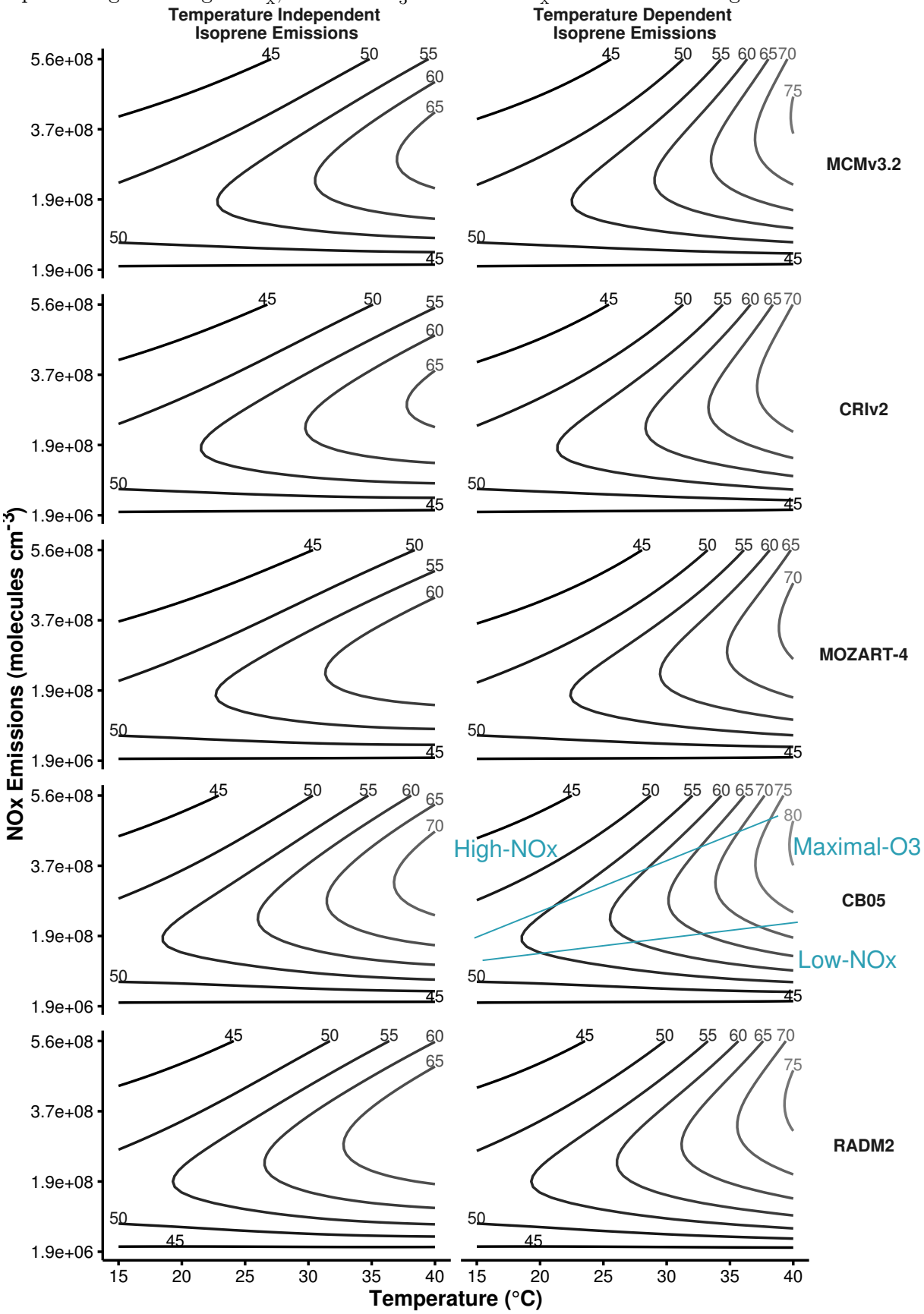
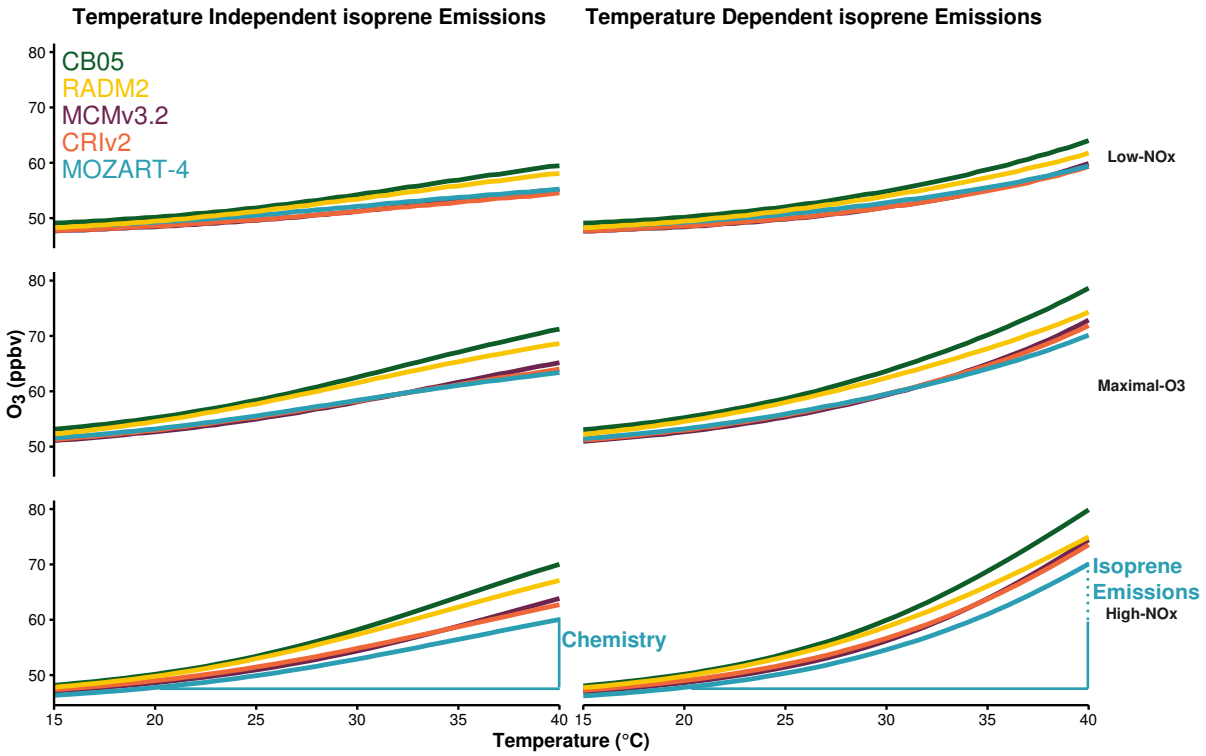


Figure 3: Mean ozone mixing ratios (ppbv) at each temperature are allocated to the different $NO_x$-regimes of Fig. 2. The differences in ozone mixing ratios due to chemistry (solid line) and isoprene emissions (dotted line) are represented graphically for MOZART-4 with High-$NO_x$ conditions. Table 2 details the differences for all chemical mechanisms and $NO_x$-conditions.




for the non-linear relationship of ozone with VOC and $NO_x$. The Low-$NO_x$ regime corresponds with regions with little increase in ozone with temperature, also called the $NO_x$-sensitive regime. The High-$NO_x$ (or $NO_x$-saturated) regime is when ozone levels increase rapidly with temperature. The contour ridges correspond to regions of maximal ozone production; this is the Maximal-$O_3$ regime. Pusede et al. (2014) showed that temperature can be used as a proxy for VOC, thus we assigned the ozone mixing ratios from each box model simulation to a $NO_x$ regime based on the ratio of $HNO_3$ to $H_2O_2$. This ratio was used by Sillman (1995) to designate ozone to $NO_x$ regimes based on $NO_x$ and VOC levels. The Low-$NO_x$ regime corresponds to $H_2O_2$:$HNO_3$ ratios less than 0.5, the High-$NO_x$ regime corresponds to ratios larger than 0.3 and ratios between 0.3 and 0.5 correspond to the Maximal-O3 regime.

The peak ozone mixing ratio from each simulation was assigned to a $NO_x$ regime based on the $H_2O_2$:$HNO_3$ ratio of that simulation. The ozone mixing ratios assigned to each $NO_x$ regime at each temperature were averaged, this is illustrated in Figure 3 for each chemical mechanism and each type of isoprene emissions (temperature independent and temperature dependent). We define the absolute increase in ozone from 20 °C to 40 °C due to faster reaction rates as the difference

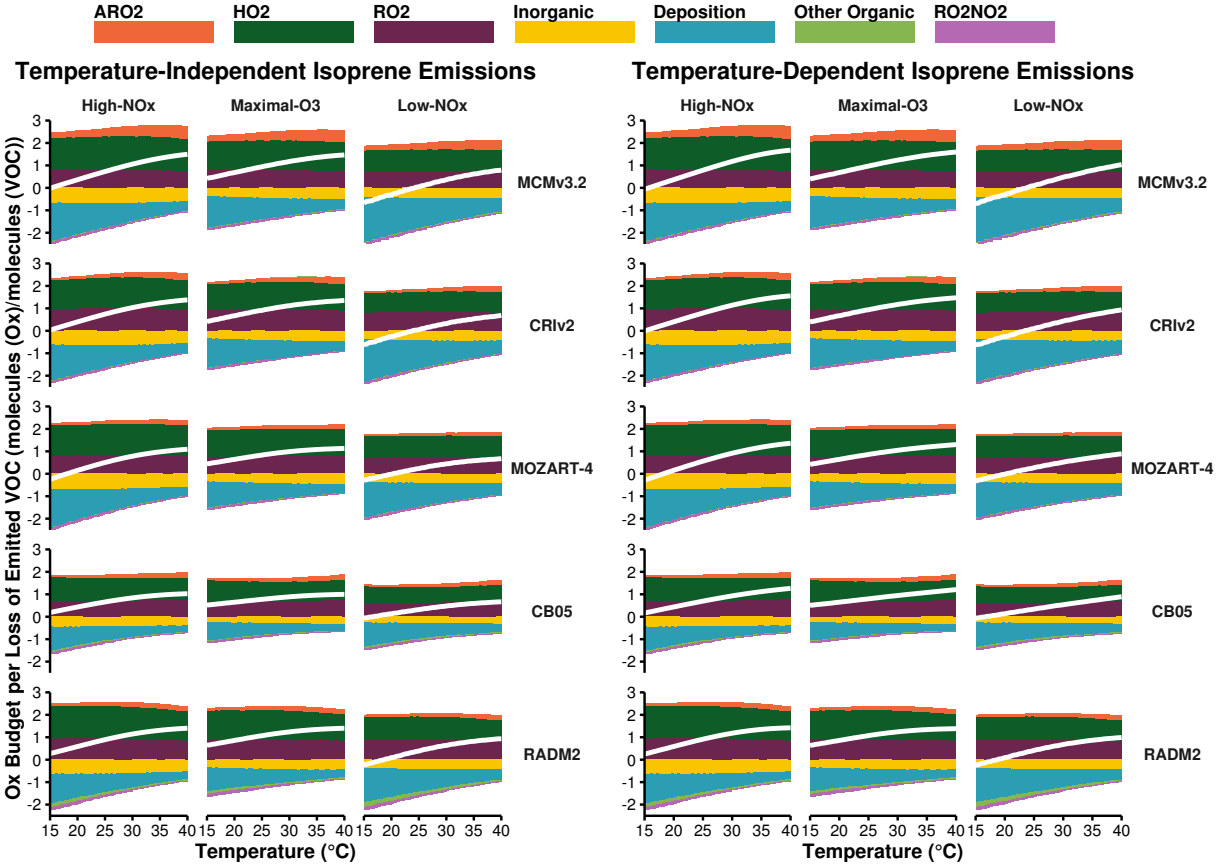
between ozone mixing ratios from 20 °C to 40 °C when using a temperature-independent source of isoprene emissions. When using a temperature-dependent source of isoprene emissions, the difference in ozone mixing ratios from 20 °C to 40 °C minus the increase due to faster chemistry, gives the absolute increase in ozone mixing ratios from increased isoprene emissions. These differences are represented graphically in Fig. 3 and summarised in Table 2.

Table 2 shows that the absolute increase in ozone with temperature due to chemistry (i.e. faster reaction rates) is larger than the absolute increase in ozone due to increased isoprene emissions for each chemical mechanism and each $NO_x$ regime. In all cases the absolute increase in ozone is largest under High-$NO_x$ conditions and lowest with Low-$NO_x$ conditions (Fig. 3 and Table 2). The increase in ozone mixing ratio from 20 °C to 40 °C due to faster reaction rates with High-$NO_x$ conditions is almost double that with Low-$NO_x$ conditions. In the Low-$NO_x$ regime, the increase of ozone with temperature using the reduced chemical mechanisms (CRIv2, MOZART-4, CB05 and RADM2) is similar to that from the MCMv3.2. Larger differences occur in the Maximal-$O_3$ and High-$NO_x$ regimes.

All reduced chemical mechanisms except RADM2 have similar increases in ozone due to increased isoprene emissions as the MCMv3.2 (Table 2). RADM2 produces 3 ppbv less ozone than the MCMv3.2 due to increased isoprene emissions in each $NO_x$ regime, indicating that this difference is due the representation of isoprene degradation chemistry in RADM2.

The Tagged Ozone Production Potential (TOPP) defined in Butler et al. (2011) is a measure of the number of molecules of ozone produced per molecule of VOC emitted. Coates and Butler (2015) compared ozone production in different chemical mechanisms to the MCMv3.2 using TOPPs and showed that less ozone is produced per molecule of isoprene emitted using RADM2 than with MCMv3.2. The degradation of isoprene has been extensively studied and it is well-known that methyl vinyl ketone (MVK) and methacrolein are signatures of isoprene degradation (Atkinson, 2000). All chemical mechanisms in our study except RADM2 explicitly represent MVK and methacrolein (or in the case of CB05, a lumped species representing both these secondary degradation products). RADM2 does not represent methacrolein and the mechanism species representing ketones (KET) is a mixture of acetone and methyl ethyl ketone (MEK) (Stockwell et al., 1990). Thus the secondary degradation of isoprene in RADM2 is unable to represent the ozone production from the further degradation of the signature secondary degradation products of isoprene, MVK and methacrolein. Updated versions of RADM2, RACM (Stockwell et al., 1997) and RACM2 (Goliff et al., 2013), sequentially included methacrolein and

Figure 4: Day-time budgets of $O_x$ normalised by the total oxidation rate of emitted VOC in the $NO_x$-regimes of Fig. 3. The white line indicates net production or consumption of $O_x$. The net contribution of reactions to $O_x$ budgets are allocated to categories of deposition, inorganic reactions, peroxy nitrates (RO2NO2), reactions of NO with HO2, alkyl peroxy radicals (RO2) and acyl peroxy radicals (ARO2). All other reactions are allocated to the 'Other Organic' category.




MVK and with these updates the TOPP value of isoprene approached that of the MCMv3.2 (Coates and Butler, 2015).

## 3.2 Ozone Production Budgets

In order to understand the temperature dependence of the ozone production directly, we examine the day-time production and consumption budgets of $O_x$ ($\equiv O_3 + NO_2 + O + O(^1D)$) normalised by the total rate of oxidation of the emitted VOC (Fig. 4). The $O_x$ budgets are assigned to each $NO_x$ regime for each chemical mechanism and type of isoprene emissions. The budgets are allocated to the net contribution of major categories, where 'HO2', 'RO2', 'ARO2' represent the reactions of NO with $HO_2$, alkyl peroxy radicals and acyl peroxy radicals respectively. 'RO2NO2' represents the net effects of peroxy nitrates, 'Deposition' represents ozone deposition, 'Inorganic' is all other inorganic contributions to $O_x$ production. Any other remaining organic reactions are included in the 'Other Organic' category. Figure 4 also illustrates the net production or consumption of $O_x$ in each case.

With each chemical mechanism, $NO_x$-condition and type of isoprene emissions (temperature dependent and temperature independent), the net $O_x$ production efficiency increases from 20 °C to 40 °C by $\sim$ 1 molecule of $O_x$ per molecule of VOC oxidised (Fig. 4). The net $O_x$ production per oxidation of emitted VOC increases with temperature due to a decrease in ozone deposition per molecule of VOC oxidation with temperature. Ozone deposition is treated as a temperature-independent process in the boxmodel with a deposition velocity of 0.5 cm $s^{-1}$. The deposition velocity ($v_d$) is typically modelled using a resistance model with

$$v_d = (R_a + R_b + R_c)^{-1}$$

where $R_a$, $R_b$ and $R_c$ are the aerodynamic, quasi-laminar boundary layer and canopy resistances, and the canopy resistance is influenced by meteorological variables such as temperature (Mészáros et al., 2009). Thus in reality, the temperature dependence of ozone deposition over vegetation all may also influence ozone production.

As the production efficiency of $O_x$ remains constant with temperature ($\sim$ 2 molecules of $O_x$ per molecule of VOC oxidised, Fig. 4), the rate of $O_x$ production is controlled by the oxidation of VOCs. Faster oxidation of VOCs with temperature speeds up the production of peroxy radicals increasing ozone production when peroxy radicals react with NO to produce $NO_2$. The review of Pusede et al. (2015) acknowledged the importance of organic reactivity and radical production to the ozone-temperature relationship. Also, the modelling study of Steiner et al. (2006) noted that the increase in initial oxidation rates of VOCs with temperature leads to increased formaldehyde concentrations and in turn an increase of ozone as formaldehyde is an important source of $HO_2$ radicals.

Increased VOC oxidation with temperature is tied to faster reaction rates and increased levels of OH with temperature. An increase of ozone with temperature goes hand in hand with an increase of OH with temperature since ozone photolysis is the dominant source of OH radicals in the atmosphere.

$$O_3 + h\nu \rightarrow O(^1D) + O_2 \quad \text{(R1)}$$

$$O(^1D) + H_2O \rightarrow OH + OH \quad \text{(R2)}$$

Furthermore, enhanced formaldehyde production from the faster degradation of VOCs increases $HO_2$ formation speeding up the reaction rate of the reaction of $HO_2$ and NO. (R3) is responsible for both ozone production (through $NO_2$ production) and OH recycling further

Table 3: Slopes ($m_{O3\text{-}T}$, ppbv per °C) of the linear fit to MDA8 values of ozone and temperature correlations in Fig. 5, indicating the increase of MDA8 in ppbv of ozone per °C. The slope of the observational data is 2.15 ppbv/°C and the slope of the WRF-Chem output is 2.05 ppbv/°C.

| **Mechanism** | **Isoprene Emissions** | **Low-$NO_x$** | | **Maximal-$O_3$** | | **High-$NO_x$** | |
|---|---|---|---|---|---|---|---|
| | | **Mixing** | **Stagnation** | **Mixing** | **Stagnation** | **Mixing** | **Stagnation** |
| MCMv3.2 | Temperature Independent | 0.28 | 1.01 | 0.51 | 1.36 | 0.59 | 0.96 |
| | Temperature Dependent | 0.42 | 1.48 | 0.74 | 2.16 | 0.93 | 2.63 |
| CRIv2 | Temperature Independent | 0.25 | 0.93 | 0.47 | 1.27 | 0.55 | 0.88 |
| | Temperature Dependent | 0.40 | 1.44 | 0.71 | 2.09 | 0.90 | 2.52 |
| MOZART-4 | Temperature Independent | 0.25 | 0.97 | 0.44 | 1.21 | 0.49 | 0.59 |
| | Temperature Dependent | 0.38 | 1.43 | 0.65 | 1.98 | 0.81 | 2.05 |
| CB05 | Temperature Independent | 0.39 | 1.30 | 0.67 | 1.72 | 0.79 | 1.45 |
| | Temperature Dependent | 0.52 | 1.72 | 0.89 | 2.44 | 1.12 | 2.94 |
| RADM2 | Temperature Independent | 0.37 | 1.31 | 0.61 | 1.64 | 0.70 | 1.28 |
| | Temperature Dependent | 0.48 | 1.68 | 0.79 | 2.22 | 0.97 | 2.49 |

illustrating the strong coupling of ozone and OH production.

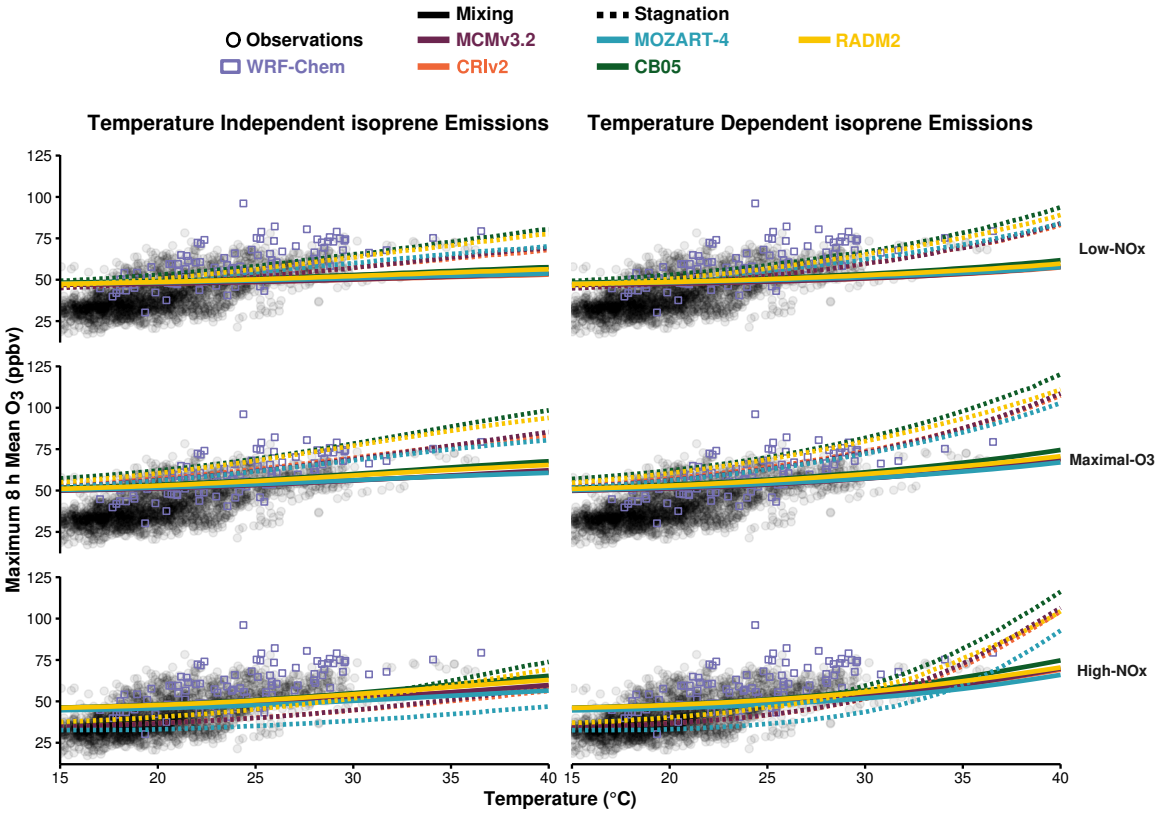
$$HO_2 + NO \rightarrow NO_2 + OH \tag{R3}$$

The net effect of peroxy nitrates on $O_x$ budgets in our study is negligible, contributing to a loss of $\sim 0.1$ molecules of $O_x$ per VOC oxidised from 20 °C to 40 °C (Fig. 4). Peroxy nitrates are produced from the reactions of acyl peroxy radicals with $NO_2$ and are an important reservoir of peroxy radicals and $NO_x$. The decomposition rate of peroxy nitrates is strongly temperature dependent and at higher temperatures the faster decomposition rate of $RO_2NO_2$ leads to faster release of peroxy radicals and $NO_2$. Dawson et al. (2007) attributed the increase in maximum 8 h ozone mixing ratios with temperature during a modelling study over the eastern US to the decrease in PAN lifetime with temperature. Steiner et al. (2006) also recognised that the decrease in PAN lifetime with temperature contributed to the increase of ozone with temperature concluding that the combined effects of increased oxidation rates of VOC and faster PAN decomposition increased the production of ozone with temperature. In our results (Fig. 4), the net effect of $RO_2NO_2$ is small since the decomposition rate is almost exactly balanced by the production from $RO_2 + NO_2$.

## 3.3 Comparison to Observations and 3D Model Simulations

This section compares the results from our idealised box model simulations to real-world observations and model output from a 3D model. Otero et al. (2016) showed that over the summer (JJA) months, temperature is the main meteorological driver of ozone production over many regions of central Europe using the interpolated observations of the maximum daily 8 h mean (MDA8) of ozone from Schnell et al. (2015) and the meteorological observational data set

Figure 5: MDA8 values of ozone from the box model simulations allocated to the different $NO_x$ regimes for each chemical mechanism with mixing (solid lines) and stagnant conditions (dashed lines). The box model ozone-temperature correlation is compared to the summer 2007 observational data (black circles) and WRF-Chem output (purple boxes).




of the ERA-Interim re-analysis. Model output from the 3D WRF-Chem regional model using MOZART-4 chemistry set-up over the European domain for simulations of the year 2007 from Mar et al. (2016) was used to further compare the box model simulations to a model including more meteorological processes than our box model.

Figure 5 compares the observational and WRF-Chem data from summer 2007 over central and eastern Germany, where summertime ozone values are driven by temperature (Otero et al., 2016), to the ozone MDA8 from the box model simulations for each chemical mechanism. Despite, a high bias in simulated ozone in WRF-Chem, the rate of change of ozone with temperature from the WRF-Chem simulations (2.05 ppbv/°C) is similar to the rate of change of ozone with temperature from the observed data (2.15 ppbv/°C). The differences in ozone production between the different chemical mechanisms with the box model are small compared to the spread of the observational and WRF-Chem data. A temperature-dependent source of isoprene with high-$NO_x$ conditions produces the most similar ozone-temperature slope to the observational data but lower than the observed ozone-temperature slope by a factor of two. In particular, the box model simulations over-predict the ozone values at lower temperatures and under-predict the ozone

values at higher temperatures compared to the observed data.

Table 3 summarises the slopes ($m_{O3\text{-}T}$) of the linear fits of the box model simulations displayed in Fig. 5 in ppbv of ozone per °C determining the rate of change of ozone with temperature. The main reason for the box model simulations being less sensitive to temperature than the observations and WRF-Chem is related to the set-up of the box model. In our simulations, we focused on instantaneous production of ozone from a freshly-emitted source of VOC and allowed transport of the chemical species into the free troposphere. This box model setup did not consider stagnant atmospheric conditions characteristed by low wind speeds slowing the transport of ozone and its precursors away from sources. Stagnant conditions have been correlated to high-ozone episodes in the summer over eastern US (Jacob et al., 1993).

Additional box model simulations were performed without vertical mixing, this box model set up simulates stagnant conditions which favours oxidant accumulation. The ozone-temperature relationship obtained with each chemical mechanism, using both a temperature-independent and temperature-dependent source of isoprene emissions and the different $NO_x$ conditions are displayed in Fig. 5. The rate of increase with ozone with temperature from these box model simulations are also included in Table 3.

In each case, simulating stagnant atmospheric conditions with the box model increased the rate of increase of ozone with temperature. The $m_{O3\text{-}T}$ results using a temperature-dependent source of isoprene and with Maximal-O3 conditions (ranging between 2.0 and 2.4 ppbv/°C) are the very similar to the slopes of the observational and WRF-Chem results (2.1 and 2.2 ppbv/°C, respectively). Stagnant conditions allowed oxidants to accumulate leading to more VOC oxidation especially at higher temperatures when the VOC oxidation rates were faster. Furthermore, the secondary degradation of the VOCs was allowed to fully evolve increasing the production of peroxy radicals which produce ozone when reacting with NO.

# 4 Conclusions

In this study, we determined the effects of temperature on ozone production using a box model over a range of temperatures and $NO_x$ conditions with a temperature-independent and temperature-dependent source of isoprene emissions. These simulations were repeated using reduced chemical mechanism schemes (CRIv2, MOZART-4, CB05 and RADM2) typically used in 3D models and compared to the near-explicit MCMv3.2 chemical mechanism.

Each chemical mechanism produced a non-linear relationship of ozone with temperature and $NO_x$ with the most ozone produced at high temperatures and moderate emissions of $NO_x$. Conversely, lower $NO_x$ levels led to a minimal increase of ozone with temperature. Thus air quality in a future with higher temperatures would benefit from dramatical reductions in $NO_x$ emissions.

Faster reaction rates at higher temperatures were responsible for a greater absolute increase in ozone than increased isoprene emissions. In our simulations, ozone production was controlled by the increased rate of VOC oxidation with temperature. Net production of $O_x$ per oxidised molecule of emitted VOC with temperature was regulated by the temperature independence of ozone deposition.

The rate of change of ozone with temperature using observational data over Europe was twice as high as when using the box model. This was consistent with our box model setup not representing stagnant atmospheric conditions that are inherently included in observational data and models including meteorology, such as WRF-Chem. When simulating stagnant conditions in the box model, the rate of increase of ozone with temperature was very similar to the rates of increase of ozone with temperature of the observational and WRF-Chem data. The sensitivity of ozone production to temperature and $NO_x$ conditions should be considered in future modelling work especially when looking at scenarios for reducing the burden of ozone pollution in the future.

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