

We would like to thank the reviewer for the review that will enable us to improve our manuscript; our responses to the review points and comments are found below.

## General Comments

**Review Point 1:** Generally, the paper needs clearer focus and to supply more model details/results/discussion pertinent to this focus. While the stated goal is to assess the temperature dependence of ozone chemistry as a function of NO<sub>x</sub>, at the end, I am not exactly sure what new has been learned. Really, what I think is missing is just a dedicated discussion section, the ‘Results’ read like a true results section, rather than combined results and discussion, so it is difficult for the reader to understand the significance of the calculations.

**Author Response:** We thank the reviewer for communicating these concerns. In order to address these concerns, we have updated each subsection in the ‘Results and Discussion’ section to include further discussion. The changes to the manuscript are detailed in the responses to the review comments on each subsection below.

**Review Point 2:** The reasons behind the ozone impacts from temperature-dependent reaction rates are unclear because it is not stated explicitly what this term includes. If I understand correctly, only temperature-dependent reaction rates,  $k(T)$ , are being tested (page 4, lines 16). Generally, the model description does not give the reader sufficient information to understand what causes the changes reported in Table 2. For example, the text states the RONO<sub>2</sub> formation is temperature dependent, at least in some mechanisms, but does this refer to the RONO<sub>2</sub> branching ratio?

**Author Response:** The reviewer is correct in saying that only temperature-dependent reaction rates,  $k(T)$ , are tested in our study. In Sect. 2.1 (Page 4, Line 2), the temperature-dependent processes included in by the chemical mechanisms is presented. We have clarified that the branching ratio of RONO<sub>2</sub> formation is not represented as a temperature-dependent process by any chemical mechanism: *Furthermore, none of the chemical mechanisms in our study represent the RONO<sub>2</sub> branching ratio as a temperature dependent process. Laboratory experiments have shown the temperature dependence of the RONO<sub>2</sub> branching ratio for some VOCs (Atkinson et al., 1987) but generally RONO<sub>2</sub> chemistry is not well known (Pusede et al., 2015) and not represented by the chemical mechanisms. Further work is required before including an adequate description of the temperature-dependence of the RONO<sub>2</sub> branching ratio is included by chemical mechanisms.*

**Review Point 3:** It is interesting to see how the five different mechanisms capture these effects; however, there is little discussion of what is learned about the different mechanisms by testing them in this way. Can the last paragraph of Sect. 3.1 be expanded? Also, I have difficulty discerning differences between panels in Fig. 2. Is there a way to highlight key differences here? Regarding Fig. 3, something has been lost in translation - there are some floating numbers in the upper left corner of the top-left panel, but the different colors are not labeled, either in the text or caption. I deduce that green is CB05, purple is MCM3.2, blue is MOZART-4, and orange and yellow are each either RADM2 or CRIv2. I would be interested to read more, not just about what causes the differences between the curves, but also the

implications for studying air quality and chemistry. Finally, the NO<sub>x</sub> regime distinction is derived from each individual model's H<sub>2</sub>O<sub>2</sub>:HNO<sub>3</sub>. Why not simply use the shape of simulated PO<sub>3</sub> versus NO<sub>x</sub>. A missing piece of information is whether maximal O<sub>3</sub> - at different temperatures - occurs at the same NO<sub>x</sub> level in each mechanism. Differences in the NO<sub>x</sub> level of maximal O<sub>3</sub> for different mechanisms have consequences for air quality decision-making.

**Author Response:** We thank the reviewer for the comments on Sect. 3.1 of the manuscript and have made changes to the manuscript based on these points. Figure 2 has been updated with a subfigure (Fig. 2b) displaying the relative differences in ozone mixing ratios in each chemical mechanism at each temperature, NO<sub>x</sub> condition from the MCMv3.2 when using a temperature-independent and temperature-dependent source of isoprene emissions. Based on this figure, a discussion of the differences between the chemical mechanisms is included in the manuscript (Page 5, Line 26): *As shown in Fig. 2b, regions of high temperatures and high NO<sub>x</sub> emissions generally lead to the largest inter-mechanism differences between ozone mixing ratios using reduced chemical mechanisms from the MCMv3.2 (up to 13 %). These differences in peak ozone mixing produced from the reduced chemical mechanisms from the MCMv3.2 in each NO<sub>x</sub> condition is consistent with Fig. 3 (described below) where RADM2 and CB05 generally produced higher ozone levels than the MCMv3.2. Also consistent with Fig. 3, CRIv2 produced the most similar amounts of ozone to the MCMv3.2 in each NO<sub>x</sub> condition whereas MOZART-4 tended to produce lower ozone mixing ratios than the MCMv3.2 in High-NO<sub>x</sub> conditions. In Fig. 3, a maximum difference of 10 ppbv between ozone mixing ratios produced using the chemical mechanisms is reached at 40 °C in the High-NO<sub>x</sub> state when using both a temperature-independent and temperature-dependent source of isoprene emissions. These differences in ozone mixing ratios between the chemical mechanisms are similar to that obtained by Coates and Butler (2015) (8 ppbv) and von Schneidemesser et al. (2016) (7 ppbv).*

The issue with display of the figures appears to be an issue with the default pdfviewer of the Chrome browser, downloading the file from the default pdfviewer in Chrome does not correct this issue. Opening the ACPD pdf using the Firefox browser and a pdfviewer extension of the Chrome browser (Kami) solves the issue and the manuscript is displayed as intended by the authors. We shall ensure that the updated manuscript does display correctly with the default pdfviewer in Chrome.

We used the H<sub>2</sub>O<sub>2</sub>:HNO<sub>3</sub> ratio to assign a NO<sub>x</sub> regime as this has been used previously in other studies for this purpose, we have updated the manuscript to include examples. Sect. 3.1 (Page 6, Line 1) was updated: *This ratio was used by Sillman (1995) and Staffelbach et al. (1997) to designate ozone to NO<sub>x</sub> regimes based on NO<sub>x</sub> and VOC levels.*

The manuscript was also updated to include further discussion on the NO<sub>x</sub> emission required for the Maximal-O<sub>3</sub> regime with each chemical mechanism: *Maximum ozone production (the contour ridges in Fig. 2a) is reached with ~ 30 % more NO<sub>x</sub> emissions using CB05 and RADM2 than the MCMv3.2 at each temperature for both a temperature-independent and temperature-dependent source of isoprene emissions. The CRIv2 and MOZART-4 chemical mechanisms require very similar NO<sub>x</sub> emissions to the MCMv3.2 at each temperature to produce maximum levels of ozone. Thus when modelling the air quality over a particular region using RADM2 and CB05, these mechanisms would predict a faster increase of ozone with temperature (i.e. more VOC-sensitive ozone production) for the same conditions. This is*

consistent with the results of Fig. 3 and the slope of the rate of increase of ozone with temperature ( $m_{O_3-T}$ ) in Table 3. The study of Mar et al. (2016) compared ozone levels produced using the 3D WRF-Chem model with MOZART-4 and RADM2 and also noted that RADM2 is in a more VOC-sensitive state than MOZART-4 for the same simulated conditions. These differences in the ozone production regime using different chemical mechanisms highlight the need for air quality studies to assess the chemical scheme used by the model, otherwise differing results from the same mitigation strategy may be produced.

**Review Point 4:** Provide a statement as to why the ozone production and consumption budget is informative for understanding the temperature dependence of ozone, i.e. what is gained compared to thinking about production alone. Also, can an equation be provided for the production and consumption budgeting? This section is in need of discussion and summary. There are many panels in Fig. 4 and it is not obvious to me what the take-away point(s) are.

**Author Response:** We agree with these review points and have provided additional information about the analysis of ozone production and consumption budgets, including equations and further description of how Fig. 4 was obtained. Section 3.2 of the manuscript (Page 7, Line 1) was updated: *Since chemical reactions contributing to both production and consumption of  $O_x$  ( $\equiv O_3 + NO_2 + O(^1D) + O$ ) have temperature-dependent rate constants, we analysed the production and consumption budgets of  $O_x$  to determine the temperature-dependent chemical processes controlling the increase of ozone with temperature in Fig. 3. The  $O_x$  budgets displayed in Fig. 4 are assigned to each  $NO_x$  regime for each chemical mechanism and source of isoprene emissions. The net production or consumption of  $O_x$  is also indicated in Fig. 4.*

Figure 4 was obtained by determining the chemical reactions producing and consuming  $O_x$  and then allocating these reactions to important categories. Reactions of peroxy radicals with NO produce  $O_x$  and the peroxy radicals are divided into ‘HO2’, ‘RO2’, ‘ARO2’ categories representing the reactions of NO with  $HO_2$ , alkyl peroxy radicals and acyl peroxy radicals respectively. Thus at each time step the  $O_x$  production rate is given by

$$k_{HO_2+NO}[HO_2][NO] + \sum_i k_{RO_2,i+NO}[RO_2,i][NO] + \sum_j k_{ARO_2,j+NO}[ARO_2,j][NO] \quad (1)$$

for each alkyl peroxy radical  $i$  and acyl peroxy radical  $j$ . The net contributions of peroxy nitrates, inorganic reactions and any other remaining organic reactions to the  $O_x$  budget are represented by the ‘RO2NO2’, ‘Inorganic’ and ‘Other Organic’ categories in Fig. 4. The net contributions of these categories to the  $O_x$  budget was calculated by subtracting the consumption rate from the production rate of the reactions contributing to each category. For example, peroxy nitrates produce  $O_x$  when thermally decomposing or reacting with OH and consume  $O_x$  when produced. Hence, at each time step the net contribution of RO2NO2 to the  $O_x$  budget was calculated by

$$\sum_k k_{RO_2NO_2,k}[RO_2NO_2,k] + \sum_k k_{RO_2NO_2,k+OH}[RO_2NO_2,k][OH] - \sum_k k_{RO_2,k+NO_2}[RO_2,k][NO_2] \quad (2)$$

for each peroxy nitrate species  $k$ . The cumulative day-time budgets were calculated by summing the net contributions of the reaction rates of each category over the day-time period. The ratio of net ozone to net

$O_x$  production was practically constant with temperature in all cases showing that using  $O_x$  budgets as a proxy for ozone budgets was suitable at each temperature in our study.

The main results of the section are summarised and discussed in the final paragraph of Sect. 3.2 (Page 8) of the updated manuscript: *Our results indicate that increased VOC reactivity and decomposition rate of peroxy nitrates are the temperature-dependent chemical processes leading to increased  $O_x$  with temperature. These results are consistent between each chemical mechanism and each  $NO_x$  condition. Of the emitted VOC, reactive VOCs such as alkenes and aromatics were responsible for the increased VOC oxidation with temperature. The study of Li et al. (2014) also showed that reactive VOC such as ethene and xylene were responsible for ozone formation in Asia. Besides decreasing  $NO_x$  emissions to address increased ozone pollution at higher temperatures, assessing the speciation of AVOC emissions and in particular the reactivity of the AVOC may also pay dividends for improved air quality, However, over longer timescales, the contribution of less-reactive VOC, such as alkanes, may also significantly contribute to ozone production (Butler et al., 2011; von Schneidmesser et al., 2016) and should be considered as part a mitigation strategy for AVOC emissions.*

**Review Point 5:** Before the authors talk about mixing, the WRF-CHEM and box MOZART- 4 results should be compared directly and discussed. The importance of atmospheric mixing appears for the first time in Section 3.3, at which time, the paper states it is the most important term in  $mO_3-T$ . At this stage in the manuscript, I am left wondering what is this paper actually about. How does Section 3.3 relate to the previous two sections? A subsequent discussion would be helpful.

**Author Response:** We thank the reviewer for these comments on Sect. 3.3, we have updated the introductory part of this section to include further details of the motivations of the analysis and also discuss the use of the metric of  $mO_3-T$  for comparing the box model, observations and WRF-Chem output. Section 3.3 (Page 8, Line 5) of the manuscript was updated: *The final step in our study was to compare how well our idealised box model simulations represent the real-world relationship between ozone and temperature. Firstly, we compared the box model simulated to the interpolated observations of the maximum daily 8 h mean (MDA8) of ozone from Schnell et al. (2015) and the meteorological observational data set of the ERA-Interim re-analysis. Using this data set, Otero et al. (2016) showed that temperature is the main meteorological driver of ozone production during the summer (JJA) months over many regions of central Europe. A further test was to compare the box model simulations to the output from a regional 3D model as 3D models include more real-world processes influencing ozone production not included in our box model. We used output from the WRF-Chem 3D model set-up over the European domain to simulate ozone production in the year 2007 using MOZART-4 chemistry, further details are described in Mar et al. (2016). Figure 5 compares the observational and WRF-Chem data from summer 2007 averaged over central and eastern Germany, where summertime ozone values are driven by temperature (Otero et al., 2016), to the MDA8 values of ozone from the box model simulations for each chemical mechanism (solid lines).*

*Directly comparing the ozone mixing ratios from our idealised box model simulations to the WRF-Chem output is difficult due to significant differences in the setups of the models. For example, WRF-Chem used gas-phase and aerosol chemistry whereas our box model setup used only gas-phase chemistry, also*

the treatment of photolysis and VOC and  $\text{NO}_x$  emissions differ between our box model and WRF-Chem. Despite these difficulties in directly comparing the ozone mixing ratios between the box model, WRF-Chem and observations, we can compare the rate of change of ozone with temperature ( $m_{\text{O}_3-T}$ ). This metric has been used to quantify future ozone pollution due to the warmer temperatures predicted by climate change (Dawson et al., 2007; Rasmussen et al., 2013) and is discussed further in the review of Pusede et al. (2015).  $m_{\text{O}_3-T}$  is calculated as the linear slope of the increase of ozone with temperature in ppbv of ozone per  $^{\circ}\text{C}$ . Table 3 summarises the calculated slopes of the box model simulations displayed in Fig. 5. Polluted areas have larger  $m_{\text{O}_3-T}$  values than rural areas corresponding to the High- $\text{NO}_x$  and Low- $\text{NO}_x$  conditions simulated in our study.

The manuscript (Page 8, Line 21) was further updated to discuss the process of mixing and its relation to the lower sensitivity in the increase of ozone with temperature from the box model simulations: *Out of the meteorological conditions not represented by our box model, stagnation is the most likely cause of the lower sensitivity of  $m_{\text{O}_3-T}$  in the box model compared to the observations and WRF-Chem output. Stagnant atmospheric conditions are characterised by low wind speeds slowing the transport of ozone and its precursors away from sources and have been correlated to high-ozone episodes in the summer over eastern US (Jacob et al., 1993). In order to investigate the sensitivity of ozone production to mixing, further box model simulations were performed without vertical mixing to approximate stagnant conditions that favour accumulation of secondary VOC oxidation products. The ozone-temperature relationship obtained with each chemical mechanism, using both a temperature-independent and temperature-dependent source of isoprene emissions and the different  $\text{NO}_x$  conditions are displayed in Fig. 5 (dotted lines).*

The final paragraph of Sect. 3.3 in the manuscript (Page 9, Line 4) was updated and additional discussion was included: *Analysis of the  $\text{O}_x$  budgets, similar to that presented in Sect. 3.2, shows an increase in absolute net production of  $\text{O}_x$  when simulating stagnant conditions compared to simulations including mixing (Fig. 4a). Moreover, the  $\text{O}_x$  budgets normalised by the chemical loss rate of VOC for the simulations without mixing show no appreciable difference to the simulations including mixing (Fig. 4b). This analysis is displayed in the supplementary material and is consistent for each chemical mechanism and each  $\text{NO}_x$  condition. Thus we conclude that the increased ozone production seen in the box model simulations with reduced mixing is due to enhanced OH reactivity from secondary VOC oxidation products.*

*A slower rate of increase of ozone with temperature with our box model was obtained compared to the rate of increase of ozone with temperature of observational and 3D model simulations. The reason for this discrepancy was that the box model did not represent stagnation conditions which are relevant to real-world conditions. The lack of mixing meant that secondary VOC oxidation products were allowed to accumulate, leading to further degradation and increased production of peroxy radicals than simulations including mixing. Thus the chemical processes driving the increase of ozone with temperature determined in Sect. 3.2 (faster VOC oxidation and peroxy nitrate decomposition) are not altered by stagnant condition but proceed at a faster rate. Thus during heatwaves, stronger reductions in  $\text{NO}_x$  are required to minimise the impact of increased ozone production at higher temperatures on the urban population.*

## Minor Comments:

**Minor Comments 1:** More information should be provided in the introduction. The three sentences in the papers first paragraph do not really follow logically. I am not familiar with the Otero paper and this single-sentence description does not stand on its own temperature was shown to be a driver of which process?

**Author Response:** We thank the reviewer for this comment and have updated the introduction (Page 1, Line 21): *In particular, heatwaves, characterised by high temperatures and stagnant meteorological conditions, are correlated with high ozone levels as was the case during the European heatwave in 2003 (Solberg et al., 2008; Vautard et al., 2005).*

**Minor Comments 2:** Fig. 2: The ozone contours are labeled left to right: 5, 50, 55, 0, 5, 0, 5. The y-axis reads: 10, 10, 30, 50. On the x-axis, the 4 of 40 has been lost.

**Author Response:** As discussed in the answer to Review Point 3, the issue with figures appears to be a problem with the default pdfviewer of the Chrome browser. We shall ensure that the updated manuscript does display correctly with the default pdfviewer in Chrome.

**Minor Comments 3:** Fig. 5: The majority of measured O<sub>3</sub> data are found at lower temperatures, so fitting the calculated O<sub>3</sub> with a straight line across the whole temperature range may not be representative.

**Author Response:** We agree with the reviewer that a linear regression may not be the best choice for the observation. However, as a metric  $m_{O_3-T}$ , defined as the linear slope of the increase of ozone with temperature, was the most appropriate with which to compare the observations not only to other studies but also to the box model and WRF-Chem output. The manuscript was updated as outlined in the response to Review Point 5 and Page 8, Line 14 was also updated to compare  $m_{O_3-T}$  of the observational data to that from other studies: *The linear slope of the observational data indicates an increase of 2.15 ppbv ozone per °C, this is comparable to the increase of ozone with temperature from other recent studies over urban areas: 2.2 ppbv/°C obtained over the Northeast US (Rasmussen et al., 2013) and Milan, Italy (2.8 ppbv/°C, Baertsch-Ritter et al. (2004)).*

**Minor Comments 4:** Fig. 5: y-axis reads 25, 50, 5, 100, 125; on the x-axis, the 4 of 40 has been lost.

**Author Response:** See response to Minor Comments 2.

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