## CHAPTER - 25 CALORIMETRY

Mass of Iron = 0.2 kg

Mass of water = 0.2 kg

Temp. of aluminium and water = 20°C = 297°k

Sp heat of Iron =  $100^{\circ}$ C =  $373^{\circ}$ k.

Sp heat of aluminium = 910J/kg-k

Sp heat of Iron = 470J/kg-k

Sp heat of water = 4200J/kg-k

Heat again =  $0.5 \times 910(T - 293) + 0.2 \times 4200 \times (343 - T)$ 

$$= (T - 292) (0.5 \times 910 + 0.2 \times 4200)$$

Heat lost = 
$$0.2 \times 470 \times (373 - T)$$

∴ Heat gain = Heat lost

$$\Rightarrow$$
 (T – 292) (0.5 × 910 + 0.2 × 4200) = 0.2 × 470 × (373 – T)

$$\Rightarrow$$
 (T – 293) (455 + 8400) = 49(373 – T)

$$\Rightarrow$$
 (T - 293)  $\left(\frac{1295}{94}\right)$  = (373 - T)

$$\Rightarrow$$
 (T – 293) × 14 = 373 – T

$$\Rightarrow T = \frac{4475}{15} = 298 \text{ k}$$

T = 298 - 273 = 25°C.

The final temp =  $25^{\circ}$ C.

2. mass of Iron = 100g

 $S_{iron} = 470 J/kg^{\circ}C$ 

water Eq of caloriemeter = 10g Let the Temp. of surface = 0°C

mass of water = 240g

Total heat gained = Total heat lost.

So, 
$$\frac{100}{1000} \times 470 \times (\theta - 60) = \frac{250}{1000} \times 4200 \times (60 - 20)$$

$$\Rightarrow$$
 47 $\theta$  – 47 × 60 = 25 × 42 × 40

$$\Rightarrow \theta = 4200 + \frac{2820}{47} = \frac{44820}{47} = 953.61^{\circ}\text{C}$$

3. The temp. of  $A = 12^{\circ}C$ 

The temp. of B =  $19^{\circ}$ C

The temp. of  $C = 28^{\circ}C$ 

The temp of  $\Rightarrow$  A + B = 16°

The temp. of  $\Rightarrow$  B + C = 23°

In accordance with the principle of caloriemetry when A & B are mixed

$$M_{CA} (16 - 12) = M_{CB} (19 - 16) \Rightarrow CA4 = CB3 \Rightarrow CA = \frac{3}{4} CB$$
 ...(1)

And when B & C are mixed

$$M_{CB} (23 - 19) = M_{CC} (28 - 23) \Rightarrow 4CB = 5CC \Rightarrow CC = \frac{4}{5}CB$$
 ...(2)

When A & c are mixed, if T is the common temperature of mixture

$$M_{CA} (T - 12) = M_{CC} (28 - T)$$

$$\Rightarrow \left(\frac{3}{4}\right) CB(T-12) = \left(\frac{4}{5}\right) CB(28-T)$$

$$\Rightarrow$$
 T =  $\frac{628}{31}$  = 20.258°C = 20.3°C