

The following table is from “Note on the spectral lines of hydrogen” by J. J. Balmer dated 1885. Numerical values are wavelengths in units of  $10^{-10}$  meter.

Investigator	$H_\alpha$	$H_\beta$	$H_\gamma$	$H_\delta$	$H_\epsilon$	$H_\zeta$	$H_\eta$	$H_\vartheta$	$H_\iota$
Van der Willigen	6565.6	4863.94	4342.80	4103.8	—	—	—	—	—
Angstrom	6562.10	4860.74	4340.10	4101.2	—	—	—	—	—
Mendenhall	6561.62	4860.16	—	—	—	—	—	—	—
Mascart	6560.7	4859.8	—	—	—	—	—	—	—
Ditscheiner	6559.5	4859.74	4338.60	4100.0	—	—	—	—	—
Huggins	—	—	—	—	—	3887.5	3834	3795	3767.5
Vogel	—	—	—	—	3969	3887	3834	3795	3769 <sup>†</sup>

(<sup>†</sup>The value given in the paper is 6769 which is an obvious typo.)

Balmer discovered the following formula for fitting the data.

$$\lambda = \frac{m^2}{m^2 - 2^2} \times 3645.6 \times 10^{-10} \text{ meter}$$

Symbol  $\lambda$  is spectral line wavelength and parameter  $m$  is from the following table.

$$m = \begin{array}{cccccccccc} H_\alpha & H_\beta & H_\gamma & H_\delta & H_\epsilon & H_\zeta & H_\eta & H_\vartheta & H_\iota \\ 3 & 4 & 5 & 6 & 7 & 8 & 9 & 10 & 11 \end{array}$$

Let  $\beta$  be the model coefficient for  $\lambda$ . Using linear regression and the above data we obtain

$$\beta = 3645.3 \times 10^{-10} \text{ meter}$$

The currently accepted value is

$$\beta = \frac{4}{R_H} = 3647.1 \times 10^{-10} \text{ meter}$$

where  $R_H$  is the Rydberg constant for hydrogen

$$R_H = 1.09677576 \times 10^7 \text{ meter}^{-1}$$

Balmer’s coefficient is within 0.04% of the modern value.

$$100 \times \frac{4/R_H - 3.6456}{4/R_H} = 0.04$$