

Bohr model

By an argument that is no longer relevant the Bohr model has for hydrogen energy levels

$$E_n = -\frac{\alpha^2 mc^2}{2n^2}$$

By the kinetic energy relation

$$v^2 = -\frac{2E_n}{m}$$

velocity v reduces to

$$v = \frac{\alpha c}{n}$$

The Bohr model quantizes orbital angular momentum as

$$mvr_n = n\hbar$$

Hence the radius is

$$r_n = \frac{n\hbar}{mv} = \frac{n^2\hbar}{\alpha mc}$$

For $n = 1$ and $m = m_e$ we have

$$E_1 = -13.6057 \text{ eV}, \quad r_1 = 5.29177 \times 10^{-11} \text{ meter}$$

For reduced electron mass

$$m = \frac{m_e m_p}{m_e + m_p}$$

the result is

$$E_1 = -13.5983 \text{ eV}, \quad r_1 = 5.29465 \times 10^{-11} \text{ meter}$$