

Let A_{nm} be the transition rate for the spontaneous emission process $\psi_n \rightarrow \psi_m$.

$$A_{nm} = \frac{e^2}{3\pi\epsilon_0\hbar c^3} \omega_{nm}^3 |r_{nm}|^2$$

The transition frequency ω_{nm} is given by Bohr's frequency condition.

$$\omega_{nm} = \frac{1}{\hbar}(E_n - E_m)$$

The transition probability (multiplied by a physical constant) is

$$|r_{nm}|^2 = |x_{nm}|^2 + |y_{nm}|^2 + |z_{nm}|^2$$

where

$$\begin{aligned} x_{nm} &= \int \psi_m^* (r \sin \theta \cos \phi) \psi_n dV \\ y_{nm} &= \int \psi_m^* (r \sin \theta \sin \phi) \psi_n dV \\ z_{nm} &= \int \psi_m^* (r \cos \theta) \psi_n dV \end{aligned}$$

Let us compute A_{21} for hydrogen. The energy levels for hydrogen are

$$E_n = -\frac{e^2}{8\pi\epsilon_0 a_0 n^2}$$

where a_0 is the Bohr radius

$$a_0 = \frac{4\pi\epsilon_0\hbar^2}{e^2 m_e} = 5.29 \times 10^{-11} \text{ meter}$$

For the transition frequency we have

$$\omega_{21} = \frac{1}{\hbar}(E_2 - E_1) = 1.55 \times 10^{16} \text{ second}^{-1}$$

To compute the transition probability $|r_{21}|^2$ we need to consider all four eigenstates for $n = 2$.

n	ℓ	m_ℓ
2	1	1
2	1	-1
2	1	0
2	0	0

The following table shows the probability for every possible transition of ψ_2 to ψ_1 .

	$\psi_{2,1,1} \rightarrow \psi_{1,0,0}$	$\psi_{2,1,-1} \rightarrow \psi_{1,0,0}$	$\psi_{2,1,0} \rightarrow \psi_{1,0,0}$	$\psi_{2,0,0} \rightarrow \psi_{1,0,0}$
$x_{21} =$	$-\frac{128}{243} a_0$	$\frac{128}{243} a_0$	0	0
$y_{21} =$	$-\frac{128}{243} i a_0$	$-\frac{128}{243} i a_0$	0	0
$z_{21} =$	0	0	$\frac{128}{243} \sqrt{2} a_0$	0
$ r_{21} ^2 =$	$\frac{32768}{59049} a_0^2$	$\frac{32768}{59049} a_0^2$	$\frac{32768}{59049} a_0^2$	0

The transition $\psi_{2,0,0} \rightarrow \psi_{1,0,0}$ has zero probability. For the remaining transitions, the probability $|r_{21}|^2$ is independent of m_ℓ .

Now that we have $|r_{21}|^2$ we can compute a numerical value for A_{21} .

$$A_{21} = \frac{e^2}{3\pi\epsilon_0\hbar c^3} \times \omega_{21}^3 \times \frac{32768}{59049} a_0^2 = 6.27 \times 10^8 \text{ second}^{-1}$$

Here is A_{21} as a product of fundamental constants.

$$A_{21} = \frac{e^2}{3\pi\epsilon_0\hbar c^3} \times \underbrace{\left(\frac{3e^4 m_e}{128\pi^2 \epsilon_0^2 \hbar^3}\right)^3}_{\omega_{21}^3} \times \underbrace{\frac{32768}{59049} \left(\frac{4\pi\epsilon_0 \hbar^2}{e^2 m_e}\right)^2}_{|r_{21}|^2} = \frac{e^{10} m_e}{26244 \pi^5 \epsilon_0^5 \hbar^6 c^3}$$

The parameters $n = 2$ and $m = 1$ contribute the following numerical factor to A_{21} .

$$\underbrace{\left(-\frac{1}{2^2} + \frac{1}{1^2}\right)^3}_{\text{from } (E_2 - E_1)^3} \times \underbrace{\frac{32768}{59049}}_{\text{from } |r_{21}|^2} = \frac{512}{2187} = \frac{2^9}{3^7}$$

Multiplying out numerical factors yields the numerical factor shown above for A_{21} .

$$\frac{1}{3} \times \underbrace{\left(\frac{1}{32}\right)^3}_{\text{from } (E_n - E_m)^3} \times \underbrace{4^2}_{\text{from } a_0^2} \times \frac{512}{2187} = \frac{1}{26244} = \frac{1}{2^2 3^8}$$

Let us analyze the units involved in computing A_{nm} . For the coefficient of A_{nm} we have

$$\frac{e^2}{3\pi\epsilon_0\hbar c^3} \propto \frac{\text{ampere}^2 \text{second}^2}{\underbrace{\epsilon_0}_{\left(\frac{\text{ampere}^2 \text{second}^4}{\text{kilogram meter}^3}\right)} \underbrace{e^2}_{\left(\frac{\text{kilogram meter}^2}{\text{second}}\right)} \underbrace{\hbar}_{\left(\frac{\text{meter}^3}{\text{second}^3}\right)} \underbrace{c^3}_{\text{meter}^3}} = \frac{\text{second}^2}{\text{meter}^2}$$

For the transition frequency we have

$$\omega_{21} = \frac{3e^4 m_e}{128\pi^2 \epsilon_0^2 \hbar^3} \propto \frac{\underbrace{e^4}_{\left(\frac{\text{ampere}^4 \text{second}^8}{\text{kilogram}^2 \text{meter}^6}\right)} \underbrace{m_e}_{\left(\frac{\text{kilogram}^3 \text{meter}^6}{\text{second}^3}\right)}}{\underbrace{\epsilon_0^2}_{\left(\frac{\text{ampere}^2 \text{second}^4}{\text{kilogram meter}^3}\right)} \underbrace{\hbar^3}_{\left(\frac{\text{kilogram}^3 \text{meter}^6}{\text{second}^3}\right)}} = \text{second}^{-1}$$

For the Bohr radius we have

$$a_0 = \frac{4\pi\epsilon_0 \hbar^2}{e^2 m_e} \propto \frac{\underbrace{\epsilon_0}_{\left(\frac{\text{ampere}^2 \text{second}^4}{\text{kilogram meter}^3}\right)} \underbrace{\hbar^2}_{\left(\frac{\text{kilogram}^2 \text{meter}^4}{\text{second}^2}\right)}}{\underbrace{e^2}_{\left(\frac{\text{ampere}^2 \text{second}^2}{\text{kilogram}}\right)} \underbrace{m_e}_{\text{kilogram}}} = \text{meter}$$

Hence

$$A_{nm} \propto \frac{\text{second}^2}{\text{meter}^2} \times \underbrace{\text{second}^{-3}}_{\omega_{nm}^3} \times \underbrace{\text{meter}^2}_{a_0^2} = \text{second}^{-1}$$

The coefficients B_{12} (absorption) and B_{21} (induced emission) can be computed from A_{21} .

$$B_{21} = \frac{c^2}{2h\nu^3} A_{21} = \frac{4.25 \times 10^{58}}{\nu^3}$$

$$B_{12} = \frac{g_2}{g_1} B_{21} = \frac{2}{1} B_{21} = \frac{8.50 \times 10^{58}}{\nu^3}$$

Symbol g_n is the multiplicity associated with energy level n .