

Representing Kekule Cells in Stable Molecules

Aaron Germuth and Alex Aravind

Department of Computer Science
University of Northern British Columbia
E-mail: (germuth,csalex)@unbc.ca

August 15, 2013

Abstract

Silicon transistors may slowly be reaching their minimum size. Possible alternatives to allow the continual decrease of circuitry size are desired. One possibility is the use of π conjugated hydrocarbons as circuitry elements. Certain polycyclic polyunsaturated hydrocarbons have already shown promise in this area. Alternating paths of double and single bonds in π conjugated hydrocarbons have to shown the ability to conduct electricity. An alternating path can be toggled by sending a soliton across it. When such a path is toggled, other alternating paths can be created or destroyed. This suggests a possible switching behavior. Kekulé theory attempts to characterize this phenomenon. The Kekulé cell captures the switching behavior in a mathematical way. All possible cells have been reduced to a classification. Therefore, if we can find a molecule for every classification, we have found a molecule for every possible switching behavior. We attempt to, given any cell, find a suitable molecule which has that cell. A genetic algorithm is used to evolve a population of graphs towards the required switching behavior. Most graphs obtained from the genetic algorithm resemble realistic graphs which in theory, could be used as circuitry elements. The genetic algorithm is capable of quickly producing results for all cells of rank 5 or lower, and most cells of rank 6.

1 Introduction

First a review of organic chemistry is given in section 1.1. A brief review of graph theory is given in Section 1.2. Section 1.3 begins to introduce Kekulé theory. Section 1.4 introduces the current problem, and our current work. Definitions for many terms are given after each subsection.

1.1 The Chemistry

Carbon is the 6th element of the periodic table and backbone of all organic molecules. It is tetravalent, meaning it has 4 valence electrons available to form chemical covalent bonds. In such bonds,

atoms share electrons relatively equally between them. For example, a neutral chlorine atom has 7 electrons in its outer shell. According to the octet rule, most atoms are stable with 8 electrons. For example, chlorine commonly dimerizes with itself, sharing one electron each way to give each chlorine 8 shared electrons in its outer shell (see Figure 1).

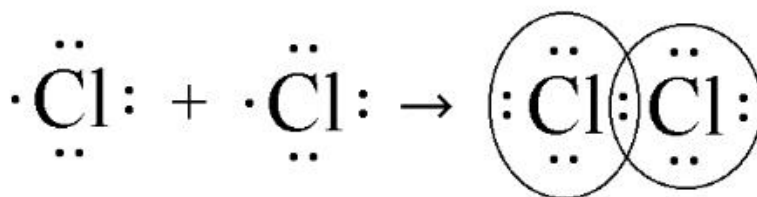


Figure 1: This shows the forming of a covalent bond between two neutral chlorine atoms. Each atom initially has 7 electrons, until they each share an electron with each other, yielding them both 8 electrons. Carbon only possesses 4 electrons in its neutral state, and therefore usually shares electrons with 4 other atoms.

Each possible bonding partner of carbon will contribute a single electron to the covalent bond, meaning carbon, in its full valency will have 8 electrons. This is considered a full octet, and the stable configuration of carbon. However, 4 bonding partners is not always needed in order to be stable. If Carbon was bonded to 3 other atoms, but one of them was able to share 2 electrons (a double bond), carbon would still have a full valency of 8 electrons (see Figure 2).

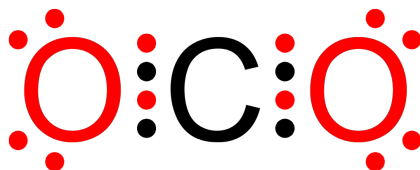


Figure 2: This figure shows the Lewis Structure of Carbon dioxide. Each oxygen is capable of sharing 2 electrons. Since carbon has 4 electrons normally, they are split evenly between the oxygen atoms. This results with every atom having 8 electrons. Carbon is said to be double-bonded to each oxygen.

According to the Valence Bond Theory, such single and double bonds have different electron structures. Single bonds are named "sigma (σ) bonds", and are caused by the head-on overlapping between atomic orbitals. Sigma Bonds are the strongest type of covalent bond. Double bonds are named "pi (π) bonds", and result from electron overlap in the nodal plane of both atoms. Double bonds are weaker, but still resilient.

Many organic molecules display alternating paths of single and double bonds. In this configuration, every atom is connected with precisely one double bond. This is named conjugation, and can increase molecular stability. Conjugation implies the possibility of different resonance structures, where every double bond shifts in one direction to obtain a different representation of the same molecule. The amount of such resonance structures is a measure of stability. The actual molecule results from the superposition of all resonance contributors.

Each of the nodal lobes involved in pi bond conjugation allow their electrons to delocalize across the molecule. Since electrical current is the flow of electrons, such alternating paths are postulated to be electrically conductive [1, 4]. When electricity (specifically a ‘soliton’) is ran through a conjugated system, all single and double bonds are interchanged [4].

This allows some molecules to display a ‘switching’ property. A channel is defined as a path through a molecule. A channel is open when the path is alternating, and closed otherwise. When some alternating paths are toggled, other paths may be opened or closed. This resembles the behavior of circuitry, and is proposed as a basis for molecular computation. This proposal is quite speculative.

1.1.1 Definitions

Organic Compound An organic compound is any member of a large class of gaseous, liquid, or solid chemical compounds whose molecules contain carbon.

Valence Electron An electron that can participate in the formation of a chemical bond;

Tetravalence The state of an atom with four electrons available for covalent chemical bonding.

Covalent Bond A chemical bond that involves the sharing of electron pairs between atoms.

Octet rule A chemical rule of thumb that states that atoms tend to combine in such a way that they each have eight electrons in their valence shells.

Lewis Structure Diagrams that show the bonding between atoms of a molecule by drawing each electron as a ‘dot’.

Double bond A chemical bond between two chemical elements involving four bonding electrons instead of the usual two.

Valence Bond Theory A basic theory developed to use to explain chemical bonding. It explains how the atomic orbitals of the atoms combine to give individual chemical bonds when a molecule is formed.

Sigma Bond The strongest type of covalent chemical bond. They are formed by head-on overlapping between atomic orbitals. Commonly referred to as a ‘single bond’.

Pi Bond Commonly referred to as a ‘double bond’.

1.2 Graph Theory

1.3 Kekulé Theory

1.4 The Problem

Moore’s Law (see Figure 3) in 1965 [5] predicted that, over the history of computing hardware, the number of transistors on a single circuit will double approximately every two years (see Figure

One alternative possibility to allow smaller circuitry is to use single molecules as computational elements. π -conjugated organic molecules have been the main focus, and there have been solutions to make them perform as wires [9], diodes [10, 11], resistors [12, 13], transistors [14, 15, 16, 17], and switches [18, 19]. Following this it was shown that a single molecule can represent an entire Boolean function or logic gate (AND[20], OR[21], IF THEN (implication)[22], NAND[23], and NOR[24]). M. H. van der Veen [3] has shown that a single molecule can facilitate all 16 fundamental logic operations. These molecules can even be combined to achieve bi-functional logic with increased complexity (another cite).

In π -logic, molecules have terminals which consist of atoms where connections outside the molecule can be attached. π -conjugated wires [9] or omniconjugated molecules [7] have been suggested to interconnect separate molecules. A pair of terminals is called a pathway. If there

exists an alternating single-double bond path from terminal to terminal, the path is deemed 'open'. Open paths have been shown to conduct electricity to a much higher extent than closed paths [8]. A path can be toggled (all single bonds become double bonds and vice versa) by sending an electrical 'soliton' over the path [4]. A molecule can contain multiple intersecting paths, and in this way, when one path is toggled, another path may be opened or closed (see Figure 4).

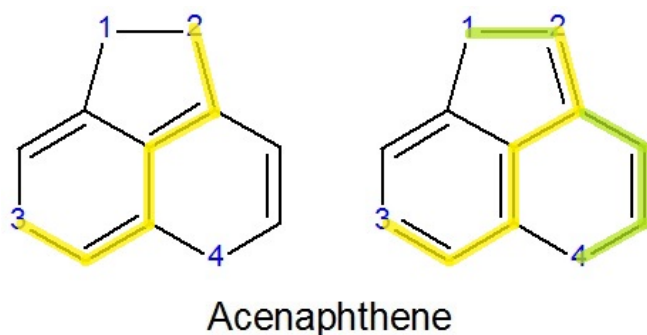


Figure 4: Acenaphthene displaying switching behavior. There exists an alternating path from terminal 2 to 3 (shown in yellow). When this path is toggled (right diagram), the path from terminal 1 to 4 (shown in green) is also opened.

To begin the analysis using π -logic, all possible double bond patterns of the molecule are deduced. Two terminals are selected for an output, and whether the path between them is open or closed represents an output of '1' or '0' respectively. The other terminals are paired, and each pair represents an input of '1' or '0', again depending on the openness of the path. From this you can observe all the different Kekulé structures of the molecule and map the inputs to the outputs, resulting with a Boolean function. If the terminals are paired in a different orientation, you may redo the process and have the same molecule represent another Boolean function.

Kekulé theory [1, 2] was introduced to investigate and systematize this qualitative switching behavior. In Kekulé theory, terminals are called ports, and pathways are called channels. A Kekulé state is a configuration of bonds such that every node (other than the ports) has precisely one double bond. Ports may or may not have a double bond. Each Kekulé state (or resonance structure) of the molecule has a so called port assignment, the set of ports which contain a double bond. A Kekulé cell then consists of the set of port assignments for every resonance structure of that molecule (see Figure 5). The rank of a cell or molecule is defined as its number of ports.

The switching behavior of a molecule is completely determined by its Kekulé cell [1, 2]. This allows us to search in the opposite direction from before. Rather than searching for any switching behavior a given molecule may exhibit, we can take a required switching behavior needed in a circuitry element, and search for a molecule who has this behavior.

Hesselink [1] has developed a method to classify all possible Kekulé cells, and has given classifications for all cells with ports ≤ 6 . Therefore, any switching behavior can be converted to a cell, and any cell can be converted to one of Hesselink's [1] classified cells.

The next step is to then, given a Kekulé cell, find a corresponding molecule which has that cell. To find molecules which have a specific cell, we abstract molecules as planar undirected graphs, where nodes represent atoms, and edges their chemical covalent bonds. These graphs

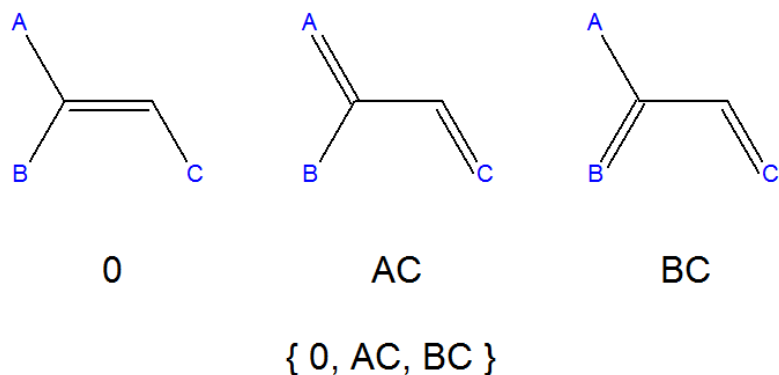


Figure 5: The Kekulé cell of a simple molecule (2-methylbutene). The 3 structures indicate the 3 possible arrangements of double bonds. The second row indicates the port assignment (which ports contain a double bond) for each structure, and the third row is the Kekulé cell for the entire molecule.

are hydrogen-suppressed as hydrogen plays no role in conjugation. Hesselink [1] has shown a recursive method which can find graphs for any cell with rank ≤ 5 and 210 out of 214 cells with rank 6.

However, Hesselink made no attempt to achieve graphs which represent realistic molecules. In regards to this, he mentions some methods which can alter a graph without changing its cell [2, 3]. [2] provides a method to split a node of high degree into two nodes of lower degree (or vice versa). It can also be used to shift edges between certain nodes. [3] describes a topological algorithm which can be used for the axiomatic construction of realistic omniconjugated molecules. However, these methods have yet to be used for this application, and may fall short. We propose an alternative, possibly easier solution for obtaining more realistic graphs. We use a genetic algorithm to create a population of graphs, and evolve them towards the desired Kekulé cell.

1.5 Overview

Section 2 consists of the restrictions used to obtain more realistic graphs, and the reasoning behind them. Section 3 explains the genetic algorithm, and how certain genetic operations are applied on graphs. It also discusses graph alterations used. Section 4 presents the results of the genetic algorithms and assesses its validity. This includes how realistic the obtained graphs are, and whether these structures can be synthesized chemically.

2 Restrictions

2.1 Degree

Carbon is tetravalent, meaning it has 4 valence electrons available to form covalent chemical bonds. This means (in most cases), carbon is only stable when bonded with 4 bonding electrons, in order to complete its octet. These 4 electrons may be provided from 4 distinct molecules, or there could be one or more molecules providing multiple electrons.

In π conjugated systems, every carbon atom contains precisely one double bond. Double bonds hold twice the amount of electrons and therefore each carbon can be at most connected to 3 distinct atoms. This is seen in all polycyclic polyunsaturated hydrocarbons. In this way, each node in our graphs has a maximum degree of 3.

The situation is slightly more complex for ports. Ports are not required to contain a double bond, and different resonance structures of the same graph may or may not contain double bonds at every port. However, all cells in Hesselink's classification, are 'flexible'. Flexible in this context, means that every port in the graph has atleast one port assignment where that port contains a double bond. And since any cell can be translated / port renamed to obtain one of the classified cells, it is safe to assume every port will contain a double bond in atleast one resonance structure, limiting ports to a degree of 3 as well.

Ports must also be able to interact with the outside world, by definition. If we assume this connection is through molecular wires, as in [9], or through an omniconjugated molecule [3], then each port must hold back one of its valence electrons to be able to connect to atoms outside the molecule. For this reason, we restrict ports to have a maximum degree of 2.

2.2 Cycles

When cycles within graphs are created, the concept of aromaticity is introduced. In Chemistry, conjugated ring formations are called aromatic, and the molecule must meet certain criteria or the ring will not be stable. This means any cycles formed within our graphs must also meet these criteria if they are to be considered stable. Aromatic molecules show increased stability compared to normal non-aromatic compounds. Anti-aromatic rings show decreased stability. In order to be considered aromatic a molecule must have:

1. A delocalized conjugated π system, resulting from alternating single and double bonds
2. Coplanar structure, where all atoms of the ring are contained in the same plane
3. Contributing atoms arranged in one or more cycles
4. The number of contributing π delocalized electrons must be even, but not a multiple of 4 ($4n + 2$, where $n = 0, 1, 2, \dots$)

All of our graphs will by default meet the first condition. The second condition is difficult to determine. Aromatic rings are planar because it is comprised of multiple conjugated bonds.

Each π conjugated bond results from the overlap in the nodal planes of each bonding atom. This creates a cloud of electron probability above and below the covalent bond. Each cloud generates a small electromagnetic field, which works to keep the ring flat (need cite here). Often non-aromatic molecules will align themselves in one plane to achieve aromatic status, whereas anti-aromatic molecules will dis-align themselves to avoid anti-aromaticity (making them non-aromatic). A consequence of planarity is that any graph of the molecule should be able to be visualized such that no edges are intersecting. (Currently we don't 100% detect whether graphs are planar or not).

2.2.1 Angle Strain

One factor attributing to whether a molecule can adopt a planar configuration is angle strain. When carbon contains only one double bond and two additional single bonds, its preferred conformation is trigonal planar. In this configuration, each atom is 120 degrees from each other and contained in the same plane. Smaller rings such as cyclobutadiene ([4] annulene) force angles of 90 between atoms if they are to remain planar. For this reason, cyclobutadiene adopts a non-planar configuration and is not aromatic. Larger rings can face the same difficulties, for example cyclodecapentene ([10] annulene) can't adopt a planar configuration as the bond angles would be 144 degrees. For this reason, cyclodecapentene is also not considered aromatic. Some larger rings such as [18] annulene can adopt an aromatic conformation, but rings of such sizes are quite rare. We limit cycles to size 5-7.

Cycles of length 5 and 7 have been shown to be stable in compounds such as azulene (see Figure 7). Cycles of length 6 represent benzene, and are generally accepted as the main contributor to aromatic rings. Any cycles of other sizes which appear can usually be artificially altered, through cycle expansion or contraction (see Section 3.4).

2.2.2 Cycle Connectivity

In regards to the third condition, groups of rings must also be considered. Instead of a single ring being aromatic, a group of rings may be deemed aromatic. In most ring systems, any two rings of a system share at most two atoms. This preserves the geometry of each cycle individually. This is seen in naphthalene, both hexagons share two common atoms. Even in complex polycyclic molecules such as pyracylene (see Figure 6), each cycle shares at most two nodes with any other cycle. Two cycles cannot share a single node (spiro-compound) as the shared node would have a degree of 4. Notice in Figure 6, the unrealistic compound has two cycles which share 4 nodes. To account for this, all cycles of each graph are found, and their intersections are computed. Unrealistic graphs are discarded.

2.2.3 Huckel's Rule

The fourth condition is the most difficult. Due to the filling of quantum-mechanical electron orbitals, only aromatic molecules with $4n+2$ electrons exhibit aromaticity. Compounds with $4n$ electrons exhibit anti-aromaticity. Groups of cycles add their delocalized electrons together to be one

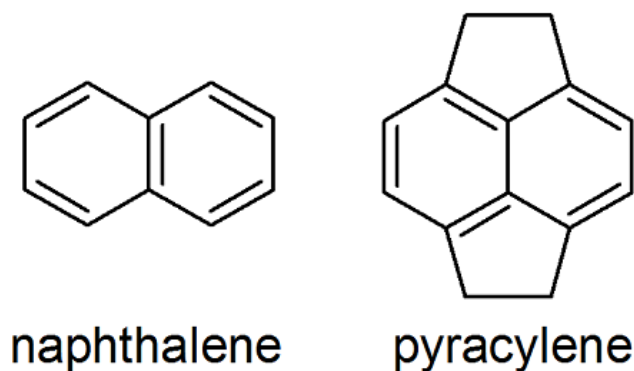


Figure 6: Naphthalene has two benzene rings with 2 carbon atoms in common. Even in complex polycyclic hydrocarbons such as pyracylene, every ring shares only 2 carbon atoms.

single π conjugated system (see Figure 7). This condition alone implies we must limit all ring systems to size of length 2, 6, or 10. However, Huckel's rule is mainly meant for monocyclic systems and has been shown to be inaccurate for many compounds containing more than three fused aromatic rings (see Figure 8) [25]. To account for this, in polycyclic systems, certain extensions to the rule have been added to better predict aromaticity [26].

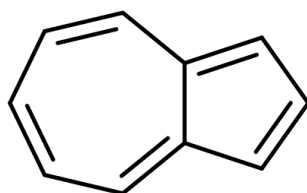


Figure 7: Azulene, a structure composed of an aromatic heptene and pentene ring attached together. The total amount of π contributing electrons is 10, 7 from heptene and 3 from pentene. Since $10 = 4n + 2$ (where $n = 2$), Huckel's rule correctly predicts this structure to be aromatic.

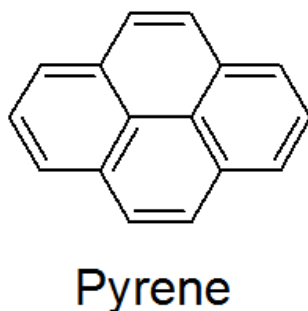


Figure 8: Although Pyrene has 16 conjugated electrons (which does not satisfy $4n + 2$), it is still an aromatic molecule.

We choose to ignore Huckel's rule and any extensions. This is because even anti-aromatic compounds such as pentalene (8 contributing electrons) can be stabilized by adding the appropriate substituent. Either electron-donating or electron-withdrawing groups can be attached to the ring to allow aromaticity (see Figure 9). Additionally, it is possible to insert heteroatoms into the structure, such as nitrogen and sulphur. In addition to possibly making the molecules easier to synthesize, the nitrogen may add electrons to the conjugated system while leaving the switching behavior undisturbed [3] (see Figure, not existent yet

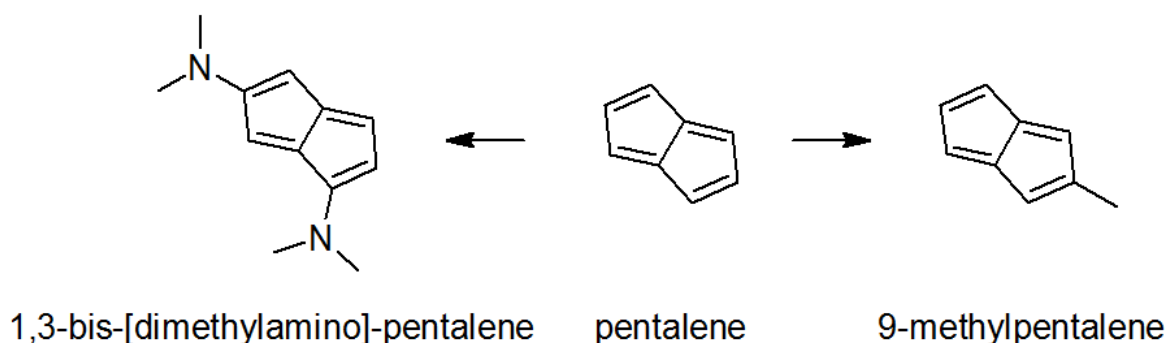


Figure 9: Pentalene can attach either the electron-donating Me [27] or N(Me)₂ [28] group to achieve aromatic structure.

Despite ignoring Huckel's rule, we cannot discard this problem completely. Since ports do not need to contain a double bond, in some resonance structures of the same molecule, the amount of contributing electrons may be different (see Figure 10 and 11). Unless every resonance structure is aromatic, or at least every structure can be made aromatic by the same substituent, the molecule will not be able to realize all port assignments of the cell and is therefore unfit.

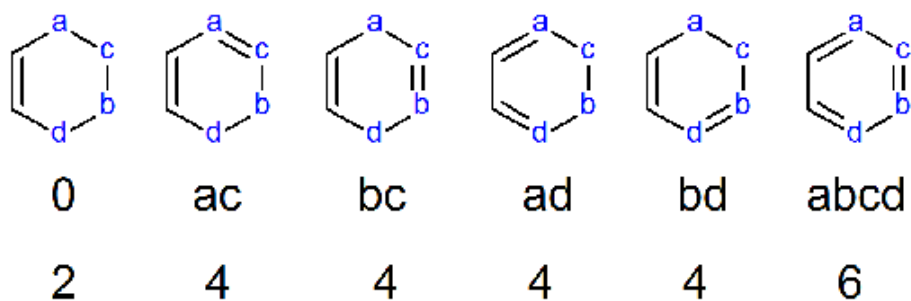


Figure 10: Below is a graph of benzene, which represents K5 of Hesselink's class 4 classification. The first line gives the structure, the second line the port assignment, and the third line the amount of electrons contributing to the ring. The problem here, is that only 2 and 6 are aromatic, and if we add substituents to stabilize the 4 electron structures, the 2 and 6 become unstable.

To account for this, we must ensure that each ring contains the same amount of electrons in every resonance structure. It is a difficult problem to determine which port placements could cause

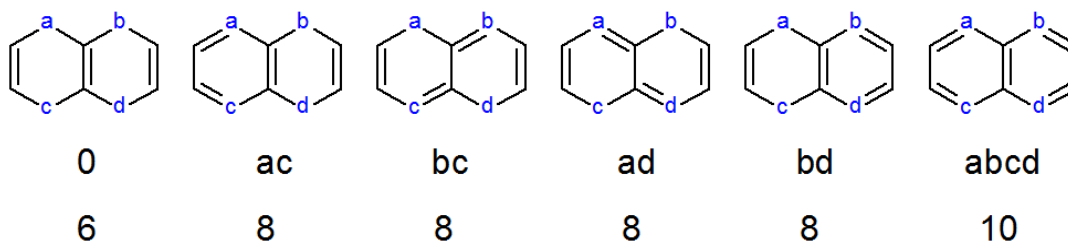


Figure 11: Naphthalene also can represent K5 of Hesselink's class 4 classification. This graph has the same problem as before. In fact, most graphs with 2 or more ports within a cycle system have this problem.

this. We have yet to discover a general solution. In any ring system, a single port is allowed. Since a double bond spans two atoms, every atom in the ring will still always contain a double bond, leaving the total number of bonds to remain constant. However, when two ports are added into a ring, or two rings are conjoined, it becomes unclear. For example, see Figure 12. To ensure all results are manageable, all rings systems are restricted to one port. This means graphs such as naphthalene, can only have a single port within the cycle. But if one graph contained two bridged naphthalenes, each could contain a single port. This unnecessarily removes some graphs which could work, but it is a costly computation to determine for each graph whether it works. All other ports within the graph must be exo-cyclic.

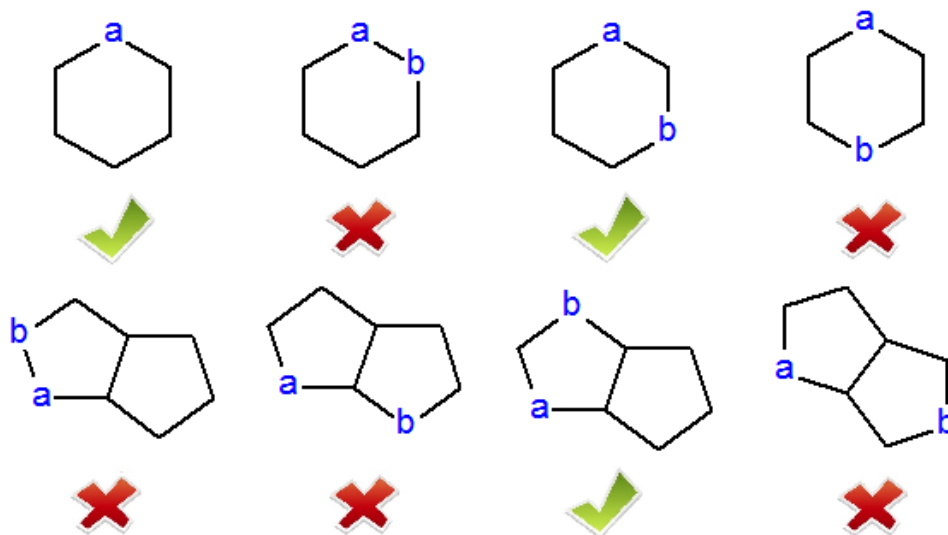


Figure 12: Benzene and Pentalene with two ports at various locations. Structures with a checkmark below indicate all resonance structures contain the same amount of double bonds. Structures with an x, indicate the opposite, and are unfit.

Finally, graphs representing molecules must be connected. A disjoint graph more accurately represents two molecules rather than one. Although these restrictions do a good job of trimming out almost all unrealistic graphs, they are not absolute. Additionally, our restrictions make no

attempt to ensure all graphs can be synthesized (more on this in Section 4).

3 Genetic Algorithm

A genetic algorithm is used to search for graphs which apply to a given cell. The algorithm begins by randomly generating a population of graphs. Each starting graph has anywhere from rank to 31 nodes, and rank - 1 to 20 edges. There must be at least rank - 1 edges in order to connect the minimum amount of nodes together. Graphs with fitness 0 or lower (normally arises from disconnected ports) are deleted and a new graph will take its place. This makes up the first generation.

In each iteration, the survivors of last generation's population must be chosen. This is done by selecting the best performing graphs (fitness-wise), along with a smaller amount of random graphs. Random graphs are added to ensure genetic diversity in the population. This sub-population now undergoes genetic operations such as mutation and crossover.

The genetic algorithm has two terminating conditions. Either the maximum number of iterations has been reached, or some threshold of graphs which satisfy the cell has been reached. In some smaller cases, the genetic algorithm terminates immediately after generating the initial population, as it contained at least x graphs which satisfy the cell (where x is the threshold of graphs desired).

Resulting graphs are converted to SMILES (cite here) (small molecular input line entry system) in order to be visualized. Conversion first involves finding a spanning tree of the graph. All edges in the original graph which are not in the spanning tree must be edges which complete a cycle. Therefore each node on either side of the edge is labeled as a cycle point. We then perform depth first search over the spanning tree, and append each node's label as we reach it.

The ports are labeled by atoms with a P (normally reserved for a phosphorous atom). We need not label each port individually since Kekulé cells are invariant across port renaming [1]. The open source Chemistry Development Kit (citation) is used to parse the SMILES and generate a two-dimensional structure. Other than the ports, the structure resembles the hydrogen-suppressed carbon skeleton of the molecule.

3.1 Mutation

Mutation involves randomly perturbing graphs to obtain new ones. We normally generate enough mutants so that it is likely each graph will be mutated more than once. Mutations include:

1. Addition of a node
2. Removal of a random node, and all edges to it
3. Addition of a random edge ¹

¹Edges are randomly chosen until an edge is found that can be added without breaking our restrictions on the degree of vertices.

4. Removal of a random edge

5. Port Extension

Port extension involves adding a new node in the position of every port, and then adding the port connected to that node. This will always be viable since any port can have at most degree 2, and after port extension, the new nodes will have at most degree 3. This is often useful for spreading the ports out and allowing new edges without breaking our degree restrictions. Since each port will always have a degree of 1 after extension, each port can facilitate a new bond. See Figure 13.

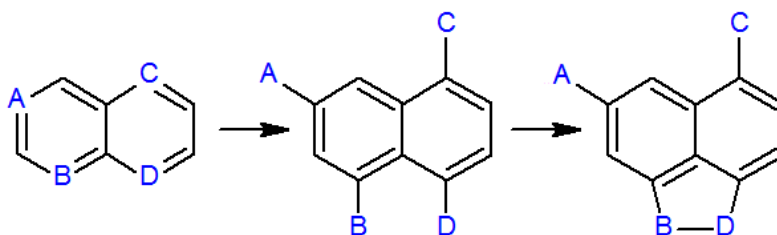


Figure 13: Naphthalene (left) undergoing a port extension (center). Notice that after the port extension, it is now feasible to add an edge between B and D, which can then create Acenaphthene (right).

3.2 Crossover

Crossover begins by randomly selecting 2 graphs from the sub-population. These graphs are combined to create a new child graph. On average, each graph produces two children. Children receive all edges with both parents share, and have a 50% chance to get any other edge found in one of the parents. This creates children which are a hybrid of both their parents.

3.3 Fitness Function

Fitness values are integers given to a graph. Any graphs cell can be calculated using Hesselink's procedure here: (explain a lot). The fitness is calculated by iterating over the graphs cell compared to the cell we are evolving towards. For every port assignment they share, the fitness is incremented by one. For every port assignment that is missing, or is extra, fitness is decremented one point. If a cell is empty, with no port assignments ², it is discarded.

²This can happen when the genetic algorithm creates a graph which has a (non-port) vertex with no edges to it (degree 0). There is no way for this graph to have a Kekulé structure since the lone vertex can never have a double bond.

3.4 Cycle Detection

Cycles are detected using an upper bounded implementation of the 'flooding' algorithm (need cite here). In the flooding algorithm, each node sends out a packet to all of its neighbors. Any time a node receives a packet, it appends its name and sends the new packet to all of its neighbors. Packets are not sent back to the node who it was just received it from. If at any time a packet has been appended by the same node twice, we have found a cycle. As soon as a packet finds a cycle, the packet is discontinued. Packets of length > 12 are also discontinued. This introduces the error that a graph could contain a cycle of length 13. The amount of computation time saved is usually worth it, considering how unlikely a cycle of length 13 or greater is.

Some optimizations can be made considering the subset of all possible graphs we are dealing with. Packets only begin on nodes with degree of 3, as this is guaranteed to find all cycles. Every point within a cycle has a degree of 2, so if that cycle is connected to any other node, then at least one of that cycles nodes will have a degree of 3. If there are no nodes with degree of 3, the graph either contains no cycles, or the graph is one big cycle. Therefore, in this case, we only start a single packet at one random node. This is still slow.

3.5 Cycle Extension and Contraction

Cycle extension is done by inserting two additional carbon atoms into the cycle at the same position (see Figure 14). This has no effect on the Kekulé cell as the two atoms are conjugated with the rest of the ring and simply expand it. This procedure is nearly identical to operation v of [3]. *Care must be taken to not select two nodes which also are a part of another cycle, as the resulting graph will contain unrealistic bonding distances and cycle connectivity. The opposite procedure can be used to contract cycles. Contraction is only possible if there are two neighboring nodes with no edges outside the ring. Graphs with un-alterable cycles outside our limits are excluded.*

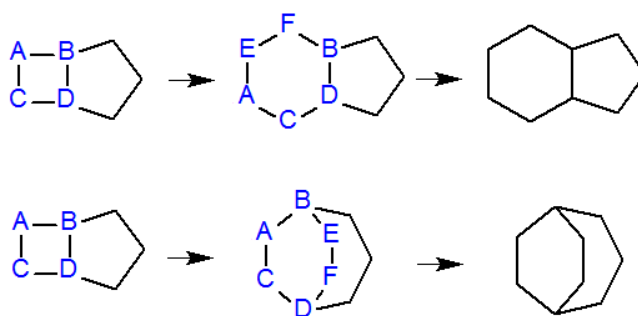


Figure 14: This shows the extension of the left 4 cycle into a proper 6 cycle. The top half shows a correct procedure, selecting to add nodes between A and B. The bottom half shows the incorrect procedure of adding two nodes in between B and D. Double bonds are excluded from this diagram for clarity.

3.6 Connecting a Disjoint Graph

A disjoint graph is not an accurate representation of a molecule. Luckily, disjoint graphs can often be connected using the following procedure. If we add internal vertices which must always conjugate within themselves, they will not interfere with the cell. For example, see Figure 15.

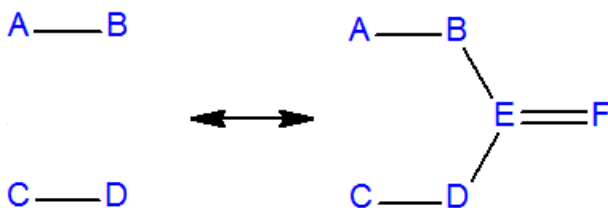


Figure 15: Process used to turn a disjoint graph to a connected graph with the same cell. Two new nodes E and F can be added which always conjugate to each other.

In the right-hand graph of Figure 15, F is not a port and so must always contain a double bond. The only other node which is available to form a double bond with F is E, so in every resonance structure of this molecule E and F will contain a double bond with each other (as shown in right diagram). Therefore d or b can never form a double bond with their new edge (that would result in two adjacent double bonds), and the Kekulé cell has not been changed. This now resembles a stable connected molecule. This procedure is similar to Operation iiiA and vi of [3].

3.7 Additional Note

It should be noted that rather than using the genetic algorithm to approximate a solution (as seen in most applications), an answer generated by it is only useful if it found an actual solution. However, for all cells of rank <6 and most cells of rank 6, the genetic algorithm can converge to multiple solutions for a given cell.

4 Results

Most graphs produced by the algorithm can be feasibly realized in stable molecules. However, results are not perfect. As stated previously, rings of sizes other than 6 are not always stable.

Planarity? The second condition is another current drawback. Since the cycle detection algorithm needs an upper bound, in rare cases, cycles above this limit can appear. This can cause other smaller cycles which share more than two atoms with the large cycle to not be detected. Additionally, some bonding distances may be unrealistic. For examples of both, see Figure 16.

In most small cases (ports <6), the genetic algorithm is able to obtain atleast 10 different graphs which satisfy the cell immediately. For some of the later classifications of rank 5, and cells of rank 6 or higher, it can take longer. Some cells of rank 6 it is unable to find a graph for. As it

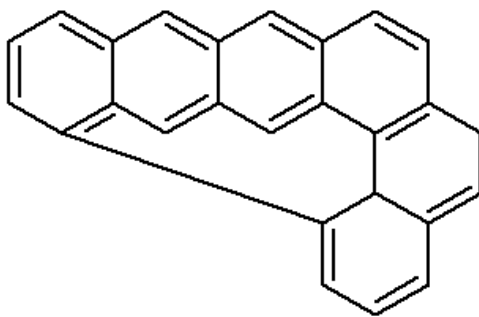


Figure 16: An example of an unrealistic molecule that could be generated by the genetic algorithm. There is a long stretched bond from the first benzene (top left) to the last benzene (bottom right). This stretched bond completes a cycle that has a length of 9. If the cycle detection algorithm had a limit of 8, this graph could be a result. Not only is the stretched bond unrealistic, but the cycle it creates shares up to 3 atoms with some rings.

is possible that no such realistic graph exists for some cells, it is difficult to determine whether a solution exists. The number of nodes is currently maxed at 31 nodes, as nodes are stored in a bit array, which would overflow the integer at 32.

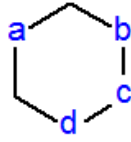
Below is graphs for all cells of rank 4 and 5. Double bonds are excluded from the graphs shown below, as only one resonance structure could be seen at a time. Some graphs may have cells different than the title below them. This is because their cells can be port translated to achieve the classified cell. This is acceptable since switching behaviour is invariant across port translation [2]. As stated previously, the genetic algorithm is capable of coming to multiple solutions for each cell, and the ones shown below are simply ones selected by us. Notice the trend that as the Kekulé cell number increases, the Kekulé cell becomes longer. This means more resonance structures of each graph, and more connections between atoms. The higher amount of connectivity required in each cell makes it so the next graph is usually more difficult to find than the last one. Looking at Figures 18 and 19, it is easy to see that the graphs get bigger and more complex as the cell number increases.

Finding graphs for cells of higher ranks builds on this challenge. Not only is every graph larger, but there must be an additional atom with degree of 2. The number of combinations increases drastically, with 3 classes of rank 3, 7 classes of rank 4, 24 classes of rank 5, and 214 classes of rank 6. When attempting to find graphs for all cells of rank 6, we were able to find graphs for X out of 214. All graphs found can be viewed here: (website).

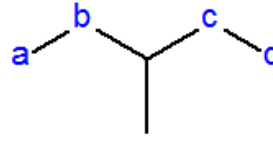
Increasing the number of nodes to above 32 will likely yield more results in some cases, however, it will lower performance. It is possible many graphs do not have a realistic graph which matches it's cell. Discuss more about results and problems encountered.

Even if all graphs produced represent stable molecules, there are additional problems. Each hydrocarbon must be able to be easily synthesized.

Polycyclic aromatic hydrocarbons are commonly produced as byproducts of fuel burning, and have been classified as carcinogenic, mutagenic, and teratogenic. This could pose a challenging



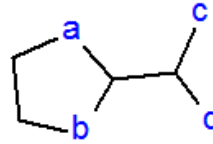
$$K_1 = \{\emptyset, ab, ac, ad\}$$



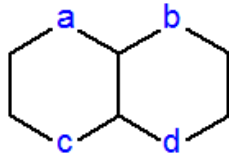
$$K_2 = \{\emptyset, bc, ad, abcd\}$$



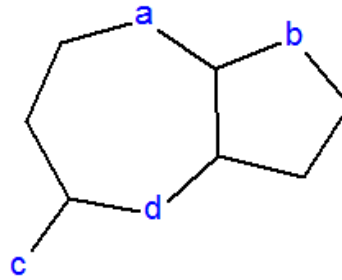
$$K_3 = \{\emptyset, ac, bc, ad, bd\}$$



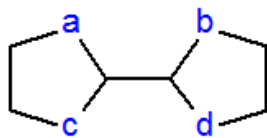
$$K_4 = \{\emptyset, ab, ac, bc, ad, bd\}$$



$$K_5 = \{\emptyset, ac, bc, ad, bd, abcd\}$$



$$K_6 = \{\emptyset, ab, ac, bc, ad, bd, cd\}$$



$$K_7 = \{\emptyset, ba, ca, cb, da, db, dc, dcba\}$$

Figure 17: Realistic Graphs for each cell in Hesselink's classification of all cells with rank 4. K_1 resembles benzene, K_5 naphthalene, and K_6 azulene.

problem if we wish to use large aromatic hydrocarbons in circuitry.

There are many possible improvements in this area. The number of nodes could be increased to allow larger graphs. Although this would increase the amount of computation required, perhaps it is possible to implement a more efficient cycle detection algorithm, or even calculate fitness on every 2nd iteration of the genetic algorithm. These changes may allow even more results to be

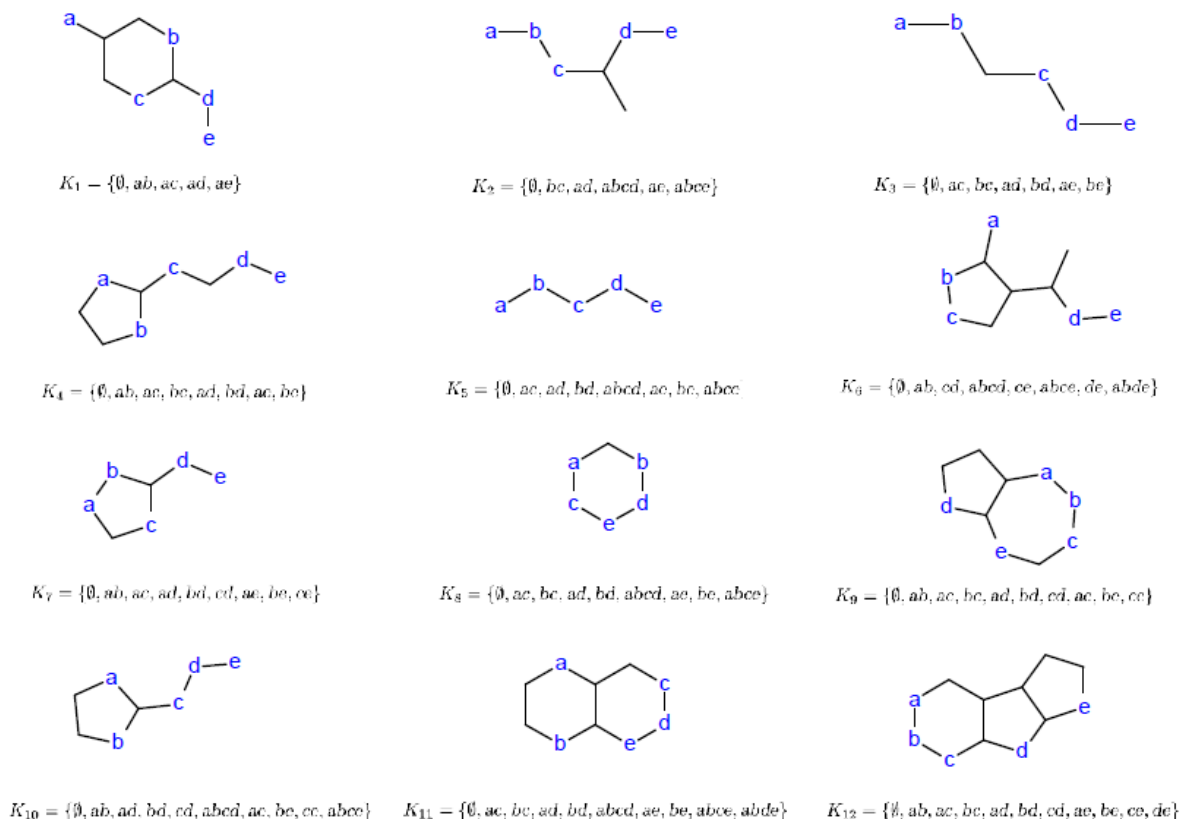


Figure 18: Realistic Graphs for the first 12 cells in Hesselink's classification of all cells with rank 5. K_1 and K_8 resembles benzene, K_{11} naphthalene, and K_9 azulene.

yielded from a heuristic genetic algorithm based approach.

References

- [1] W. H. Hesselink, *Graph Theory for alternating hydrocarbons with attached ports. Indagationes Mathematicae, Elsevier, 24:115141, 2013.*
- [2] W.H. Hesselink, J.C. Hummelen, H.T. Jonkman, H.G. Reker, G.R. Renardel de Lavalette, M.H. van der Veen, *Kekule Cells for Molecular Computation. Cornell University Library Online, 2013.*
- [3] M.H. van der Veen. *π -Logic. PhD thesis, University of Groningen, May 2006.*
- [4] A.J. Heeger, S. Kivelson, J.R. Schrieffer, and W.-P. Su. *Solitons in conducting polymers, Rev. Mod. Phys., 60:781, 1988.*
- [5] G. E. Moore, *Cramming more components onto integrated circuits. Electronics Magazine, p. 4. Retrieved 2006-11-11.*

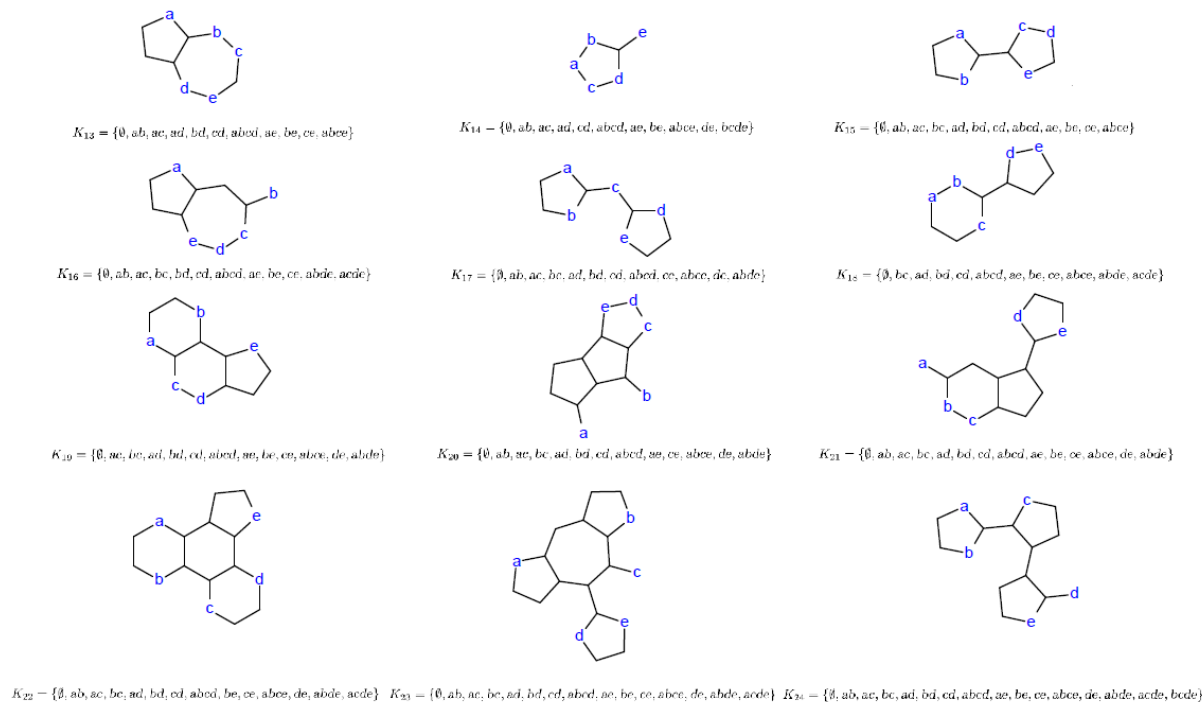


Figure 19: Realistic Graphs for the second half of cells in Hesselink's classification of all cells with rank 5. K₁₃, K₁₆ azulene both resemble azulene.

- [6] C. A. Mack, *Fifty Years of Moore's Law. IEEE Transactions on semiconductor manufacturing*, 24:2 2011.
- [7] M. H. van der Veen, M. T. Rispens, H. T. Jonkman, and J. C. Hummelen, *Molecules with Linear π -Conjugated Pathways between all Substituents: Omniconjugation. Adv. Function Mater*, 14:3, 2004.
- [8] S.N. Yalirahi, M.A. Ratner, *Interplay of topology and chemical stability on the electronic transport of molecular junctions*, Ann. New York Acad. Sci. 960 (2002) 153.
- [9] J. Reichert, R. Ochs, D. Beckmann, H. B. Weber, M. Mayor, H. von Lhneysen, *Phys. Rev. Lett.* 2002, 88, 176804.
- [10] A. Aviram, M. A. Ratner, *Chem. Phys. Lett.* 1974, 29, 277.
- [11] D. B. Strukov, K. K. Likharev, *Nanotechnology* 2005, 16, 137.
- [12] Y. Karzazi, J. Cornil, J. L. Brdas, *J. Am. Chem. Soc.* 2001, 123, 10076.
- [13] J. M. Tour, M. Kozaki, J. M. Seminario, *J. Am. Chem. Soc.* 1998, 120, 8486.
- [14] A. Aviram, *J. Am. Chem. Soc.* 1988, 110, 5687.
- [15] H. W. Ch. Postma, T. Teepen, Z. Yao, M. Grifoni, C. Dekker, *Science* 2001, 293, 76.

- [16] C. Joachim, J. K. Gimzewski, *Chem. Phys. Lett.* 1997, 265, 353.
- [17] T. D. Anthopoulos, C. Tanase, S. Setayesh, E. J. Meijer, J. C. Hummelen, P. W. M. Blom, D. M. de Leeuw, *Adv. Funct. Mater.* 2004, 16, 2174.
- [18] G. M. Tsivgoulis, J.-M. Lehn, *Chem. Eur. J.* 1996, 2, 1399.
- [19] J. J. D. de Jong, L. N. Lucas, R. M. Kellogg, J. H. van Esch, B. L. Feringa, *Science* 2004, 304, 278.
- [20] A. P. de Silva, H. Q. N. Gunaratne, C. P. McCoy, *Nature* 1993, 364, 42.
- [21] F. M. Raymo, S. Giordani, *J. Org. Chem.* 2003, 68, 4158.
- [22] K. Rurack, A. Koval'chuk, J. L. Bricks, J. L. Slominskii, *J. Am. Chem. Soc.* 2001, 123, 6205.
- [23] D. Parker, J. A. G. Williams, *Chem. Commun.* 1998, 245.
- [24] A. P. de Silva, I. M. Dixon, H. Q. N. Gunaratne, T. Gunnlaugsson, P. R. S. Maxwell, T. E. Rice, *J. Am. Chem. Soc.* 1999, 121, 1393.
- [25] Roberts, J. D., Streitwieser Jr, A., and Regan, C. M. *Small-Ring Compounds. X. Molecular Orbital Calculations of Properties of Some Small-Ring Hydrocarbons and Free Radicals. Journal of the American Chemical Society*, 74(18), 4579-4582. 1952.
- [26] Kruszewski, J., and Krygowski, T. M. *An Extension of the Hckel $4N+2$ Rule to Polycyclic Non-alternant Conjugated Hydrocarbons. Canadian Journal of Chemistry*, 53(6), 945-951. 1975.
- [27] Hafner, K., Donges, R., Goedecke, E. and Kaiser, R. *Concerning Pentalene, 2-Methylpentalene, and 1,3-Dimethylpentalene. Angew. Chem. Int. Ed. Engl.*, 12: 337-339. 1973.
- [28] Hafner, K., Bangert, K. F. and Orfanos, V *1,3-Bis(dimethylamino)pentalene. Angew. Chem. Int. Ed. Engl.*, 6: 451452. 1967.