NAME:	Lewis Dot Structures Activity 1					
Lewis Dot Structures (	Electron Dot Structures)					
<b>Lewis Dot Structures</b> are a tool for representing in chemical substances. We are going to use Lewis between covalent bonds and ionic bonds. First let and ions:	the arrangement of valence electrons around atoms dot structures to understand the difference us make Lewis Dot Structures for common atoms					
a. H	b. Li					
c. H <sup>1+</sup>	d. Li <sup>1+</sup>					
c. H <sup>1+</sup>	a. Li <sup>2</sup>					
e. CI	f. Cl <sup>1-</sup>					

h. Ne

g. O<sup>2-</sup>

**Lewis Dot Structures for Substances Containing Ionic Bonds:** For each of the hypothetical chemical reactions below please draw Lewis Dot Structure for each of the participants.

 Li	+	F	$\rightarrow$	Li <sup>1+</sup>	+	F <sup>1-</sup>

Na	+	Cl	$\rightarrow$	Na <sup>1+</sup>	+	Cl <sup>1-</sup>

Ca	+	S	$\rightarrow$	Ca <sup>2+</sup>	+	S <sup>2-</sup>

2 K	+	0	$\rightarrow$	2 K <sup>1+</sup>	+	O <sup>2-</sup>

Al	+	Р	$\rightarrow$	Al <sup>3+</sup>	+	P <sup>3-</sup>