

1. Simple chemical formula:  $\text{H}_2\text{O}$
2. Ionic compounds:  $\text{NaCl}$
3. Isotopes:  $^{14}\text{C}$
4. Charges:  $\text{SO}_4^{2-}$
5. Hydrates:  $\text{CuSO}_4 \cdot 5 \text{H}_2\text{O}$
6. Reaction arrows:  $\text{H}_2 + \text{O}_2 \longrightarrow \text{H}_2\text{O}$
7. Equilibrium arrows:  $\text{CO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3$
8. States of matter:  $\text{H}_2\text{O}(\text{l}) \longrightarrow \text{H}_2\text{O}(\text{g})$
9. Stoichiometry:  $2 \text{H}_2 + \text{O}_2 \longrightarrow 2 \text{H}_2\text{O}$
10. Acids and bases:  $\text{HCl} + \text{NaOH} \longrightarrow \text{NaCl} + \text{H}_2\text{O}$