Balancing acid base reactions using mhchem

- 1. $HCl + NaOH \longrightarrow NaCl + H_2O$
- 2. $HNO_3 + KOH \longrightarrow KNO_3 + H_2O$
- 3. $H_2SO_4 + 2 NaOH \longrightarrow Na_2SO_4 + 2 H_2O$
- 4. $H_3PO_4 + 3 KOH \longrightarrow K_3PO_4 + 3 H_2O$
- 5. $CH_3COOH + NaOH \longrightarrow CH_3COONa + H_2O$
- 6. $HBr + KOH \longrightarrow KBr + H_2O$
- 7. $HF + NaOH \longrightarrow NaF + H_2O$
- 8. $HClO_4 + KOH \longrightarrow KClO_4 + H_2O$
- 9. $H_2CO_3 + 2 KOH \longrightarrow K_2CO_3 + 2 H_2O$
- 10. $HSO_4^- + NH_4^+ \longrightarrow NH_4HSO_4$