

Q12-4

(a) for CSTR,

$$\tau = \frac{C_{A0} X_A}{(-r_A)}$$

$$-r_A = k_1 C_A - k_2 C_B = k_1 C_{A0} (1 - X_A) - k_2 (C_{A0} X_A)$$

$$\tau = \frac{C_{A0} X_A}{C_{A0} [k_1 (1 - X_A) - k_2 X_A]}$$

$$k_1 \text{ at } 280 \text{ K} :- 10^3 \exp\left(\frac{-2500}{280}\right) \text{ s}^{-1} = 0.1325 \text{ s}^{-1}$$

$$K = 7.8 = \frac{k_1 \text{ at } 280 \text{ K}}{k_2 \text{ at } 280 \text{ K}}$$

$$k_2 \text{ at } 280 \text{ K} = 0.0169$$

$$\tau = \frac{X_A}{0.1325(1 - X_A) - 0.0169 X_A}$$

$$600 = \frac{X_A}{0.1325 - 0.1494 X_A}$$

$$79.5 - 89.64 X_A = X_A$$

$$\boxed{X_A = 0.877} \text{ @ } 280 \text{ K}$$

Von't Hoff equation :-

$$\ln \left(\frac{k_2}{k_1} \right) = \frac{-\Delta H_{\text{rxn}}}{R} \left[\frac{1}{T_2} - \frac{1}{T_1} \right]$$

$$\ln \left(\frac{k_T}{k_{280}} \right) = \frac{+10 \frac{\text{Kcal}}{\text{mol}}}{8.314 \text{ J/mol K}} \left[\frac{1}{480} - \frac{1}{280} \right]$$

$$\Delta H_{\text{rxn}} = -10 \frac{\text{Kcal}}{\text{mol}} = -10 \times 4.184 \times 1000 \text{ J/mol} \\ = -41840 \text{ J/mol}$$

$$\ln \left(\frac{k_T}{k_{280}} \right) = \frac{+41840}{8.314} \left[\frac{-1}{672} \right] = -7.488$$

$$k_T = 0.0043 \text{ at } 480 \text{ K}$$

$$k_1 \text{ at } 480 \text{ K} = 10^3 \exp \left(\frac{-2800}{480} \right) = 5.47 \text{ s}^{-1}$$

$$k_T = \frac{k_1}{k_2} = 0.0043$$

$$k_2 = \frac{5.47}{0.0043} = 1272.09 \text{ s}^{-1} \text{ at } 480 \text{ K}$$

$$\tau = 600 = \frac{X_A}{k_1(1-X_A) - k_2(X_A)}$$

$$600 = \frac{X_A}{5.47(1-X_A) - 1272.09 X_A}$$

~~$$3282 - 766536 X_A = X_A$$~~

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$$\boxed{X_A = 0.0042}$$

b)

Energy balance:-

$$0 = \dot{Q} + F_{A0} \sum_{i=1}^n \theta_i (H_{i0} - H_i) - \Delta H_{rxn}(T_f) F_{A0} X$$

$$C_p = \frac{1000 \text{ cal}}{\text{kg} \cdot \text{K}}$$

$$\frac{C_p \rho}{C_{A0}} = \frac{1000 \text{ cal}}{\text{kg} \cdot \text{K}} \times \frac{1 \text{ kg}}{\text{L}} \times \frac{1 \text{ mol}}{5 \text{ kg}}$$

$$\frac{C_p \rho}{C_{A0}} = 0.2 \frac{\text{Kcal}}{\text{mol} \cdot \text{K}}$$

$$0 = \dot{Q} + F_{A0} \times 0.2 \times 10^3 (320 - 280) - (-10^4) \times F_{A0} \times 0.877$$

$$\dot{Q} = -F_{A0} [8000 + 8770] \text{ K}$$

$$\dot{Q} = -\frac{C_{A0} V}{\tau} \times 16770$$

$$\frac{\dot{Q}}{V} = -139.75 \frac{\text{cal}}{\text{L} \cdot \text{s}}$$

Heat removed per liter of reactor volume to maintain at 280 K operating temperature.