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Answer the following items in the space provided. 5. Which isotope defines the atomic mass unit, and how is the atomic mass unit defined? 6. Why is a mole used to count atoms? 7. What is the relationship between an atom's atomic mass and one mole of that atom? Name Class Date Concept Review Skills Worksheet

Skills Worksheet Concept Review - Forestville Middle-High ...

Concept Review: Avogadro's Number and Molar Conversions 1. 2.41 10²⁴ atoms Al 2. 9.33 31023 atoms Na 3. 44.0 mol F 4. 70.7 mol H 2 5. 291 mol K 6. 877 g Sb 7. 119 g U 8. 167 g Fr

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Concept Review Skills Worksheet. Concept Reviews SECTION: ATOMIC STRUCTURE 1. Check students' drawings. Drawings ... used to count atoms because atoms are too small to count individually. 4. a. 40.08 g/mol c. 32.07 g/mol b. 58.93 g/mol d. 16.00 g/mol 5. List the given and unknown values.

Concept Review - Manchester High School

Concept Review: Avogadro's Number and Molar Conversions 1. 2.41 10²⁴ atoms Al 2. 9.33 31023 atoms Na 3. 44.0 mol F 4. 70.7 mol H 2 5. 291 mol K 6. 877 g Sb 7. 119 g U 8. 167 g Fr

Skills Worksheet Concept Review

Concept Review: Covalent Bonds 1. A covalent bond forms when two or more valence electrons are attracted by the positively charged nuclei of two atoms and thus are shared between both atoms. 2. The H₂ molecule is stable because each hydrogen atom now has a shared pair of electrons and has achieved a stable noble gas configuration. 3.

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Holt Chemistry 14 Atoms and Moles Section: Electron Configuration In the space provided, write the letter of the term or phrase that best answers the question. ____ 1. Which of the following atomic models could be referred to as the "solar-system" model of the atom? a. Thompson's plum-pudding model b. Rutherford model c. Bohr model d ...

Assessment Quiz - Woodland Hills School District

Chapter One Test Review. Chapter Five. Chapter One Leftovers. Chapter Five Assignments. Fill in Worksheet. ... Counting Atoms Practice Quiz. Practicing Molar Mass. Calculating Percent Composition. Study Trifecta. ... Count the atoms in the following molecules. 1. H₂SO₄ - how many H? ____ how many S? ____ how many O? ...

Counting Atoms Practice Quiz - Mr. Ballard's HS Science

For purposes of counting, the objects behave as though they were all identical, as though they each actually had the average mass. We count atoms in exactly the same way. Because atoms are so small, we deal with samples of matter that contain huge numbers of atoms. Even if we could see the atoms, it would not be possible to count them directly.

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