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Facoltà di Ingegneria dei Processi Industriali
Degree in Materials Engineering and Nanotechnology



ATOMIC LAYER DEPOSITION OF
FERROELECTRIC COMPLEX
OXIDE THIN FILMS

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For my parents...

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Abstract

Comments on application of ferroelectric materials.

130 As devices based on ferroelectric films become more commonplace, a commercially viable process for fabricating the material is needed; low cost and high volume are implied necessities for such a process. This work focuses on the application of atomic layer deposition (ALD) to this task. ALD is a standard fabrication process in the electronics industry, valuable for its high uniformity
135 across large surfaces, capability to control film thickness with very high resolution, and conformality across three dimensional structures. The various tasks in designing and optimizing an ALD process, as well as characterization methods for analysis of the produced films, will be discussed in detail.

Work supported by ARO under Grant #??.

¹⁴⁰ **Sommario**

A detailed summary of the document to be presented solely in Italian.

Chapter 1

Introduction

1.1 Scope of this work

¹⁴⁵ Chapter 2

Lead Titanate

2.1 Structure

Lead titanate (PbTiO_3 , PTO) naturally orders into the tetragonal perovskite crystal structure at room temperature (figure 2.1). The structure can
150 be affected by compositional changes, temperature, or strain (primarily in thin-film systems), allowing a transition to a cubic phase. In the perovskite crystal structure, the central cation (Ti^{4+} in the case of PbTiO_3) is encapsulated in a octahedral cage of anions (O^{2-}), with the remaining cations (Pb^{2+}) situated in the eight corners of the unit cell. If the material was doped (as in a mixed
155 solid-solution), some of the cations would be replaced with the dopant ions, for example Zr^{4+} would be randomly distributed in Ti^{4+} sites in the $\text{PbTi}_{1-x}\text{Zr}_x\text{O}_3$ (PZT) system.

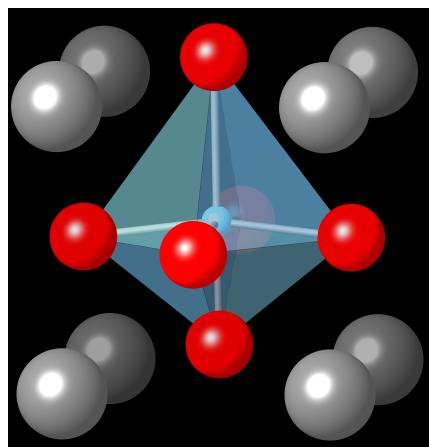


Figure 2.1: Tetragonal perovskite structure of PbTiO_3 .
Grey, red, and blue spheres refer to Pb^{2+} , Ti^{4+} , or O^{2-} , respectively

2.1.1 Effect of Temperature

The transition from tetragonal to cubic perovskite is highly dependent on
160 temperature. The critical temperature at which this transition occurs is referred to as the Curie temperature (T_C). If the material cools through this

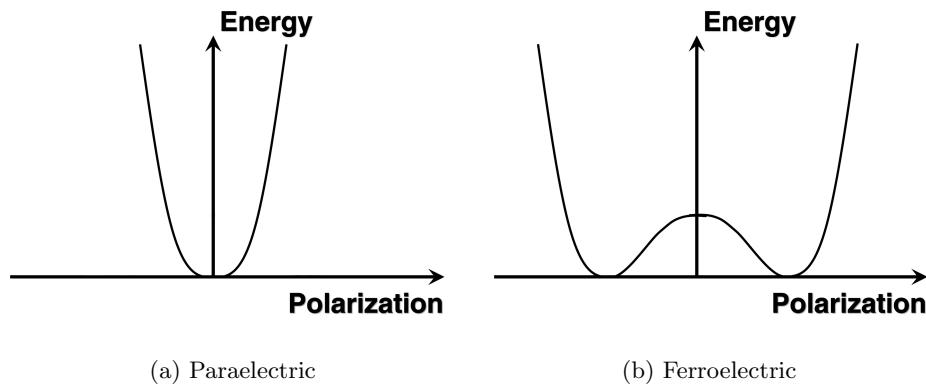


Figure 2.2: Example plots of the energy required to polarize a material. Ferroelectric materials (b) have non-zero polarization at the energy minima. Above T_C all ferroelectric materials transition to a paraelectric phase (a).

temperature, a lengthening of the ‘c’ axis of the unit cell spontaneously occurs. This creates anisotropy in the structure and allows for an anisotropic charge distribution to develop. In lead titanate this is caused by the shifting of the
₁₆₅ titanium ion, along with a slight shift of some of the oxygen ions as well (visible in figure 2.1 on the previous page). Thus, a permanent dipole is created whose magnitude increases as the system cools further from T_C . This permanent dipole allows the system to exhibit ferroelectricity, implying an ability to semi-permanently switch the orientation of the dipole in the material. This
₁₇₀ switching can be reversed, but this will not occur spontaneously.

2.2 Ferroelectricity

Ferroelectricity is the capability of a material to exhibit spontaneous electric polarization that requires external influence, such as an applied electric field, to be reversed. This is different from paraelectric (or even dielectric) materials,
₁₇₅ where there is no polarization without external field being applied. This can be seen in a plot of energy vs. polarization (fig. 2.2) for the two types of materials. In a ferroelectric material, the energy minima are found at non-zero levels of polarization.

One of the most important effects of ferroelectric behavior is the exhibition
₁₈₀ of hysteresis in the polarization state of the material. **Suboptimal explanations...** Because additional energy is required to make the transition from one state to another, the original state endures for some amount of opposing field

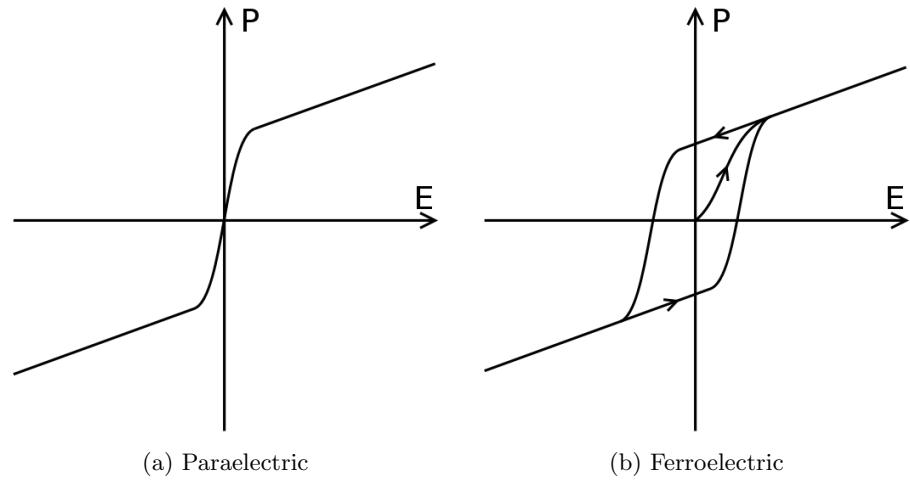


Figure 2.3: Example plots of the energy required to polarize a material. Ferroelectric materials (a) have non-zero polarization at the energy minima. Above T_C all ferroelectric materials transition to a paraelectric phase (b).

(see fig. 2.3b).

Chapter 3

¹⁸⁵ Synthesis Methods

Synthesis of perovskite oxides has been demonstrated using a wide range of techniques. These range from solution-based processing methods (sol-gel approach), to physical vapor methods (molecular beam epitaxy and pulsed laser deposition), and gas phase chemical methods (chemical vapor deposition and atomic layer deposition). This review will briefly discuss sol-gel, physical vapor deposition, as well as CVD methodology, but will focus in more depth on films deposited via ALD.

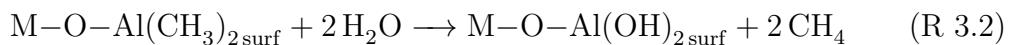
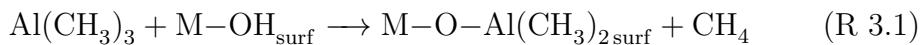
3.1 Sol-Gel Processing

3.2 Physical Vapor Deposition

¹⁹⁵ 3.3 Metallorganic Chemical Vapor Deposition

3.4 Atomic Layer Deposition

Atomic Layer Deposition (ALD) is a modification on standard CVD processes, with a few major differences. The defining aspect of an ALD process is the separation of the overall reaction into two steps: first the precursor is allowed to react with the substrate surface (see reaction 3.1), excess reactant is purged from the chamber and an oxidizer is introduced to complete the reaction (see reaction 3.2). These reactions show a very simple ALD reaction between trimethylaluminum (TMA, $\text{Al}(\text{CH}_3)_3$) and water.



In this example, it is seen that the first stage allows the TMA to react with the hydrated substrate surface to form part of a layer of alumina (Al_2O_3), liberating a molecule of methane as a byproduct. In the next step, the remaining ligands are stripped away from the bound TMA molecule and replacing them

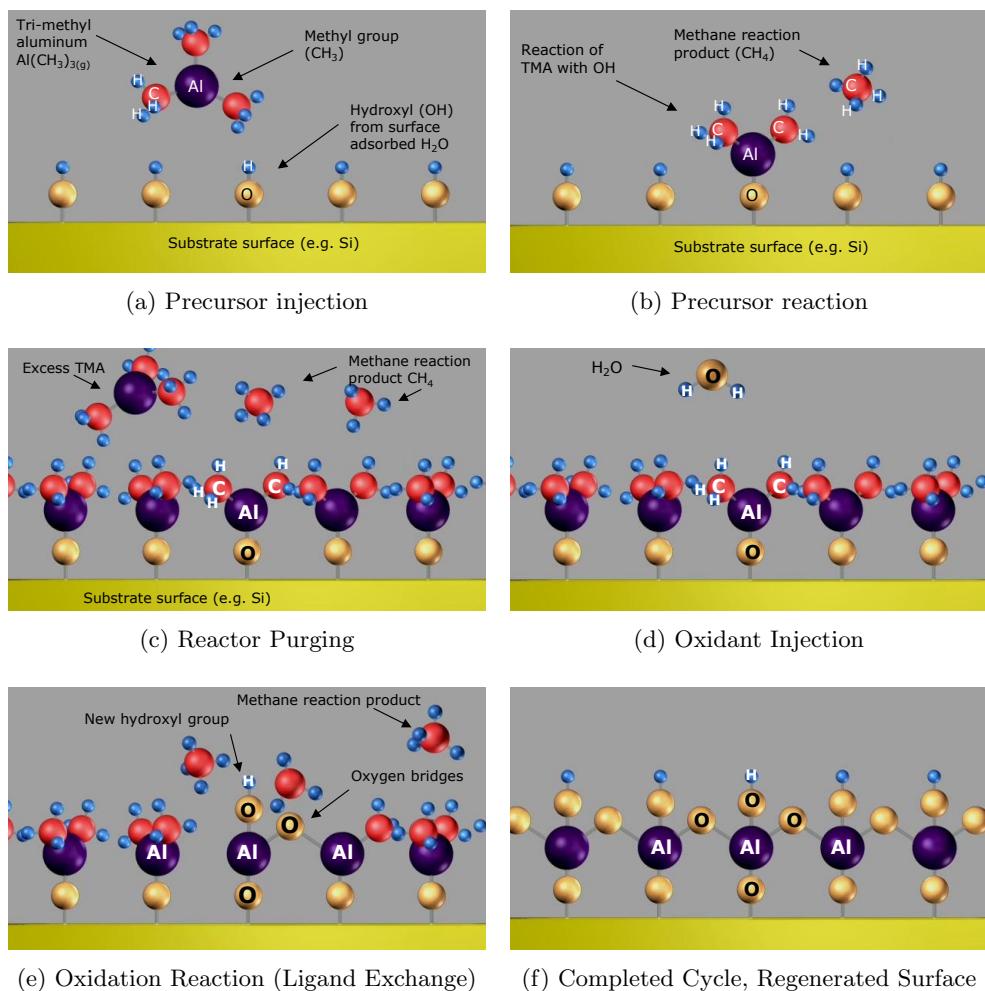


Figure 3.1: Example schematic of the process of an ALD deposition cycle. This example illustrates the reaction of TMA and water to form alumina (Al_2O_3).

Graphic courtesy of Cambridge NanoTech, inc.

with hydroxyl groups. This returns the system to the initial state — where the surface is presenting sites available to react with more TMA — and the cycle is completed. A graphical example of this process can be found in figure 3.1.

Having only surface reactions be permitted, as opposed to CVD where gas-phase interactions dominate, affords ALD a number of unique characteristics. One of these is the concept of the “self-limiting” growth mode. This behavior arises from the limited number of available reaction sites; when all of these have either been reacted with or made unavailable by a blocking mechanism such as stearic hindrance from other local chemisorbed precursor the reaction can no longer proceed. At this point, additional available precursor is not going to be utilized, and instead will be removed and treated as waste material.

The system is then evacuated, and a inert purge gas such as dry nitrogen or
220 argon (at UHP grade) is flowed through the reactor. The purge gas serves both to push any remaining gases out of the reactor as well as to help desorb physisorbed species from the surface. If these are allowed to remain adsorbed they would react with the oxidant and negate the surface-limited aspects of
225 ALD. The system would then again be evacuated, and the oxidant introduced and then pumped away to complete the cycle.

In the implementation of most ALD systems, the purge gas is also used as a carrier gas for the precursors. Thus a constant flow of gas is passed through the system, instead of having it occasionally fully evacuated, and the precursor is able to be delivered from its source to the reactor more effectively. For some
230 precursor compounds, in particular those with a low vapor pressure, having carrier-assisted transportation can greatly improve the behavior of the system.

Because of the self-limiting behavior, each deposition cycle is limited to a theoretical maximum of one monolayer of material (in practice a much lower coverage per cycle is attained), which is far less than a unit cell. Generally
235 per cycle growth rates range between 0.03–1.5 Å, with the rate being nearly invariable during most of the deposition. This gives the second defining characteristic of ALD: very high (Å level) thickness resolution. The downside of this aspect is that growths are generally much slower than other types of depositions; ALD is generally slower by an order of magnitude or more than a
240 similar CVD process, as an example. This has proved invaluable in many processes where high precision is critical, such as electronics manufacturing. Intel, for example, uses ALD to deposit extremely thin layers of a high- κ dielectric (such as hafnia, HfO₂) for use as the gate oxide in its transistors, with layer thickness generally less than 2 nm.

245 This method will produce a layer of a binary oxide material (AO_x), if more complex materials are desired the method must be changed. The basic principles remain the same; one would perform the procedure for depositing a cycle of a binary oxide and then change the precursor and deposit another cycle of a different oxide material. For example, if one wished to deposit PbTiO₃,
250 one would begin by depositing a layer of TiO₂ and then depositing a layer of lead oxide (PbO). Repeating this set of cycles — a super-cycle — would eventually form the PbTiO₃ film.

However, deposition of complex oxides is not this simple in practice. In many cases, running each oxide cycle in a 1:1 ratio will deposit a non-stoichiometric
255 material. This makes it necessary to modify the method to deposit more of one

type of oxide than the other. For example, if a material is Ti-rich the super-cycle ratio would be modified to increase the number of lead oxide cycles as compared to the titania cycles. Needs more here.

ALD reactions are rather sensitive to a number of factors, such as temperature.
²⁶⁰ The temperature must be high enough that the reactants have sufficient energy to drive the surface reaction but not so high as to allow undesirable reactions to activate (e.g. precursor cracking or surface material desorption). Precursor selection is also very important, for similar reasons. The precursors must also be incapable of reacting with themselves, to allow the self-limiting mechanism to work properly. This section needs more work.

²⁶⁵

Chapter 4

Characterization Methods

4.1 Compositional Analysis

4.1.1 Energy-Dispersive X-Ray Spectroscopy

270 Energy dispersive X-ray spectroscopy (EDXS) is a commonly used analysis technique for determining the composition of a sample. In this process, a sample is bombarded with high-energy electrons (2–30 keV) which interact with the sample. Some of these electrons will cause a core electron of an atom in the sample to be ejected. This leaves a vacant orbital in the inner shell, which a higher energy electron will fill. In the process of filling the vacancy, the electron will emit an X-ray photon equal to the energy difference between the two states. These energies are referred to using the common X-ray spectroscopy nomenclature (e.g. K_{α} , K_{β} , L_{α}). An illustration of this process can be found in figure 4.1 on the following page.

280 Since the energies of the emitted photons are very specific to each element, the procedure can be used to identify the presence of the element in the sample. With some calibration, EDXS can also be used to quantify the relative amounts of each element in a sample using the different number of collected photons. However, this can sometimes be difficult due to some elements have overlapping spectra as the peaks are not sharp and different elements can have similar energies for some transitions.

One of the downsides of EDXS is that the interaction volume of electrons is much larger deeper into the sample, so the surface sensitivity of the technique is smaller than with some other techniques.

290 **4.1.2 X-Ray Fluorescence Spectroscopy**

X-ray fluorescence spectroscopy (XRFS) is a similar technique to EDXS. In XRFS the excitation used is X-ray photons (often from a Cu K_{α} source with $\lambda = 1.54 \text{ \AA}$), as opposed to energetic electrons. In other considerations the techniques are equivalent.

295 XRFS often has a lower noise floor than EDXS, allowing smaller signals

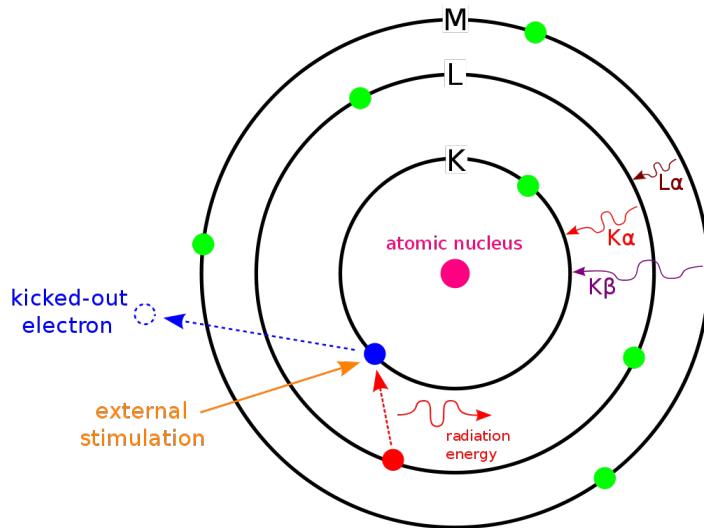


Figure 4.1: Graphic illustrating the basic mechanism for EDXS, along with the commonly used notation for the various energies. The external stimulation would be a high energy electron.

to be more easily identified (such as in ultra-thin films). It does suffer the same disadvantage of having overlapping peaks. This resolution issue can be alleviated to some degree by using wavelength dispersive XRFs (WD-XRFS), which uses diffraction techniques to analyze the emitted x-ray spectrum.

³⁰⁰ 4.1.3 Rutherford Backscattering Spectroscopy

4.2 Thin Film Characterization

4.2.1 Variable Angle Spectroscopic Ellipsometry

Ellipsometry is a powerful non-destructive optical technique that allows for the determination of a large number of properties of complex thin film structures. The basic tenet of ellipsometry relies on the analysis of the change in polarization state of a reflected light beam after interaction with the sample. The incident beam is generally linearly polarized, but upon reflection becomes elliptically polarized due to a phase shift in the components of the beam in the s- and p-plane, as well as a change in their relative amplitudes. The phase shift is correlated to the ellipsometric parameter Δ , while the amplitude change is given by $\tan \Psi$ (Ψ is the angle between the s-plane and the major axis of the ellipse). The last major parameter is the incident angle, denoted by Φ . A schematic diagram illustrating these parameters can be seen in figure 4.2 on the next page.

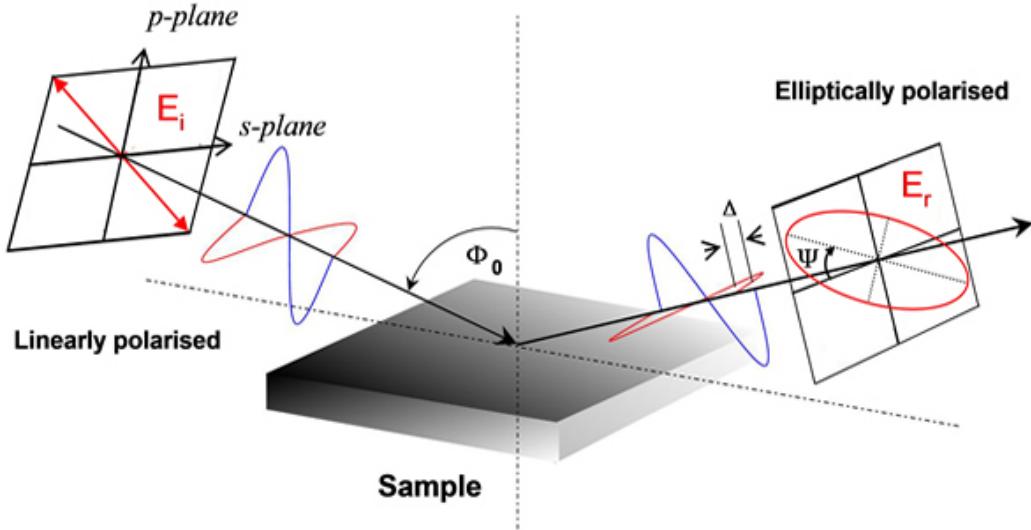


Figure 4.2: Schematic of the beam path during an ellipsometric measurement, critical parameters are indicated

From these parameters, one can directly determine the ratio between the reflectance in the p-plane (r_p) and the reflectance in the s-plane (r_s) from the fundamental ellipsometric relation (eqn. 4.1). Once this relationship is known, the Fresnel equations (eqn. 4.2) can be used to numerically determine the value of the complex index of refraction at the specific wavelength of the incoming beam. The complex index of refraction (eqn. 4.3 on the following page) describes the nominal index of refraction but additionally includes an imaginary term to describe absorption of light in the material (commonly referred to as the extinction coefficient, κ).

$$\rho = \frac{r_p}{r_s} = \tan(\Psi)e^{i\Delta} \quad (4.1)$$

$$r_p = \frac{\tilde{n}_1 \sqrt{1 - \left(\frac{\tilde{n}_1}{\tilde{n}_2} \sin \Phi \right)} - \tilde{n}_2 \cos \Phi}{\tilde{n}_1 \sqrt{1 - \left(\frac{\tilde{n}_1}{\tilde{n}_2} \sin \Phi \right)} + \tilde{n}_2 \cos \Phi} \quad (4.2a)$$

$$r_s = \frac{\tilde{n}_1 \cos \Phi - \tilde{n}_2 \sqrt{1 - \left(\frac{\tilde{n}_1}{\tilde{n}_2} \sin \Phi \right)}}{\tilde{n}_1 \cos \Phi + \tilde{n}_2 \sqrt{1 - \left(\frac{\tilde{n}_1}{\tilde{n}_2} \sin \Phi \right)}} \quad (4.2b)$$

$$\tilde{n} = n + i\kappa \quad (4.3)$$

This type of analysis is sufficient for thick, isotropic samples without any
 325 surface layers (e.g. surface oxides or adsorbed gases), and can directly provide the value of \tilde{n} . However, once layers are stacked upon one another, the system becomes very difficult to analyze directly due to interference effects between the layers. It becomes necessary to use modeling techniques to determine the correct values of \tilde{n} and thickness (t) for each layer.

The power of ellipsometry as a high-resolution optical analysis technique stems from the use of phase and polarization changes. This allows the analysis to overcome the diffraction limit, and can be accurate down to angstroms. Properly modeling the system is critical for this analysis to be as precise as possible. Thus, there have been refinements of the ellipsometric method to
 335 greatly increase the amount of experimental data points, allowing the overall system to be over-determined and thus letting all of the systems parameters to be calculated.

Variable angle spectroscopic ellipsometry (VASE) is one of these variants. Spectroscopic ellipsometry differs from single-wavelength ellipsometry by utilizing a broad-band light source as opposed to a monochromatic source. By performing ellipsometric analysis at each of the wavelengths, one can determine the wavelength (and thus photon-energy) dependence of n and κ . This not only helps to improve data analysis (as it can generally be safely assumed that the values of n and κ are smooth functions of λ), but allows for the determination of many other properties of the material. Of specific importance is the complex dielectric function ($\tilde{\epsilon}$), which is related to \tilde{n} by the relation shown in equation 4.4. Knowing this function can allow for determination of electronic properties such as the bandgap energy, the absorption coefficient, amongst others. Finally, by obtaining spectra at a number of different incident angles, one directly provides additional data points across the entire wavelength spectrum. Even a small number of additional angles can quickly provide sufficient data points for the system to be over determined.

$$\tilde{\epsilon} = \epsilon_1 + i\epsilon_2 = \tilde{n}^2 \quad (4.4)$$

During this project, a VASE M-2000U system (figure 4.3 on the facing page) built by J.A. Woollam, inc. was used to collect all of the ellipsometric data.
 355 In addition, data analysis was performed using the WVASE32[©] package also

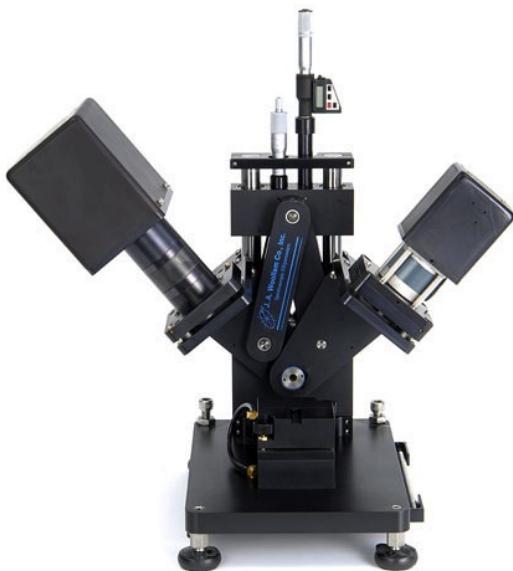


Figure 4.3: Photograph of the J.A. Woollam M-2000U variable angle spectroscopic ellipsometer (VASE)

provided by J.A. Woollam, inc. The system utilizes a rotating compensator and a CCD detector to greatly decrease data collection time by collecting data across the entire spectrum simultaneously. More information on this system is available from the J.A. Woollam, inc. webpage.¹

360 4.3 Phase Identification

4.3.1 X-Ray Diffraction

4.3.2 Grazing Incidence X-Ray Diffraction

4.4 Thermal Analysis

4.4.1 Thermogravimetric Analysis

365 Thermogravimetric analysis (TGA) is a very useful tool when attempting to determine the viability of a precursor in an ALD process. It allows for estimation of vaporization rate at various temperature rates as well as indications of chemical breakdown (i.e. thermalization) which would negate the precursor's usefulness.

370 At its core, TGA is a measurement of mass loss as a function of temperature or time. A small sample (1–10 mg) of material is placed in a microgram balance pan and suspended inside a furnace. The furnace is then heated at a specified



Figure 4.4: Photograph of the thermal analysis instrumentation used during this study, made by T.A. Instruments, inc.

rate while the sample mass is carefully monitored. For the experiments used in this study (evaluation of thermal vaporization and thermal degradation) it is
375 important to ensure that the testing environment is inert. This is accomplished by using a platinum pan in the microgram balance and constantly purging the furnace with a small flow of dry nitrogen gas. The heating rate can be varied according to a pre-determined program to provide more information at various individual temperatures.

380 This technique was used to evaluate various precursor candidates for the lead oxide half of the PbTiO_3 deposition procedure. The instrument used was a Q50 TGA device (fig. 4.4a). A detailed discussion of TGA procedures and the investigated chemicals can be found in subsequent chapters (see 6.1 on page 25 and 7.1 on page 33).

385 4.4.2 Differential Scanning Calorimetry

Differential scanning calorimetry (DSC) is a technique that allows for the determination of various critical temperatures for a material, and also can highlight changes in chemical structure due to degradation or other thermal processes.

390 DSC is the analysis of energy absorption as a function of temperature, which is the essence of calorimetry. DSC uses a sample and reference system

to isolate the energy absorbed by the sample from that of the holder pan. Sample sizes usually range from 0.3–2 mg of material; as the samples used in this study are volatile the sample pans are hermetically sealed to prevent
395 mass loss. The sample and reference pans are then placed inside a thermally insulated chamber. The temperature of each is carefully monitored, and differing amounts of heat are applied to negate the temperature difference between the sample and reference. The difference in absorbed heat as a function of temperature is then given as the result. In general, experiments include both
400 heating and cooling curves to gain a complete understanding of the different energies.

DSC was used to analyze the behavior of precursor chemicals around their evaporation and reaction temperatures. The main goal was to determine if the material underwent any thermally-activated degradation processes at either of
405 these two temperature ranges. At the evaporation temperature, the sample was generally cycled multiple times to simulate actual use in the ALD. These measurements were taken using a Q2000 DSC system (fig. 4.4b on the facing page) made by T.A. Instruments, inc.

Chapter 5

⁴¹⁰ Sample Fabrication

5.1 Precursor Selection

5.1.1 Titanium Source

The source of titanium that was used was titanium(IV) isopropoxide ($\text{Ti}-i\text{-Pr}$, $\text{Ti}(\text{OCH}(\text{CH}_3)_2)_4$). This compound is very commonly used in ALD literature. It is a liquid precursor with a high vapor pressure and reacts easily with most oxidizers; the most commonly used oxidant for this reaction is water vapor, similar to the TMA-H₂O reaction (see section [3.4 on page 7](#)).

5.1.2 Lead Source

Lead sources investigated included tetraphenyl lead (Ph_4Pb), bis(2,2,6,6-tetramethyl-3,5-heptanedionato)lead(II) (Pb(TMHD)_2), and lead(II) hexafluoroacetylacetone (Pb(HFAC)_2). These three compounds had been referenced in literature, were widely available from multiple chemical suppliers, and were expected to have the highest vapor pressure of the choices available at the time.

⁴²⁵ 5.1.3 Oxidizer

5.2 Substrate Preparation

Fabrication and preparation of substrates was an important part of the deposition process. Some substrates were purchased and simply cleaned, others needed to be fabricated or otherwise processed prior to cleaning and use in deposition. Three main types of substrates were used: thermally oxidized single crystal silicon (100) wafers, silicon wafers that have a thin layer of platinum deposited on the surface, and strontium titanate (100) single crystal substrates.

5.2.1 Si(100)

The silicon substrates were prepared simply. 4" silicon wafers with 200 nm of thermally grown oxide were diced into 1.5 cm x 1.5 cm pieces. When a

sample was to be used for deposition, it was cleaned by sonicating in acetone, isopropanol (each for one minute), and subsequently 5 minutes of sonication in deionized (DI) water. These were then air dried with dry nitrogen. Finally, the substrates were cleaned in a oxygen plasma cleaning system to remove any
440 remaining organic residues.

5.2.2 Platinized Si(100)

Platinized silicon substrates were prepared in a similar manner to the Si(100) samples. For the initial platinization, a large piece of silicon wafer (5 x 5 cm²) was prepared in the manner shown above. Then a 15 nm layer of
445 platinum was deposited via ALD (deposition recipe can be found in the appendix A.4 on page 42). The substrates were then cleaved into smaller pieces for later use.

If the samples are stored, it is recommended to again clean the samples in the standard procedure (see 5.2.1 on the previous page).

450 5.2.3 STO(100) and Nb:STO(100)

Oxide crystal substrates were prepared in such a fashion as to promote the formation of atomically flat terraces. This has the advantage of promoting a uniform surface species across the entire sample — the etchant used in this process leaves the sample titania-terminated.

455 To prepare these samples, the substrates are first pre-cleaned in a three step sonication process. The samples were cleaned for five minutes in each of acetone, methanol, and isopropyl alcohol. Subsequently, the samples were sonicated for fifteen minutes in DI water. Need to find reference and exact timings. Email Eric.
460 The substrates were then dipped into buffered hydrofluoric acid to etch for 35 seconds, then removed and flushed with copious DI water. Once the sample is thoroughly rinsed, the samples are dried using dry nitrogen.

After the etching process, the samples are annealed at 1050°C for two hours. Once the samples are cooled, they are ready for immediate use. AFM can be used to confirm the presence of atomic terraces.

465 If the samples are stored, it is recommended to again clean the samples in the standard procedure (see 5.2.1 on the preceding page).

5.3 Deposition Parameters

There are four main parameters that can affect the behavior of an ALD deposition.

⁴⁷⁰ **5.3.1 Growth Temperature**

The temperature of the growth chamber has a strong effect on reaction behavior. ALD reactions are sensitive to temperature, and will only proceed properly within a certain range known as the ‘ALD window.’ Outside of this range, the reaction enters one of a number of different regimes; these are determined by comparing the growth rate of the deposition to that of a reaction in the self-limiting saturated “ALD mode.”

If the growth temperature is less than the lower bound of the ALD window, the two regimes are condensation limited and activation energy limited. Condensation limited growth occurs when the substrate temperature is low enough that precursor condenses onto the surface without reacting with the presented sites. This causes higher than expected growth rates, and a lack of self-limiting behavior. If the reaction instead proceeds into the activation energy limited regime, molecules of precursor lack sufficient energy to react with the surface. This is characterized by lower deposition rates.

Conversely, if the reactor temperature is excessive the reaction again become anomalous. Decomposition limited growth, characterized by excessive deposition, is a result of thermal cracking of the precursor materials. This reaction is not limited to the surface, and accounts for the extra material being deposited. Lower deposition rates indicate that the temperature is sufficient to cause desorption of previously-reacted material from the sample.

For an ALD run to be successful, the acceptable temperature window for all of the reactions should overlap in some temperature range. This can become difficult with reactions requiring multiple metal precursors (e.g. PbTiO_3 , a combination of TiO_2 and PbO), as these can have widely varying ALD windows for their respective reactions.

5.3.2 Precursor Dosage

The dosage of precursor or oxidant to the surface is another parameter of critical importance. An ALD reaction requires a minimal amount of precursor to sufficiently saturate the surface, while it is beneficial to minimize any excess

500 precursor as it will be a wasted byproduct (minimizing costs, environmental impact, etc.).

The vaporization behavior of the precursor can have a dramatic impact on how simple or difficult it is to deliver a saturating dose to the surface. Some materials have readily available precursors with high vapor pressures; 505 titanium isopropoxide and trimethylaluminum (TMA) are both liquids, and tetrakis(dimethylamido)hafnium ($\text{Hf}\{\text{N}(\text{CH}_3)_2\}_4$) is a low-melting temperature solid. These are commonly used precursors for depositing their respective oxides. This vapor pressure becomes an important consideration when choosing a potential compound for use in ALD (as discussed in section 5.1 on 510 page 19).

Insufficient dosing is apparent in a deposition run by a slower than average growth rate, or also as a non-uniform deposition rate across the sample. However, overdosing is not readily apparent in an ALD-mode deposition. The dose must be lowered to a point where the dose is insufficient, and then increased 515 back to a saturating level.

Controlling the dose is dependent on injection time (which is the time the valve between the process line and the precursor storage vessel is open), precursor temperature, and the cycle duration (time between precursor injections). By increasing either the injection time or the precursor's temperature the dose 520 is increased, except in some cases with low vapor pressure materials. In this case, it can sometimes be found that the evaporation kinetics are slow and it takes additional time to build up a sufficient amount of vapor to provide a dose to the reactor.

If necessary, multiple doses of precursor can be delivered to the sample 525 during each cycle to increase the total delivered dose.

5.3.3 Purge Time

5.3.4 Exposure Time

5.4 Post-Deposition Annealing

Two types of annealing procedures were used in this study. Oven annealing, 530 with the simple use of a furnace in ambient atmosphere; and rapid thermal annealing (RTA), characterized by very high heating and cooling rates and performed in an inert atmosphere (dry N_2).

5.4.1 Oven Annealing

5.4.2 Rapid Thermal Annealing

⁵³⁵ **Chapter 6**

Data Collection and Analysis

6.1 Thermogravimetric Analysis

6.1.1 Data Collection

6.1.2 Analysis Methods

⁵⁴⁰ **6.2 Differential Scanning Calorimetry**

6.2.1 Data Collection

6.2.2 Analysis Methods

6.3 VASE and Modeling

⁵⁴⁵ Ellipsometry was used extensively to determine a variety of properties of the material. However, the primary goal of ellipsometric analysis was to determine the film thickness, in order to be able to determine the film growth rate (in terms of Å per deposition cycle) of the process.

6.3.1 Data Collection

⁵⁵⁰ In order to collect the experimental data, the following series of steps were followed:

1. Optics alignment
2. Ambient light compensation (DC offset)
3. Data collection at multiple angles

⁵⁵⁵ Alignment of the optics of the system is performed in the manner described in the ellipsometer's manual.² The system can have focusing optics installed which diminish the spot size of the analysis, which is useful if inhomogeneity is expected in the sample as this is a major problem for analysis (two of the main assumptions made by ellipsometric models are that the layers have consistent thicknesses and optical behavior across the analysis area). This is

560 done by manually adjusting the sample stage height and the sample surface plane. The system is designed so that when the incoming signal is maximized the sample is properly aligned with respect to the p- and s-planes defined by the equipment.

Once the system is aligned, the signal that is due to ambient light (not produced by the light source) must be compensated for. The M-2000U defines this as the “DC offset.”² The offset is calibrated automatically by the system by blocking the light source and measuring the signal from the surroundings. As the light present in the room is generally randomly polarized, the signal will be invariant to the polarizer settings. Correctly setting this value greatly decreases the uncertainty during the analysis phase; it mainly affects the degree of light depolarization measured by the system. The ellipsometer includes the depolarization in its calculation of the confidence interval for the final measurement. If the degree of ambient unpolarized light is not determined before the measurement, the depolarization will be nearly completely unrelated to the actual depolarization by the sample. In addition, the depolarization can be used by non-idealized models to determine such parameters as layer thickness variation, or internal interface roughness. This process will not be used for the remainder of this discussion, but more information can be found in the manual for the M-2000U.²

580 After the calibration steps have been completed, data collection can be performed. Three different incident angles were used for the data collection: 55°, 60°, and 65°. At each angle, the data was averaged over three hundred revolutions of the compensator to minimize noise in the experimental data. The system was set up to collect depolarization data simultaneously with the 585 ellipsometric parameters.²

If the sample is expected to be inhomogeneous, the focusing optics can be used and data collected at several different locations on the sample. This can provide data on how the growth process behaves spatially, such as if there is abnormal growth near the edges of the sample but homogeneous deposition as 590 one moves nearer to the center of the substrate.

6.3.2 Model Definition

The definition of the model is a critical part of the analysis procedure. The model dictates how the software package will perform its various calculations to predict the overall optical behavior, which it iteratively compares to the

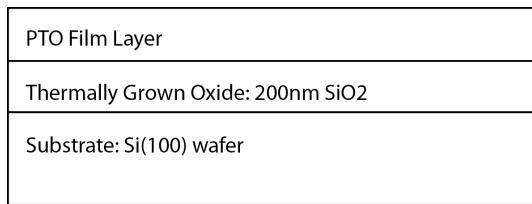


Figure 6.1: A simple graphical example of the model used for analysis of the film stack in the Si(100) samples. The parameters are t and the spectroscopic values of \tilde{n} for the PTO layer.

595 experimentally determined Ψ and Δ .

Simply put, the model is defined as a bulk (semi-infinite) substrate layer, with a nominal number of nano- to micrometer thick layers stacked upon it. Each layer is modeled with a prediction of optical constants at each test wavelength. These optical constants can be provided as a table of experimentally 600 determined results, which are available for many commonly used materials such as silicon, silica, titania, amongst others; they can also be predicted using a variety of different models. These can be empirical predictors, such as the Cauchy dispersion, or based upon physical properties of the material, oscillator-based models for example. The model types relevant to this work 605 are discussed in more detail in subsequent sections.

The four different substrates require different material layer stacks to properly represent them, and each poses individual challenges for characterization. The Si(100) substrate that was most commonly used for this work was the simplest to model. It can be represented as a substrate layer of silicon, with 610 a 200 nm layer of silica on top. The deposited film would be layered above the SiO_2 layer (see fig. 6.1 for a schematic representation). A large number of these substrates were analyzed for their oxide layer thickness, where the only parameter to be fit was the layer thickness. It was found that the nominal oxide layer was 200 ± 5 nm thick. This was consistent enough that 615 200 nm could be used for the initial thickness estimate for all samples using this substrate, and after the ALD layer was analyzed this thickness could also be included in the fit to confirm the true dimensions of the oxide layer. The substrate with thermally-grown oxide was preferred in comparison to silicon with only native 620 oxide layer; this is because the thicker layer of transparent oxide helps to generate large oscillations in Ψ and Δ , which assists in the analysis (particularly the thickness, where the fringes are very closely related to this parameter).

6.3.3 Analysis Procedure

Once the data was collected, a specific series of steps was followed in order to obtain the highest degree of accuracy from the model. All steps were performed on the PTO layer. The procedure went as follows:

1. High- λ Cauchy Model
2. Direct Calculation of n and κ
3. Conversion to Oscillator Model
4. Refinement of Oscillator Layer Parameters

This first step takes advantage of the transparency of the film at high wavelengths (low energies) where the photon energy is below the optical bandgap of the material. In this region, the Cauchy model can be used. The Cauchy model is empirical rather than physically descriptive, and best used for amorphous materials such as polymeric films, however the assumptions required for reasonable accuracy are met when absorption in the film layer is minimized (therefore $\kappa(\lambda) \approx 0$). The equations used in the Cauchy model are shown in equation 6.1. Generally, analysis during this step was performed in the spectral region where $\lambda = 600 - 1000$ nm ($E_{ph} = 2.06 - 1.24$ eV).

$$n(\lambda) = A_n + \frac{B_n}{\lambda^2} + \frac{C_n}{\lambda^4} + \dots \quad (6.1a)$$

$$\kappa(\lambda) = A_\kappa e^{B_\kappa \left(\frac{hc}{\lambda}\right) - C_\kappa} \quad (6.1b)$$

Once reasonable estimates for n , κ , and the film layer thickness are obtained at the higher wavelengths, the second step of the analysis is to generate values of \tilde{n} for the rest of the spectrum. The film thickness parameter is fixed at the value determined from the Cauchy model. The values of n and κ are allowed to be determined freely without the use of a model (i.e. directly determined by use of the Fresnel relations). This type of modeling is not physical, but assists in the generation of the oscillator-based model in the next step. The software package is then instructed to run a point-by-point fit of the data from highest- λ to lowest- λ , attempting to minimize the change in n or κ between adjacent data points.

This model is then inputted into a oscillator model. For the analysis of these films, a Tauc-Lorentz oscillator model was utilized. The oscillator models

used by WVASE32 are expressed in terms of the complex dielectric function $\tilde{\epsilon}$, which relates to \tilde{n} via the relationship in equation 6.2. The Tauc-Lorentz model changes the Lorentzian model by allowing for some absorption below the fundamental bandgap energy, which would be due to defect states and other intra-band transition mechanisms. The Tauc-Lorentz model uses the parameterization shown in equation 6.3.^{2,3} ϵ_1 is provided here in a condensed version (eq. 6.3a); the full expanded version, and its derivation via Kramers-Kronig integration (whose relations are shown in equation 6.4) from ϵ_2 , has been presented by Jellison and Modine.³

$$\tilde{\epsilon} = \epsilon_1 + i\epsilon_2 = \tilde{n}^2 \quad (6.2a)$$

$$\epsilon_1 = n^2 - \kappa^2 \quad (6.2b)$$

$$\epsilon_2 = 2n\kappa \quad (6.2c)$$

$$\epsilon_1 = \frac{2}{\pi} P \int_{E_g}^{\infty} \frac{\xi \epsilon_2(\xi)}{\xi^2 - E^2} d\xi \quad (6.3a)$$

$$\begin{cases} \epsilon_2(E) = \frac{AE_0C(E - E_g)^2}{(E^2 - E_0^2)^2 + C^2E^2} \cdot \frac{1}{E} & E > E_g \\ \epsilon_2(E) = 0 & E \leq E_g \end{cases} \quad (6.3b)$$

$$(6.3c)$$

660

$$\epsilon_1(\omega) - 1 = \frac{2}{\pi} P \int_0^{\infty} \frac{\omega' \epsilon_2(\omega')}{\omega'^2 - \omega^2} d\omega' \quad (6.4a)$$

$$\epsilon_2(\omega) = -\frac{2\omega}{\pi} P \int_0^{\infty} \frac{\epsilon_1(\omega') - 1}{\omega'^2 - \omega^2} d\omega' \quad (6.4b)$$

The WVASE32 software package allows one to use a graphical interface to provide initial guesses for the various fit parameters. At times this required multiple oscillators to best fit the predicted ϵ_2 function. Once this has been set, all of the parameters affecting ϵ_2 (A, E_0, C, E_g) are marked to be included in the fit. The software is then instructed to perform a best-fit of the oscillator to ϵ_2 . Once this operation completes, the software is set to fit to ϵ_1 and only the value of the ϵ_1 offset is allowed to be fit. Finally, the software is set to optimize vs both ϵ_1 and ϵ_2 , and all parameters are included. This completes the initial setup of the oscillator model.

670

Finally, the model is set to also allow the layer thickness to be fit and a general fit to the entire experimental dataset is performed. This provides the best guess to the physical values of the film. The thickness calculated by this procedure matches very closely to measurements performed by other methods
675 (e.g. SEM imaging or AFM measurement of a lithographically created step). The bandgap parameter also matches well to literature values (when using standard samples of well defined materials, such as a thin layer of titania).

If similar deposition parameters are utilized, it is possible to save the parameterized model for later use. This allows the analysis to be streamlined
680 when the material is expected to remain constant, for example if tests of deposition at different layer thicknesses are performed. In this case, the material would have optical behavior very similar to the initial sample, and the oscillator model would be sufficiently close to valid parameters to be used directly for fits. All that would need to be adjusted initially would be the estimated layer
685 thickness. If the fit fails to produce useful data, such as unreasonable values for any of the parameters or very large confidence intervals, it is recommended to proceed with the entire standard analysis procedure.

Experimental data sets and the resulting fitted model for selected samples are presented in appendix [A.6 on page 42](#).

690 6.4 Composition Analysis

In order for the desired phase to be preferred, without any impurity phases precipitating, it is important to be able to control the stoichiometry of the produced film. Previous reports have shown that there is a close relationship between the composition of the film and the final resultant phase (see figure 6.2
695 [on the next page](#)).

6.4.1 Energy-Dispersive X-Ray Spectroscopy

6.4.2 X-Ray Fluorescence Spectroscopy

6.5 X-Ray Diffraction

6.5.1 Grazing Incidence XRD

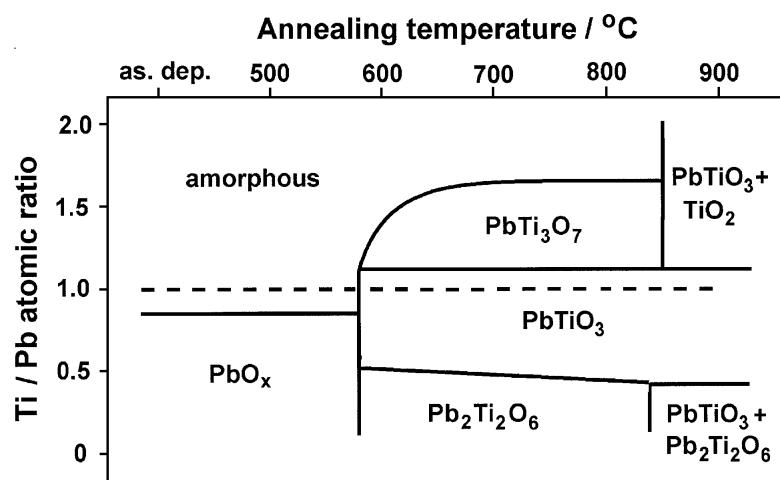


Figure 6.2: Graphic illustrating preferred phase of an annealed film at a range of stoichiometric ratios and annealing temperatures. A slight excess of Pb in the system is expected to help stabilize the perovskite PbTiO_3 phase.⁴

700 **Chapter 7**

Results

As discussed, the goal of this study was to determine methods for atomic layer deposition of ferroelectric oxides. In the process of realizing this goal there were a number of different areas of study. The first is the analysis of thermal and chemical behavior of the various potential precursors, during which TGA and DSC were primarily used (see [6.1 on page 25](#) and [6.2 on page 25](#)). Secondly, the analysis of the film growth behavior under various conditions, this was primarily measured and analyzed using the ellipsometric techniques discussed earlier (see [6.3 on page 25](#)). Third, the film deposition needed to be tuned to produce films with a stoichiometric composition, as this was expected to produce films which would crystallize into the desired perovskite phase, see section [6.4 on page 30](#) for the methods used for this characterization. Fourth, the phase of the crystallized film was analyzed in detail to determine behavior of the films post-annealing. XRD was used extensively for this task (see [6.5 on page 30](#)).

7.1 Thermal Analysis

While a viable titanium precursor was well identified in literature as well as experimentally, as was the oxidizers that were used, there was no such universally accepted chemical used in ALD to provide a source of lead. The primary issue was either a lack of chemical stability or a undesirably low volatility in the compounds that currently were being used. TGA and DSC was performed on a number of potential candidates (see [5.1 on page 19](#) for more details) in order to gauge the performance of these materials.

7.1.1 Thermogravimetric Analysis

Thermogravimetric analysis allows for the estimation and comparison of volatility between multiple samples. It was used in this study to compare $\text{Pb}(\text{HF}\text{Ac})_2$ and $\text{Pb}(\text{TMHD})_2$, primarily their individual volatilities — by analyzing the mass loss at various temperatures — and whether or not there are

indications of imperfect evaporation. The data collected for these materials
730 can be found in section A.5.1 on page 42.

By analyzing the TGA curve for $\text{Pb}(\text{HFAc})_2$, found in figure A.2a on
page 43, there are a number of features that are immediately noticeable. First
of these is the multiple stages of evaporation present in the curve. These occur
at approximately 170, 190, and 230°C. A TGA curve for a material that is
735 purely evaporative, e.g. pure water, will have a single step in the curve. Additional steps indicate that other processes are activating, and causing changes
to the compound affecting the mass loss.

More detail can be seen by performing a derivation on the TGA curve,
giving the mass loss rate. This plot can be found in figure A.2b on page 43.
740 This plot shows a large number of peaks, particularly around 200°C, indicative
of numerous features on the original plot.

In comparison

7.1.2 Differential Scanning Calorimetry

7.2 Ellipsometry

745 7.3 Composition

7.4 X-Ray Diffraction

Chapter 8

Conclusions

8.1 Future Work

- 750 Optimize consistency
- Doped materials
- Other material systems (BFO, BST, etc.)

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⁷⁶⁵ Appendix A

Supplemental Information

A.1 List of Chemicals

	Ph ₄ Pb
	Ti—o—i—Pr
⁷⁷⁰	Pb(TMHD) ₂
	Pb(HFAc) ₂
	Buffered HF
	N ₂
	O ₂
⁷⁷⁵	O ₃
	H ₂ O
	IPA
	Acetone

A.2 List of Samples

Table A.1: A list of samples used during the course of this project.

Temperature (°C)	Pb:Ti Deposition Ratio	Run Number	Substrate Type	Super- cycles	Total cycles	Anneal Type	Anneal Temp (°C)	Anneal Time	XRD	Notes
200	1:1	3	Si	250	500				N	Blah
200	1:2	2	Si	250	750				N	
200	3:1	30	- Si - Pt	160	640				N	
250	1:1	0	Si Si	625	1250	Oven - Amb - Oven - Amb -	650 900	2h 2h	Y Y	Green sample, no apparent change Sample turned a purplish blue color, still shiny but not as shiny as untreated Small chip, no longer shiny, odd structures present after annealing (via opt. micro), appears to have formed nano-rods after anneal (see SEM images). Made more samples in RTA at 600, 700, 800, 900 °C, similar results on all but 600 °C test
250	1:1	1	Si	475	950				N	
250	1:2	6	Si	250	750				N	
250	3:1	13	Si	250	1000				N	Tried 60s between Pb pulses, needs a little more. Sample appeared uniform except for one spot (effect of tape?). Yellow spot has different thickness (23.5nm) and different osc. parameters.
250	3:1	16	Si	150	600	RTA	650	1m	N	
250	3:1	19	- Si - Pt	100	400				N	Replaced precursor prior to this run.
250	3:1	20	Si Pt	200	800	- Fast Oven	650	1m	Y	Runs appear to be growing well. Spectacular colonization on the reactor surface where platinum was deposited.
250	3:1	21	- Si - Pt	150	600	- Fast Oven	650	1m	N	
250	3:1	22	- Si - Pt	150	600	- Fast oven	650	1m	Y	Discolored, not homogeneous across sample..

Table A.1: A list of samples used during the course of this project.

Temperature (°C)	Pb:Ti Deposition Ratio	Run Number	Substrate Type	Super- Cycles	Total Cycles	Anneal Type	Anneal Temp (°C)	Anneal Time	XRD	Notes
250	3:1	23	- Si - Pt	200	800	- Fast oven	-	-	-	-
250	3:1	24	- Si - Pt	200	800	-	-	-	N Y	
250	3:1	25	- Si - Pt	60	240	-	-	-	N -	
250	2:1	26	- Si - Pt	80	240	-	-	-	N -	
250	5:2	27	- Si - Pt	110	770	-	-	-	N -	
250	3:1	28	- Si - Pt - STO	120	480	-	-	-	N -	
250	3:1	29	- Si - Pt	50	200	-	-	-	N -	
250	2:1	18	Si	200	600				N	Might be running low on Pb(TMHD)2 precursor, going to add some more this afternoon
300	1:1	4	Si	250	500				N	
300	1:2	5	Si	200	600				N	
300	2:3	9	Si	175	875				N	May have gotten recipe timing wrong, might have to re-do this test...
300	3:2	10	Si	50	250				N	
300	3:1	11	Si	150	450				N	

⁷⁸⁰ **A.3 ALD Reactor Diagram**

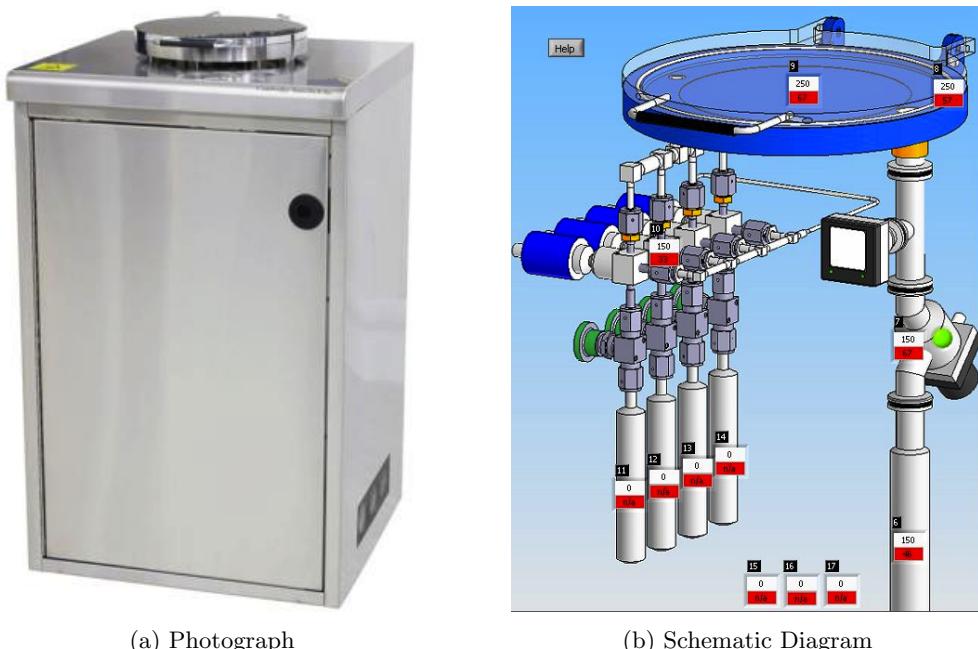


Figure A.1: Cambridge NanoTech, inc. Savannah S100 ALD reactor. Precursors are stored in heated cylinders, flow up to the reaction zone, and byproducts are pumped out of the vacuum line on the right side. Each zone can be individually temperature controlled.

A.4 Recipes for S100 ALD System

Some of these recipes (HfO_2 , Pt, and TiO_2) were provided courtesy of Cambridge NanoTech, inc.⁵

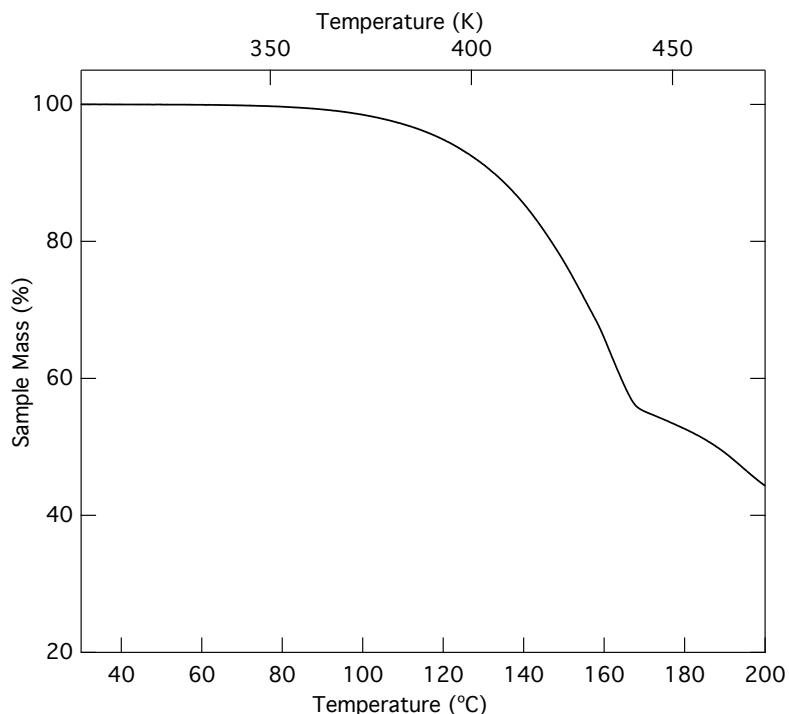
A.5 Thermal Analysis Results

⁷⁸⁵ **A.5.1 Thermogravimetric Analysis**

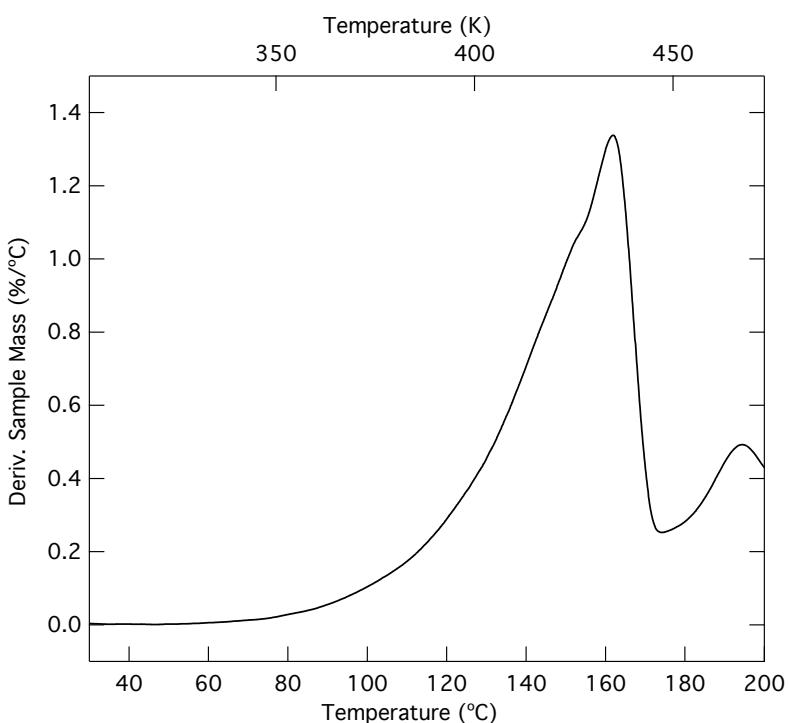
A.5.2 Differential Scanning Calorimetry

A.6 Ellipsometry Results

A.7 XRD Results

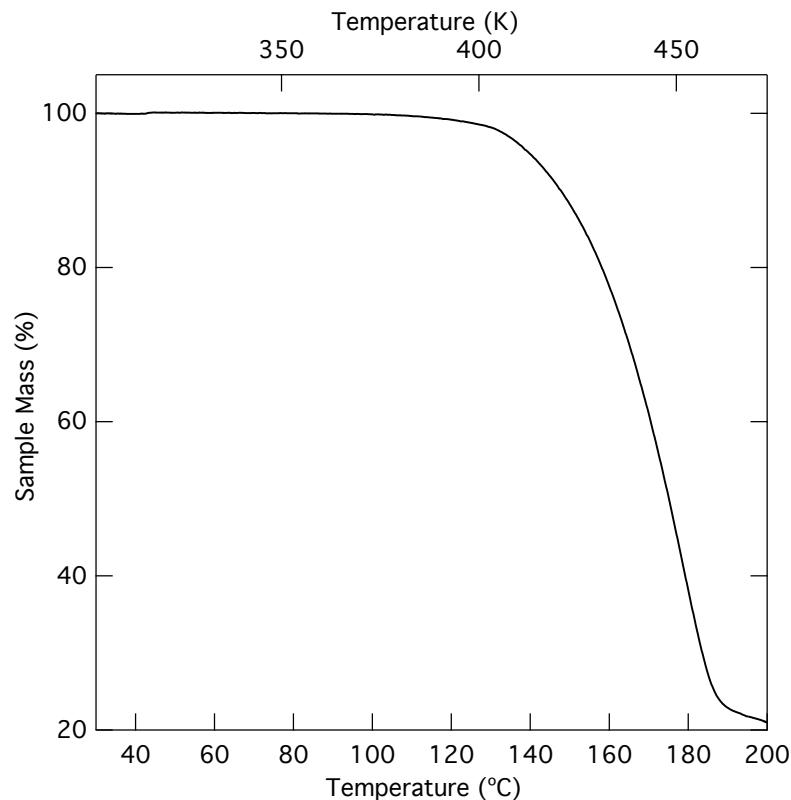


(a) Mass vs. Temperature

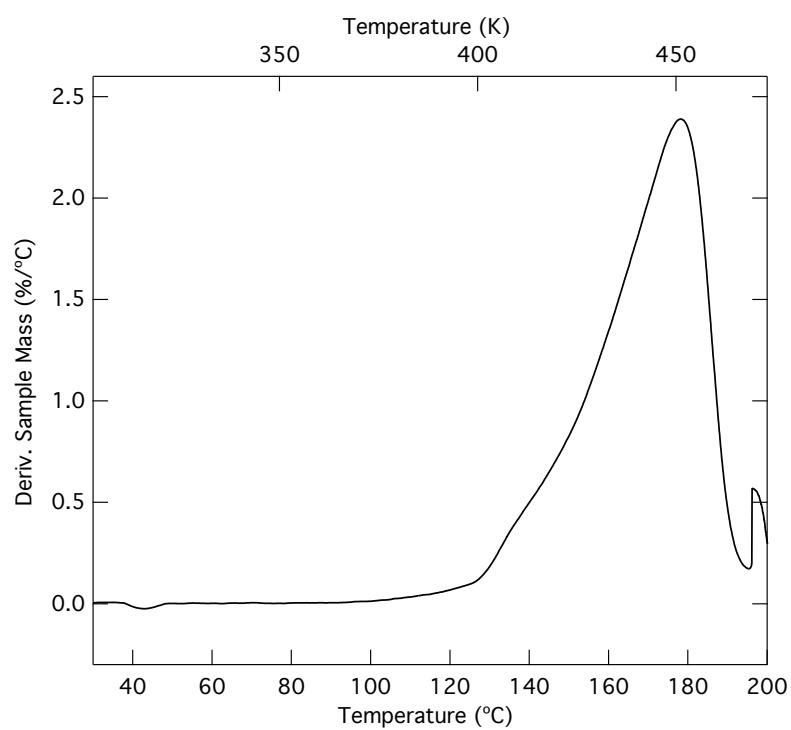


(b) Derivative of Mass vs. Temperature

Figure A.2: Plots of the results from TGA experiments on $Pb(HF_2)_2$. The plot shown in (a) gives the raw data showing the current mass as a function of temperature. (b) gives the same data, transformed to show the derivative of mass. Thus (b) shows the rate of mass loss at a given temperature. Initial sample mass: 6.092 mg



(a) Mass vs. Temperature



(b) Derivative of Mass vs. Temperature

Figure A.3: Plots of the results from TGA experiments on Pb(TMHD)_2 . As in figure A.2 on the preceding page, (a) presents the actual mass as a function of temperature, while (b) gives the derivative of that function. Initial sample mass: 3.719 mg

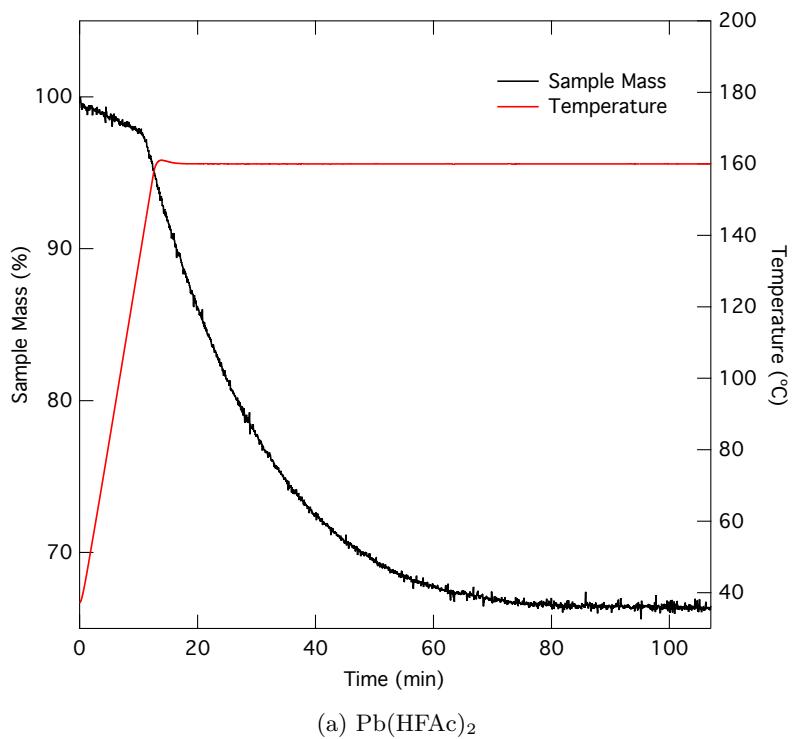
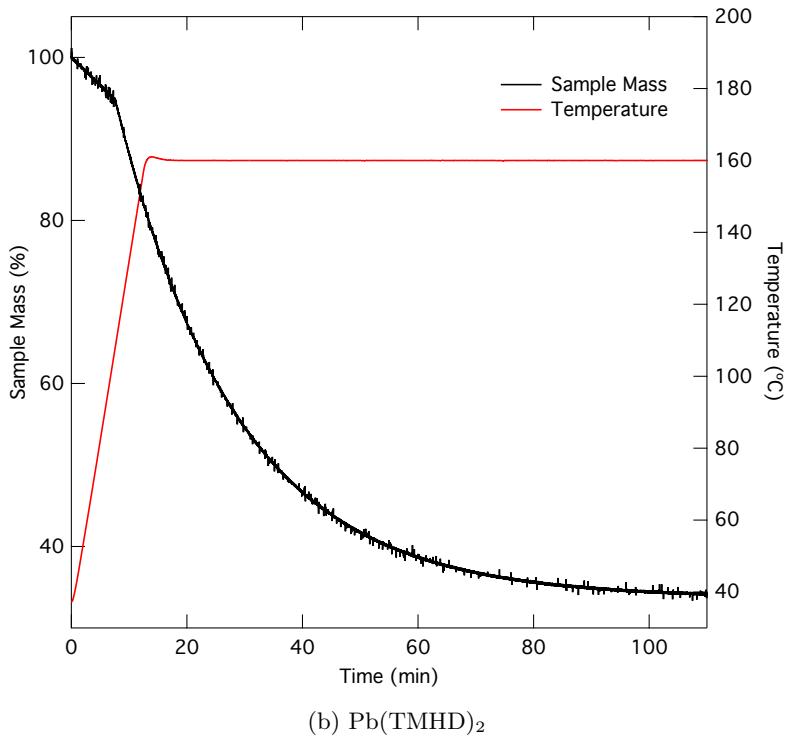
(a) $\text{Pb}(\text{HFAc})_2$ (b) $\text{Pb}(\text{TMHD})_2$

Figure A.4: Plots of the results from ramp-and-hold TGA experiments designed to investigate residue formed after complete evaporation at a given temperature. From the TGA experiments seen above (figs. A.2 on page 43 and A.3 on the facing page), a common temperature of 160°C was chosen for this experiment. Sample sizes were 3.921 mg and 4.381 mg for $\text{Pb}(\text{HFAc})_2$ and $\text{Pb}(\text{TMHD})_2$ respectively.

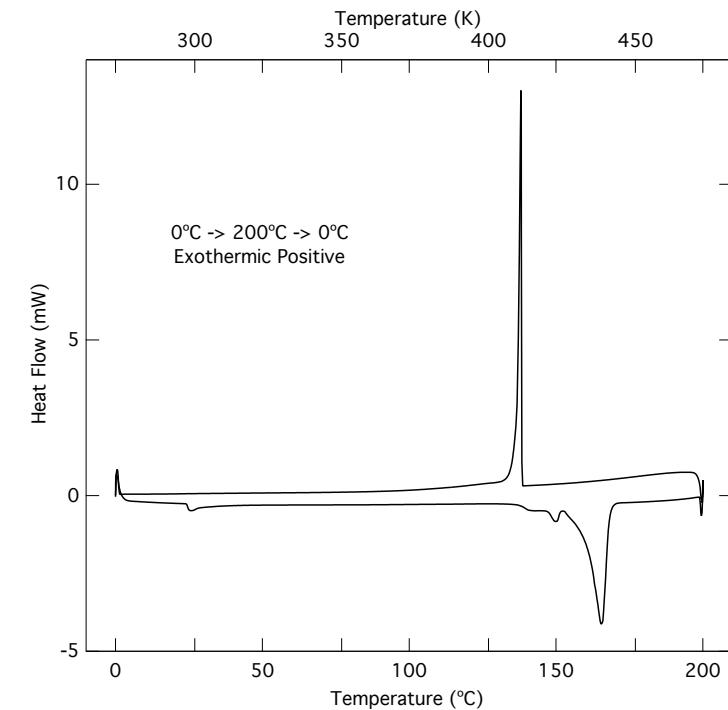
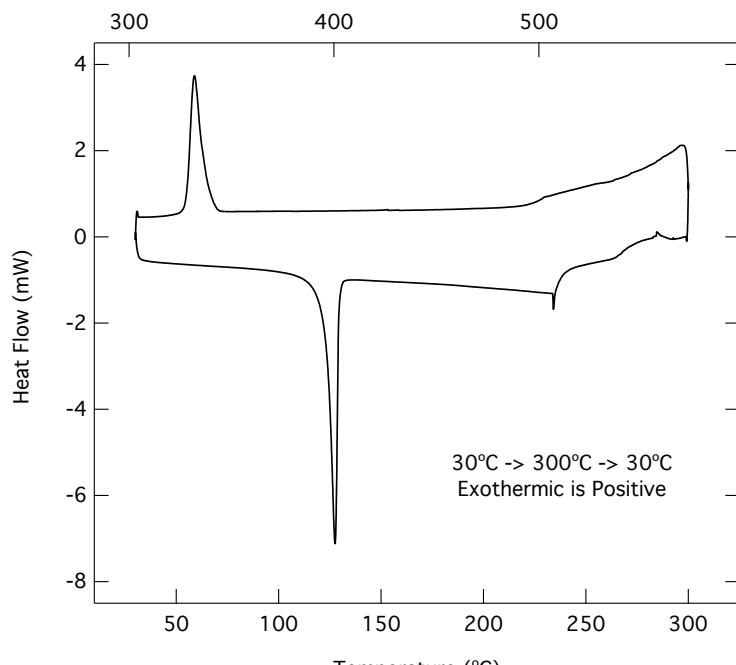
(a) $\text{Pb}(\text{HFAC})_2$ (b) $\text{Pb}(\text{TMHD})_2$

Figure A.5: This plot shows results from DSC experiments carried out during the course of this project. (a) shows the results from analyzing the $\text{Pb}(\text{HFAC})_2$ compound. (b) gives the data from $\text{Pb}(\text{TMHD})_2$. Exothermic (heat release) flow is positive in both plots.