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The generalized dilution equation is: $C_1V_1 = C_2V_2$ The concentration x volume before dilution = the concentration x volume after dilution Let $C_2 = 250 \text{ mL}$ and $V_2 = 0.60 \text{ M Al}_2(\text{SO}_4)_3$ Let $C_1 = 2.0 \text{ M Al}_2(\text{SO}_4)_3$ $C_1V_1 = C_2V_2$ $2.0 \text{ M} \times V_1 = 0.60 \text{ M} \times 250 \text{ mL}$ $V_1 = 0.60 \text{ M} \times 250 \text{ mL} / 2.0 \text{ M} = 75 \text{ mL}$ Add 75 mL of 2.0 M $\text{Al}_2(\text{SO}_4)_3$ solution to water to make a total volume of 250 mL.

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