

## *Section 103 Reactions In Aqueous Solutions Answers*

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11.3 Reactions in Aqueous Solution Reactions that occur in aqueous solutions are double-replacement reactions. The products are precipitates, water, or gases. Lesson Summary Net Ionic Equations Net ionic equations show what species present in solution actually are part of the chemical reaction.

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Section 11.3 Reactions in Aqueous Solution 343 An ion that appears on both sides of an equation and is not directly involved in the reaction is called a spectator ion. When you rewrite an equation leaving out the spectator ions, you have the net ionic equation. The net ionic equation is an equation for a reaction in solution that shows

**11.3 Reactions in Aqueous Solution - Evaluation 2016**

Chapter 4 Practice Worksheet: Reactions in Aqueous Solutions 1. List the three general classes of chemical reactions: precipitation, acid-base neutralization, and redox reactions 2. How can you identify each of the three reaction types above (e.g., what characteristic defines each one?)?

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The reaction of aqueous solutions of silver nitrate and sodium chloride to form solid silver chloride and aqueous sodium ... 11.3 Reactions in Aqueous Solution 27 > By using the general rules for solubility of ionic compounds, you can predict the formation of a precipitate.

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In a(n) reaction, a compound breaks down into two or more products. 8. In a(n) reaction, the reactants are a compound and an element, and the products are a different compound and element. 9. A(n) reaction takes place between two compounds in aqueous solution. 10. Oxygen is always a reactant in a(n) reaction. 11.3 Reactions in Aqueous Solution

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