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The generalized dilution equation is:  $C_1V_1 = C_2V_2$  The concentration  $\times$  volume before dilution = the concentration  $\times$  volume after dilution Let  $C_2 = 250 \text{ mL}$  and  $V_2 = 0.60 \text{ M Al}_2(\text{SO}_4)_3$  Let  $C_1 = 2.0 \text{ M Al}_2(\text{SO}_4)_3$   $C_1V_1 = C_2V_2$   $2.0 \text{ M} \times V_1 = 0.60 \text{ M} \times 250 \text{ mL}$   $V_1 = 0.60 \text{ M} \times 250 \text{ mL} / 2.0 \text{ M} = 75 \text{ mL}$  Add 75 mL of 2.0 M  $\text{Al}_2(\text{SO}_4)_3$  solution to water to make a total volume of 250 mL.

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