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Colligative properties depend on the number of dissolved particles, so the solution with the greater number of particles in solution will show the greatest deviation. When NaCl dissolves, it separates into two ions,  $\text{Na}^+$  and  $\text{Cl}^-$ .

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In all solutions, whether gaseous, liquid, or solid, the substance present in the greatest amount is the solvent, and the substance or substances present in lesser amounts are the solute(s). The ...  
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Solutions Colligative Properties • Changes in colligative properties depend only on the number of solute particles present, not on the identity of the solute particles. • How would the ionic compounds and covalent compounds behave.....

### Chapter 13 Properties of Solutions

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Molarity. Molarity is the number of moles of solute per liter of solution. It is abbreviated with the symbol M, and is sometimes used as a unit of measurement, e.g. a 0.3 molar solution of HCl. In that example, there would be 0.3 moles of HCl for every liter of water (or whatever the solvent was).

### General Chemistry/Properties of Solutions - Wikibooks ...

172 Guided Reading and Study Workbook SECTION 16.3 COLLIGATIVE PROPERTIES OF SOLUTIONS (pages 487–490) This section explains why a solution has a lower vapor pressure, an elevated boiling point, and a depressed freezing point compared with the pure solvent of that solution. Vapor Pressure Lowering (pages 487–488) 1.

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a solution that holds more dissolved solute than is required to reach equilibrium at a given temperature Henry's law at a given temperature the solubility of a gas in a liquid is directly proportional to the pressure of the gas above the liquid

### 16.1 properties of solutions Flashcards | Quizlet

Practice-Problem-6660020 book December 22, 2010 8:37 Chapter 16: Colligative Properties of Solutions 45 16-4. The mole fraction of  $(\text{NH}_4)_2\text{SO}_4(\text{aq})$  is given by  $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$  Because  $(\text{NH}_4)_2\text{SO}_4(\text{aq})$  is a strong electrolyte, it dissociates completely into  $\text{NH}_4^+(\text{aq})$  and  $\text{SO}_4^{2-}(\text{aq})$  ions. Assume a one ...

### CHAPTER 16. Colligative Properties of Solutions

Molarity ( M ) is the number of moles of solute dissolved in one liter of solution. To calculate the molarity of a solution, divide the moles of solute by the volume of the solution. Note that the volume involved is the total volume of the resulting solution, not the volume of the solvent alone.

### 16.1 Properties of Solutions 16

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their chemistry and physical properties Most of the chemistry we deal with in the world (and in our bodies) takes place in solution, so it is important to know what factors influence the solubility of a substance, and to understand the physical properties of the resulting mixture.

### **Essential Background for General Chemistry**

This is "Properties of Solutions", section 9.4 from the book Introduction to Chemistry: General, Organic, and Biological (v. 1.0). For details on it (including licensing), click here . This book is licensed under a Creative Commons by-nc-sa 3.0 license.

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The relative level of acidity of alkalinity of a solution can be shown by using indicators. Arrhenius acids yield  $\text{H}_3\text{O}^+$  (hydronium ions) as the only positive ions in an aqueous solution. Arrhenius bases yield  $\text{OH}^-$  (hydroxide ions) as the only negative ions in an aqueous solution. Given properties, identify substances as Arrhenius acids or ...

### **Unit 6 - Solutions Workbook - Dr. G's Chemistry**

Workbook If your instructor recommends the Active Learning Workbook, do Questions, Exercises, and Problems 53-65. Chapter 16-Assignment E: Summary and Review Learn well the language of solutions and use it correctly. In addition to the terms at the beginning of the chapter, know exactly what is meant by percentage and molarity, and, if

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