

***Holt Chemfile Problem Solving Workbook Limiting Reactants  
Answers***

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Holt ChemFile: Problem-Solving Workbook 51 Mole Concept Name Class Date Problem Solving continued Sample Problem 2 A student needs 0.366 mol of zinc for a reaction. What mass of zinc in grams should the student obtain? Solution ANALYZE What is given in the problem? amount of zinc needed in moles What are you asked to find? mass of zinc in grams ...

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Holt ChemFile: Problem-Solving Workbook 127 Percentage Yield Name Class Date Problem Solving continued Sample Problem 2 Acetylene,  $C_2H_2$ , can be used as an industrial starting material for the production of many organic compounds. Sometimes, it is first brominated to form 1,1,2,2-tetrabromoethane,  $CH_2Br_2$ , which can then be

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Holt ChemFile: Problem-Solving Workbook 191 Stoichiometry of Gases Name Class Date Problem Solving continued Rearrange the ideal-gas-law equation to solve for the unknown quantity,  $V$ .  $PV = nRT$   $V = \frac{nRT}{P}$  COMPUTE EVALUATE Are the units correct? Yes; units canceled to give liters of  $SO_2$ . Is the number of significant figures correct?

**Holt ChemFile Problem Solving Workbook 191 Stoichiometry ...**

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Holt ChemFile: Problem-Solving Workbook 261 pH Name Class Date Problem Solving continued Sample Problem 1 A HCl solution has a concentration of 0.0050 M. Calculate  $[OH^-]$  and  $[H_3O^+]$  for this solution. HCl is a strong acid, so assume it is 100% ion-ized. Solution ANALYZE What is given in the problem? the molarity of the HCl solution, and the fact

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Problem Solving continued Sample Problem 1 Ammonia is made industrially by reacting nitrogen

and hydrogen under pressure, at high temperature, and in the presence of a catalyst. The ... Holt ChemFile: Problem-Solving Workbook 102 Stoichiometry Name Class Date Problem Solving continued

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Holt ChemFile: Problem-Solving Workbook 280 Titrations Name Class Date Problem Solving continued Sample Problem 3 A supply of NaOH is known to contain the contaminants NaCl and MgCl<sub>2</sub>. A 4.955 g sample of this material is dissolved and diluted to 500.00 mL with water. A 20.00 mL sample of this solution is titrated with 22.26 mL

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Holt ChemFile: Problem-Solving Workbook 58 Mole Concept Name Class Date Problem Solving continued CONVERTING NUMBER OF ATOMS OF AN ELEMENT TO MASS Sample Problem 4 uses the progression of steps 1 → 2 → 3 to convert from the mass of an element to the number of atoms.

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