KINETICS I:

THERMODYNAMICS DESCRIBES <u>DIRECTIO</u>N OF REACTION. KINETICS DESCRIBES TIME-DEPENDENCE.



VOCAB:

ORDER OF REACTION:

BULK is MICROSCOPIC

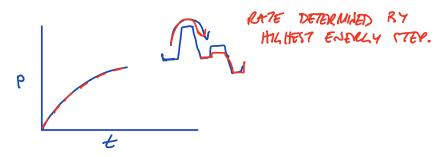
ZERUTH: RATE ~ K
FIRST: RATE ~ K(A)
SECOND. RATE ~ K·CAJCBJ

- RATE LIMITING STEP
MICROSCOPIC

CONTROLS RATE.

KEY I PBAS:

- 1. YOU CAN (USUALLY) ONLY SEE BULK RATES, NOT MICRISCOPIC RATES.
- 2. YOU CAN'T SEE STOPS AFTER LATE-LIMITISH STEP.



3. YOU HAVE TO BE CALGEUL INFERLING MILROSCOAL MECHANISTIC INFORMATION FROM MACROSCOPIC PATES.

RATE LAWS:

$$\frac{dA}{[A]} = -kdt$$

$$h(A) = -kt + c$$
 $ATt = \emptyset$, $A = A$.

$$A7t=\emptyset, A=A.$$

$$l_{\Lambda}(A) = -kt + \lambda(A_{\bullet})$$

$$h(A) - h(A) = -kt$$

$$A/A_0 = e^{-kt}$$

$$A = A_0 e^{-kt}$$

LET'S ADD A BACK REACTION: WILL THIS CHANGE THE ORDER? NO.

$$A \stackrel{k_{+}}{\rightleftharpoons} B \qquad \frac{dA}{dk} = -k_{+} (A) + k_{-} (B)$$

$$\frac{dB}{dt} = k_{+}(A) - k_{-}(B) \qquad -\frac{dA}{dt} = \frac{dB}{dt}$$

$$-\frac{dA}{dt} = \frac{dB}{dt}$$

$$[A]_o + (B]_o = (A)_e + (B)_e = (A)_e + (B)_e$$

REARRANGE/SUBSTITUTE:

$$\frac{d\Omega}{dt} = k_{+}(R) - k_{-}(A_{e} + k_{e}, A_{e} - A)$$

$$\frac{dt}{dt} = k_{+}(A_{-}) - k_{-}A_{e} - k_{-}k_{e} + k_{-}A$$

$$= k_{+}A - k_{+}A_{e} - k_{-}A_{e} + k_{-}A$$

$$= k_{+}A - k_{+}A_{e} - k_{-}A_{e} + k_{-}A$$

$$= k_{+}(A - A_{e}) - k_{-}(A - A_{e})$$

$$\frac{dB}{dt} = (A - A_{e})(k_{+} - k_{-})$$

$$\frac{dA}{dt} = (A - A_{e})(k_{+} - k_{-})$$

$$\frac{dA}{dt} = (k_{+} - k_{-})dt$$

$$A - A_{e} = (k_$$

YOU CAN ALSO WRITE AS PATE MATRIX: